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Surface tension of spherical drops from surface of tension

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The determination of surface tension of curved interfaces is a topic that raised many controversies during the last century. Explicit liquid-vapor interface modelling (ELVI) was unable up to now to reproduce interfacial behaviors in drops due to ambiguities in the mechanical definition of the surface tension. In this work, we propose a thermodynamic approach based on the location of surface of tension and its use in the Laplace equation to extract the surface tension of spherical interfaces from ELVI modelling. © 2014 AIP Publishing LLC. [<http://dx.doi.org/10.1063/1.4862149>]

I. INTRODUCTION

Molecular simulation has become a powerful method in the last decade to quantitatively predict the surface tension of complex fluids (γ).^{1,2} Whereas the calculation of γ for planar geometries is now under control, many controversies remain about the dependence of the surface tension on the drop size, and the validity of the mechanical definition of the local pressure^{3–7} resulting from the arbitrariness of the choice of contour to define the pairwise interactions. This leads to an ambiguity in the definition of the local elements of the pressure tensor in the spherical geometry. Despite this ambiguity, two mechanical methods are widely used in molecular simulation of explicit liquid-vapor interface (ELVI): the Irving-Kirkwood (IK)⁸ and the Harasima (H)⁹ approaches. On the contrary, the density functional theory (DFT) calculations allow an accurate and rigorous estimation of $\gamma(R)$ with respect to the drop size (R).^{5,7} From the IK method, the following tendencies were found: (i) a monotonic increase of γ with respect to the drop size while a non-monotonic behavior was observed from DFT; (ii) the values of γ significantly disagree with the DFT results; and (iii) a positive Tolman length (δ) was found while DFT calculations predict a negative value. These disagreements were systematically observed over the last 30 years. Up to now, no molecular simulation of the explicit liquid-vapor interface was able to recover the DFT trends.¹⁰ Indeed, Binder *et al.*¹¹ used a thermodynamic analysis of two-phase coexistence in finite boxes at fixed total density to extract the surface tension. Although the DFT method allows a direct and unambiguous calculation of γ , its application to realistic systems with complex interactions is more difficult. Conversely, molecular simulation is capable to fully take into account the complexity in the interactions, but the calculation of surface tension from molecular simulation of the ELVI remains ambiguous.

In this work, we present a general thermodynamic route to calculate the surface tension of spherical drops and Tol-

man length from molecular simulations of the ELVI. Our route is based on the location of the surface of tension (which is the radius at which the tension acts³) and its use in the Laplace equation. Indeed, if the radius of surface of tension is known, Gibbs,¹² Hill¹³, and Rowlinson and Widom³ have clearly established that the calculation of surface tension of spherical interfaces is possible through the Laplace relation $\gamma_s = \frac{R_s(\Delta P)}{2}$, where R_s is the radius of surface of tension, γ_s the surface tension at R_s , and ΔP the pressure difference between the vapor and liquid phases separated by the surface of tension. Up to now, the surface of tension was determined from the Laplace equation once the surface tension was calculated. In this work we follow the thermodynamics development of Hill¹³ and present a new approach to calculate γ_s based on the extraction of R_s from molecular simulations.

II. THERMODYNAMICS AND STATISTICAL MECHANICS BACKGROUNDS

As shown in Figure 1 the system is modeled by a sphere of liquid (α) surrounded by a vapor phase (β).

The spherical drop is centered in a cubic box of length L . The two phases are deemed to be separated by a dividing surface of radius R . Furthermore, the density profile is assumed to reach uniform values at some distance prior to R_α and R_β . The normal pressures acting on the surface at R_α and R_β are constant through the liquid and vapor phases, respectively. Using this model, and following, for example, Rowlinson and Widom,³ the total differential of the Helmholtz free energy can be expressed as

$$dF = -p_\alpha dV_\alpha - p_\beta dV_\beta + \gamma dA + C dR - S dT + \sum_i \mu_i dN_i, \quad (1)$$

where p_α and p_β are the pressures of the α and β phases, A is the area located at R , γ is the surface tension at R , C is the coefficient of the curvature terms, μ_i is the chemical potential of particle i , S is the entropy, and T is the temperature. From

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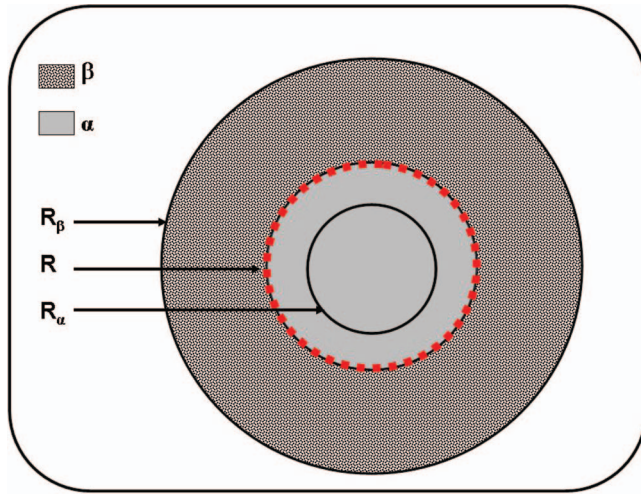


FIG. 1. Schematic picture of the system under consideration, with some key variables.

the model defined in Figure 1, the following relations can be expressed as $V_\alpha = \frac{4\pi(R^3 - R_\alpha^3)}{3}$, $V_\beta = \frac{4\pi(R_\beta^3 - R^3)}{3}$, and $A = 4\pi R^2$. Thus by differentiating V_α , V_β , and A with respect to R_α , R_β , Eq. (1) can be written as

$$dF = +R_\alpha^2 4\pi p_\alpha dR_\alpha - R_\beta^2 4\pi p_\beta dR_\beta - SdT + \sum_i \mu_i dN_i + [C + 8\pi R\gamma + 4\pi p_\beta R^2 - 4\pi p_\alpha R^2]dR. \quad (2)$$

From Eq. (2) we define the following unambiguous partial derivative:

$$\left[\frac{\partial F}{\partial R} \right]_{R_\alpha, R_\beta, N, T} = C + 4\pi R[2\gamma + (p_\beta - p_\alpha)R]. \quad (3)$$

The thermodynamic equilibrium can be deduced from the minimality criterion such that

$$\left[\frac{\partial F}{\partial R} \right]_{R_\alpha, R_\beta, N, T} \Big|_{R=R_s} = 0, \quad (4)$$

where R_s corresponds to the minimum of the free energy. One of the solutions of Eq. (3) and Eq. (4) is

$$C(R_s) = 0 \text{ and } \gamma_s = \frac{R_s(p_\alpha - p_\beta)}{2}. \quad (5)$$

Thus, we recover the well-known dividing surface condition expressed by Hill *et al.*¹³ based on the so-called surface of tension of Gibbs¹² and defined as $C(R_s) = 0$, where R_s is the position of the surface of tension. At R_s the surface of tension is noted γ_s and the Laplace equation ($\gamma_s = \frac{R_s(p_\alpha - p_\beta)}{2}$) must be recovered.³ This development provides a thermodynamic evidence of Laplace equation and of the dividing surface condition expressed by Hill and Gibbs. Thus, the problem is reduced to the determination of R_s . We define to this end one possible transformation, compatible with the canonical ensemble, to evaluate the derivative given in Eq. (3). This transformation is indexed by a test radius of tension (denoted by R_t). We then evaluate all the corresponding free energy derivatives and look for the value R_s where the derivative vanishes. We approximate the derivative by a finite difference based on

the Free Energy Perturbation (FEP) method,^{14–16} where two states are defined: (i) a reference state (indexed by 0) where the position of the surface of tension is R_{t0} ; and (ii) a perturbed state (indexed by 1) where the perturbed radius of tension is noted R_{t1} . For a small parameter ε , we consider the following volume preserving transformation, which leaves the boundaries of the system invariant (R_α , R_β): a particle at Cartesian coordinates $\{x_i, y_i, z_i\}$ at distance $R_t = \sqrt{x_i^2 + y_i^2 + z_i^2}$ from the origin is displaced in the radial direction as

$$\begin{cases} R_{t1} = R_{t0}(1 + \varepsilon(R_\alpha - R_t)) = R_{t0}\Delta\alpha & \text{if } R_\alpha \leq R_{t0} < R_t, \\ R_{t1} = R_{t0}(1 + \varepsilon(R_\beta - R_t)) = R_{t0}\Delta\beta & \text{if } R_t \leq R_{t0} \leq R_\beta, \end{cases} \quad (6)$$

with $(1 + \varepsilon(R_\alpha - R_t)) = \Delta\alpha$ and $(1 + \varepsilon(R_\beta - R_t)) = \Delta\beta$. Note that one subsystem is expanded while the other one is contracted. Indeed, for $R_\alpha \leq R_{t0} < R_t$, $R_\alpha - R_{t0} \leq 0$ while for $R_t \leq R_{t0} \leq R_\beta$, $R_\beta - R_{t0} \geq 0$. The absolute free energy reads

$$F = -k_B T \ln Q_{NVT}, \quad Q_{NVT} = \frac{V^N}{\Lambda^{3N} N!} \int e^{-\beta U} d\tau^N, \quad (7)$$

where $U(\mathbf{r}^N)$ is the configurational energy of the system and $d\tau = dr_x dr_y dr_z$, the infinitesimal volume. r_x , r_y , and r_z are the positions of particles. Our transformation divides the system in three parts: (1) the space of phase which is invariant by perturbation (Q_I is the associated partition function), (2) the contracted space of phase (Q_{II} is the associated partition function), and (3) the expanded phase (Q_{III}). The total partition function in the reference state is then expressed as $Q_{NVT}^0 = Q_I^0 Q_{II}^0 Q_{III}^0$. The free energy derivative is then approximated as

$$\left[\frac{\partial F}{\partial R} \right]_{R=R_t} \simeq \frac{\Delta F(R_t)}{R_t \varepsilon} = -\frac{k_B T}{R_t \varepsilon} \ln \frac{Q_{NVT}^1}{Q_{NVT}^0} \quad (8)$$

$$= -\frac{k_B T}{R_t \varepsilon} \ln \frac{Q_I^1 Q_{II}^1 Q_{III}^1}{Q_I^0 Q_{II}^0 Q_{III}^0}. \quad (9)$$

The ratio of partition function can be expressed as function of $d\tau_I$, $d\tau_{II}$, and $d\tau_{III}$

$$\frac{Q_I^1 Q_{II}^1 Q_{III}^1}{Q_I^0 Q_{II}^0 Q_{III}^0} \propto \frac{\int [d\tau_I^1]^{N_I} [d\tau_{II}^1]^{N_{II}} [d\tau_{III}^1]^{N_{III}}}{\int [d\tau^0]^N}, \quad (10)$$

with $[d\tau^0]^N = [d\tau_I^0]^{N_I} [\tau_{II}^0]^{N_{II}} [d\tau_{III}^0]^{N_{III}}$. N_I , N_{II} , and N_{III} are the number of particles of regions (I), (II), and (III), respectively, such that $N = N_I + N_{II} + N_{III}$. Given the expression in Eq. (6) the Cartesian components of position can be written as $r_\eta^1 = r_\eta^0 \Delta_\xi$ with $\eta = x, y, z$ and $\xi = \alpha, \beta$. Thus, by differentiating we obtain $dr_\eta^1 = dr_\eta^0 \Delta_\xi$, i.e., $d\tau_I^1 = d\tau_I^0$, $d\tau_{II}^1 = d\tau_{II}^0 \Delta_\alpha^3$, $d\tau_{III}^1 = d\tau_{III}^0 \Delta_\beta^3$ and Eq. (10) can be expressed as

$$\frac{Q_I^1 Q_{II}^1 Q_{III}^1}{Q_I^0 Q_{II}^0 Q_{III}^0} \propto \frac{\int [d\tau^0]^N \Delta_\alpha^{3N_{II}} \Delta_\beta^{3N_{III}}}{\int [d\tau^0]^N}, \quad (11)$$

TABLE I. Physical properties of nanodroplets. N is the number of particles, R_e the Gibbs radius, R_s the radius of surface of tension, and Δp the pressure difference between liquid and vapor phases. The number in brackets denotes the uncertainty in the last digit.

R_e	R_s	N	L	p_l	p_g	Δp	$\delta(R) = R_e - R_s$	γ_s
3.96	4.52(2)	484	20	0.20(1)	0.0210(2)	0.178	-0.54	0.4023(6)
5.94	6.42(2)	960	20	0.15(2)	0.0183(1)	0.133	-0.46	0.4269(4)
6.60	7.11(2)	1662	30	0.13(1)	0.0178(1)	0.118	-0.52	0.4195(5)
10.02	10.16(4)	3658	30	0.10(2)	0.0166(2)	0.081	-0.08	0.4115(9)
16.22	16.32(3)	15 660	50	0.071(8)	0.0156(3)	0.049	-0.12	0.4036(7)

which involves

$$\begin{aligned} \left. \frac{\partial F}{\partial R} \right|_{R=R_t} &\simeq \frac{\Delta F(R_t)}{R_t \varepsilon} = -\frac{k_B T}{R_t \varepsilon} \ln \frac{Q_{NVT}^1}{Q_{NVT}^0} \\ &= -\frac{k_B T}{R_t \varepsilon} \ln \left\langle \Delta_\alpha^{3N_{II}} \Delta_\beta^{3N_{III}} \right. \\ &\quad \left. \times \exp \left(- \left[\frac{U(\mathbf{r}^1(N)) - U(\mathbf{r}^0(N))}{k_B T} \right] \right) \right\rangle. \end{aligned} \quad (12)$$

In practice, we average over the estimated derivatives obtained with $\pm \varepsilon$ in order to reduce statistical errors. To summarize, R_s corresponds to the value of R_t such that $\Delta F(R_s) = 0$. The extraction of the surface tension from Laplace's equation follows once $\Delta p = p_l - p_v$ is computed. We use the non-exponential thermodynamic perturbation method to compute the local pressure.¹⁷ As shown in Ref. 18, the ambiguity on the location of pairwise interactions only influences the interfacial region while the bulk liquid and vapor phases are not impacted.

Very interestingly, our approach can also be used to extract the Tolman length (δ), which is a measure of the deviation of the surface tension with respect to the planar limit. It can be expressed as the limit in the planar case ($R \rightarrow +\infty$) of the difference between the surface of tension R_s and the Gibbs equimolar dividing surface R_e :

$$\delta = \lim_{R \rightarrow \infty} \delta(R) \approx z_e - z_s, \quad \delta(R) = R_e - R_s, \quad (13)$$

where z_e and z_s , respectively, are the positions of the Gibbs equimolar surface and surface of tension of planar interface (upon considering that the normal at the interface is along the z axis). While z_e can be obtained by a fit using an error function,¹ z_s is more difficult to extract. A straightforward modification of our previous theoretical framework for planar interfaces allows us to obtain the Tolman length from

$$\lim_{R \rightarrow \infty} \left(\frac{\partial F}{\partial R} \right)_V \Big|_{R=R_s} \approx \left(\frac{\partial F}{\partial z} \right)_V \Big|_{z=z_s} = 0. \quad (14)$$

The value z is a test value for the planar interface, and we consider the following transformation, in analogy with Eq. (6):

$$\begin{cases} z_{i1} = z_{i0}(1 + \varepsilon(z_\alpha - z_t)) & \text{if } z_\alpha \leq z_{i0} < z_t, \\ z_{i1} = z_{i0}(1 + \varepsilon(z_\beta - z_t)) & \text{if } z_t \leq z_{i0} \leq z_\beta. \end{cases} \quad (15)$$

III. COMPUTATIONAL PROCEDURE

A. Lennard-Jones model

In order to consistently compare our results to published works,⁷ we perform Monte Carlo (MC) simulations of the liquid-vapor spherical interface of Lennard-Jones (LJ) particles from a shifted truncated spherically potential (STS) at $T = 0.8$ and $r_c = 2.5$:

$$U_{\text{STS}}(r) = 4 \left(\left(\frac{1}{r} \right)^{12} - \left(\frac{1}{r} \right)^6 \right), \quad (16)$$

$$U_{\text{STS}}(r) = U_{\text{STS}}(r) - U_{\text{STS}}(r_c) \text{ if } r \leq r_c, \quad (17)$$

$$U_{\text{STS}}(r) = 0 \text{ if } r > r_c, \quad (18)$$

where r is the reduced distance, and T and r_c are the reduced temperature and the cutoff radius, respectively. The simulation cell is cubic. Box lengths (L) are given in Table I. The liquid-vapor nanodrops are carved from an equilibrated liquid-bulk configuration. Number of particles is given in Table I. Periodic boundary conditions are applied in the three directions. All the results of this paper are given in reduced units. MC simulations of liquid-vapor are carried out in NVT ensemble using 5×10^9 MC steps of equilibration and 5×10^9 MC steps of acquisition. The maximum displacement is adjusted during the equilibration phase to give an acceptance ratio of 0.4. Calculations were carried out with $\varepsilon = 1 \times 10^{-5}$.

B. Water liquid-vapor and water-octane spherical interfaces

To compare with the results obtained by Zakharov *et al.*¹⁹ from the mechanical route the TIP4P water model was considered.²⁰ The Transferable Potentials for Phase Equilibria (TraPPE) force field²¹ was used to model octane. The description of the interactions can be found in Refs. 2 and 21. Molecular dynamics simulation of the liquid-vapor water nanodroplets was performed in the canonical statistical ensemble using the Hoover's thermostat²² with periodic boundary conditions at 298 K. Spherical water-octane interface was generated by surrounding a spherical water droplet with octane. Simulation of water-octane phase was carried out in the NpT ensemble from the Nose-Hoover barostat.²³ We set up our simulation box in the x , y , z directions of 50, 50, and 50 Å,

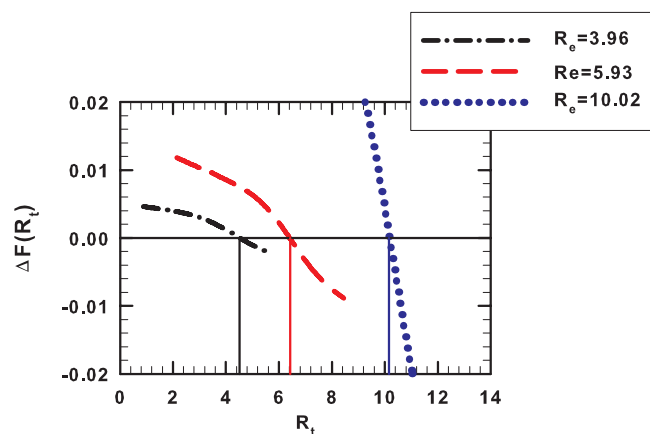


FIG. 2. Free energy difference (ΔF) as function of the trials radii R_t for various drops of Gibbs equimolar radius R_e . The radius of surface of tension R_s is the value such that $\Delta F(R_s) = 0$. The droplets are made of LJ particles. Results are given in reduced units.

respectively. A cutoff of 12 \AA was used to model the Lennard-Jones interactions. The electrostatic interactions were computed from the Ewald summation.²³ The equations of motion were integrated using the velocity Verlet algorithm²³ with a time step of 2 fs. Data analysis was performed for the last 10 ns of simulations after 5 ns of equilibration. All simulations were carried out from a modified DL-POLY package.²⁴

IV. RESULTS AND DISCUSSION

We report in Figure 2 the free energy difference as function of R_t . Independently of the drop size, an intersection between $\Delta F(R_t)$ and 0 is observed. This clearly establishes a numerical evidence of our theoretical framework.

In Figure 3 we report the surface tension of Lennard-Jones particles with respect to the drop radius by using different definitions. We observe that the fundamental measure theory (FMT) DFT method provides a surface tension higher than the planar limit (γ_∞) from $R_e = 4$ and a non-monotonic curvature dependence in γ . From the mechanical definition of the surface tension (IK definition), the curvature dependence in surface tension is monotonic. Additionally, this approach

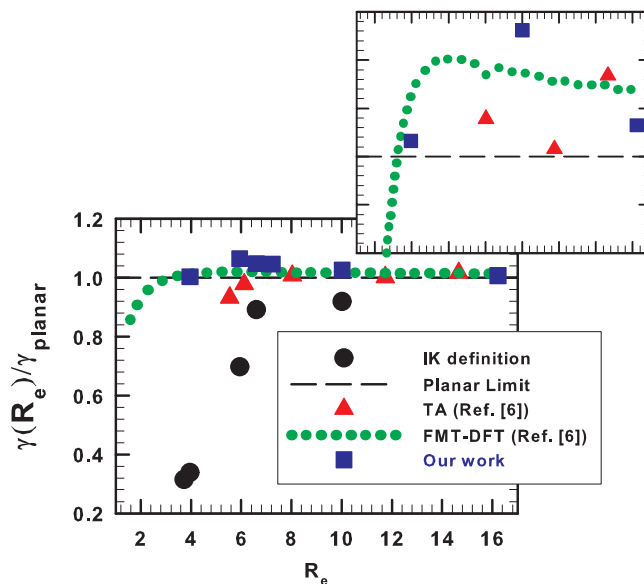


FIG. 3. Normalised surface tension of Lennard-Jones particles with truncated shifted potential ($T = 0.8$, cutoff $r_c = 2.5$) with respect to the drop radius and different definitions. γ_{planar} corresponds to the planar surface tension. TA corresponds to the test area method⁶ and FMT-DFT is a non-local density functional theory. In upper part an enlargement is given to highlight the non-monotonic behavior. Results are given in reduced units.

predicts $\gamma < \gamma_\infty$. The method of Sampayo *et al.*⁶ based on elliptic test-area transformation provides a non-monotonic behavior and $\gamma(R_e) < \gamma_\infty$ while the absolute values in surface tension are slightly underestimated with respect to the FMT-DFT values. From our approach a similar tendency to FMT-DFT is found. Indeed, γ is larger than γ_∞ and a non-monotonic curvature dependence in γ is observed. This result shows that the surface tension can be computed from a direct molecular simulation. Interestingly our results are in line with those of Block *et al.*,²⁵ van Giessen and Blokhuis,²⁶ and Blokhuis and van Giessen²⁷ which obtained surface tension from a thermodynamic analysis of two phases coexistence in finite boxes at fixed total density.

The error bar of surface tension calculation was computed by differentiating Eq. (5) and is based on uncertainty of liquid and vapor pressures and the location of surface of tension. As shown in Table I uncertainties on pressure, surface tension, and location of surface of tension are small. Error

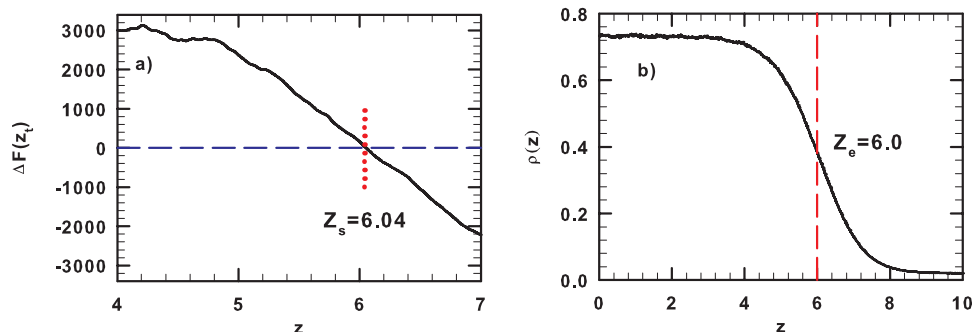


FIG. 4. (a) Free energy difference as function of z . The value z_s corresponds to $\Delta F(z_s) = 0$. (b) Density profile as a function of z for a liquid-vapor system of Lennard-Jones particle at $T = 0.8$ with a cutoff radius of 2.5. Results are given in reduced units.

TABLE II. Physical properties of water nanodroplets. N is the number of particles, R_e the Gibbs radius, R_s the radius of surface of tension, $\delta(R) = R_e - R_s$ is the Tolman length, and Δp the pressure difference between the liquid and the vapor phase. Surface tension of TIP4P water liquid-vapor planar interface is 56.2 mN m^{-1} without including the long range corrections. Pressure in liquid and vapor phases was computed from the thermodynamic route¹⁷ and validated from a comparison with the mechanical definition.¹⁸

R_e (Å)	R_s (Å)	N	Δp (Mpa)	$\delta(R)$ (Å)	γ (mN m ⁻¹)
Water liquid-vapor					
7.2	7.3	71	188.1	-0.1	68.6
8.4	10.3	137	114.3	-1.9	58.9
11.6	12.0	235	84.3	-0.4	50.6
Water-octane					
9.6	9.8	137-409	116.2	-0.2	55.7

bars on pressures of liquid and vapor phases were computed from the radial profile and are reported in Table I. This accuracy is due to a longer simulation time (5×10^9 MC steps) which allowed us to get a stable radial pressure in bulk phases and an accurate location of surface of tension.

To confirm the correctness of our approach, we compute the Tolman length of a planar interface based on Eq. (14). Details of computational procedure of planar simulations can be found in Ref. 28. We report in Figure 4(a) the free energy difference (ΔF) as function of the trials position of surface of tension. The intersection between $\Delta F(z_t)$ and 0 occurs at $z_s = 6.04$. In Figure 4(b) we provide the density profile and the position of the Gibbs equimolar surface obtained from a fit with an error function ($z_e = 6.0$). Therefore, the Tolman length is $\delta = -0.04 \pm 0.01$. This value is in fair agreement with the DFT calculation obtained by van Giessen and Blokhuis²⁶ ($\delta = -0.1$). Let us note that even with the uncertainty the Tolman length is always negative. The important result is not the absolute value of δ but rather its sign. Indeed, only Sampayo *et al.*⁶ have been capable to reproduce a negative Tolman length from molecular simulation of ELVI.

Eventually, we test our calculation on water nanodrops including the dispersion-repulsion and electrostatic interactions. We report in Table II the surface tension and the radius of surface of tension of the water liquid-vapor and water-octane droplets. The values of R_s and R_e are very similar, as expected in realistic calculations. Furthermore, we observe a decrease in the surface tension of water drops immersed in liquid octane. This result is consistent with those obtained for planar interfaces. All of these results confirm the soundness of our approach. Thus this method can be easily transferred to multicomponent systems.

V. CONCLUDING REMARKS

To summarize, we developed in this work a new thermodynamic path to compute the surface tension of spherical interfaces from the determination of the position of the surface of tension. Indeed, (i) a non monotonic behavior of surface tension is recovered; (ii) the surface tension of nanodrops is in good agreement with the DFT calculations; and (iii) the Tolman length is in line with theoretical route, in sign, and in absolute value.

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- ¹S. Senapati and M. L. Berkowitz, *Phys. Rev. Lett.* **87**, 176101 (2001).
- ²A. Ghoufi, F. Goujon, V. Lachet, and P. Malfreyt, *J. Chem. Phys.* **128**, 154716 (2008).
- ³J. S. Rowlinson and B. Widom, *Molecular Theory of Capillarity* (Clarendon Press, Oxford, 1982).
- ⁴P. Schofield and J. R. Henderson, *Proc. R. Soc. London, Ser. A* **379**, 231 (1982).
- ⁵E. M. Blokhuis and D. Bedeaux, *J. Chem. Phys.* **97**, 3576 (1992).
- ⁶J. G. Sampayo, A. Malijevský, E. A. Müller, E. de Miguel, and G. Jackson, *J. Chem. Phys.* **132**, 141101 (2010).
- ⁷A. Malijevsky and G. Jackson, *J. Phys.: Condens. Matter* **24**, 464121 (2012).
- ⁸J. G. Kirkwood and F. P. Buff, *J. Chem. Phys.* **17**, 338 (1949).
- ⁹A. Harasima, *Adv. Chem. Phys.* **1**, 203 (1958).
- ¹⁰Y.-A. Lei, T. Bykov, S. Yoo, and X. Zeng, *J. Am. Chem. Soc.* **127**, 15346 (2005).
- ¹¹A. Troster, M. Oettel, B. Block, P. Virnau, and K. Binder, *J. Chem. Phys.* **136**, 064709 (2012).
- ¹²J. Gibbs, *The Collected Works of J. Willard Gibbs* (Yale University Press, New Haven, 1948).
- ¹³T. Hill, *J. Chem. Phys.* **56**, 526 (1952).
- ¹⁴W. Zwanzig, *J. Chem. Phys.* **22**, 783 (1954).
- ¹⁵A. Ghoufi and P. Malfreyt, *Mol. Phys.* **104**, 2929 (2006).
- ¹⁶A. Ghoufi, C. Bonal, J. P. Morel, N. Morel-Desrosiers, and P. Malfreyt, *J. Phys. Chem. B* **108**, 11744 (2004).
- ¹⁷A. Ghoufi and P. Malfreyt, *J. Chem. Phys.* **136**, 024104 (2012).
- ¹⁸S. M. Thompson and K. E. Gubbins, *J. Chem. Phys.* **81**, 530 (1984).
- ¹⁹V. Zakharov, E. Brodskaya, and A. Laaksonen, *J. Chem. Phys.* **107**, 10675 (1997).
- ²⁰W. Jorgensen and J. Madura, *Mol. Phys.* **56**, 1381 (1985).
- ²¹M. Martin and J. Siepmann, *J. Phys. Chem. B* **102**, 2569 (1998).
- ²²W. Hoover, *Phys. Rev. A* **31**, 1695 (1985).
- ²³M. P. Allen and D. J. Tildesley, *Computer Simulation of Liquids* (Oxford University Press, 1987).
- ²⁴T. R. Forester and W. Smith, *DLPOLY, CCP5 Program Library*, Daresbury Lab., U.K., 1004.
- ²⁵B. Block, S. Das, M. Oettel, P. Virnau, and K. Binder, *J. Chem. Phys.* **133**, 154702 (2010).
- ²⁶A. E. van Giessen and E. M. Blokhuis, *J. Chem. Phys.* **131**, 164705 (2009).
- ²⁷E. M. Blokhuis and A. E. van Giessen, *J. Phys.: Condens. Matter* **25**, 225003 (2013).
- ²⁸A. Ghoufi and P. Malfreyt, *Mol. Simul.* **39**, 603 (2013).