



HAL
open science

Local order around rare earth ions during the devitrification of oxyfluoride glasses.

Maurício A. P. Silva, Geraldine Dantelle, Michel Mortier, André Monteil, Sidney J. L. Ribeiro, Younès Messaddeq, Valérie Briois, Marcel Poulain

► **To cite this version:**

Maurício A. P. Silva, Geraldine Dantelle, Michel Mortier, André Monteil, Sidney J. L. Ribeiro, et al.. Local order around rare earth ions during the devitrification of oxyfluoride glasses.. *Journal of Chemical Physics*, 2008, 128 (24), pp.244516. 10.1063/1.2943672 . hal-00864261

HAL Id: hal-00864261

<https://hal.science/hal-00864261>

Submitted on 23 Sep 2013

HAL is a multi-disciplinary open access archive for the deposit and dissemination of scientific research documents, whether they are published or not. The documents may come from teaching and research institutions in France or abroad, or from public or private research centers.

L'archive ouverte pluridisciplinaire **HAL**, est destinée au dépôt et à la diffusion de documents scientifiques de niveau recherche, publiés ou non, émanant des établissements d'enseignement et de recherche français ou étrangers, des laboratoires publics ou privés.

Local order around rare earth ions during the devitrification of oxyfluoride glasses

Maurício A. P. Silva,^{1,a)} Geraldine Dantelle,^{2,b)} Michel Mortier,² André Monteil,³ Sidney J. L. Ribeiro,⁴ Younès Messaddeq,⁴ Valérie Briois,⁵ and Marcel Poulain⁶

¹*Departamento de Química, Universidade Federal de Juiz de Fora, Campus Universitário S/N, Martelos, Juiz de Fora, 36036-900 Minas Gerais, Brazil*

²*Laboratoire de Chimie de la Matière Condensée de Paris, UMR 7574, ENSCP, 11 rue Curie, F-75005 Paris, France*

³*POMA, UMR 6136, Université d'Angers, 2 Bd. Lavoisier, F-49045 Angers, France*

⁴*Instituto de Química, Universidade Estadual Paulista-UNESP, C.P. 355, Araraquara, 14801-970 Sao Paulo, Brazil*

⁵*Synchrotron SOLEIL, L'Orme des Merisiers Saint-Aubin, P.O. Box 48, 91192 Gif-sur-Yvette, France*

⁶*Laboratoire des Matériaux Photoniques, Université de Rennes I, Bât. 10B, Campus de Beaulieu, 35065 Rennes, France*

(Received 1 February 2008; accepted 20 May 2008; published online 27 June 2008)

Erbium L_3 -edge extended x-ray absorption fine structure (EXAFS) measurements were performed on rare earth doped fluorosilicate and fluoroborate glasses and glass ceramics. The well known nucleating effects of erbium ions for the crystallization of cubic lead fluoride (based on x-ray diffraction measurements) and the fact that the rare earth ions are present in the crystalline phase (as indicated by Er^{3+} emission spectra) seem in contradiction with the present EXAFS analysis, which indicates a lack of medium range structural ordering around the Er^{3+} ions and suggests that the lead fluoride crystallization does not occur in the nearest neighbor distance of the rare earth ion. Molecular dynamics simulations of the devitrification process of a lead fluoride glass doped with Er^{3+} ions were performed, and results indicate that Er^{3+} ions lower the devitrification temperature of PbF_2 , in good agreement with the experimental results. The genuine role of Er^{3+} ions in the devitrification process of PbF_2 has been investigated. Although Er^{3+} ions could indeed act as seeds for crystallization, as experiments suggest, molecular dynamics simulation results corroborate the experimental EXAFS observation that the devitrification does not occur at its nearest neighbor distance. © 2008 American Institute of Physics. [DOI: 10.1063/1.2943672]

I. INTRODUCTION

The devitrification trends of erbium-doped oxyfluoride glasses containing PbF_2 have attracted the interest of scientific community due to the well known ease of control of the crystalline phase formation.^{1–5} This interest can be attributed to two factors. On one hand the erbium ions seem to ease the selective devitrification of β - PbF_2 nanoparticles in the transparent glassy matrix, decreasing the amorphous to crystalline phase transition, giving rise to transparent glass ceramics. On the other hand, the rare earth ions have been found inside these nanocrystals. The optical properties of the rare earth in the crystalline matrix, with well defined emission lines (in contrast to the inhomogeneously broadened emission spectra of these ions in an amorphous matrix), are largely favored in a heavy metal fluoride crystal due to the low phonon energy found in this media. Optical amplifiers, lasers, and other optoelectronic devices are examples of technologies that explore the potential use of these systems.

Due to the decrease of the devitrification temperature induced by the presence of rare earth ions in the glass com-

position, associated with the fact that these ions seem to be concentrated inside the crystalline phase, it has been proposed that the rare earth ion acts as a nucleating agent for the β - PbF_2 phase. Recent research⁶ has indicated that the tendency of formation of PbF_2 :Er solid solutions can be attributed as a possible explanation for the devitrification trends and it has also been observed that other halides do not act in a similar way as fluorides.⁷ Nevertheless the experimental observations gathered up to now were not capable to describe unambiguously why and how some rare earth fluorides act in the devitrification dynamics. Recently,⁸ molecular dynamics (MD) simulations have shown that the presence of erbium ions in a PbF_2 glass leads to a decrease on the devitrification temperature, in accord with experimental evidences. Moreover, an important conclusion could be drawn from these simulations: the devitrification does not start in the close vicinity of the rare earth ion, i.e., in its first or second coordination shell. Instead, the devitrification was found to begin in further distances from the rare earth. In this work we present new experimental and theoretical evidences that, in fact, the devitrification in the close vicinity of the rare earth is less complete and that considerable ordering effect is observed only at distances larger than 6 Å from the rare earth ions.

^{a)}Author to whom correspondence should be addressed. Electronic mail: mauricio.maps@gmail.com.

^{b)}Present address: Department of Materials, University of Oxford, Parks Road, Oxford, OX1 3PH, United Kingdom.

TABLE I. Composition, heat treatment, and labels used for the fluorosilicate and fluoroborate glasses and glass ceramics.

Composition (mol)	Heat treatment	Label
0.4SiO ₂ –0.3PbF ₂ –0.3CdF ₂	As quenched glass	S
	440 °C, 30 min	ST1
	440 °C, 60 min	ST2
(0.4SiO ₂ –0.3PbF ₂ –0.3CdF ₂) _{0.99} (ErF ₃) _{0.01}	As quenched glass	SEr
	440 °C, 30 min	SErT1
	440 °C, 60 min	SErT2
0.4B ₂ O ₃ –0.3PbF ₂ –0.3CdF ₂	As quenched glass	B
	470 °C, 20 min	BT1
	470 °C, 60 min	BT2
(0.4B ₂ O ₃ –0.3PbF ₂ –0.3CdF ₂) _{0.99} (ErF ₃) _{0.01}	As quenched glass	BEr
	470 °C, 20 min	BErT1
	470 °C, 60 min	BErT2

II. EXPERIMENTAL

A. Glasses and glass ceramic preparation

Pure and erbium-doped oxyfluoride glasses of molar compositions 0.4SiO₂–0.3PbF₂–0.3CdF₂ (“S” sample), (0.4SiO₂–0.3PbF₂–0.3CdF₂)_{0.99}(ErF₃)_{0.01} (“SEr” sample), 0.4B₂O₃–0.3PbF₂–0.3CdF₂ (“B” sample), and (0.4B₂O₃–0.3PbF₂–0.3CdF₂)_{0.99}(ErF₃)_{0.01} (“BEr” sample) were obtained by quenching of the melts on brass molds, as reported in detail in our previous works.^{1,2} The vitrification process was confirmed by the absence of x-ray diffraction (XRD) Bragg peaks and the occurrence of glass-transition temperature on differential scanning calorimetry analysis. Fluorosilicate and fluoroborate glass ceramics were obtained by heat treatment at 440 °C for 30 and 60 min and at 470 °C for 20 and 60 min, respectively.^{1,2} The heat-treated samples will hereafter be labeled with “T1” and “T2” according to the time of heat treatment, “T1” used for the samples treated for shorter time, and “T2” for those treated for longer time. Table I summarizes the compositions, heat treatments, and labels used in this work.

B. EXAFS

Extended x-ray absorption fine structure (EXAFS) spectra were obtained in the D44 beamline of the DCI storage ring (1.85 GeV, 250 mA) at LURE, France, at the Er *L*₃ edge. The x-ray absorption spectra were collected at ambient temperature in the fluorescence mode, using a monoelement fluorescence detector Si(Li). Spectra were obtained with an energy step of 2 eV and a counting time of 5 s. EXAFS data analysis was performed with the Michalowicz’s “EXAFS POUR LE MAC” software.⁹ After absorption background subtraction by a linear function from the rough data, EXAFS signal was calculated using the Lengeler–Eisenberger method¹⁰ and converted to a function of wave vector $k\chi k$. The data were Fourier transformed over the k range of 2.5–13 Å⁻¹, using k^3 weight and a Kaiser window with $\tau = 2.5$ leading to a spectrum scaled in distances (angstroms). Then the peaks in the Fourier transforms corresponding to a coordination shell were filtered and transformed back to

k -space. The resulting EXAFS-filtered signal was treated as a sum of sinusoidal wave functions using the single scattering approximation¹¹

$$k\chi(k) = S_0 \sum_i \frac{N_i}{R_i^2} e^{-2k^2\sigma_i^2} e^{2R_i\Gamma/k} f_{ij}(\pi, k) \sin[2kR_i + \Phi_i(k)], \quad (1)$$

where N_i is the number of atoms in the coordination shell at average interatomic distance R_i from the absorbing atom, σ_i is the Debye–Waller factor, which takes into account the static and thermal structural disordering, $\lambda(k)$ is the photoelectron mean free path ($\lambda = k/\Gamma$), and $f_{ij}(\pi, k)$ and $\Phi_i(k)$ are the amplitude and the phase functions for this coordination shell, which have been obtained experimentally from the reference ErF₃ compound. Finally, a least squares fitting was implemented on the Fourier-filtered signal obtained for the glasses and glass ceramics, leading to the determination of the structural parameters R , N , and σ corresponding to the respective coordination shells of the absorbing atoms. Typical accuracy in the determination of these parameters is 10% for N , 20% for σ , and 1% for the interatomic distances. The reliability of the fit is given by the value of the residual function defined as

$$\rho = \frac{\sum_i k_i [\chi_{\text{expt}}(k_i) - \chi_{\text{th}}(k_i)]^2}{\sum_i k \chi_{\text{expt}}(k_i)^2}. \quad (2)$$

III. MD SIMULATIONS

The classical MD simulations of the Er:PbF₂ system, performed using the DL_POLY program,¹² were based on a model where the interactions between two ions i and j were described by the Buckingham potential $U(r)$ as follows:

$$U(r) = \frac{Z_i Z_j e^2}{r_{ij}} + A_{ij} \exp\left(\frac{-r_{ij}}{\rho_{ij}}\right) - \frac{C_{ij}}{r_{ij}^6}, \quad (3)$$

where r_{ij} represents the distance between the ions i and j , A_{ij} stands for the short range repulsive coefficient, C_{ij} is an interaction parameter, and ρ_{ij} is an adjustable constant. The parameters concerning the interactions Pb–Pb, Pb–F, Er–F, and F–F were obtained from previous studies.^{13–16} For the long range electrostatic interactions, the Ewald sum technique was used.

The simulation details were described in Ref. 8. Neutrally charged face-centered cubic Er:PbF₂ systems of about 2590 particles, containing one and four erbium atoms (called PFE1 and PFE4, respectively) and some fluorine or lead vacancies, were melted at 3000 K and thermalized during a 50 ps run. The cells were then cooled down to 300 K at the rate of $q = -20$ K/ps and the glasses obtained were thermalized during a 50 ps run. Devitrification was attained by heating the glass at a rate of $q = +0.2$ K/ps from 300 to 490 K. Total enthalpy, radial distribution functions, XRD curves, and diffusion coefficients were calculated, and simulation cell snapshots were monitored.⁸

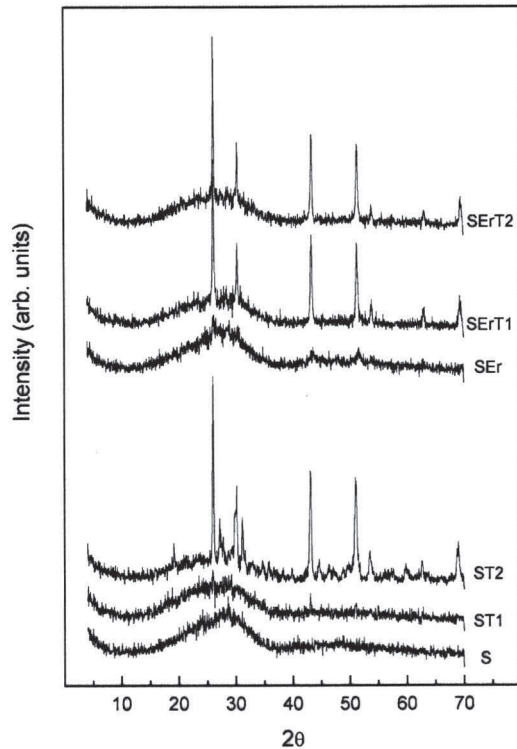


FIG. 1. XRD patterns for the fluorosilicate and fluoroborate glasses and glass ceramics (see Table I for compositions and heat treatments).

IV. RESULTS

A. Experimental

Figure 1 shows the XRD curves for the samples of the fluorosilicate systems, the results from the fluoroborates being comparable.^{1,2} Cubic lead fluoride [β -PbF₂ (Ref. 17)] is the main crystalline form obtained from heat treatment of the samples. The presence of Er³⁺ ions in the glass matrix increases the crystallization trend. XRD curves indicate that the crystalline phase is observed only under a longer heat treatment for the nondoped samples, ST2, while, in the doped sample, the presence of β -PbF₂ crystallites was detected on the SEr glass. Under the first heat treatment, SErT1, the devitrification is already well accomplished.

The presence of erbium in the crystalline phase can be proposed based on the changes of the Bragg peak positions, leading to a larger PbF₂ cubic cell parameter. Additionally the near infrared Er³⁺ emission spectra change from the $^4I_{13/2} \rightarrow ^4I_{15/2}$ broad transition, due to the statistical distribution of sites in the glasses, to well resolved lines characteristic of the glass ceramics.^{1,2}

Figure 2(a) shows the EXAFS $k\chi(k)$ spectra obtained at the Er L_3 edge of the samples. Figure 2(b) shows the corresponding Fourier transform $|F(r)|$ curves of the EXAFS signals. The peaks around 1.9 Å on the $|F(r)|$ curves, uncorrected for phase shift, correspond to the first Er coordination shell composed of fluorine anions, as will be discussed below. These $|F(r)|$ curves, which express the short range ordering, do not show for all samples any structure indicating a medium range ordering.

Quantitative simulations of the inverse-Fourier trans-

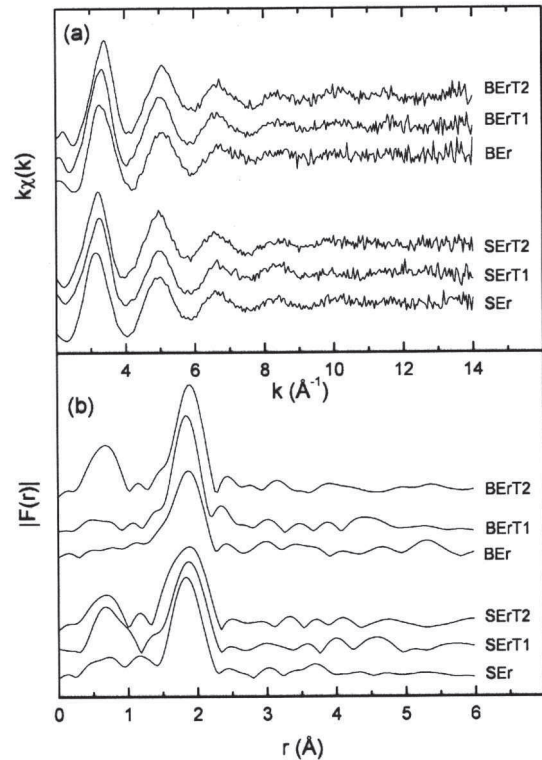


FIG. 2. (a) EXAFS $k\chi(k)$ spectra obtained at the Er L_3 edge of the samples and (b) the corresponding Fourier transform $|F(r)|$ curves of the EXAFS signals.

formed mean peak in $|F(r)|$, filtered between 1.40 and 2.35 Å for the fluorosilicate samples and 1.47–2.27 Å for the fluoroborate samples, are reproduced in Fig. 3. For example, simulations for the SEr and BEr samples are shown in this figure, and the structural parameters obtained from the simulations of all samples are listed in Table II. In the glasses and glass ceramics studied, the first Er coordination shell does not change considerably from that of the reference compound ErF₃, where the erbium atoms are surrounded by eight fluorine neighbors at a distance of 2.29 Å.¹⁸

B. MD simulations

The calculation of the total enthalpy, radial distribution functions, XRD curves, diffusion coefficients, and simulation cell snapshots of the systems PFE1 and PFE4 showed that the crystallization temperature of PbF₂ decreases with the increase of Er³⁺ ions.⁸ Moreover, in the simulation model, the devitrification of the PFE4 system occurs in a temperature range between 380 and 390 K.

The simulation cell of the devitrified PFE4 system, at 390 K, is shown in Fig. 4(a). The cell was divided in eight subcells, labeled by (+++), (++-), (+-+), and so forth, depending on their position with respect to the origin of the cell, as indicated in Fig. 4(b). The diffraction intensity at each subcell of PFE4 has been calculated separately, using the Debye formula,⁸ and the results are reported in Fig. 4(c). Despite the fact that each subcell is at least partly crystallized, indicated by the diffraction peaks, the diffracted intensity varies from one subcell to another one. The homoge-

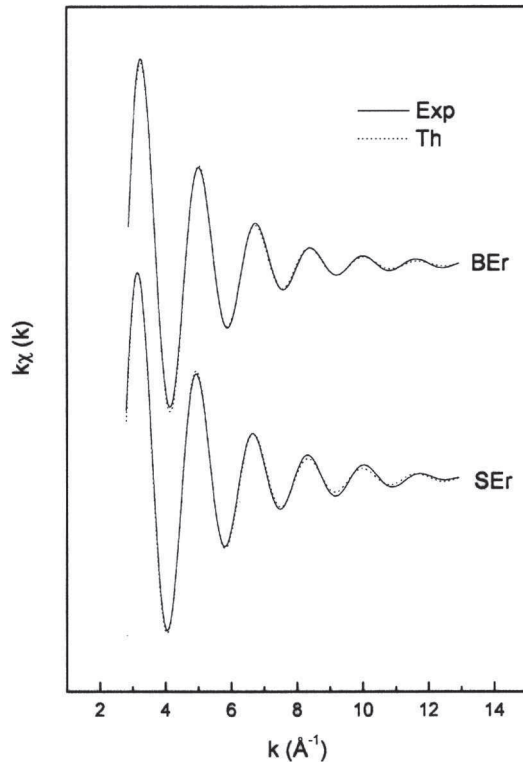


FIG. 3. Simulations of the inverse-Fourier transformed mean peak in $|F(r)|$ (first coordination shell) for the fluoroborate (BEr) and fluorosilicate (SEr) glassy samples.

neous crystallization of the system is not related to the location of Er^{3+} ions inside the subcells since the subcells containing Er^{3+} ions are those with less ordering.

Partial radial distribution functions, $g_{\alpha\beta}(r)$, were calculated for each crystallization stage, starting from the glass at 370 K. As shown in our previous work, Pb–F radial distribution functions confirmed the devitrification of the cubic $\beta\text{-PbF}_2$ structure.⁸ Figure 5 shows the Er–Pb radial distribution functions for the PFE4 glass at 370 K and for the devitrifying system at 380, 390, and 400 K. The first peak at around 4 Å indicates the first Pb neighbors surrounding the rare earth ions. It is clear from this figure that the first Er–Pb shell does not change significantly from the glass to the crystallized material. On the other hand, crystallization is evidenced at longer distances by the intensification, upon devitrification, of the peaks at around 5.8 and 7.2 Å.

From the atomic coordinates provided by the MD simulations, it is possible to simulate the respective EXAFS spectra. MD/EXAFS spectra are obtained using the FEFF6 program,¹⁹ in which the MD coordinates are used in the

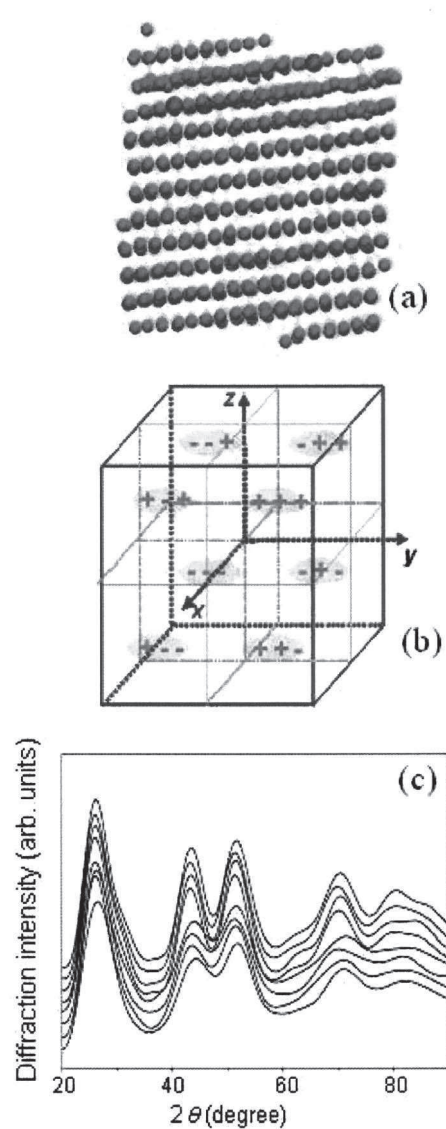


FIG. 4. (a) Snapshot (x - y axis) of PFE4 after the devitrification process at 390 K. (b) Representation of the simulation cell divided in eight subcells. (c) Diffraction intensity on each subcell at 390 K from high to low: $(---)$, $(--+)$, $(+-+)$, $(+++)$, $(-+-)$, $(-++)$, $(+--)$, and $(+++)$. The subcells $(-+-)$, $(-++)$, and $(+++)$ are the ones containing at least one Er atom.

ATOMS card of the feff.inp input file. To obtain the local environment of the absorbing atom, each of the four Er atom in the MD simulation shell was chosen as the central position $\{0,0,0\}$ and the surrounding atom positions were determined by the MD configuration with respect to the central absorbing atom. The FEFF program calculates the scattering geom-

TABLE II. Structural parameters obtained from the Er L_3 -edge EXAFS analysis.

	SEr	SErT1	SErT2	BEr	BErT1	BErT2
N	8.3	7.5	8.3	7.5	8.2	7.2
σ (Å)	0.109	0.104	0.111	0.112	0.106	0.100
R (Å)	2.25	2.26	2.26	2.25	2.27	2.27
ΔE_0 (eV)	0.7	1.1	1.1	-0.9	-0.4	0.1
ρ	1.8×10^{-3}	8.8×10^{-4}	2.4×10^{-3}	9.8×10^{-4}	5.1×10^{-3}	1.8×10^{-3}

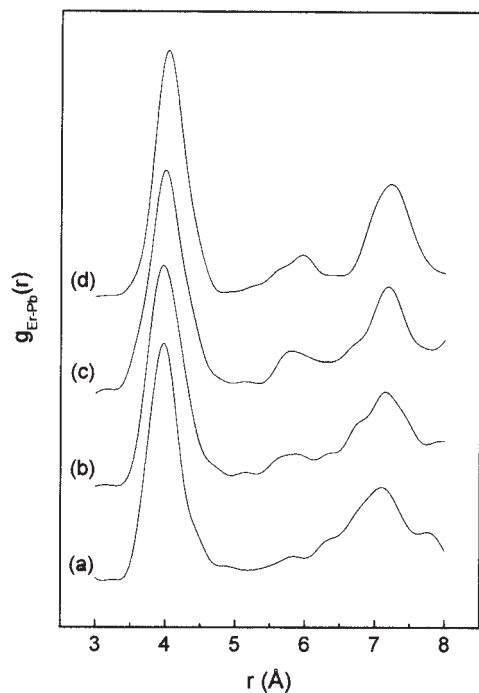


FIG. 5. Er–Pb radial distribution functions for the PFE4 glass at 370 K (a) and for the devitrifying system at 380 K (b), 390 K (c), and 400 K (d).

eries within a given radius around the absorbing atom, considering all possible single and multiple scattering paths. The total $\chi(k)$ function is then computed as a sum over all scattering paths. The $\chi(k)$ function obtained for each Er atom in the MD simulation shell, at a given temperature, was then averaged and the MD/EXAFS data, Fourier transformed using k^3 weight and a Kaiser window with $\tau=2.5$ leading to a spectrum scaled in distances (angstroms) in the same way as the experimental ones. Figures 6(a)–6(d) show the magnitude of the Fourier transformed MD/EXAFS signal obtained for the PFE4 glass at 370 K [Fig. 6(a)] and for the devitrifying system at 380 K (b), 390 K (c), and 400 K (d). As in the case of the experimental data, these $|F(r)|$ curves does not show, for all samples, any structure indicating a significant medium range ordering, such as a second coordination shell surrounding the Er atoms. In order to make clear what kind of EXAFS signal would be expected if the environment of the Er ions was crystalline, a MD simulation was performed in a fcc β -PbF₂ structure with a lead atom position occupied by an erbium atom. The 50 ps MD simulation run was performed at 300 K in the *NVT* (constant particle number, volume, and temperature) ensemble and thermalized during another 50 ps in the *NVE* (constant particle number, volume, and energy) ensemble. The Fourier transformed MD/EXAFS signal obtained for this short range crystalline Er environment is shown in Fig. 6(e). The effects of the cationic ordering in the close vicinity of the Er atom are reflected by the peak centered at 3.9 Å, i.e., the second Er coordination shell, formed by the lead atoms.

V. DISCUSSION

Erbium ions are often described as nucleating agents for the β -PbF₂ crystallization (based on XRD experiments, as

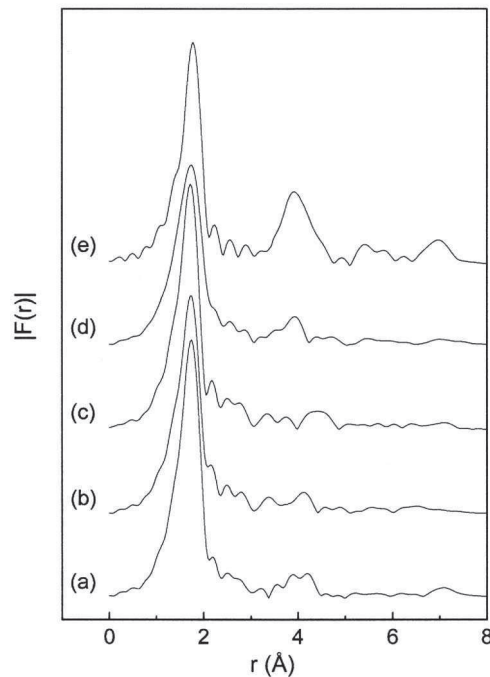


FIG. 6. The magnitude of the Fourier transformed Er L_3 -edge EXAFS signal calculated from the MD configurations obtained for the PFE4 glass, averaged over the four Er atoms, at 370 K (a), and for the devitrifying system at 380 K (b), 390 K (c), and 400 K (d). The Fourier transformed EXAFS signal in (e) is related to a single Er atom occupying a lead crystallographic position in the fcc crystalline β -PbF₂ matrix.

shown above), explaining their presence in the crystalline phase (as, for instance, indicated by Er³⁺ emission spectra experiments). The description of Er³⁺ as a nucleating agent seems to imply that the nucleation of PbF₂ starts around Er³⁺ ions. However, EXAFS analysis indicates that except from the first coordination shell of the Er atoms composed by fluorine ions, no structural ordering is observed as a second coordination shell.

In fact, if the devitrification occurred in a very close distance from the rare earth ion, i.e., the Er ions acting as a nucleating agent for the heterogeneous crystallization, retrodiffusion of Pb scatterers would have been observed on the EXAFS spectra, as a second peak on the Fourier transformed signal [see Fig. 6(e)]. Figure 2 does not show this feature, which indicates that the lead fluoride crystallization does not occur in the close vicinity of the rare earth ion.

The simulation results⁸ indicate that while the devitrification temperature T_c depends on the number of Er³⁺ incorporated in the glasses (higher Er³⁺ content, lower T_c), the devitrification seems to be less efficient around the doping ions. In fact, the evolution of the $g_{\text{ErPb}}(r)$ curves during devitrification, obtained by MD simulations and depicted in Fig. 5, shows clearly that the short range order around the doping ions ($R < 5$ Å) does not reveal any structural change upon crystallization. On the other hand, the main structural features concerning crystallization around the Er ions are located at distances larger than 6 Å, the main peak appearing at distances larger than 7 Å. Due to the finite mean free path of the photoelectron that limits the EXAFS technique to local order configurations typically shorter than 6 Å, the distances

from the rare earth ions at which the crystallization occurs do not allow the detection of this structural evolution. Our MD simulations were able to reproduce this behavior with accuracy, as shown in Fig. 5.

A close report between MD simulations and EXAFS technique could be drawn by the use of the MD configurations to simulate EXAFS $\chi(k)$ functions. Despite the fact that some ordering is indeed observed around 4 Å in the MD/EXAFS data of Figs. 6(a)–6(d), related to the lead atoms around the central Er atoms, the agreement with the experimental result, as shown in Fig. 2(b), is noteworthy.

As it is well known, the damping effects of the thermal motions on the EXAFS equation [represented, in part, by the Debye–Waller factor σ_i in Eq. (1)] limit the backscattering effects perceived by the absorbing atoms, which make the detection of the second coordination shell in the Fourier transformed EXAFS $\chi(k)$ signal difficult. However, it must be kept in mind that the MD configuration used in the MD/EXAFS calculations represents an instantaneous, static picture of the system at that temperature, in comparison to the dynamical, ambient temperature system, from which the experimental data were extracted. In other words, the MD/EXAFS spectra of Figs. 6(a)–6(e), obtained via MD configurations, are not affected by thermal vibrations. Our MD/EXAFS results show that the thermal effect is not the factor that limits the experimental observation of structural ordering around the Er atoms only to the first fluorine coordination shell. In this case, the intensity degree of the peak at around 4 Å, representing the second coordination shell in the Fourier transformed EXAFS signals [Fig. 2, 5, and 6(e)], is dependent mainly on the statistical disorder of the scatterer Pb atoms. Actually, the peak observed at around 4 Å in the Fourier transformed EXAFS signal of Fig. 6(e) corroborates our assumptions that, had the devitrification started in the close vicinity of the Er atom and the cationic crystallographic positions statistically occupied around the rare earth, the presence of a second coordination shell formed by the lead atoms would be identified by the EXAFS analysis. Indeed, this MD/EXAFS data treatment reinforces the conclusions drawn by the experimental EXAFS observations and corroborates the lack of structural ordering in the close vicinity of Er atoms computed by the MD simulations.

Several factors are yet unclear and must be evaluated to draw a precise conclusion about the genuine role of Er³⁺ ions in the devitrification process of PbF₂. One of these factors, which is yet to be tested experimentally and theoretically, is the possibility of an amorphous-amorphous phase separation induced by the rare earth ion, creating nanosized PbF₂-rich regions around the rare earth. In this case, the lack of the structural evolution during crystallization could be attributed to the fact that the devitrification itself occurs not in the immediate neighborhood of the rare earth ion but in the interface between the two amorphous phases, leaving the rare earth trapped inside the crystal, but with a considerable disorder in its close vicinity due to the direction of the crystalline growth.

VI. CONCLUSION

In this work, MD simulations were performed on the devitrification properties of Er-doped lead fluoride glass, and the results were unambiguously confirmed by Er *L*₃-edge EXAFS measurements. The fact that the crystallization of β-PbF₂ does not occur at the immediate neighborhood of the rare earth ion is in contradiction with the widely spread theory that these trivalent ions act as a nucleating agent for this crystalline phase. Experimental and theoretical results indicate that the erbium ions indeed act on the crystallization behavior of fluoride and oxyfluoride glasses, but simply inducing a decrease of the crystallization temperature that should reflect the tendency to form a PbF₂–ErF₃ solid solution. The nucleation role of these ions on the heterogeneous crystallization of the glasses was not confirmed by the EXAFS measurements and by the MD simulations presented due to the lack of cationic structural ordering around the rare earth. The use of MD configurations to simulate EXAFS spectra is well suited for examining the convergence of both theoretical and experimental techniques and was used in this work to confirm that the lack of structural short-range ordering around the Er atoms, observed by MD and EXAFS, is not an issue of thermal damping effects.

ACKNOWLEDGMENTS

The financial support of FAPEMIG, FAPESP, CNPq, and CAPES (Brazilian agencies) is gratefully acknowledged.

- ¹M. A. P. Silva, V. Briois, M. Poulain, Y. Messaddeq, and S. J. L. Ribeiro, *J. Phys. Chem. Solids* **64**, 95 (2003).
- ²M. A. P. Silva, S. J. L. Ribeiro, Y. Messaddeq, V. Briois, and M. Poulain, *J. Braz. Chem. Soc.* **13**, 200 (2002).
- ³L. A. Bueno, P. Melnikov, Y. Messaddeq, and S. J. L. Ribeiro, *J. Non-Cryst. Solids* **247**, 87 (1999).
- ⁴G. Dantelle, M. Mortier, G. Patriarche, and D. Vivien, *J. Solid State Chem.* **179**, 1995 (2006).
- ⁵M. Mortier and F. Auzel, *J. Non-Cryst. Solids* **256–257**, 361 (1999).
- ⁶I. I. Buchinskaya and P. P. Fedorov, *Russ. Chem. Rev.* **73**, 371 (2004).
- ⁷M. Mortier, C. Chateau, M. Génotelle, and N. Gardant, *J. Non-Cryst. Solids* **326–327**, 287 (2003).
- ⁸G. Dantelle, M. Mortier, A. Monteil, S. Chaussedent, and M. A. P. Silva, *J. Chem. Phys.* **127**, 094509 (2007).
- ⁹A. Michalowicz, *EXAFS pour le Mac* (Société Française de Chimie, Paris, 1991), p. 102.
- ¹⁰B. Lengeler and P. Eisenberger, *Phys. Rev. B* **21**, 4507 (1980).
- ¹¹B. K. Teo, *EXAFS: Basic Principles and Data Analysis*, Inorganic Chemistry Concepts, Vol. 9 (Springer, Berlin, 1986).
- ¹²W. Smith and T. R. Forester, *DL_POLY_2 User Manual*, (The Council for the Central Laboratory of the Research Councils, Daresbury Laboratory at Daresbury, Nr. Warrington, 2003).
- ¹³A. B. Walker, M. Dixon, and M. J. Gillan, *J. Phys. C* **15**, 4061 (1982).
- ¹⁴M. E. G. Valerio, R. A. Jackson, and J. F. Lima, *J. Phys.: Condens. Matter* **12**, 7727 (2000).
- ¹⁵M. A. P. Silva, A. Monteil, Y. Messaddeq, and S. J. L. Ribeiro, *J. Chem. Phys.* **117**, 5366 (2002).
- ¹⁶M. A. P. Silva, J. P. Rino, A. Monteil, S. J. L. Ribeiro, and Y. Messaddeq, *J. Chem. Phys.* **121**, 7413 (2004).
- ¹⁷JCPDS Card No. 6-0251.
- ¹⁸R. W. Wyckoff, *Crystal Structures*, 2nd ed. (Krieger, Malabar, Florida, 1982), Vol. 1, p. 241.
- ¹⁹S. I. Zabinsky, J. J. Rehr, A. Ankudinov, R. C. Albers, and M. J. Eller, *Phys. Rev. B* **52**, 2995 (1995).