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**Influence of mesostructuration on the reactivity of
bioactive glasses in biological medium:
a PIXE-RBS study.**

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Abstract

Building mesostructured biomaterials is a challenging and exciting task that has attracted much attention because of their use as drug carriers or drug delivery systems. In the case of bioactive materials, the mesostructuration can also deeply impact their physico-chemical properties and the reactivity. In this study, we show how highly ordered mesoporosity influence the early steps of biomineralization process and the reactivity in binary ($\text{SiO}_2\text{-CaO}$) and ternary ($\text{SiO}_2\text{-CaO-P}_2\text{O}_5$) bioactive glasses. Conventional porous sol-gel glasses were synthesized using a classical route, while mesostructured glasses were elaborated using a non-ionic surfactant. Textural properties of these materials have been characterized. The in vitro biomineralization process was followed, using Particle Induced X-ray Emission (PIXE) associated to Rutherford Backscattering, Spectrometry (RBS), which are efficient methods for a highly sensitive multi-elemental analysis. Elemental maps of silicon, calcium and phosphorus were obtained at a micrometer scale and revealed for the first time a bulk reactivity for mesostructured glasses. This is a major advantage over conventional glasses, for which the first steps of biomineralization are limited to the periphery of the material. Their enhanced bioactivity combined with their possible use as drug-delivery systems make them promising candidates for bone regeneration.

1. Introduction

Highly ordered mesoporous materials have been widely studied as drug delivery systems during the past decade (antitumorals¹, antibiotics², anti-inflammatory³). In the case of bone reconstruction, implantation of such systems has several advantages over conventional materials. Indeed, the main interest of these materials is the control of kinetics and the spatiality of molecules release, which implies a higher efficiency of delivery, a reduced toxicity, and a continuous action. Several kinds of materials like mesoporous carbons, bioceramics or polymers have been studied, but the most described drug delivery system is silica.

For bone regeneration, an attractive solution is the synthesis of mesostructured bioactive glasses via sol-gel routes. When implanted, these materials are able to bond to living bone through the formation of an interfacial apatite-like layer at the glass surface, which composition is close to the mineral phase of bone⁴. This newly-formed biomimetic layer improves the osteo-integration properties of the implant, while the degradation products from the glass promote the bone tissue regeneration. Synthesis of binary ($\text{SiO}_2\text{-CaO}$), ternary ($\text{SiO}_2\text{-CaO-P}_2\text{O}_5$) or even doped mesostructured bioactive glasses has been described in some publications^{5,6,7}. Their *in vitro* drug release behaviors have also been evaluated⁸. Moreover several research groups investigated the influence of the mesostructuration on physico-chemical reactions (dissolution, diffusion, ionic exchange and precipitation)^{9,10,11}. These studies demonstrated that the increase of the specific surface area (higher silanol concentration¹²) and pore volume of bioactive glasses has an effect on the kinetics of reaction and on the structural evolution of the calcium phosphate phases formed at the glass interface. The present study aspires to add spatial information about the physico-

chemical reactions occurring when bioactive glasses are in contact with biological mediums. Differences between conventional porous sol-gel glasses and mesostructured glasses will be explained using Particle Induced X-ray Emission coupled to Rutherford Backscattering Spectrometry (PIXE-RBS). These nuclear techniques allow monitoring the spatial distribution and diffusion of elements within bioactive glasses with an excellent (a few ppm) sensitivity, through the recording of quantitative chemical maps of glass grains after interaction with biological fluids.

2. Materials and method

2.1 Synthesis

Binary ($\text{SiO}_2\text{-CaO}$) and ternary ($\text{SiO}_2\text{-CaO-P}_2\text{O}_5$) mesostructured glasses have been synthesized through a sol-gel route, by using commercially available nonionic surfactant Pluronic F127 (BASF) as structure-directing agent. These materials are labeled B75-F127 and B67.5-F127 depending on their chemical composition (Table 1). Tetraethylorthosilicate ($\text{Si}(\text{OC}_2\text{H}_5)_4$) TEOS, Triethylphosphate ($\text{PO}(\text{OC}_2\text{H}_5)_3$) TEP, and Calcium nitrate tetrahydrate ($\text{Ca}(\text{NO}_3)_2 \cdot 4\text{H}_2\text{O}$) CaNT (Aldrich) were used as SiO_2 , CaO and P_2O_5 sources. The amounts of reactants are indicated in Table 1.

In a typical synthesis of mesostructured glasses (MG), F127 is dissolved in ethanol (18 ml). In the same time, stock solution is prepared by mixing TEOS, TEP, CaNT, H_2O , HCl (2M) and EtOH. After stirring them separately for 1 h, both solutions are mixed and vigorously stirred together for another 4 h. The resulting sol is introduced into a Petri dish to undergo an

Evaporation-Induced Self-Assembly (EISA) process¹³ during 48h. The dried gel is then calcined at 600°C for 12h to eliminate the surfactant and to obtain the final glass powder. The complete elimination of the surfactant was checked with FTIR spectroscopy. For comparison, conventional porous glasses (CG), named B75 and B67.5 and with the same compositions (Table 1) have also been synthesized by a similar sol-gel process but without surfactants.

2.2 Materials characterization

Nitrogen gas sorption analyses were performed to characterize the glasses textural properties. The samples were vacuum outgassed at 120°C for 12 hours to remove physically adsorbed molecules from the pores. The adsorption/desorption isotherms were recorded on a Quantachrom Autosorb-1 MP apparatus. The instrument determined isotherms volumetrically by a discontinuous static method at 77K. The surface areas were obtained by applying the BET method to the N₂ isotherm using 7 points in the 0.05-0.35 p/p₀ range. The pore size distribution was determined by applying the BJH method on the desorption branch. Total pore volume was measured at a relative pressure P/P₀=0.995. The average pore diameter was calculated using surface area and total pore volume and considering cylindrical pores.

Powder X-ray diffraction experiments were performed with an X'Pert Pro diffractometer equipped with Cu K α radiation (wavelength 1.5406 Å). XRD patterns were collected in the 2 Θ range between 0,6° and 3,1° with a step size of 0.02° and counting time of 5 s per step.

Transmission electron microscopy (TEM) was carried out with a Hitachi H-7650 microscope, operating at 120 kV (resolution 2Å). Images were recorded using a CCD camera (Hamamatsu HR, 1024 x 1024 pixels)

2.3 In vitro studies and samples preparation

For studying the glass reactivity in biological conditions, glass powders were immersed at 37°C for 15 minutes, 30 minutes, 1, 6, 12 hours and 1, 2 days in 40 mL of a standard Dulbecco's Modified Eagle Medium (DMEM, Biochrom AG, Germany) for which concentrations in Inorganic salts is close to those of human plasma¹⁴. Mandel et al.¹⁵ demonstrated that DMEM can be regarded as a feasible alternative to using Simulated Body Fluid (SBF) solutions in for *in vitro* bioactivity testing of synthetic biomaterials. Unlike SBF, DMEM contains amino acids, vitamins and glucose. Because of amino acids, lower rates for the materials dissolution and a subsequent delay in surface layer formation are observed in DMEM when compared to soaking in SBF¹⁶. Indeed amino acids from DMEM are charged species that can be attracted by the negative glass surface and coat it with a film. In terms of simulating the *in vivo* environment, DMEM can be a better choice as it also contains other components present in *in vivo* systems besides inorganic salts¹⁷. For the same reasons, DMEM is commonly used in cell culture^{18,19,20}, contrary to SBF (because of its deficiency of nutrients²¹). Then we decided to use DMEM in the present study in view of confronting with further biological results.

To simulate homeostasis of a real biological system, interactions in dynamic conditions are better adapted. Indeed, it has been demonstrated²², that results of dynamic and static procedures differ for several parameters (constant concentration of interesting elements,

constant pH...), and that dynamic is closer to in vivo conditions. Despite of these advantages, static procedure has been classically chosen in this work, because of its easier implementation. Moreover, its use allows also direct comparison of the different mesostructured glasses with samples described previously in the literature for which static conditions were mainly used.

For each sample, the powder weight to DMEM volume ratio was fixed to allow studying the influence of the textural effects only, on physico-chemical properties. After interaction, part of the DMEM was sampled to analyze its chemical composition by Inductively Coupled Plasma-Atomic Emission Spectroscopy (ICP-AES), while the glass particles were removed from the solution and air dried. Before characterization with PIXE-RBS nuclear microprobe, glass particles were embedded in resin (AGAR,Essex, England). 1000 nm thin sections of these samples were prepared by mean of a Leica EM UC6 Ultramicrotome, and inserted in 50 mesh copper folding grids, which were placed on a Mylar film with a hole of 3 mm in the centre. Measurements were performed on the area of the section placed over the hole.

2.4 PIXE-RBS analysis

PIXE and RBS methods are used simultaneously. The PIXE method permits the identification and the quantification of elements in sections of biomaterials grains after interaction with biological medium²³. RBS is used to determine the electric charge received by the samples during irradiation, which is necessary for PIXE spectra quantification. Analyses of our materials were carried out using nuclear microprobes at the CENBG (Centre d'Études Nucléaires de Bordeaux-Gradignan, France). The experimental characteristics of the CENBG microbeam line have been published previously^{24,25}. For PIXE-RBS analyses, we chose a proton scanning microbeam of 1.5 MeV energy and 50 pA in intensity. The beam size was

nearly 1 μm . Such parameters resulted in higher ionization cross sections for light elements ($Z < 20$) and thus in a better sensitivity for PIXE analysis by using a detector without filter. Furthermore, weak intensities and the choice of protons as the ion beam allowed the target degradation to be minimized during irradiation. However, the intensities were sufficient to permit measurement duration below 1 h. An 80 mm^2 Si(Li) detector was used for X-ray detection, orientated at 135° with respect to the incident beam axis, and equipped with a 12 μm thick beryllium window. PIXE spectra were treated with the software package GUPIXWin. For RBS measurements, a silicon particle detector placed at 135° from the incident beam axis provided us with the number of protons that interacted with the sample. Data were treated with the simNRA code.

3. Results

3.1 Textural properties

Figure 1a shows the N_2 sorption isotherms for the mesostructured glasses (MG) and conventional sol-gel glasses (CG). Contrary to CG, MG curves can be identified as type IV isotherms characteristic of porous materials. B75-F127 and B67.5-F127 show type H1 hysteresis loops in the mesopore range, which are characteristic of cylindrical pores open at both ends. The pore size distributions are shown in Figure 1b. MG present a single-modal pore size distribution centered around 4 nm, whereas CG pore size distributions are not very well defined and exhibit much bigger pore size. Table 2 collects the specific surface area, mesopore volume, and pore size measured for the four samples. Specific surface areas are significantly higher for the mesostructured glasses than those obtained for conventional sol-gel glasses of analogous compositions (from x4 up to x15). This increase is essentially due to the specific surface area associated to the mesoporosity. In the same way, pore volume is also higher for MG compared to that CG. The modal pore diameter and average

pore diameter are close for the two mesostructured materials, which confirm the hypothesis of cylindrical pores and consequently the successful templating.

Figure 2 presents the images of Transmission Electron Microscopy (TEM) of B75-F127 and B-67.5-F127 glasses. After calcination, organized networks of mesopores are present for both materials. The observed hexagonal structure is confirmed in other articles^{26,27} for similar syntheses in the same domains of the ternary diagram H₂O/EtOH/F127. The results of X-ray diffraction at small angles abound in this direction, as the identified peaks (h=1, k=0) and (h=1, k=1) (Fig.3) are consistent with a P6mm hexagonal symmetry of cylindrical pores. These peaks, whose positions are given in Table 3, permitted to calculate the lattice parameters using Bragg's law. These parameters are close for both glasses. Distances estimated using average measurements on TEM images confirm a similar order of magnitude.

3.2 Elemental mapping

Several elemental maps for each time of interaction with DMEM were recorded. In this paper, we only present a restricted selection of these maps. The observed distributions correspond to the intensity of X-rays locally emitted by the sample under proton irradiation. Figure 4 shows evolutions of chemical mapping for binary glasses. Before immersion in DMEM (0 day), elemental distributions of silicon and calcium are homogenous in the grains of both B75 and B75-F127. After 1h of immersion distributions evolve in two different ways. For B75 glass, a layer is formed at the periphery. This layer is rich in calcium and phosphorus. In the same time, a loss of calcium is observed in the inner region. For longer interaction time (2 days) the phosphocalcic layer disappears and only silica remains as discussed in previous work²⁸. For B75-F127 glass, distributions and concentrations of calcium seem stable comparing to the material before interaction. Phosphorus from the biological medium is integrated in the volume of the grain and not only at the periphery as for B75. After 2 days

of interaction, a persistence of silica matrix is observed for the two glasses. However, unlike the B75 glass, calcium and phosphorus are still present within the mesostructured material intimately mixed with silicon oxide.

Ternary glasses (Figure 5) maintain a homogeneous elemental distribution from 0 day to 1 hour for CG and MG. However behaviors of both materials differ thereafter. Indeed after 2 days, the core/layer spatial discrimination observed for B67.5 is not applicable for B67.5-F127. The mesostructured glass is then mainly composed of calcium and phosphorus and the silicon concentration seems very low.

For both materials, the main trend highlighted by the chemical mapping is the spatial discrimination of elemental distribution observed for conventional sol-gel glasses (formation of a phosphocalcic layer at the periphery) and not for mesostructured glasses.

3.3 Evolution of the concentrations in the grains

The PIXE-chemical maps were divided into various regions of interest using the SUPAVISIO analysis software. Whenever the Ca-P-rich layers were detected, areas of measurement were created at the periphery and in the inner region of grains, focusing on the X-ray spectra of these user-defined regions of interest. Even if layers are not observed for mesostructured glasses, the same methodology has been used. Then elemental concentrations were calculated in these areas and thanks to the elemental concentrations measured for each delay, the temporal evolution of concentrations for each region can be monitored. The results correspond to the average concentrations calculated in several identical regions of interest. These regions of interest were defined over various samples in order to be ensured of measurements reproducibility. Each point represents an average of 5 measurements.

Errors on elemental concentrations are associated to the statistical uncertainty due to the differences of concentrations for all the measurement used to obtain the average. The uncertainty associated to each point is then calculated thanks to an empirical standard deviation. These errors, which are not represented on the graphs for a better reading, are below 5% for Si, 10% for Ca and 7% for P. Taking into account the uncertainties, general and reliable trends can still be observed.

Evolutions of concentrations at the periphery and in the inner region of mesostructured grains are the same. Consequently B75-F127 and B67.5-F127 curves represent global evolution of the grains.

Binary glasses

The overall concentrations evolution of mesostructured glass is compared to peripheral (Figure 6) and internal (Figure 7) evolutions of conventional sol-gel glass. The evolution of the interface of B75 glass (Figure 6) is in agreement with mechanisms described for this kind of glass²⁹. First there is the formation of a Ca-P layer between 1 hour and 6 hours. Indeed, during this period a significant increase in concentrations of calcium and phosphorus is observed. This phenomenon is logically associated with a decrease of silicon. Subsequently, this layer dissolves from 6 hours to 2 days (loss of calcium and phosphorus and predominance of silicon). The core of the B75 glass (Figure 7) undergoes a release of Ca^{2+} and despite an incorporation of phosphorus (between 0 and 1 day), its main component is silicon.

In the case of the mesostructured glass, the decrease of calcium before 1 hour (opposite effect of silicon) is due to the dealcalinization of the material. The evolution of this process

over the time is close to the one at the center of the grain of glass B75 (Ca and P evolutions, Figure 7). Moreover, while calcium phosphate is only formed on the surface for the B75 glass, it is present within the total volume of B75-F127 grain, as demonstrated by consistent increases in Ca and P (Figure 7). This calcium phosphate also appears more stable or still in formation, because it does not dissolve, contrary to the phosphate layer of conventional sol-gel glass (low Ca and P between 6 hours and 2 days). The last remark concerns the important decrease of the silicon concentration for the mesostructured material, whereas that of B75core remains relatively stable. This phenomenon does not automatically mean increased dissolution of silica network, but may reflect a decrease in its mass contribution to the benefit of the phosphate phase.

Ternary glasses

As for binary glasses, the global evolution of the mesostructured material is compared with those of the conventional sol-gel material. Curves corresponding to the elemental evolution in the total volume of B67.5-F127, and to the evolution at the periphery of B67.5 glass are presented in Figure 8. The first important point is the similarity of the two curves for each of the three elements. In both cases the decrease of calcium concentration (and respectively the increase of silicon concentration), reflecting the network dealcalinization takes place between 0 and 1 hour. Subsequently, the concomitant increases in calcium and phosphorus can be interpreted by the formation of calcium phosphate precipitates within the grain for the mesostructured glass and at the periphery for the conventional sol-gel glass. Given these results, it is clear that mechanisms occurring in the two cases are similar (kinetics, intensity).

Consequently the mesostructured glass allows not only a surface, but a three-dimensional formation of a phosphocalcic phase.

3.5 Elemental evolution of the biological medium.

Binary glasses

The evolutions of silicon, calcium and phosphorus concentrations in the biological medium (Figure 9) confirm some trends outlined by local measurements in grains of B75 and B75-F127 glasses. First, variations of calcium concentrations between 0 and 1 hour support the theory of very similar kinetics of dealkalinization for both materials. After this maximum, the calcium concentration decreases much faster for the B75 glass than for the B75-F127 glass between 1 hour and 2 days. The amplitude and the kinetics of the phosphorus decrease are also higher for the conventional sol-gel glass. The curves representing silicon, for which evolutions are almost the same for B75 and B75-F127, invalidates the hypothesis of a larger dissolution of the glassy matrix for the B75-F127 glass.

Ternary glasses

The concentration changes of DMEM during interaction with B67.5 and B67.5-F127 glasses are presented in Figure 10. Similarly to the binary glass, the same trends of silicon evolution indicate that the kinetics and intensity of the degradation of the silica matrix are very similar for both glasses. The kinetics of calcium release are also very close (maximum between 0 and 1 h for B67.5 and between 0 and 6 h for B67.5-F127), although the amount of released calcium into the biological medium is more important for the mesostructured glass. This difference may potentially be explained by reactivity within the volume for B67.5-F127.

Finally, as for the binary glass, the decrease in phosphorus concentration is much less important for the mesostructured glass.

4. Discussion

Changes in elemental concentrations are powerful indications on how mesostructured glasses react in an acellular medium. The specific surface area is considerably increased from conventional sol-gel glasses to mesostructured glasses (from x4 for ternary glasses to x15 for binary glasses), thanks to homogeneous and interconnected nanoscaled porosity. This large specific surface area allows a three-dimensional homogeneous reactivity for dealkalinization, dissolution of silicate network and biomineralization.

4.1 Compositional evolution of phosphocalcic phases

Ca-P phases formed at the beginning of the biomineralization process are partially crystallized and embedded in amorphous silica, which makes XRD unusable to follow structural changes. NMR or vibrational spectroscopies which are powerful techniques to understand structural evolutions, could be helpful but they are either global (NMR, FTIR) or surface sensitive (Raman) techniques. Although PIXE-RBS does not directly probes phases structure, it allows to add complementary informations about the three-dimensional evolution of phosphocalcic phases nature, thanks to the evolution of Ca/P atomic ratio over time .

Figure 11 presents SEM images of a B75- F127 glass before immersion in the biological medium and after 12 hours. These pictures clearly show that new phases have been formed

during the interaction. According to elemental maps obtained with PIXE, these phases are calcium phosphates. Three phosphocalcic phases are mainly formed in biological fluids such as SBF or DMEM³⁰: dicalcium phosphate and octacalcium phosphate which are metastable phases and hydroxyapatite which is the most thermodynamically stable phase and therefore the less soluble (Table 4).

The nature of the phase was monitored via the Ca / P atomic ratio measured at the glass / biological medium interface for conventional sol-gel materials and within the grain for mesostructured materials. Curves representing these evolutions as a function of interaction time are shown in Figure 12 for binary and ternary glasses.

For the binary glasses, the Ca / P ratio reaches a minimum value for both materials after 1 hour of interaction. However these minima differ on several points: the minimum is lower (2.85 versus 5.7) for the B75 glass. It is also reached in a shorter time for the B75 sample and is already close to 23 after 6 hours because of the dissolution of the layer. This quite low ratio and the dissolution of the phase indicate the formation of metastable phases mentioned above. Although they are less stable than apatite, precipitation kinetics of these phases is faster³¹. For the conventional B75 glass, the surface reactivity (limited release of calcium) only allows the formation of these phases, where the Ca / P ratio is between 1 and 1, 5 (Table 4). On the contrary, release of calcium is enhanced for the mesostructured B75-F127 glass thanks to its bulk reactivity. The precipitation of an apatitic phase (Ca / P = 1.67) is then possible. This assumption is confirmed by the increased stability of Ca-P phase for the B75-F127 glass demonstrated by chemical mapping (Figure 4) and evolutions of Ca / P atomic ratios (Figure 12). Moreover for the B75-F127 glass, the relatively high Ca/P ratio (compare to apatite) between 12 hours and 2 days, could be explained by the presence of

amorphous or crystallized phases rich in calcium but phosphorus-free³² such as Calcium oxide CaO, Portlandite Ca(OH)₂ or Calcium carbonate CaCO₃. The formation of these phases may be explained by the combination of a massive and rapid release of calcium in the biological medium with a limited source of available phosphorus, which prevent Ca-P phases formation.

The problem developed above for binary glasses does not occur for ternary glasses, because they initially contain phosphorus. Both curves (conventional et mesostructured B67,5) present a similar shape: an exponential decay to a value close to 2.3. This ratio, which is higher than 1.67 (Hydroxyapatite), can be due to calcium-rich phases as seen previously.

4.2 Surface versus volume precipitation

From the PIXE-RBS mapping obtained , it is clear that Ca-P phases appear within the grains for mesostructured glasses and only at the periphery for the conventional sol-gel glasses.

The first requisite for a bulk precipitation is a high interconnection of pores, in order to facilitate the diffusion of the chemical elements (calcium, phosphorus) involved in the biomineralization process. While mesostructured glasses provide this ideal interconnection, conventional sol-gel glasses do not. Indeed, although the sol-gel naturally generates mesopores, these latter are quite randomly distributed and poorly connected, as evidenced by differences of specific surfaces, which are more than four times higher for materials synthesized using a surfactant (Table 2).

The pore size might also influence the formation of Ca-P phases (particularly Hydroxyapatite). Deng et al.³³ investigated the effect of pore size on the growth of hydroxyapatite (HA) from mesoporous CaO–SiO₂ substrates. In this work the heterogeneous

nucleation of hydroxyapatite is theoretically studied and particularly the critical nuclei size r^* . For low values of r , the nucleus is thermodynamically unstable and dissolved. For larger values of r , the nucleus is thermodynamically stable and persists over time. Nucleation of hydroxyapatite in pores with diameter smaller than $2r^*$ is consequently not allowed because the nuclei are not thermodynamically stable. The calculated range is $0,8 \text{ nm} < 2r^* < 8 \text{ nm}$. This critical radius is consistent with pores diameters of our mesostructured glasses ($2r$ equals to 8.6 and 7.8 for B75-F127 and B67.5-F127 respectively) and could explain why hydroxyapatite and more generally calcium phosphate phases could be formed within the grains of our materials. Moreover Yan et al. **Erreur ! Signet non défini.** concluded that small pore size (1.9 nm in their study) could slow down the growth of hydroxyapatite. Indeed, when it occurs, the nucleation stage could easily block small pores and then inhibit the diffusion of Ca^{2+} and HPO_4^{2-} ions which are required to form hydroxyapatite. With these respect , our materials seem to gather all pre requisite for easy CaP formation making them valuable materials for the targetted applications.

5. Conclusions

Three chronological reactions take place during the immersion of $\text{SiO}_2\text{-CaO}$ and $\text{SiO}_2\text{-CaO-P}_2\text{O}_5$ glasses in DMEM: calcium release, dissolution of the silica network and precipitation of Ca-P rich phases. A key result of our study is the demonstration that ion beam techniques are powerful to study the effect of mesostructuration over the biomineralization process in sol-gel derived glasses. Indeed, thanks to the use of PIXE-RBS techniques the bulk reactivity of mesostructured glasses is clearly highlighted. This bulk reactivity is a major advantage over conventional glasses, for which the first steps of biomineralization are limited to the outer surface of the material. Two conditions are necessary to allow this bulk reactivity: a high interconnection of pores and pore diameters larger than the critical size for HA precipitation. These two criteria are fully satisfied by our mesostructured glasses explaining their peculiar behavior. Such demonstrated reactivity of mesostructured glass throughout the material and not only on the outer surface makes them promising materials for bone substitution and combined drug delivery applications.

Sample	Nominal composition (wt%)	TEOS (g)	CaNT (g)	TEP (g)	HCl (2M) (g)	H ₂ O (g)	EtOH (total amount, g)	F127 concentration (wt%)
B75-F127	75 SiO ₂ - 25 CaO	6,47	2,63	0	1	3,15	19,73	7,95
B67.5-F127	67.5 SiO ₂ - 25 CaO - 7.5 P ₂ O ₅	5,82	2,63	0,48	1	3,08	19,73	7,88
B75	75 SiO ₂ - 25 CaO	6,47	2,63	0	1	3,15	7,89	0
B67.5	67.5 SiO ₂ - 25 CaO - 7.5 P ₂ O ₅	5,82	2,63	0,48	1	3,08	7,89	0

Table 1 : Nominal compositions and amounts of reactants for the synthesis of mesostructured and conventional sol-gel B75 and B67.5 glasses

	B75	B67.5	B75-F127	B67.5-F127
BET surface area (m ² .g ⁻¹)	30	112	442	410
BJH modal pore diameter (nm)	4,6	8,9	4,3	3,9
Average pore diameter (nm)	8,4	13,2	4,3	4,8
Total pore volume (cm ³ .g ⁻¹)	0,062	0,289	0,444	0,490

Table 2 : Textural properties of conventional sol-gel and mesostructured B75 and B67.5 glasses

	XRD		a (nm)	TEM	XRD/BET
	2θ (°)			a (nm)	Wall thickness (nm)
	h = 1, k = 0	h = 1, k = 1			
B75-F127	0,84	1,49	12,12	11,2	7,8
B67.5-F127	0,9	1,6	11	9,8	7,1

Table 3 : Lattice parameters and wall thickness measured by XRD and TEM

Name	Formula	Ca/P atomic ratio
Dicalcium phosphate	CaHPO ₄ .2H ₂ O	1
Octacalcium phosphate	Ca ₈ (HPO ₄) ₂ (PO ₄) ₄ .5H ₂ O	1,33
Hydroxyapatite	Ca ₁₀ (OH) ₂ (PO ₄) ₆	1,67

Table 4 : Main Ca-P crystalline phases formed in SBF or DMEM

Figures and Tables captions

Figure 1 : Nitrogen adsorption-desorption isotherm plots (a) and pore size distribution curves (b) of calcined glasses : B75-F127(○), B67.5-F127(●), B75 (△), B67.5(▲).

Figure 2 : TEM images for mesostructured glasses : B75-F127 (a) and B67.5-F127 (b).

Figure 3 : Small Angle XRD patterns for B75, B75-F127 and B67.5-F127 glasses.

Figure 4 : Representative elemental maps of binary glasses grains before interaction and after 1 hour and 2 days of interaction with biological fluids.

Figure 5 : Representative elemental maps of binary glasses grains before interaction and after 1 hour and 2 days of interaction with biological fluids.

Figure 6 : Evolution of elemental concentrations at the periphery of B75 grains (■) and in the volume of B75-F127 grains(□).

Figure 7 : Evolution of elemental concentrations in the inner region of B75 grains (■) and in the volume of B75-F127 grains (□).

Figure 8 : Evolution of elemental concentrations at the periphery of B67.5 grains (●) and in the volume of B67.5-F127 grains (○).

Figure 9 : Evolution of elemental concentrations in biological fluids for B75 (■) and B75-F127 (□) glasses.

Figure 10 : Evolution of elemental concentrations in biological fluids for B67.5 (●) and B67.5-F127 (○) glasses.

Figure 11 : SEM images of B75-F127 glass before (a,c) and after 12 hours (b,d)of immersion in DMEM.

Figure 12 : Evolution of Ca/P atomic ratio for B75 (■), B75-F127 (□), B67.5 (●) and B67.5-F127 (○) glasses.

Table 5 : Nominal compositions and amounts of reactants for the synthesis of mesostructured and conventional sol-gel B75 and B67.5 glasses.

Table 6 : Textural properties of conventional sol-gel and mesostructured B75 and B67.5 glasses .

Table 7 : Lattice parameters and wall thickness measured by XRD and TEM.

Table 8 : Main Ca-P crystalline phases formed in SBF or DMEM.

References

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- ¹ S. Myungkoo, L. Ho-Jin, P. Jin-Young, U.H. Lee, K. Young-Uk, K. Deug Joong, *ChemPhysChem* **2008**, *9*, 1402-1408.
- ² A. Rámila, B. Muñoz, J. Pérez-Pariente, M. Vallet-Regí, *J. Sol-Gel Sci. Tech.* **2003**, *26*, 1199-1202
- ³ C. Charnay, S. Bégu, C. Tourné-Péteilh, L. Nicole, D.A. Lerner, J.-M. Devoisselle, *Eur. J. Pharm. Biopharm.* **2004**, *57*, 533-540.
- ⁴ L.L. Hench, *J. Am. Ceram. Soc.* **1998**, *81*, 1705-1728.
- ⁵ X. Yan, C. Yu, X. Zhou, J. Tang, J.; D. Zhao, *Angew. Chem. Int. Edit.* 2004, *43*, 5980-5984.
- ⁶ A. Lopez-Noriega, D. Arcos, I. Izquierdo-Barba, Y. Sakamoto, O. Terasaki, M. Vallet-Regí, *Chem. Mat.* 2006, *18*, 3137-3144.
- ⁷ A.J. Salinas, S. Shruti, G. Malavasi, L. Menabue, M. Vallet-Regí, *Acta Biomater.*, In Press, doi:10.1016/j.actbio.2011.05.033.
- ⁸ L. Zhao, X. Yan, X Zhou, L. Zhou, H Wang, H Tang, C. Yu, C. *Microporous and Mesoporous Mat.* **2008**, *109*, 210-215.
- ⁹ X. Yan, G. Wei, L. Zhao, J. Yi, H. Deng, L. Wang, G. Lu, C. Yu, *Microporous and Mesoporous Mat.*, *132*, 282-289
- ¹⁰ X. Yan, X. Huang, C. Yu, H. Deng, Y. Wang, Z. Zhang, S. Qiao, G. Lu, D. Zhao, *Biomaterials* **2006**, *27*, 3396-3403.
- ¹¹ M. Mozafari, F. Moztarzadeh, M. Tahriri, *J. Non-Cryst. Solids*, **2010**, *356*, 1470-1478.
- ¹² M. Vallet-Regí, *Dalton Transaction*, **2006**, 5211-5220.
- ¹³ C. J. Brinker, Y. F. Lu, A. Sellinger, H. Y. Fan, *Adv. Mater.* **1999**, *11*, 579.
- ¹⁴ Lee, J. T. Y.; Leng, Y.; Chow, K. L.; Ren, F.; Ge, X.; Wang, K.; Lu, X. *Acta Biomaterialia*, *7*, 2615-2622.
- ¹⁵ Mandel, S.; Tas, A. C. *Materials Science and Engineering: C*, *30*, 245-254.
- ¹⁶ Clupper, D. C.; Gough, J. E.; Hall, M. M.; Clare, A. G.; LaCourse, W. C.; Hench, L. L. *Journal of Biomedical Materials Research Part A* **2003**, *67A*, 285-294.
- ¹⁷ Lee, J. T. Y.; Leng, Y.; Chow, K. L.; Ren, F.; Ge, X.; Wang, K.; Lu, X. *Acta Biomaterialia*, *7*, 2615-2622.
- ¹⁸ Webster, T. J.; Ejiogor, J. U. *Biomaterials* **2004**, *25*, 4731-4739
- ¹⁹ Gough, J. E.; Notingher, I.; Hench, L. L. *Journal of Biomedical Materials Research Part A* **2004**, *68A*, 640-650.
- ²⁰ Ong, J. L.; Villarreal, D. R.; Cavin, R.; Ma, K. *Journal of Materials Science: Materials in Medicine* **2001**, *12*, 491-495
- ²¹ Kundu, P.; Waghode, T.; Bahadur, D.; Datta, D. *Medical and Biological Engineering and Computing* **1998**, *36*, 654-658
- ²² A. Rámila, M. Vallet-Regí, *Biomaterials* **2001**, *22*, 2301-2306.
- ²³ J. Lao, J.-M. Nedelec, P. Moretto, E. Jallot, *Nucl. Instrum. Meth. B* **2006**, *245*, 511-518.

-
- ²⁴ S. Incerti, C. Habchi, P. Moretto, J. Olivier, H. Seznec, *Nucl. Instrum. Meth. B* **2006**, *249*, 738-742.
- ²⁵ S. Incerti, Q. Zhang, F. Andersson, P. Moretto, G.W. Grime, M.J. Merchant, D.T. Nguyen, C. Habchi, T. Pouthier, H. Seznec, *Nucl. Instrum. Meth. B* **2007**, *260*, 20-27.
- ²⁶ W. Xia, J. Chang, *J. Control Release* **2006**, *110*, 522-530.
- ²⁷ H.S. Yun, S.E. Kim, Y.T. Hyun, *Solid State Sci.* **2008**, *10*, 1083-1092.
- ²⁸ Lao, J.; Nedelec, J. M.; Moretto, P.; Jallot, E. *Surface and Interface Analysis* **2008**, *40*, 162-166.
- ²⁹ Lao, J.; Nedelec, J. M.; Moretto, P.; Jallot, E. *Nuclear Instruments and Methods in Physics Research Section B* **2008**, *266*, 2412-2417.
- ³⁰ X. Lu, Y. Leng, *Biomaterials* **2005**, *26*, 1097-1108.
- ³¹ X. Lu, Y. Leng, *Biomaterials* **2005**, *26*, 1097-1108.
- ³² M. Bini, S. Grandi, D. Capsoni, P. Mustarelli, E. Saino, L. Visai, *J. Phys. Chem. C*, **2009**, *113*, 8821-8828.
- ³³ Y. Deng, X. Li, Q. Li, Q. *Ind. Eng. Chem. Res.* **2009**, *48*, 8829-8836.