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Phase relations and equation of state of a natural MORB: Implications for the density profile of subducted oceanic crust in the Earth's lower mantle

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[1] The phase relations and density of a natural mid-ocean ridge basalt (MORB) were investigated from 28 to 89 GPa and 1600 to 2700 K by in situ X-ray diffraction measurements and chemical analysis of the quenched samples using transmission electron microscopy (TEM). We observed an assemblage of five phases up to 50 GPa, namely an aluminum-bearing magnesium perovskite phase, a calcium perovskite phase, a stishovite phase, the new aluminum-rich (NAL) phase, and a calcium ferrite-type phase. The NAL phase was no longer observed above 50 GPa. The phase proportions were obtained by Rietveld refinement of the in situ X-ray diffraction patterns. After the disappearance of the NAL phase beyond 50 GPa, the proportion of each phase remains constant up to 89 GPa. The density of MORB was calculated using the measured volumes, phase proportions, and chemical compositions of the coexisting phases. The thermoelastic parameters of the MORB sample were estimated from the fit of the measured densities at various pressure and temperature conditions. Resulting MORB density profiles were calculated for different subducting slab temperature profiles. MORB densities are 0.5% to 2% greater than those of the surrounding mantle over the entire lower mantle range, suggesting MORB likely subducts to the core–mantle boundary.

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1. Introduction

[2] The fate of the subducted slabs in the Earth's deep mantle is an important issue for mantle dynamics and has been studied for many decades using different techniques (for a review, see King [2007]). *Irfune and Ringwood* [1987, 1993] presented pioneering high-pressure experimental work on the behavior of mid-ocean ridge basalt (MORB) up to 27 GPa. Since then, several studies have focused on the fate of basaltic crust around the 660 km discontinuity [*Hirose et al.*, 1999; *Litasov et al.*, 2004; *Litasov and Ohtani*, 2005]. The issue of slabs trapped at the 660 km discontinuity has been studied with other techniques

such as geodynamical modeling [e.g., *King*, 2001] and tomographic studies [e.g., *Fukao et al.*, 2001].

[3] Tomographic images show the presence of the subducted slab in the lower mantle beneath the western Pacific and Central America [e.g., *Albarède and van der Hilst*, 2002]. However, the behavior of slabs in the deep lower mantle is still uncertain. Several estimates of MORB density at lower mantle pressures indicate that subducted slabs may become buoyant in the lower mantle at depths below 1500–2000 km [*Kesson et al.*, 1998; *Ono et al.*, 2001]. More recent experimental studies suggest that MORB are always denser than the mantle and can reach the core–mantle boundary (CMB) [*Ono et al.*, 2005; *Hirose et al.*, 2005].

[4] Compared to the average mantle, MORB has a chemical composition enriched in incompatible elements (silicon, aluminum, calcium, and sodium) and depleted in compatible elements (magnesium) relative to a pyrolitic composition. Consequently, MORB assemblage at lower mantle pressure-temperature (P-T) conditions does not contain ferropericlase but displays high-pressure SiO₂ phase and additional aluminum-rich phases. A number of candidate phases have been proposed: the K-hollandite phase, the calcium aluminosilicate (CAS) phase, the new aluminum-rich (NAL) phase, and the calcium ferrite (CF)-type phase.

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Table 1. Chemical Composition of our Starting Material Along With Other Compositions From the Literature^a

Wt%	SiO ₂	TiO ₂	Al ₂ O ₃	FeO	MgO	CaO	Na ₂ O	K ₂ O	P ₂ O ₅	MnO	SO ₂	Total
Starting material (this study)	49.74 (0.49)	1.33 (0.13)	15.89 (0.23)	9.73 (0.28)	8.46 (0.12)	11.74 (0.19)	2.69 (0.09)	0.1 (0.04)	0.11 (0.03)	0.18 (0.06)	0.14 (0.05)	100.1
<i>Irifune and Ringwood</i> [1993], <i>Kesson et al.</i> [1994], <i>Funamori et al.</i> [2000]	50.39	0.57	16.08	7.68	10.49	13.05	1.87	-	-	-	-	100.13
<i>Hirose et al.</i> [1999, 2005], <i>Hirose and Fei</i> [2002]	49.64	1.64	14.88	11.43	8.51	10.55	2.9	0.12	-	-	0.18	99.85
<i>Ono et al.</i> [2001]	50.3	1.7	15.9	10	7.9	11.3	2.7	0.2	-	-	-	100
<i>Ono et al.</i> [2005]	51	1.7	15.7	9.9	7.7	11.2	2.4	0.2	-	-	-	99.8
<i>Litasov et al.</i> [2004], <i>Litasov and Ohtani</i> [2005]	51.09	1.5	15.7	9.81	7.74	11.51	2.48	0.17	-	-	-	100
<i>Hofmann</i> [1988]	50.45	1.615	15.255	10.426	7.576	11.303	2.679	883.7 (ppm)	-	-	-	99.304

^aNumbers in parentheses are 1 σ error.

K-hollandite has been shown to be stable up to 95 GPa [Tutti *et al.*, 2001]. However, K-hollandite has never been reported in any high-pressure assemblage of MORB composition but only in a sedimentary composition with ~2 wt% of potassium [Irifune *et al.*, 1994]. The CAS phase, first synthesized in a high-pressure assemblage of a sedimentary composition [Irifune *et al.*, 1994], was reported in MORB assemblage near transition zone pressures and only at very high temperature [Hirose and Fei, 2002; Litasov *et al.*, 2004]. The NAL phase is the major phase resulting from garnet decomposition at pressures of the transition zone and disappears at higher pressure because of increasing solubility of aluminum in Mg-perovskite [Miyajima *et al.*, 1999, 2001]. The NAL phase appears near the transition zone in MORB composition [Litasov *et al.*, 2004; Hirose and Fei, 2002] and is not present at pressures higher than those corresponding to the mid lower mantle [Perrillat *et al.*, 2006; Ricolleau *et al.*, 2008]. The CF-type phase was the first aluminum-rich phase reported in MORB high-pressure assemblage [Irifune and Ringwood, 1993]. Its thermodynamic parameters have been studied up to 70 GPa for the end-member compositions [e.g., Tutti *et al.*, 2000; Dubrovinsky *et al.*, 2002] and for a composition relevant to MORB [Ono *et al.*, 2002; Guignot and Andrault, 2004]. This phase has been observed in a MORB assemblage up to 130 GPa [Hirose *et al.*, 2005; Ono *et al.*, 2005].

[5] Here we present an experimental study that contributes to further understanding the fate of the oceanic crust in the Earth's lower mantle. In situ X-ray diffraction (XRD) patterns were collected on a natural MORB sample at high pressures between 28 and 89 GPa and at temperatures between 1600 K and 2700 K. We obtained the molar volumes and phase proportions from Rietveld refinement of the XRD patterns. Chemical compositions of quench phases were measured by analytical transmission electron microscopy (TEM) [Ricolleau *et al.*, 2008] on the recovered samples. The volumes, phase proportions, and chemical compositions of the coexisting phases are used to calculate the densities of the whole assemblage at different pressure-temperature conditions. The inversion of the present density data set gives the thermoelastic parameters of this assemblage, allowing us to calculate density profiles of MORB along different geotherms. These density profiles are compared with the average mantle density profiles. The

implications of the present data set for the behavior of subducted MORB in the Earth's lower mantle are discussed.

2. Experiments

2.1. Starting Material

[6] A fresh natural MORB glass from the East Pacific Rise was used as starting material (for details, see Perrillat *et al.* [2006]) (see Table 1). Although the composition of the actual oceanic crust differs from a fresh MORB glass by incorporation of variable amounts of water and carbon and by slight variations in major element composition during alteration, the present glass is generally representative of the basaltic oceanic crust bulk composition. Indeed, there are numerous indications that slabs might be dehydrated before penetrating the transition zone [e.g., Poli and Schmidt, 1995; Ono, 1998; Hacker *et al.*, 2003a]. Some of the earthquakes observed along the slab at different depths in the upper mantle could be related to oceanic crust dehydration [e.g., Meade and Jeanloz, 1991; Hacker *et al.*, 2003b]. Electrical conductivity measurements are also in agreement with the extraction of water before the transition zone [Tarits *et al.*, 2004]. Carbon behavior in subducted slabs is less well understood [e.g., Hammouda, 2003; Dasgupta *et al.*, 2004]. In some cases, carbonates are the major phases in the sedimentary layer entrained in the subduction zone as in Central American trench [Plank and Langmuir, 1998]. However, it is still unclear whether or not carbonates can be buried deeper than the transition zone [Hammouda, 2003]. Therefore, carbonates and water were not included in the present study. The chemical composition of our starting material is comparable to compositions used in other high-pressure MORB studies [e.g., Irifune and Ringwood, 1993] and to N-MORB [Hofmann, 1988] (see Table 1).

2.2. Experimental Setup

[7] Samples were loaded in 200 μm thick rhenium metal gaskets indented to a thickness of 40 to 20 μm and drilled to a diameter of 150 to 70 μm according to the desired pressure. Diamond anvil cells (DAC) with a large optical aperture [Chervin *et al.*, 1995] were used and equipped with diamond culets of 300 to 150 μm in diameter for the high-pressure experiments. Neon was used as a pressure-

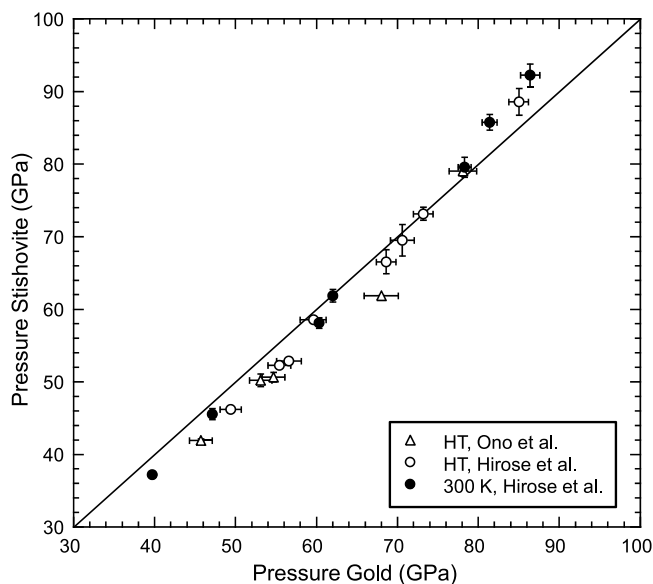


Figure 1. Pressure obtained from gold unit-cell volumes in the studies of *Ono et al.* [2005] and *Hirose et al.* [2005] compared with pressure derived from stishovite unit-cell volumes in the same studies calculated with the equation of state of *Liu et al.* [1999].

transmitting medium in runs MORB#1, #4, #5, and #6. MORB#3 was loaded with argon; MORB#2 and MORB#8 were loaded without medium. Neon pressure-transmitting media was loaded using a 2000 bar high-pressure, gas-loading apparatus [Couzinet *et al.*, 2003]. Argon was cryogenically loaded into the sample chamber.

2.3. In situ X-Ray Diffraction

[8] In situ X-ray diffraction experiments were performed at the high-pressure beamline ID30 (now relocated at ID27) at the European Synchrotron Radiation Facility (ESRF, Grenoble). The incident monochromatic X-ray beam at 0.3738 Å was focused by two mirrors in Kirkpatrick-Baez geometry into an X-ray spot $8 \times 9 \mu\text{m}^2$ in size (full width at half maximum in vertical and horizontal direction, respectively), as obtained by scanning the X-ray beam with a sharp edge.

[9] Two yttrium/aluminum/garnet (YAG) lasers (40 W each) were focused through the diamond anvils onto the sample in a double-sided, laser-heating geometry. The hot spot was about 30–40 μm in diameter. Thermal emission of the sample was recorded on one side by a spectrometer during the X-ray diffraction acquisition (for details, see *Schultz et al.* [2005]). Typical X-ray diffraction patterns were acquired for 200 s. At least five thermal emission spectra of the sample were recorded during the data acquisitions to check the temperature stability, which was kept to within ± 50 K. Angle-dispersive X-ray patterns were collected on a MAR® 345 image plate. The X-ray diffraction patterns were recorded as a function of temperature; they increased step by step from 1600 K to 2700 K at a given loading pressure. Ambient temperature patterns were collected after high-temperature cycles. Two-dimensional diffraction patterns were integrated with the Fit2d program [Hammersley *et al.*, 1996].

[10] Because of the large number of diffraction peaks in X-ray patterns, we did not add an internal pressure marker. To derive pressure conditions, we used variations in the unit-cell volume of stishovite. The validity of this approach was assessed in several ways. First, chemical analysis of stishovite in our samples showed Al_2O_3 contents less than 1 wt% (see Table 3 and *Ricolleau et al.* [2008] for more details). Second, at room temperature, pressures measured from the equation of state (EoS) of stishovite were in good agreement with pressures obtained from the equations of state of neon and argon (*Hemley et al.* [1989] and *Jephcoat* [1998], respectively) with the calculated pressure differences less than 1 GPa (2–3%) by using EoS parameters of stishovite from *Liu et al.* [1999]. Third, we tested this set for high-temperature data as follows. Similar studies on MORB composition by *Ono et al.* [2005] and *Hirose et al.* [2005] used gold as a pressure calibrant. Pressures given in latter studies agree within 5% to the pressures calculated from the thermal equation of state of stishovite given by *Liu et al.* [1999] (Figure 1). This validates the use of stishovite as an internal pressure scale in the present study. Stishovite transform from rutile-type structure to CaCl_2 -type structure after 55 GPa [e.g., *Andrault et al.*, 1998]. Since *Andrault et al.* [2003] observed that the CaCl_2 form is only slightly more compressible than the rutile-type structure (i.e., 1.5 GPa at 90 GPa), and since there are no thermal parameters for the CaCl_2 -type structure, we consider that using *Liu et al.* [1999] EoS after 55 GPa is still valid. The transition pressure of stishovite will be discussed in the following section.

[11] Rietveld refinements of the in situ XRD patterns were performed with the GSAS package [*Larson and Von Dreele*, 2004] to determine the volumes and proportions of phases. After convergence, values of the residual, R_{wp} (function of the observed intensity minus the calculated intensity; for more details, see *Rietveld* [1969] and *Larson and Von Dreele* [2004]), were always lower than 0.03. The crystal structure of the phases, atomic positions and thermal parameters from the literature were used as fixed parameters to refine the unit-cell volumes and phase proportions. The structure of phases used to refine our XRD patterns are rutile-type structure, space group $\text{P4}_2/\text{mmn}$ for stishovite [*Ross et al.*, 1990], orthorhombic structure, space group Pbnm for Mg-perovskite [*Jephcoat et al.*, 1999], cubic structure, space group $\text{Pm}3\text{m}$ for Ca-perovskite, orthorhombic structure, space group Pbnm for the CF-type phase [*Yamada et al.*, 1983] and hexagonal structure, space group $\text{P6}_3/\text{m}$ for the NAL phase [*Gasparik et al.*, 2000]. A limited subset of the in situ XRD results discussed in section 3 were previously published in *Perrillat et al.* [2006], namely MORB#1, MORB#2, MORB#3, MORB#4, and MORB#5 at room temperature and at about 2000 K to 2100 K. Here, we added higher pressure runs and refined the entire high temperature data set of all runs using the Rietveld method with the exception of MORB#1, which is too noisy for good refinement due to the intense neon peaks. MORB#1, MORB#4, MORB#5, and MORB#8 were also previously used for analytical TEM work [*Ricolleau et al.*, 2008].

2.4. Analytical Transmission Electronic Microscopy Investigation

[12] The recovered samples were placed on a conducting carbon tape, and prepared for TEM observations. We used

Table 2. Volume and Proportion of Phases and Whole Assemblage Density as a Function of *P-T* Conditions^a

MORB#	Temperature (K)	Pressure (GPa)	Mg-Perov Volumes (Å ³)	Proportions (wt%)	Ca-Perov Volumes (Å ³)	Proportions (wt%)	Stishovite Volumes (Å ³)	Proportions (wt%)	CF Volumes (Å ³)	Proportions (wt%)	NAL Volumes (Å ³)	Proportions (wt%)
MORB#2	1628	28.7 (1.2)	159,117 (131)	22.70	41,956 (9)	30.07	43,783 (10)	21.85	221,921 (478)	10.71	164,729 (224)	14.68
	1791	30.2 (1.2)	158,770 (159)	21.35	41,920 (10)	28.59	43,749 (9)	22.32	221,343 (435)	11.77	165,587 (153)	15.97
	1931	30.4 (1.3)	159,521 (107)	17.24	41,947 (9)	32.66	43,856 (11)	22.33	221,270 (741)	11.91	165,961 (207)	15.86
	2077	32.2 (1.3)	158,918 (163)	21.65	41,918 (10)	31.77	43,778 (10)	23.37	221,561 (285)	9.73	165,461 (114)	13.48
	2229	32.7 (1.3)	159,654 (153)	22.42	42,008 (10)	30.08	43,863 (11)	22.82	221,868 (762)	10.06	165,428 (223)	14.62
	2398	34.1 (1.4)	159,709 (154)	22.96	42,065 (11)	26.88	43,87 (13)	23.30	221,938 (457)	10.42	165,440 (194)	16.43
	300	22.7 (0.1)	158,534 (165)	20.82	41,79 (9)	30.03	43,612 (10)	21.64	220,326 (610)	12.11	164,353 (256)	15.40
	300	21.4 (0.1)	159,011 (171)	20.18	41,921 (10)	30.25	43,749 (9)	22.73	221,600 (454)	10.15	165,018 (197)	16.69
	1779	34.4 (1.2)	157,779 (98)	31.53	41,571 (9)	26.30	43,276 (11)	15.70	217,612 (261)	10.09	165,226 (133)	16.38
	1954	35.2 (1.3)	158,276 (111)	33.36	41,621 (10)	26.16	43,34 (11)	15.32	218,289 (243)	9.21	165,432 (119)	15.94
MORB#3	2149	36.2 (1.3)	158,493 (95)	37.31	41,716 (10)	24.44	43,395 (12)	14.82	218,244 (242)	8.56	166,025 (131)	14.86
	2350	36.8 (1.4)	158,753 (116)	29.01	41,824 (10)	26.44	43,506 (13)	16.98	219,059 (254)	10.62	165,587 (124)	16.96
	2526	37.8 (1.4)	158,371 (112)	27.36	41,86 (10)	33.62	43,564 (10)	17.58	219,359 (433)	4.65	165,863 (100)	16.78
	300	29.2 (0.1)	155,546 (111)	30.59	41,131 (8)	28.67	42,952 (11)	14.10	215,074 (424)	8.07	163,332 (135)	18.58
	300	28.9 (0.1)	155,467 (126)	30.57	41,127 (9)	27.63	42,976 (11)	14.44	216,704 (421)	9.41	163,77 (134)	17.95
	300	29.2 (0.1)	156,184 (158)	34.68	41,137 (9)	23.26	42,947 (11)	13.53	213,512 (959)	10.31	163,476 (141)	18.21
	300	0.0	168,342 (76)				46,681 (5)		243,949 (149)		183,017 (146)	
	1993	43.2 (1.4)	152,274 (96)	30.32	40,513 (15)	32.13	42,543 (20)	18.44	212,011 (938)	12.86	163,175 (419)	6.26
	2104	41.8 (1.4)	153,222 (143)	28.29	40,991 (26)	31.24	42,77 (24)	19.40	218,218 (1,168)	13.82	163,709 (210)	7.26
	2225	43.6 (1.4)	153,918 (115)	35.44	41,029 (18)	29.68	42,685 (15)	19.62	214,191 (727)	9.81	164,081 (249)	5.44
2291	43.6 (1.4)	154,674 (154)	35.97	41,118 (20)	33.23	42,739 (18)	19.64	212,748 (727)	8.64	163,818 (368)	2.51	
2391	44.8 (1.5)	154,483 (189)	38.82	41,138 (26)	25.58	42,701 (22)	20.06	210,394 (453)	9.44	163,775 (265)	6.10	
300	35.5 (0.1)	151,297 (81)	40.78	40,26 (10)	25.26	42,357 (13)	16.72	211,914 (594)	9.57	161,125 (228)	7.67	
300	36.0 (0.2)	151,511 (98)	39.14	40,349 (13)	24.32	42,314 (18)	19.53	209,736 (855)	9.22	161,054 (597)	7.79	
300	35.1 (0.2)	151,343 (102)	33.59	40,157 (18)	25.31	42,396 (16)	18.52	214,979 (583)	15.36	161,148 (126)	7.21	
MORB#5	2055	52.4 (1.3)	148,661 (81)	49.46	39,247 (9)	29.57	41,742 (13)	13.34	206,252 (174)	7.63		
	2064	49.8 (1.4)	149,937 (105)	46.86	39,146 (11)	30.20	41,977 (17)	14.85	207,289 (395)	8.09		
	2217	52.6 (1.4)	150,459 (110)	45.07	39,191 (13)	26.90	41,828 (15)	17.03	207,136 (264)	11.00		
	2270	53.8 (1.4)	148,981 (86)	45.20	39,234 (9)	33.26	41,759 (13)	14.33	206,440 (160)	7.21		
	2314	54.7 (1.4)	148,214 (57)	46.77	39,160 (9)	27.83	41,713 (13)	15.71	207,269 (340)	9.69		
	2500	54.2 (1.4)	150,123 (107)	43.31	39,427 (16)	31.35	41,897 (18)	16.28	207,656 (229)	9.06		
	2550	55.2 (1.4)	149,825 (126)	41.29	39,322 (15)	31.30	41,841 (17)	15.73	208,121 (339)	11.68		
	2693	55.4 (1.4)	149,62 (90)	44.59	39,152 (10)	31.16	41,933 (14)	14.64	207,298 (311)	9.60		
	300	42.3 (0.2)	148,35 (59)	46.68	39,105 (9)	29.55	41,776 (13)	14.69	206,858 (323)	9.08		
	300	43.4 (0.1)	147,583 (50)	42.76	39,079 (8)	31.12	41,68 (11)	15.64	206,616 (252)	10.48		
300	43.2 (0.1)	147,521 (45)	47.32	39,139 (8)	28.69	41,698 (11)	14.92	206,909 (282)	9.07			
300	43.0 (0.1)	147,562 (44)	44.85	39,159 (7)	30.64	41,714 (12)	14.70	207,107 (306)	9.81			
300	43.5 (0.1)	147,675 (47)	40.03	39,156 (7)	31.04	41,672 (10)	16.94	206,978 (254)	11.98			
300	42.7 (0.1)	148,219 (52)	46.51	39,074 (9)	30.91	41,743 (12)	14.35	206,815 (280)	8.23			
300	0.0	169,840 (230)				46,848 (7)		244,637 (277)				
MORB#6 ^c	1783	58.1 (1.6)	145,403 (102)		38,278 (21)		41,097 (44)		200,105 (466)			
	2023	61.8 (2.1)	145,033 (128)		38,256 (16)		40,938 (76)		201,538 (756)			
	2137	63.1 (1.8)	145,150 (88)		38,246 (12)		40,899 (51)		202,236 (903)			
	2183	62.4 (1.9)	145,527 (139)		38,361 (15)		40,984 (56)		203,297 (794)			
	2337	64.2 (2.1)	144,931 (132)		38,186 (27)		40,933 (69)		204,347 (1,262)			

Table 2. (continued)

Temperature (K)	Pressure (GPa)	Mg-Perov Volumes (Å ³)	Proportions (wt%)	Ca-Perov Volumes (Å ³)	Proportions (wt%)	Stishovite Volumes (Å ³)	Proportions (wt%)	CF Volumes (Å ³)	Proportions (wt%)	NAL Volumes (Å ³)	Proportions (wt%)
2430	61.5 (1.7)	146.336 (122)		38.497 (15)		41.212 (38)		202.446 (847)			
2433	64.2 (1.8)	145.386 (98)		38.325 (12)		40.996 (45)		203.752 (769)			
2543	63.7 (2.0)	146.458 (141)		38.485 (16)		41.106 (63)		204.311 (922)			
300	55.7 (1.4)	144.189 (134)		37.955 (23)		40.739 (100)		201.256 (571)			
300	54.8 (0.8)	144.289 (132)		37.974 (18)		40.802 (61)		201.284 (699)			
MORB#8 ^d											
1796	80.3 (1.3)	139.208 (77)	41.15	37.083 (9)	32.86	39.504 (23)	14.87	192.370 (150)	11.12		
1876	80.4 (1.3)	139.397 (76)	42.60	37.128 (9)	30.31	39.531 (21)	14.28	192.542 (145)	12.81		
1963	79.7 (1.4)	139.961 (80)	44.21	37.2 (9)	29.70	39.613 (22)	14.35	192.975 (136)	11.74		
1973	80.5 (1.3)	139.433 (74)	43.39	37.165 (8)	30.77	39.568 (19)	14.34	192.738 (125)	11.50		
2028	80.8 (1.4)	139.862 (67)	45.29	37.098 (11)	29.79	39.573 (22)	12.41	192.650 (268)	12.51		
2121	80.2 (1.4)	140.142 (79)	44.34	37.241 (9)	29.64	39.655 (22)	14.55	193.229 (132)	11.47		
2263	80.5 (1.5)	140.112 (76)	43.37	37.276 (10)	30.91	39.703 (22)	14.56	193.451 (136)	11.16		
2294	83.8 (1.4)	138.538 (77)	41.06	37.182 (8)	31.60	39.502 (16)	14.35	192.584 (231)	12.99		
2433	81.1 (1.5)	140.239 (74)	43.54	37.324 (9)	30.46	39.753 (22)	14.72	193.766 (130)	11.28		
2580	81.2 (1.5)	140.727 (74)	44.21	37.387 (10)	29.57	39.822 (23)	15.02	194.282 (130)	11.21		
300	77.6 (0.4)	138.269 (83)	41.34	36.847 (10)	32.95	39.314 (22)	13.95	191.067 (166)	11.75		
300	79.1 (0.4)	137.994 (82)	40.39	36.776 (10)	32.89	39.226 (26)	15.40	190.653 (167)	11.32		
300	80.0 (0.3)	137.261 (73)	42.80	36.706 (9)	32.48	39.173 (19)	14.35	190.085 (183)	10.37		
300	76.3 (0.4)	138.235 (73)	43.18	36.973 (10)	31.24	39.395 (21)	14.23	191.359 (195)	11.36		
1813	87.0 (1.3)	136.146 (56)	43.16	36.715 (7)	30.65	39.092 (18)	14.98	193.063 (281)	11.21		
1967	87.3 (1.3)	136.296 (52)	41.86	36.769 (7)	29.82	39.130 (18)	15.30	193.343 (285)	13.03		
2115	87.7 (1.4)	137.040 (68)	43.25	36.796 (7)	29.75	39.170 (19)	15.29	190.266 (176)	11.70		
2132	88.2 (1.4)	136.523 (59)	43.02	36.778 (7)	29.84	39.149 (21)	15.31	193.710 (385)	11.84		
2134	87.7 (1.4)	136.577 (55)	42.57	36.823 (7)	29.03	39.181 (20)	15.43	193.732 (329)	12.97		
2219	88.2 (1.4)	136.671 (59)	43.78	36.823 (7)	28.53	39.182 (20)	15.57	193.943 (336)	12.11		
2248	88.1 (1.4)	136.688 (54)	40.56	36.832 (7)	30.56	39.204 (18)	15.51	194.015 (357)	13.37		
2399	88.9 (1.4)	136.786 (57)	44.73	36.844 (7)	27.87	39.223 (20)	15.76	194.108 (336)	11.64		
300	84.0 (0.3)	135.322 (52)	36.78	36.523 (10)	35.47	38.947 (18)	14.33	191.494 (290)	13.42		
300	84.5 (0.3)	135.798 (77)	37.92	36.497 (9)	35.23	38.920 (19)	15.07	188.648 (231)	11.78		
300	85.1 (0.4)	135.812 (83)	39.43	36.469 (9)	34.09	38.885 (21)	15.52	188.542 (212)	10.96		
300	85.2 (0.4)	135.763 (84)	40.94	36.460 (9)	32.90	38.880 (21)	15.18	188.386 (202)	10.98		
300	85.4 (0.4)	135.770 (83)	40.96	36.448 (9)	33.28	38.870 (20)	14.86	188.338 (194)	10.91		

^aNumbers in parentheses are 1σ error in last digits.

^bIn run MORB#4, the error bars are larger than the other runs because of an important proportion of our transmitting pressure medium, i.e., neon.

^cIn run MORB#6, the proportion of phase were not refined because of a much too strong proportion of neon in the XRD patterns.

^dIn run MORB#8, two sets of pressures are presented at high-pressure and high-temperature. The first set was initially collected at 80 GPa before raising the pressure up to 89 GPa.

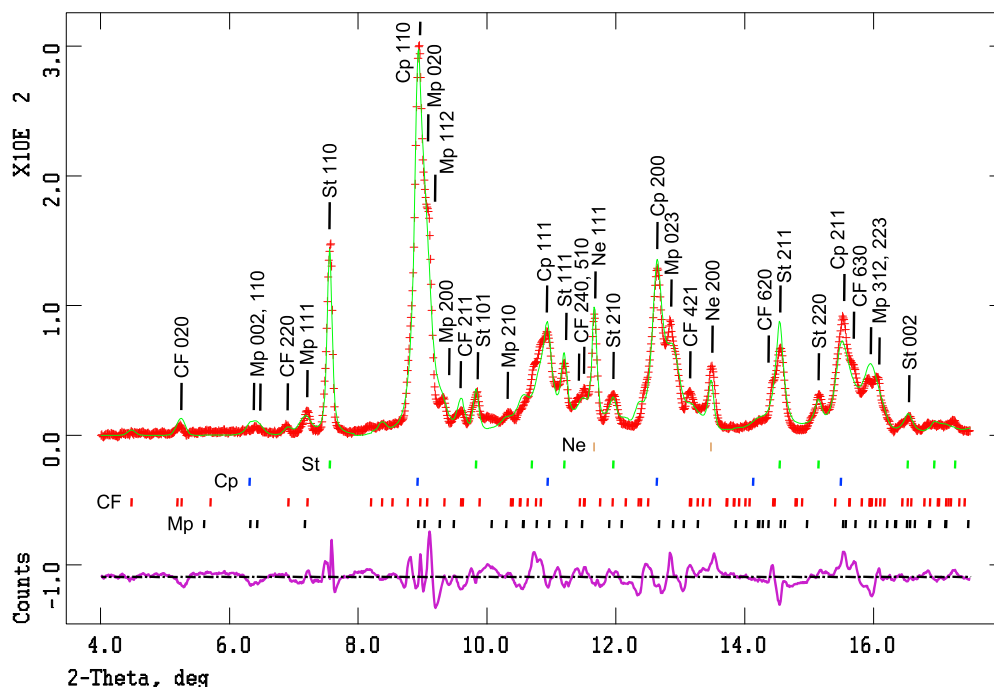


Figure 2. In situ X-ray diffraction pattern acquired at 53 GPa and 2200 K (red line) and its refined pattern (green line). Tick marks stand for diffraction peaks of Mg-perovskite, CF-type phase, Ca-perovskite, stishovite and neon, from bottom to top. The main diffraction peaks of each phase are indexed on top of the pattern.

the focused ion beam (FIB) “lift-out” technique [Giannuzzi and Stevie, 1999] to extract thin film from the laser-heated sample area. The preparation was performed with a FEI Model 200 TEM FIB system at the University Aix-Marseille III. The FIB system uses a gallium liquid metal ion source, which allows milling. At the end of the milling, the thin film of $\sim 15 \mu\text{m}$ length and $\sim 5 \mu\text{m}$ width with a thickness of $\sim 100 \text{ nm}$ was removed from the recovered sample and deposited onto a TEM grid.

[13] Selected area electron diffraction (SAED) patterns were obtained on several grains for phase identification using a 200 kV JEOL-2100F TEM at the Institut de Minéralogie et Physique des Milieux Condensés (Paris, France). The chemical compositions of the phases were measured with a 300 kV Philips CM30 transmission electron microscope equipped with a Noran Instruments energy-dispersive analytical system (EDS) at Laboratoire de Structure et Propriétés de l’Etat Solide in Lille (for details, see Ricolleau *et al.* [2008]).

3. Results

3.1. Unit-Cell Volume of MORB High-Pressure Phases

[14] Six runs, conducted at pressures between 28 GPa and 89 GPa, and at high temperature, show resolvable XRD patterns (MORB#1 was only used for TEM work, MORB#7 is absent because of poor-quality XRD patterns). Pressure-temperature conditions and the results of XRD patterns refinement are presented in Table 2. Figure 2 shows an example of an X-ray pattern obtained at 53 GPa and 2200 K and the corresponding refined pattern. Diffraction lines were assigned to stishovite, Ca-perovskite, aluminum-bearing Mg-

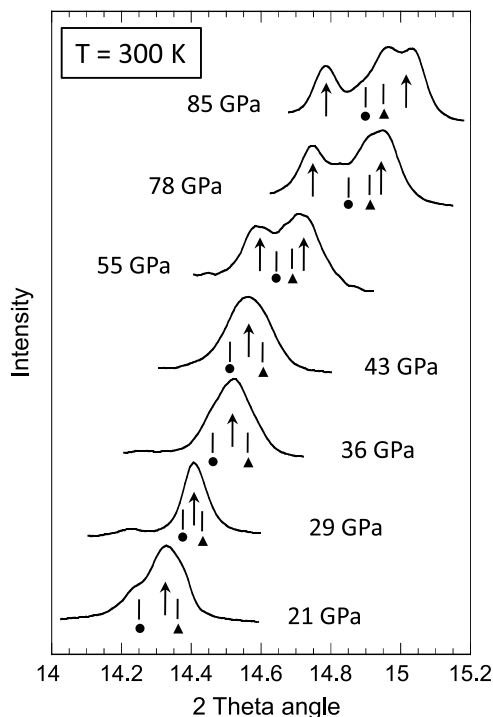


Figure 3. Evolution of X-ray diffraction patterns acquired at 300 K with pressure. The arrows point the peak of stishovite, (211) in rutile-type SiO_2 ($P4_2/mnm$). After 55 GPa, the peak split into two peak, (121) and (211), due to the transition to the CaCl_2 -type SiO_2 ($Pnmm$). Tick marks indicated by circles and triangles are peaks (620) and (131) of CF-type phase and Mg-perovskite, respectively.

Table 3. Chemical Composition Obtained by Analytical Transmission Electronic Microscopy on MORB Samples Recovered From High-Pressure High Temperature Experiments^a

	Mgpv	St	Capv	CF	NAL
33 GPa/MORB#1					
SiO ₂ (wt%)	45.4 (1.9)	98.8 (0.3)	51.6 (0.1)	26.0 (1.3)	26.5 (1.5)
TiO ₂ (wt%)	3.6 (0.0)	0.1 (0.1)	1.3 (0.4)	0.4 (0.2)	0.3 (0.2)
Al ₂ O ₃ (wt%)	13.4 (0.7)	0.5 (0.2)	3.3 (0.4)	37.8 (0.9)	38.9 (2.2)
FeO (wt%)	14.0 (0.9)	0.1 (0.2)	3.5 (1.3)	14.0 (0.5)	8.7 (1.0)
MgO (wt%)	20.8 (2.0)	0.0 (0.1)	1.4 (1.1)	12.7 (1.7)	17.7 (1.2)
CaO (wt%)	2.7 (1.5)	0.4 (0.3)	38.8 (2.5)	1.3 (0.3)	2.2 (0.4)
Na ₂ O (wt%)	0.0 (0.0)	0.0 (0.0)	0.0 (0.0)	7.9 (1.5)	5.9 (0.4)
Proportions	23 (5)	17 (2)	28 (2)	22 (5)	10 (5)
Weight (g/mol)	109.26	60.17	115.56	152.03	148.91
Weight after correction	110.96	60.17	115.56	153.10	150.37
44 GPa/MORB#4					
SiO ₂ (wt%)	42.3 (1.4)	98.7 (0.4)	51.9 (0.4)	28.5 (2.1)	23.3 (0.4)
TiO ₂ (wt%)	2.5 (0.3)	0.1 (0.1)	1.3 (0.2)	0.5 (0.3)	0.3 (0.1)
Al ₂ O ₃ (wt%)	21.2 (1.4)	0.8 (0.2)	3.5 (0.8)	39.8 (1.9)	45.0 (0.7)
FeO (wt%)	11.3 (1.4)	0.1 (0.1)	4.4 (0.9)	9.7 (2.4)	4.7 (0.4)
MgO (wt%)	18.7 (1.2)	0.0 (0.1)	3.2 (1.5)	11.0 (3.3)	18.9 (0.6)
CaO (wt%)	3.5 (0.7)	0.1 (0.1)	35.7 (1.8)	1.3 (0.4)	1.9 (0.4)
Na ₂ O (wt%)	0.5 (0.3)	0.0 (0.0)	0.1 (0.1)	9.2 (1.1)	5.9 (0.2)
Proportions	38 (5)	15 (2)	30 (2)	12 (3)	6 (2)
Weight (g/mol)	107.88	60.12	114.88	149.38	146.16
Weight after Correction	110.44	60.12	114.88	151.53	147.94
55 GPa/MORB#5					
SiO ₂ (wt%)	40.9 (0.6)	98.5 (0.4)	52.1 (0.2)	26.7 (0.4)	
TiO ₂ (wt%)	2.1 (0.2)	0.1 (0.1)	1.2 (0.1)	0.2 (0.1)	
Al ₂ O ₃ (wt%)	23.6 (0.9)	0.9 (0.1)	2.5 (0.7)	43.2 (1.3)	
FeO (wt%)	10.1 (0.5)	0.1 (0.1)	3.5 (0.6)	6.2 (1.1)	
MgO (wt%)	19.1 (0.6)	0.1 (0.1)	3.9 (1.1)	12.9 (1.7)	
CaO (wt%)	3.3 (0.3)	0.4 (0.3)	36.8 (2.5)	1.7 (0.6)	
Na ₂ O (wt%)	0.8 (0.2)	0.0 (0.0)	0.1 (0.1)	9.1 (0.2)	
Proportions	40 (4)	16 (1)	29 (2)	15 (2)	
Weight (g/mol)	107.05	60.16	114.63	147.02	
Weight after Correction	110.35	60.16	114.63	149.39	

^aMgpv, magnesium perovskite; St, stishovite; Capv, calcium perovskite; CF, calcium ferrite-type phase; NAL, new aluminum-rich phase. Proportions in wt% obtained from least squares mass balance calculation. Numbers in parentheses are the standard deviation (1σ).

perovskite, and CF-type phase. These phases were observed at all pressures between 28 and 89 GPa. Experiments carried out at pressures below 45 GPa are characterized by the occurrence of the NAL phase, which disappears at higher pressure [Perrillat *et al.*, 2006]. The observed mineral assemblages are in good agreement with previous DAC studies on MORB [Kesson *et al.*, 1994; Ono *et al.*, 2005; Hirose *et al.*, 2005]. In runs MORB#6 and MORB#8 (i.e., at 60 and 89 GPa), we observed splitting of reflections 101 and 211 of stishovite into 011–101 and 121–211 doublets, corresponding to the phase transition from the rutile-type structure into the CaCl₂-type structure of SiO₂ [e.g., Andrault *et al.*, 1998]. Figure 3 shows the evolution of the room temperature XRD patterns around the 211 peak of stishovite with increasing pressure, with splitting of the peak after 55 GPa clearly visible. Recent studies by Bolfan-Casanova *et al.* [2009] and Lakshatanov *et al.* [2007] show that the transition starts around 25 GPa in stishovite containing about 4 wt% Al₂O₃. Lakshatanov *et al.* [2005] observed no transition up to at least 44 GPa and a possible beginning of the transition at 49 GPa by studying stishovite containing 1.8 wt% Al₂O₃. Our higher-pressure transition is in good agreement with our chemical analysis

on stishovite giving less than 1 wt% Al₂O₃ (see Table 3). In the Rietveld refinement, we subsequently used the CaCl₂-type structure from Haines *et al.* [1996] at higher pressures. The transition in stishovite structure occurs between 55 and 62 GPa at high temperature in our sample, which is in good agreement with the study of Hirose *et al.* [2005] where it is observed between 58 and 68 GPa.

[15] For Ca-perovskite, recent studies have shown a phase transition that occurs at about 600 K in pure CaSiO₃ between tetragonal and cubic structures [e.g., Shim *et al.*, 2002, Komabayashi *et al.*, 2007]. This phase transition is a second-order phase transition without any volume change [Komabayashi *et al.*, 2007]. We observed a splitting of the 200 peak only in few of our room temperature XRD patterns. More frequently, a shoulder is just barely observable. Therefore, we used the cubic structure to refine the volume of Ca-perovskite. Kurashina *et al.* [2004] observed a transition between orthorhombic and cubic structures with increasing temperature in an aluminum-bearing Ca-perovskite. In our XRD patterns, the presence of multiple phases including orthorhombic Mg-perovskite makes such observation very difficult. Nevertheless, the intensity of the cubic Ca-perovskite peaks does not change from high temperature to room tem-

Table 4. Summary of P - V - T Equation of State Parameters for MORB High-Pressure Phases Compared With Previous Studies^a

	V_0 (Å ³)	$K_{0,T}$ (GPa)	$K'_{0,T}$	$(dK/dT)_P$ (GPaK ⁻¹)	α_0 (10 ⁻⁵ K ⁻¹)	α_1 (10 ⁻⁸ K ⁻²)	
Mg-perovskite	170.0 (1)	238 (2)	4 (fixed)	-0.044 (2)	2.82 (35)	1.65 (33)	This study
	170.0 (1)	239 (3)	3.98 (7)	-0.044 (2)	2.83 (35)	1.62 (34)	This study
	169.7 (2)	225 (5)	4.72 (11)	-0.048 (6)	0.80 (55)	3.61 (71)	This study, after NAL disappearance
	169.5 (fixed)	217 (2)	4 (fixed)				Hirose et al. [2005]
	163.51 (6)	232	4 (fixed)				Yagi et al. [2004], En90
	164.85	253.4 (6.2)	4 (fixed)				Walter et al. [2004], En75
CF-type phase	243.3 (2)	197 (1)	4 (fixed)	-0.007 (2)	1.96 (22)		This study
	244.2 (2)	174 (3)	4.86 (12)	-0.013 (3)	2.56 (27)		This study
	244.6 (3)	168 (3)	5.09 (15)	-0.004 (5)	1.63 (51)		This study, after NAL disappearance
	239.7 (1.5)	214 (8)	4 (fixed)				Hirose et al. [2005]
	242.7 (3)	193 (2)	4 (fixed)				Guignot and Andraut. [2004], exp. 6
	245.4 (6)	200 (5)	4 (fixed)				Guignot and Andraut [2004], exp. 7
	244.07 (55)	253 (14)	3.6 (6)				Ono et al. [2002]
Ca-perovskite	45.58 (fixed)	223 (1)	4 (fixed)	-0.011 (2)	2.21 (18)		This study
	45.58 (fixed)	209 (1)	4.8 (fixed)	-0.021 (2)	2.86 (20)		This study
	45.58 (fixed)	194 (1)	5.78 (05)	-0.029 (2)	3.51 (20)		This study
	45.6 (2)	245 (6)	4 (fixed)				Hirose et al. [2005]
	45.58 (4)	232 (8)	4.8 (3)	-0.033 (8)	3.0		Wang et al. [1996]
	45.58 (4)	236 (4)	3.9 (2)	-0.028 (11)	2.2 (3)		Shim et al. [2000]
	ρ (g/cm ³)	$K_{0,T}$ (GPa)	$K'_{0,T}$	$(dK/dT)_P$ (GPaK ⁻¹)	α_0 (10 ⁻⁵ K ⁻¹)		
MORB density	4.25 (1)	246 (5)	4 (fixed)	-0.025 (5)	2.83 (27)		This study
	4.25 (1)	245 (5)	4 (fixed)	-0.022 (5)	2.59 (40)		This study, after NAL disappearance

^aThermal expansion coefficient, α is described as $\alpha_0 + \alpha_1 T$. Numbers in brackets are 1σ error in last digits.

perature. The proportions of Ca-perovskite obtained by the Rietveld refinement remain constant at all temperatures (see Table 2).

3.2. Chemical Composition of Phases

[16] Samples from runs MORB#1 MORB#4, and MORB#5 were recovered and prepared for analytical TEM analysis. Run MORB#1 ($P = 33$ GPa), showed weak diffraction patterns making the Rietveld method of analysis difficult to perform: some analyzed patterns were previously reported in Perrillat et al. [2006], and the recovered sample was studied by TEM. Every phase observed in the in situ X-ray diffraction patterns was confirmed by selected area electron diffraction patterns. Stishovite, Mg-perovskite, Ca-perovskite, CF-type phase, and the NAL phase were observed at 33 GPa (MORB#1) and 44 GPa (MORB#4). At 55 GPa (MORB#5), the NAL phase has disappeared. In all samples, both perovskites are amorphous upon recovery; nevertheless, they can be distinguished by their different contrast in backscattered and transmitted electron images and by their different chemical composition [Ricolleau et al., 2008]. The protocol used for quantifying chemical compositions is described in Ricolleau et al. [2008]. The compositions of each phase for the preceding three pressures are shown in Table 3. The Al₂O₃ content of Mg-perovskite is 13 wt% in MORB#1 and increases with pressure up to 24 wt% in MORB#5 synthesized at 55 GPa. Similarly, the Al₂O₃ content of CF-type phase increases with pressure, from 38 wt% at 33 GPa to 43 wt% once the NAL phase disappears at pressure exceeding 55 GPa. Thus, the Al increase

in Mg-perovskite and CF-type phase might be linked to the NAL phase disappearance. On the other hand, the Al₂O₃ content in stishovite never exceeds 1 wt% at any pressure. The modal mineralogy was estimated using mass balance and by assuming that the bulk chemical composition of the samples is equal to the starting composition (Table 3).

[17] Chemical mass balance showed that laser-heated areas were affected by 15%–40% iron loss and commensurate magnesium enrichment. We observed that the iron loss is about 16%, 27%, and 38% in samples synthesized at 33, 44, and 55 GPa, respectively. At the same time, magnesium enrichment is about 15%, 24%, and 27%, respectively. Iron diffusion has been observed in laser-heated DAC samples [e.g., Sinmyo et al., 2008; Hirose et al., 2005] due to the large thermal gradient, which is known as the Soret effect.

[18] Mass balance calculation shows a loss of sodium of about 14%, 38%, and 36% in our samples synthesized at 33, 44, and 55 GPa, respectively. FIB milling use gallium ions bombardment, which leads to significant Ga implantation in the slice. Ga L α line is very close in energy to the Na K α line, the two peaks are then correlated. This lack of sodium is mainly due to a correlation of peaks that cannot be solved and is not constant through samples.

3.3. Equation of State of MORB High-Pressure Phases

[19] The thermoelastic parameters of the observed phases were deduced from our dataset using a numerical code based on the generalized nonlinear inversion technique [Tarantola and Valette, 1982] and a third-order Birch-Murnaghan

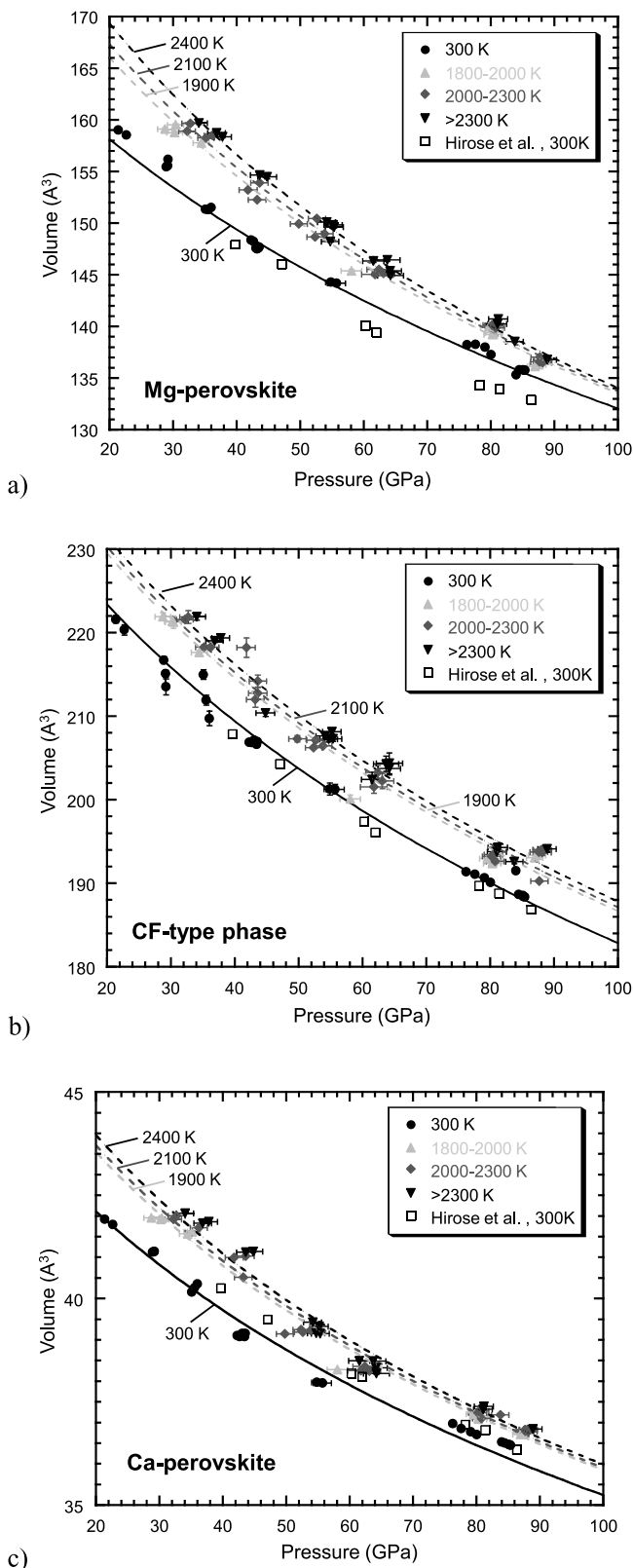


Figure 4. Variation of volume as a function of pressure at room temperature and at high temperature for Mg-perovskite, CF-type phase and Ca-perovskite. Isotherms are calculated at temperatures between 300 K and 2400 K using refined thermoelastic parameters.

equation of state. The best-fit elastic parameters are reported in Table 4, for either a fixed $K' = 4$ or a refined K' . As unit-cell volumes of the NAL phase could not be obtained over a wide enough pressure range, its elastic parameters could not be constrained.

[20] To check for significant changes of the behavior of the phases due to NAL phase disappearance, we refitted the thermoelastic parameters of Mg-perovskite and CF-type phase using only the volumes measured at pressures above 50 GPa (Table 4). The elastic parameters of Ca-perovskite after NAL disappearance are not given because the fit convergence was poor (i.e., $\chi^2 > 2$). The thermoelastic parameters obtained for Mg-perovskite and CF-type phase indicate a change in the thermoelastic behavior of these two phases before and after the NAL phase disappearance. However, it should be noted that the second sets of parameters were obtained over a limited pressure range. The change in thermoelastic behavior is likely due to the change in chemical composition of these phases as the NAL phase disappears [Ricolleau et al., 2008]. The Al₂O₃ content in the Mg-perovskite and CF-type phase increase with pressure as mentioned in a previous section.

[21] Table 4 also shows the elastic parameters available in the literature for comparison. The bulk modulus obtained on the aluminum-bearing Mg-perovskite is significantly lower than that of the MgSiO₃ end-member. This is in good agreement with several works that studied the behavior of Mg-perovskite with incorporating aluminum; these works showed that increasing aluminum content in Mg-perovskite decreased its bulk modulus [e.g., Nishiyama et al., 2007]. The bulk modulus of CF-type phase and Ca-perovskite are also slightly lower than those given in the literature for end-member compositions. The bulk modulus of the CF-type phase obtained on a similar composition by Guignot and Andrault [2004] is in agreement (Table 4). Concerning Ca-perovskite, we assume that its aluminum content could cause the difference in compressibility with CaSiO₃ end-member.

[22] Figure 4 displays the changes of the unit-cell volume of Mg-perovskite, CF-type phase, and Ca-perovskite as a function of pressure at different temperatures, compared with the isotherms calculated using the fitted thermal EoS parameters. The volumes of each phase at room temperature obtained by Hirose et al. [2005] on a MORB composition are plotted for comparison. We note that the volume of Ca-perovskite and the CF-type phase are in good agreement with our measurements. The volumes of Mg-perovskite given in the study of Hirose et al. [2005] deviate from our volume at pressures above 60 GPa by about 3%. Our Mg-perovskite contains more aluminum than the one in the study of Hirose et al. [2005], which might explain this difference. As shown in Figure 4a, the volume of Mg-perovskite below 35 GPa cannot be fitted well by the EoS parameters obtained on the entire range. As we study an assemblage where phase chemistry is not constant though pressure, this observation can be attributed to a large change in Mg-perovskite composition. Indeed, at 35 GPa the NAL phase is disappearing, and the Al₂O₃ content of Mg-perovskite nearly doubles from 13 wt% to 24 wt% (Table 3). Compression curves for the other phases do not show the effect seen for Mg-perovskite, most likely because their chemistry changes less drastically at NAL phase disappearance.

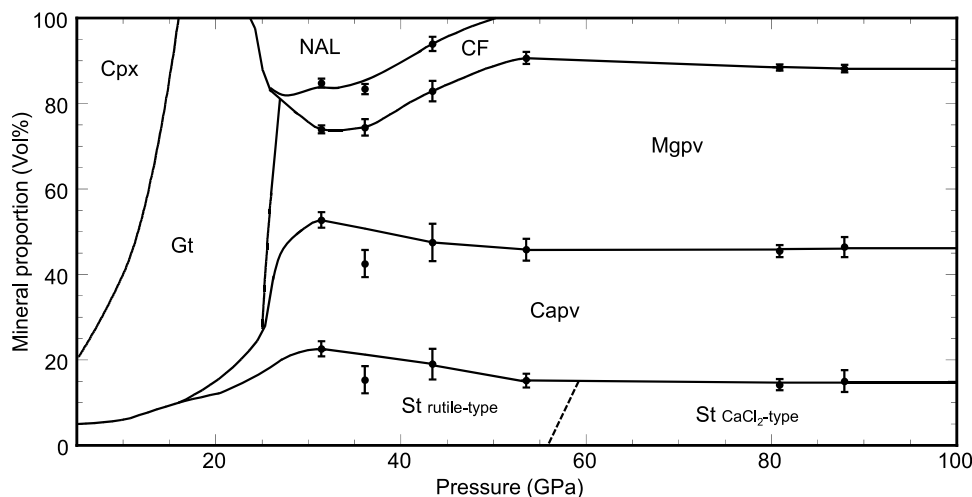


Figure 5. Variation of mineral proportion obtained by Rietveld refinement on XRD patterns as a function of depth. Cpx, clinopyroxene; Gt, garnet; St, stishovite in rutile-type structure and in CaCl_2 -type structure; NAL, new Al-rich phase; CF, calcium ferrite-type phase; Capv, calcium perovskite; Mgpv, aluminum-bearing magnesium perovskite. Phase proportions at pressures less than 30 GPa are taken from multianvil experiments [Irfune *et al.*, 1986, Irfune and Ringwood, 1993; Litasov and Ohtani, 2005].

3.4. Proportion of Phases

[23] The relative amounts of each phase (Table 2) were obtained by Rietveld refinement of in situ XRD patterns [see Perrillat *et al.*, 2006 for further details]. The chemical compositions obtained by analytical TEM were used to specify the unit-cell composition of the phases. Only samples MORB#1, MORB#4, and MORB#5 (synthesized at 33, 44, and 55 GPa, respectively) were successfully analyzed by analytical TEM, in terms of chemical compositions of the coexisting phases. Consequently, we used the compositions obtained on MORB#1 for MORB#2 and MORB#3 (performed at 33 and 36 GPa, respectively) because these runs were performed in the same range of pressure and temperature. Because no phase change was observed between 55 and 89 GPa in our X-ray diffraction experiments, we assume that chemical composition changes were minor and would not significantly modify our density calculations. Therefore, the compositions determined from sample MORB#5 (55 GPa) were used for MORB#6 and MORB#8 (performed at 64 and 89 GPa, respectively).

[24] Mineral proportions determined from Rietveld refinement of the XRD patterns were also compared with mineral proportions determined from mass balance calculations of the chemical compositions of phases. The two methods are in agreement within the error bars (see Tables 2 and 3 for comparison).

[25] In each run, we compared phase proportions obtained from our XRD refinements at different temperatures and did not observe any significant change in mineral proportions, within the uncertainty. This might indicate that mineral abundances in MORB at subsolidus conditions do not change significantly with temperature. In the following section, we assumed no change of chemical composition with temperature at a given pressure.

[26] Average proportions obtained for each run are reported as a function of pressure in Figure 5. At the lowest

pressure, around 30 GPa, the major phase is Ca-perovskite, followed by Mg-perovskite and stishovite, the NAL phase, and finally the CF-type phase. At 40 GPa, Mg-perovskite is the most abundant phase, followed by Ca-perovskite, stishovite, CF-type phase, and finally the NAL phase. At higher pressure, this order does not change, but the NAL phase disappears. At pressures between 30 and 50 GPa, the variation in the proportions of the different phases, namely the decrease of the NAL phase and the increase of Mg-perovskite content can be explained as a response to the changes in phase compositions (i.e., the aluminum solubility increase in Mg-perovskite). At pressures of 50 and 80–88 GPa, the proportion of each phase remains constant within experimental error. In run MORB#6 at 60 GPa, the high intensity of neon peaks in the XRD patterns precluded refinement of phase proportions. Because no significant mineralogical changes were observed from 50 to 89 GPa, proportions obtained at 50 GPa (MORB#5) were used to obtain the densities for run MORB#6 at 60 GPa.

3.5. MORB Density

[27] We obtained experimentally the unit-cell volumes, proportions, and compositions of each phase to be able to calculate the density of the whole assemblage at each pressure-temperature condition investigated. As already described, our chemical compositions show a loss of iron, which is an element of major importance for density calculation. Since we observe this iron loss commensurate with magnesium enrichment, we decided to apply a correction involving adding iron and subtracting magnesium from our phases. Mg-perovskite, CF-type phase, and NAL phase are the major hosts of Fe and Mg. To be able to reduce iron loss and magnesium enrichment to less than 5%, we added 3, 4.6, and 6 wt% FeO accompanied by the same subtraction of MgO in Mg-perovskite for MORB#1, MORB#4, and MORB#5, respectively. By keeping the partition coefficient

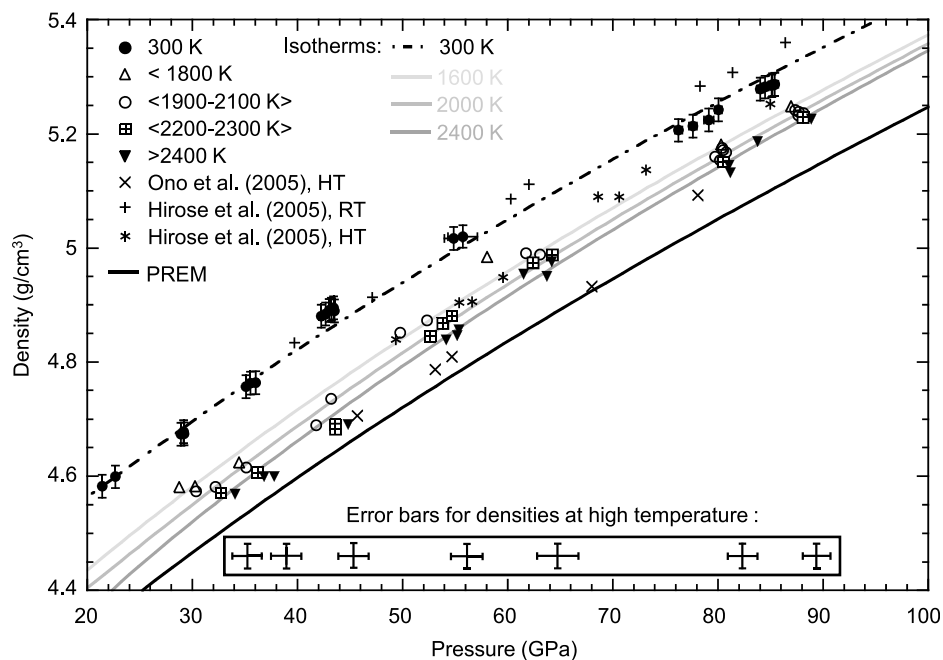


Figure 6. Density of MORB at high-pressure high-temperature compared to PREM. Symbols represent densities calculated in the present study using data obtained at 300 K (solid circles), less than 1800 K (open triangles), between 1900 K and 2100 K (open circles), between 2200 K and 2300 K (cross squares), and to 2600 K (inverse solid triangles). Calculated isotherms at 1600 K, 2000 K, and 2400 K with our thermoelastic parameters from Table 4 are shown. Density data obtained by *Ono et al.* [2005], and by *Hirose et al.* [2005] at room temperature and at high temperature are indicated by \times , $+$, and stars, respectively.

of Fe and Mg between these phases unchanged (within 15% of the initial value) and maintaining the mass balance requirement, we added 1.4, 3, 3.3 wt% FeO in the CF-type phase, respectively, and 2 and 2.5 wt% FeO in NAL phase for MORB#1 and MORB#4, respectively, accompanied by the same subtraction of magnesium. Table 3 shows the weight obtained for each phase before and after this correction.

[28] The unit-cell volumes, proportions, and corrected compositions of each phase were used to calculate the density of the whole assemblage at each pressure-temperature condition investigated. The density values are displayed as a function of pressure and temperature in Figure 6, along the PREM density profile [Dziewonski and Anderson, 1981]. The density values from *Ono et al.* [2005] and *Hirose et al.* [2005] using similar starting compositions are also plotted for comparison. Below 60 GPa, our result is similar to the result of *Hirose et al.* [2005]. The densities determined by *Ono et al.* [2005] are lower than ours, and those determined by *Hirose et al.* [2005] are higher after 60 GPa. The discrepancy observed at pressures above 60 GPa might be due to the difference observed in the compression curve of Mg-perovskite from the two studies (Figure 4a) and the difference in pressure scale. The discrepancy between our result and those of *Ono et al.* [2005] is difficult to determine since they have few data points obtained exclusively at high temperatures. The difference might be due to a difference in pressure scale (i.e., they use gold as a pressure calibrant with the pressure scale of *Jamieson et al.* [1982]) and a difference in determined volumes of phases. Mass balance calculation

also shows a loss of iron of about 15% in *Ono et al.* [2005] and *Hirose et al.* [2005] studies. Correcting this loss would not change significantly the observed discrepancy.

[29] We used our zero-pressure densities and the one given by *Ono et al.* [2001] combined with our high-pressure data set at room temperature and at high temperature to determine the thermoelastic parameters using a generalized inversion as describe in previous section. Zero pressure density was obtained using X-ray diffraction data from samples MORB#3 and MORB#5 quenched to ambient conditions (see Table 2). Since Ca-perovskite is amorphous at ambient pressure, we used the value of $V_0 = 45.58 \text{ \AA}^3$ from *Wang et al.* [1996] to obtain its room conditions density. Thermoelastic parameters were obtained by fitting our data set with the pressure derivative of the bulk modulus, K' , fixed to 4. The set of parameters obtained is reported in Table 4. The influence of the disappearance of the NAL phase on the equation-of-state parameters was checked by fitting the densities using only the data obtained above 50 GPa; the zero pressure densities are also used to constrain the fit. The fitting results, listed in Table 4, showed that the disappearance of the NAL phase does not significantly affect the EoS parameters of MORB. This is rather different from the effect of the NAL disappearance on the volume variation of each individual phase, especially for the Mg-perovskite EoS. This difference in behavior of the bulk density EoS and the EoS of each individual phase can be explained by the compositional variations of phases. Indeed, aluminum-bearing Mg-perovskite contains more aluminum

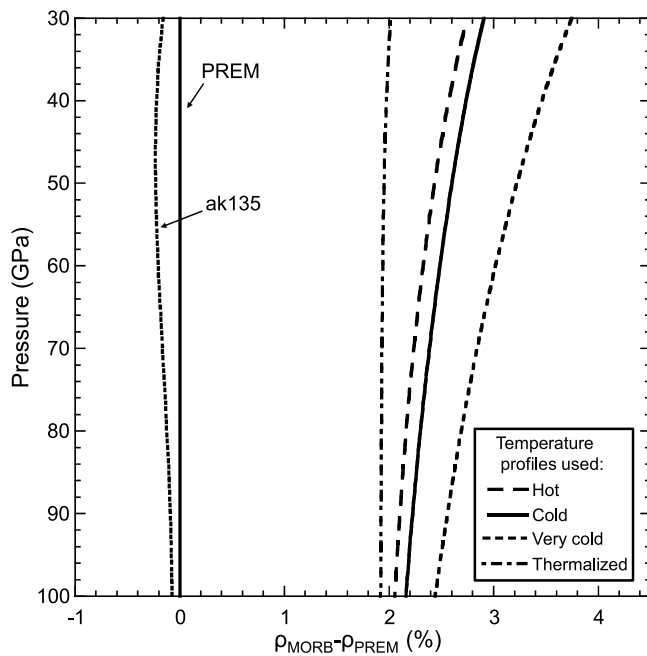


Figure 7. Density differentials of MORB samples compared to PREM calculated along the temperature profiles of a cold and a hot slab (solid and long-dashed lines), and calculated along the temperature profile of a very cold slab given in the literature and the geotherm of *Brown and Shankland* [1981] (short-dashed and dot-dashed lines). The density differential between the densities given by ak135 [Kennett *et al.*, 1995] minus densities of PREM are reported for reference.

after the NAL phase has disappeared. The changes in composition and volume compression of Mg-perovskite seem to compensate any potential density jump.

[30] The density profiles of MORB along isotherms of 1600 K to 2400 K calculated using the EoS parameters are shown in Figure 6. With the EoS parameters, we calculated the density profiles of MORB along different temperature profiles. First, we used two thermal profiles corresponding to a hot and a cold slab as calculated by *Ricard et al.* [2005] for an oceanic lithosphere with velocities of 3 and 10 cm/yr, respectively, which includes the most current range of plate velocities [Gripp and Gordon, 2002]. Moreover, temperature profiles of the subducting slabs have been calculated using different thermal models [e.g., Meade and Jeanloz, 1991; Kirby *et al.*, 1996]. These models gave very similar temperature profile for a hot slab. For a cold slab, the previous models predicted much lower temperature, as low as 900 K at 670 km depth. We calculated the density profile of MORB with this very low temperature profile as well. We also calculated the density of MORB in a subducting slab thermally equilibrated with the surrounding mantle. The adiabatic geotherm from *Brown and Shankland* [1981] was used. Figure 7 shows the density differentials between MORB and the surrounding mantle with the MORB density profiles calculated using slab temperature profiles for a cold and a hot slab and using the very cold temperature profile given in the literature and the geotherm of *Brown and Shankland* [1981]. Depending

on the temperature profile, the density differential is between 2% and almost 4% at depths beyond the transition zone and near 2% at the bottom of the lower mantle.

[31] For a better comparison, the difference between the two average density profiles of the Earth's mantle (i.e., PREM and ak135) [Kennett *et al.*, 1995] is shown in Figure 7. Nevertheless, the error bar given for PREM is $\pm 0.05\%$ [Silver *et al.*, 1988], which is larger than the difference between the PREM and ak135 density profiles. The conclusion does not change by taking into account the error bar.

[32] Since we obtained the EoS of each phase present after 50 GPa, we were able to obtain the density profile of each individual phase and to calculate the MORB density profile by using the phase proportions. The MORB density profile calculated using the thermoelastic parameters obtained on each phase is very close to the density profile obtained by using the elastic parameters. These two profiles are plotted in Figure 8, using shaded and solid lines, respectively. This gives an error on our profiles of about 0.2%. We attempted a possible sodium correction on our compositions and ended up with a difference of about 0.1% on the density profile. This makes a final error of about 0.3% on the density differential. Thus, in all cases, MORB is always denser than

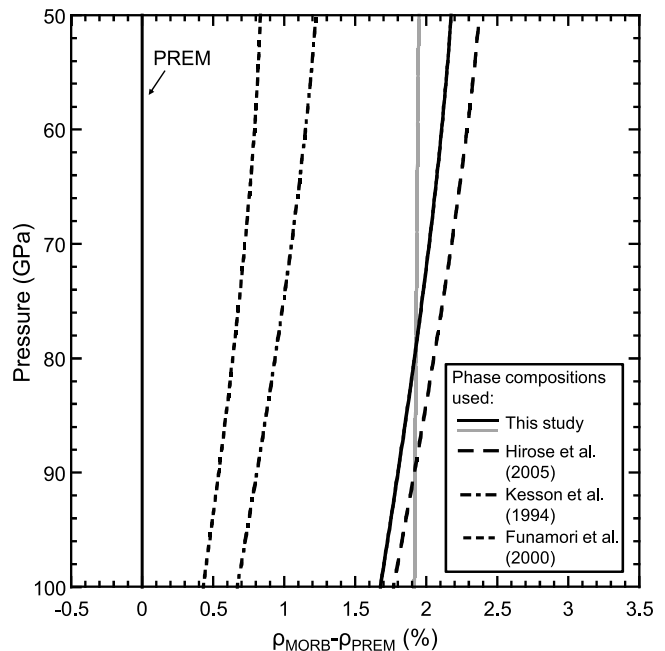


Figure 8. Density differential of MORB sample calculated using different phase compositions from the literature compared to PREM along the temperature profile of *Brown and Shankland* [1981] (geotherm). Density profile has been obtained by using EoS parameters of each individual phase (Table 4). Profile obtained by using the EoS of the entire MORB assembly is plotted for comparison (shaded line) along with the profile calculated using the EoS of each phase (solid line). Chemical compositions from the study of *Hirose et al.* [2005], *Kesson et al.* [1994], *Funamori et al.* [2000] are used to calculate the density differential shown as a dashed line, dot-dashed line, and dotted line, respectively.

the surrounding mantle up to the bottom of the lower mantle (Figure 7).

4. Discussion and Conclusions

[33] The density profile of MORB calculated along different temperature profiles demonstrates that the crustal part of the slab is denser than the surrounding mantle in the Earth's lower mantle and thus can reach the bottom of the mantle. The present results confirm the conclusions drawn by *Ono et al.* [2005] and *Hirose et al.* [2005] that MORB may reach the core–mantle boundary.

[34] For a thermally equilibrated slab, the MORB's density is also higher than average mantle density. Figure 7 shows the maximum density difference due to thermal buoyancy by comparing a thermalized crust and a very cold crust. The difference is almost 2% at the top of the lower mantle and decreases to about 0.5% at the bottom of the mantle.

[35] For a depth between 1200 and 1800 km, *Ono et al.* [2005] calculated a density difference between the slab and the average mantle of about $+0.05 \text{ g/cm}^3$ ($\sim 1\%$). From our calculated density profile, at 1000 km depth, the density difference is about +3% and decreases to about +2% at ~ 2000 km depth for a nonthermalized crust (see Figure 7). In the case of a very cold slab, the value is as high as 3.5% at 1000 km depth. At the bottom of the lower mantle, the density differential is about 2% for every temperature profiles used. The slope of the density profile is different due to the different temperature gradients of the thermal profiles used. The density difference obtained for a thermalized crust is almost constant with depth.

[36] We did not extrapolate our results beyond 100 GPa because our experimental data were acquired up to 90 GPa; after 100 GPa, two phase transitions take place. Around 110 GPa, the transition of perovskite to post-perovskite, a higher density phase, occurs in basaltic compositions [*Hirose et al.*, 2005; *Hirose*, 2006]. As shown by *Murakami et al.* [2004] and *Ono and Oganov* [2005], the density increase in Mg-perovskite associated with this transition is most likely to be up to 1%. As the proportion of magnesium perovskite in MORB is about 40%, we would obtain a total increase in MORB density of about 0.4%. The transition into the α - PbO_2 structure for SiO_2 [*Murakami et al.*, 2003] has also been observed in MORB assemblage after 100 GPa [*Hirose et al.*, 2005].

[37] The slab negative buoyancy could also be affected by the compositions of the oceanic crust, in particular, the iron content and the chemical composition of high-pressure phases. To observe this effect, we plotted the density differential of MORB profile by using the EoS parameters of each phase with different sets of phase compositions from the literature. We used the composition given by *Hirose et al.* [2005], *Kesson et al.* [1994], and *Funamori et al.* [2000], obtained on quench samples synthesized at 60, 45, and 43 GPa, respectively, where the NAL phase is absent. Figure 8 shows that the density profile of MORB depends mostly on phase compositions. The profile obtained with compositions from *Hirose et al.* [2005] is very similar to the one obtained with the compositions from *Ono et al.* [2005] obtained at 30–40 GPa. The density profile obtained with the compositions

from *Kesson et al.* [1994] and *Funamori et al.* [2000] are lower. This is mainly due to a different repartition of iron and aluminum between Mg-perovskite and the CF-type phase. The Al_2O_3 content of Mg-perovskite and the CF-type phase is higher and the content of iron smaller in studies of *Funamori et al.* [2000] and *Kesson et al.* [1994] compared to the study of *Hirose et al.* [2005]. In every study, proportions of phase have been obtained by mass balance calculation, which shows an iron loss of about 15%. Nevertheless, *Hirose et al.* [2005] used a starting composition containing 11.4 wt% FeO and the other studies used one with 7.8 wt% FeO (see Table 1). The density differential obtained with the compositions of *Funamori et al.* [2000] is very similar to ours before applying iron correction. The FeO content in natural MORB varies from ~ 6.5 wt% up to ~ 11.5 wt% [*Klein and Langmuir*, 1987]. Overall, Figure 8 reflects the density differentials for the range of MORB composition, our profile depicting an average MORB composition with 9.7 wt% FeO. Thus, depending on the MORB composition used, the density differential in the lower mantle is about 2%–3% down to 0.5%–1%.

[38] Geophysical studies of the fate of the subducting slabs in the deep mantle are not entirely consistent among different studies. *Van Der Hilst and Kárason* [1999] observed no tomographic anomaly at depths greater than 1800 km. This leads us to propose that slabs could lose their excess negative (thermal) buoyancy and be assimilated in the middle of the lower mantle at a depth of ~ 1800 km. Other studies using numerical convection modeling [*Christensen and Hofmann*, 1994; *Coltice and Ricard*, 1999] suggested a segregation of subducted crust at the bottom of the mantle, with the crustal part denser than the surrounding mantle and a possible re-entrainment in mantle plumes. Our study shows that for the crustal part of the lithosphere, the density is higher than that of the lower mantle up to the bottom of the lower mantle even if it is at the same temperature as the surrounding mantle, which would support such models.

[39] To develop a realistic model of slab behavior, it is important to consider the lithospheric (harzburgite or lherzolite) component of the slab. *Irifune and Ringwood* [1987] studied a harzburgitic composition to the transition zone pressure, and showed that its density profile is similar to that of the mantle. Recently, *Ricard et al.* [2005] took into account the thermal contribution and calculated density profiles of MORB, harzburgite, and pyrolite along the calculated temperature profiles. They computed the equilibrium mineralogy of mantle assemblages by Gibbs energy minimization using the thermodynamic parameters of end-members. They showed that the pyrolite and harzburgite profiles converge at the CMB with the harzburgite density profile still higher than that of pyrolite. However, their results are limited by the available thermodynamic parameters of high-pressure phases especially for the Al-rich phase; they used the corundum phase due to the lack of equation-of-state parameters for the high-pressure Al-rich phase.

[40] Finally, most of the previous density profile comparisons have been made between experimental data and seismological models, which do not have the same type of error and imprecision. Thus, for a better comparison, the same kind of experiments could be extended to average mantle composition such as the pyrolitic or chondritic model

as well as to compositions representative of sediments and the lithospheric part. *Ricolleau et al.* [2009] determined the density profile of a pyrolite lower mantle along the geotherm of *Brown and Shankland* [1981]. They showed that the pyrolite density profile is about 1.3% lower than the PREM density profile at the 660 km depth and 0.4% higher at the bottom of the lower mantle. Comparing the density profile of pyrolite with the MORB density profile will lead to a different differential density profile between MORB and the surrounding mantle, but the crustal part of a slab will still be denser than a pyrolytic lower mantle.

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