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Cytosine - cytosinium dimer behavior in a co-crystal with a decavanadate anion as a function of the temperature

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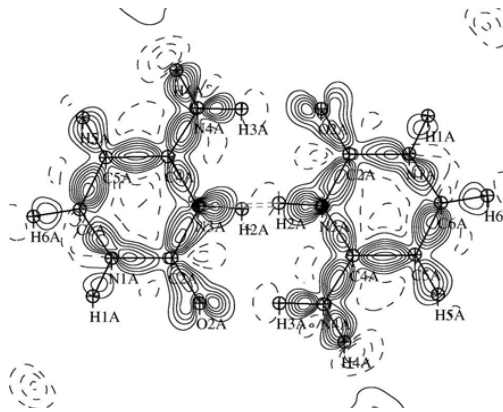
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Abstract

We have performed X-ray diffraction measurements on single crystals of $\text{Na}_3 [\text{V}_{10}\text{O}_{28}] (\text{C}_4\text{N}_3\text{OH}_5)_3 (\text{C}_4\text{N}_3\text{OH}_6)_3 \cdot 10 \text{H}_2\text{O}$ as a function of the temperature. On cooling, from room temperature to 100 K, we have observed additional peaks well indexed in $P1$, while the phase at room temperature crystallizes in $P\bar{1}$. The molecular structure at 210 K indicates that the center of inversion is located between two cytosine molecules, formally described with a charge of +0.5. When heating this crystal to room temperature, and reindexed the structure in $P\bar{1}$, some peaks remained unindexed. A protonation - deprotonation process gives rise to additional diffraction peaks at lower temperature than 210 K. The triply bridged hydrogen bond cytosine - cytosinium dimer is discussed according to the results of the charge density analysis and topological analysis at 210 K. The structure at 100 K has been completely solved based on a comparative study with other compounds containing cytosine-cytosinium dimer. This description could be considered as a reference for such dimer. It could help for discrimination between cytosine and cytosinium molecule for any new structure containing a cytosine-cytosinium pair and for which the quality doesn't allow a precise determination of the hydrogen localization.

Graphical abstract

We have performed X-ray diffraction measurements on single crystals of $\text{Na}_3 [\text{V}_{10}\text{O}_{28}] (\text{C}_4\text{N}_3\text{OH}_5)_3 (\text{C}_4\text{N}_3\text{OH}_6)_3 \cdot 10 \text{H}_2\text{O}$ as a function of the temperature. A protonation - deprotonation process gives rise to additional diffraction peaks at lower temperature than 210 K. The triply bridged hydrogen bond cytosine - cytosinium dimer is discussed according to the results of the charge density analysis and topological analysis.



Les paramètres requis sont manquants ou erronés.

Introduction

Nucleobases are the parts of deoxyribonucleic acid (DNA) and ribonucleic acid (RNA) that may be involved in pairing. The main ones are cytosine, guanine, adenine (DNA and RNA), thymine (DNA) and uracil (RNA). Nucleobases possess multi hydrogen bonded sites and various tautomers such that they can form abundance of aggregates through hydrogen bonds. Hydrogen bonding between naturally occurring nucleic acid bases of the purine and pyrimidine type is of great biological importance for secondary structures of DNA and RNA [1-3]. Nucleobases exist in solution as equilibrium of two or more tautomers. The relative stability of these tautomers is important for the structure and functioning of DNA and RNA. Specifically, the tautomeric equilibrium between the keto-amino and enol forms of cytosine has been investigated in aqueous solution (scheme 1). Experimental [4-6] and theoretical [7-14] studies have shown that cytosine exists primarily in the two most stable tautomeric forms, i.e. the aminooxo (canonical) and aminohydroxo forms. Aleman indicates that cytosinium cation is less stable than cytosine using *ab initio* calculations [15]. Using UV-UV and IR-UV double resonance spectroscopy Nir *et al* have reported the first spectroscopic data for dimer [16]. Quantum chemical calculations [17-19] of the structure and stability of dimer are reported. From a crystallographic point of view, examination of the Cambridge Structure Database (CSD) [20] give 73 compounds containing cytosine or cytosinium (scheme 1). For 13 compounds, the solid state structures present a cytosinium - cytosine entity is linked with triply hydrogen bonded dimer [21].

Many poloxometalate (POM's) are biologically active. POM's interest recently increased due to their medicinal (antiviral and antitumoral activities) [22-24]. Polyoxovanadate (POV's) interact with biomolecules with various and versatile activity (enzymes inhibitor or activator) [25-26]. Decavanadate is formally a $[V_{10}O_{28}]^{6-}$ polyanion with a cage structure giving rise to

fascinating supramolecular architectures in the solid state [27-28]. Since the biological properties result from interactions with viral enzymes or with viral cell envelope, the understanding of these interactions at a molecular level is essential for the interpretation and the development of potent compounds with selective enzymatic affinity. In order to get a better understanding on the biochemical interactions of such polyanions, we have synthesized a co-crystal of decavanadate with cytosine. The charge density of $\text{Na}_3 [\text{V}_{10}\text{O}_{28}] (\text{C}_4\text{N}_3\text{OH}_5)_3(\text{C}_4\text{N}_3\text{OH}_6)_3 \cdot 10 \text{H}_2\text{O}$ at 210 K has been determined and the interactions between the $[\text{V}_{10}\text{O}_{28}]^{6-}$ anion and its environment have been discussed [29]. In this paper, we will describe the structural behavior of this compound as function of the temperature and focus the discussion on the cytosine - cytosinium dimer according to the result of the charge density at 210 K which have not been developed in the first paper [29] and the new X-ray crystal structure as determined at 100 K.

Computational method.

Theoretical calculations have been performed with DFT calculation methods using the Gaussian 03 program package [30]. The coordinates used in calculations are those obtained from the multipole refinement [29]. The calculation of the electrostatic potential have been performed in a vacuum using B3LYP functional with 6-31++G(2d,2p).

Experimental

Crystallographic measurements were carried out on Bruker-SMART three-circle diffractometer equipped with a CCD area detector using $\text{MoK}\alpha$ radiation ($\alpha = 0.71073 \text{ \AA}$) from 300 K to 100 K, and summarized in Table 1. The absorption correction was applied

using SADABS [31]. Data were sorted and merged using SORTAV [32]. The structure was solved by SIR92 [33]. $\text{Na}_3 [\text{V}_{10}\text{O}_{28}] (\text{C}_4\text{N}_3\text{OH}_5)_3 (\text{C}_4\text{N}_3\text{OH}_6)_3 \cdot 10\text{H}_2\text{O}$ crystallizes in the centrosymmetric space group $\text{P}\bar{1}$ at temperature greater than 180 K. All structural calculations (Table 2) were performed by using the SHELXL-97 [34] program of the WINGX software package [35]. Thermal ellipsoid plots were obtained using the program ORTEP3 [36]. Synthesis, crystalization, data collection, data reduction and refinements of the structure at 210 K have already been described [29]. The methodology used for this structure and some pertinent parameters are given in footnote [37].

Results and discussion

Diffraction behaviour in cooling and heating

The following comments underline some experimental steps described in Table 1.

We have tried two ways for cooling the single crystal until 100 K. The first way was an ultra fast freezing, the second way was a very slow cooling speed of 10 K/h, from 300 K to 100 K. Before first cooling a room temperature measurement has been done. The experimental conditions and refinement parameters are summarized in table 2. Surprisingly, for the two processes of cooling, the determination of the orientation matrix at 100 K and the unit cell, gives the same $\text{P}\bar{1}$ space group as at room temperature, but some diffraction peaks remain unindexed in that space group (figure 1d). They were indexed by doubling the unit cell volume and indexing the crystal in the $\text{P}1$ space group. We made a complete data collection at 100 K, 1.5 hour after the stabilization of the temperature. More suprisingly, when we returned at room temperature, and reindexed the crystal in the previous space group and orientation matrix determined before the cooling, some peaks remained unindexed (figure 1b). We decided to go back slowly at 210 K and performed a high resolution data collection for a

charge density study, where we have still not observed additional peaks during the first cooling. This temperature seems sufficiently low to minimize the effect of the temperature on the U_{ij} parameters and sufficiently high to avoid problems of additional peaks.

The refinements of the structure have been done for the three temperatures in the $P\bar{1}$ space group at 300 K and 210 K and in the P1 space group at 100 K (table 2). At 100 K in the non centrosymmetric space group the center of inversion is suppressed, leading to an asymmetric unit including 12 molecules of cytosine, 2 molecules of decavanadate, 6 atoms of sodium, and 40 water molecules. Due to the high number of parameters, we were not able to localize hydrogen atoms from the Fourier difference maps. At 210 K, a multipole refinement according to Hansen-Coppens model [38] has been done on crystal 2 and crystal 3 [37]. The decavanadate environment has been described at 210 K [29]. In order to interpret the change between the high and the low temperature phase, we have to go back deeply in the structure at 100 K and 210 K focusing on the dimer behaviour.

Molecular structure, crystal packing, deformation maps, topology of the total density, net atomic charges and electrostatic potential at 210 K of the cytosine – cytosinium dimers

The asymmetric unit of the title compound contains three different molecules of cytosine. The cytosine C is not protonated and is neutral. The cytosinium B is protonated, and hydrogen atom H2b is bonded to N3b leading to charge of +1 for the molecule. The cytosinium A is protonated, a half of hydrogen atoms H2a is bonded to N3a, with charge +0.5. This hydrogen lies on the crystallographic center of symmetry (Figure 2). Some interatomic distances in cytosine C (C2c-O2c, C6c-N1c) deviate significantly from the corresponding distances in cytosinium A and cytosinium B. Cytosine C presents high U_{ij} values, mainly located on N1c and C6c. These high U_{ij} have been observed on the two crystal for which high resolution X-ray diffraction have been performed. The distances for protonated cytosinium B (with H atom

on N3b) has been compared with corresponding molecules of cytosinium, retrieved from CSD [20]; 31 data extracted from 25 different structures containing cytosinium molecule, with agreement factor less than 0.05 [43]. The same study has been done for cytosine C (without H atom on N3c); 5 data extracted from 5 different cytosine structures, with agreement factor less than 0.05 [44] (table 3). Distances for cytosine A have not been taken into account, the symmetry center have implied that the structure have been modelized with half of a hydrogen atom on N3a. Table 3 shows that our distances are in good agreement with those from literature data. Most of them are in the range from minimum to maximum distance for selected bonds.

The geometry is quite similar to those found in previously reported structures of cytosine. As it shown, in previous studies [45-47], the protonation of the cytosine molecule induced a modification of the bond length. Not unexpectedly, the N3-C bonds of the protonated N3b nitrogen atom are slightly longer than those of the nitrogen atom N3c that acts as hydrogen acceptor. Accordingly, both the N3b-C4b and N3b-N2b bonds are longer than N3c-C4c and N3c-N2c (1.354(1) *versus* 1.351(1) Å and 1.365(1) Å *versus* 1.357 (1) Å), respectively. Furthermore, the angle C2-N3-C4 at protonated N3b is slightly greater (0.2°) than those of the non-protonated N3c (122.26° *versus* 122.01°). Thus, as expected, the difference between the angle in the protonated (cytosinium) and the non-protonated molecule (cytosine) are sensitive to the protonation ($\Delta_{B-\square C}(N3-C4-C5)=-1.5^\circ$) and $\Delta_{B-\square C}(N4-C4-C5) =+1.6^\circ$). All these experimental values exhibit the same trend founded in previous theoretical DFT study [46] and experimental study works [45, 47].

Two molecules of cytosine A (figure 2a) are connected over the three strong hydrogen bonds (Table 4) forming a dimeric cytosinium – cytosine cation, namely two terminal N4-H...O (N4a...O2a⁽ⁱ⁾ =2.841(1)Å and N4a-H3a-O2a⁽ⁱ⁾ =177.12(6)°) and a central N3-H...N3 (N3a...N3a⁽ⁱ⁾ =2.838(2)Å and N3a-H2a-N3a⁽ⁱ⁾ =167.72(2)° (figure 2a). In this dimeric cation,

the terminal N4a-H3a...O2a hydrogen bonds are identical. The hydrogen atom of the central N3a-H2a-N3a hydrogen bond is disordered over two equally occupied positions related by an inversion center. Two molecules of cytosinium B and C are connected via three strong hydrogen bonds. The central hydrogen bond is N3b...N3c ($N3b...N3c^{(ii)} = 2.816(1) \text{ \AA}$ and $N3b-H2b-N3c^{(ii)} = 173.39(5)^\circ$) (figure 2b). The two terminal N4-H3...O2 hydrogen bonds are different, a weaker ($N4b...O2c^{(ii)} = 2.874(2) \text{ \AA}$) and a stronger one $N4c...O2b^{(ii)} = 2.756(1) \text{ \AA}$). This is expected because protonation implies to different bond distances and bond angles (table 3). We have observed a difference in the hydrogen bond in cytosine A, B and C, in the N1 - C6 region. The hydrogen bonds which exist in cytosine C ($C5c-H5c...O22^{(iii)}$, $N1c-H1c...O2w$) are longer than those present in cytosinium B and A ($H...O = 2.6819(5)$ *versus* $2.0702(6)$ and $H...O = 2.2430(7)$ and $2.288(1)$ *versus* $1.7519(7)$ and $1.7006(7)$), respectively. Hydrogen bonds which exist in cytosine C ($C5c-H5c...O22^{(iii)}$, $N1c-H1c...O2w$) are longer than in cytosine B and C (table 4). The summation of $d(H...A)$ interatomic distances for each cytosine, gives a quantification of the strength of the hydrogen bond network and clearly indicates that cytosine C is not stabilized with strong hydrogen bond, compared to the two other cytosine molecules. This in agreement with the thermal parameters, which are greater for cytosine C and then coherent with the fact that these high U_{ij} were observed for multipole refinement at 210 K of crystal 2 and 3. The hemiprotonated cytosine dimers with triple hydrogen bonds reported in CSD [21] present the both cytosine-cytosinium entities. The B-C type exists is seven structure [21a,b,d,e,f,g,h] while the A-A type has been observed in six structures [21c,e,i,j,k].

The deformation density of the cytosine (figure 3) illustrates the covalent bonding between adjacent carbon atoms [37] with the maxima of $0.6 - 0.9 \text{ e\AA}^{-3}$, or carbon and nitrogen atoms, with the maxima of $0.5 - 0.9 \text{ e\AA}^{-3}$. These features exhibit pronounced electron density accumulations, usually attributed to covalent bonding. Figures 3 show large electron density

peaks between atoms and those corresponding to lone pair density peaks of O2 and N3c atoms. If we compare N3a, N3b and N3c, static density maps indicate very distinctly that N3a and N3b are connected to the H2 hydrogen atoms. We can also see the density peaks, which are significantly displaced toward the N3, due to the electronegativity of this atom. In the vicinity of the N3c atom (figure 3), the nonbonding density corresponds to the lone pair. One can compare these features with previous experimental or theoretical deformation density. The same behaviour around the N3 atom have been found in the previous theoretical 4-31G** calculation [48] and experimental results [49] or in experimental and theoretical 6-31G** study on a cytosine monohydrate [50]. Eisenstein [49] was the first who reports an experimental charge density study on the cytosine monohydrate in 1988. She did an *ab initio* theoretical calculation with 4-31G and 4-31G** bases because previous STO-3G study [51] was not satisfactory. She has shown that charge density maps from 4-31G** results, is much closer to the experimental map, than to maps computed with lower bases. Our experimental static deformation density of cytosine fits very well with corresponding theoretical 4-31G** deformation density in positions, shapes and heights of most peaks (figure 3). Eisenstein has also reported that the difference between the experimental and 4-31G** maps occurs in non-bonding regions, where the experimental density peaks are lower. In both studies, it has been found a non-bonding density around the N3 and O2b atoms corresponding to their lone pairs. The N-H...N hydrogen bonds serve to link two cytosine molecules in order to form a dimer (figure 3). The nitrogen lone pair is clearly engaged in the hydrogen bond in the case of the B-C dimer, while the A-A model, using half of an hydrogen does not allowed a real observation of the density of this bond.

The topological parameters of the electron density for the (3,-1) bond critical points (BCP) [37] in cytosine molecules are given in table 5. As expected, the BCP of the heteronuclear C-N and C-O bonds, is always located closer to the less electronegative carbon atom ($d_1 < d_2$).

The $\rho(r_{cp})$ and $\nabla^2\rho(r_{cp})$ values ($\nabla^2\rho(r_{cp}) < 0$ corresponding to a shared-shell interaction) at the bond critical point are characteristic of covalent bonds; a shorter bond corresponds to a higher density and a more negative Laplacian. The highest absolute value of Laplacian is found for the C-O bond, which corresponds the shortest distance. The bond belonging to the ring have higher electron density at the bond critical point with an average value of $2.1 \text{ e}\text{\AA}^{-3}$, while ellipticities at the BCPs vary from 0.06-0.42. These values indicate the π character of these bonds. The $\rho(r_c)$ values range from $1.78 \text{ e}\text{\AA}^{-3}$ for C-C bond to the relative high value of $2.84 \text{ e}\text{\AA}^{-3}$ for C-O. Inspection of table 6 indicates that the two experimental works [51] and theoretical using 6-31G** basis set [50] give topological parameters of the covalent bonds in the cytosine molecule which generally agree well. The agreement is even quantitative.

The κ coefficients and experimental atomic net charges are listed in table 7. We have found a negative charge for all O atoms ranging from $-0.23(6) e$ to $-0.46(7) e$ and a κ parameter going from $1.01 e$ (slightly contracted electron density) to $0.948 e$ (slightly expanded electron density). All N atoms carry a negative charge ranging from $-0.26(8) e$ to $-0.87(8) e$ and a positive charge on most of the carbon atoms, except for the C5 atom in all three cytosine. This carbon is between two carbons (C4 and C6), which are connected with the most negative nitrogen atoms. The H atoms are all positively charged. However, those attached to N or O atoms (H1, H2, H3, H4) possess a significantly more positive charge than the H atoms (H5 and H6) connected to a carbon atom. This trend is similar to results, which have been found previously reported in experimental studies [49, 52] and theoretical computations [48] for the cytosine molecule. However, the absolute values obtained experimentally for the cytosine molecule are generally smaller than the theoretical values. A large number of the differences between the experimental and theoretical results could be attributed to hydrogen bonding. Thus, the experimental charges of the hydrogen-bonding acceptors O2 are less negative than

theoretical values. The experimental charges for H3 and H4 are lower than the theoretical value as expected because they are involved in strong hydrogen bonds. For H5 and H6, which participate only in weak hydrogen bond, the difference between theory and experimental is smaller.

Figure 4 shows the theoretical and experimental electrostatic potential (EP) projected on the 3D isodensity surface (0.007 eA^{-3}) of the cytosine molecules and the two dimers. The most negative value is found in the vicinity of O2c atom (figure 4a). The deepest minimum of electrostatic potential is found close to the O2a oxygen atom (-0.002 eA^{-1}). The second deepest minimum is located in the vicinity of N3a (eA^{-1}). In both molecules (A and B), the vicinity of the hydrogen atom H4 and H5 display a minimum electrostatic potential value (figure 4b and 4c). For cytosine C the blue region correspond to 0.42 eA^{-1} and the red value to -0.45 eA^{-1} . The deepest minimum of EP is found close to the O2c oxygen atoms (-0.45 eA^{-1}) and around N3c (-0.28 eA^{-1}). It means, that there is an extensive electronegative region around O2 and N3 and an electropositive region around the N4 amino group and C5. Similar features were also observed [50, 52]. The EP projected on the cytosinium – cytosine dimer surface BC dimer is given in Figure 4d to 4g. The differences observed between the experimental and theoretical values are due to the crystal packing effect. The theoretical calculations have been performed in the gas phase while the experimental plot represent the results at the molecule level extracted from the crystal. The differences observed between figure 4f and 4g are due to the hydrogen bond of O2c connected to the water molecule O2c...H3wO3w which is not in the B-C plane. Such a dissymmetry has already been observed in quite similar crystal packing of a salicylaldehyde thiosemicarbazone compound [53]. The H...O interatomic values (figure 4f) are generally coherent with the EP, i.e. a highest positive EP value on the molecular surface corresponds to a shorter H...O distance except for the C6-H6...O.

Molecular structure at 100 K

The structure at 100 K, solved in the P1 space group, shows six dimers in the asymmetric unit instead of three at higher temperature. The difference between the low and the high temperature structure lie in the lost of the symmetry center between the two cytosine A, and the unit cell determine in P1 consists in a complicated transformation from the centrosymmetric space group. The quality of the structure at 100 K does not allow a determination of the hydrogen positions. Therefore, an indirect determination has been done in two steps.

The first step consists in an examination of the crystal packing in order to determine from which cytosine types (A, B or C) are originated the twelve cytosine molecules at low temperature. Table 8, summarizes in a simplified nomenclature the crystal packing of the structures at 210 and 100 K while figure 5 presents the cytosine – cytosinium interactions in the 6 dimers. In order to discriminate between the initial A-A or B-C dimers, one have to observed the O2 cytosine atom. In the case of an O2b or O2c the oxygen is linked to a water molecule, while the O2a takes place to the coordination polyhedra of the sodium atom (figure 5). The assignation of the original cytosine from the B-C dimer could be easily done by looking at the hydrogen bond connected to the C5 atom. A terminal decavanadate oxygen (O1x) is linked to the cytosine B while it is linked to a O2x oxygen type in the case of the cytosinium C. According to this table, we can indicate in the line “assignation” a letter indicating from which type of cytosine (A, B or C) the dimer corresponds in the high temperature structure.

The second step consists in the determination of the real type of the cytosine (B, C or A) at 100 K. That can be done through the observation of the angles around the C4 and the N3 atoms. The difference between the different angles from B-C is reported in Table 9. The sign

and the absolute value of these differences help in the discrimination between a cytosine C and a cytosinium B molecule, despite the fact that there is no hydrogen atom determination. In case of cytosine of A type (a proton in equilibrium between the two molecules of the dimer), one can assume that the differences has to be the lowest. These observations imply the assignation (A, B, C type) of the different cytosine molecules (last line of table 8). Some comments have to be done from observation of Table 9: i) the trend clearly indicates that the higher values of the parameter reported here concerned the pure cytosine-cytosinium dimer while the lowest values concerned the intermediate situation (A-A dimers); ii) the values obtained at 300 or 210K seems to be abnormally small, but it is due to a better determination of the nucleus position from a charge density study [55-56]. Therefore one can clearly assign the cytosine type to the twelve cytosine of the asymmetric unit (table 9). Four dimers conserve the same type from their originated situation at 210 K and the last two change.

Back to the interpretation of the behavior of the single crystal in heating and cooling under X-ray diffraction

The hydrogen atom is statistically equally shared between the two cytosine at high temperature inducing that the X-ray “see” in average half an hydrogen at the position determined at 300 K and 210 K. At these high temperatures a dynamic of proton exchange exists, one can say that the proton can “jump” between the two molecules forming for a considered molecule alternatively a cytosine and a cytosinium molecule leading to a symmetry center on the middle of the distance N...N (Figure 4d). At low temperature, this dynamic is blocked, and this proton is “freeze” on one cytosine, leading to the suppression of the inversion center and the doubling of the unit cell volume. It means that the X-ray “see” one hydrogen atom which is located on the one cytosine. The dimer type is conserved for four

dimers and changed for two. In heating the proton dynamic exists, but it is not any more perfectly statistically distributed, that means that the proton could not be modelized at the middle of the N...N distance. This is not-reversible dynamic cytosine - cytosinium conversion, which is responsible of the additional non-indexed peaks at room temperature after a cooling.

This conversion has been extensively studied due to the fact that the base-base interactions play an important role in the DNA structure. For example, *ab initio* calculations [19] of different cytosine dimers lead to an electronic dissociation energy of 15.7 kcal/mole (5510 cm⁻¹) by Nir while Sponer *et al* [17] has obtained 20.0 kcal/mole. The quantum chemical calculation on the DFT level of theory [46] show that the interaction energy was calculated to be 42.4 K cal/mole and revealed a high stability of the hydrogen bond in a methylcytosine – methylcytosinium dimer.

Conclusion

During the crystal cooling, below 200 K, we have observed a phase transition in the [V₁₀O₂₈] (C₄N₃OH₅)₃(C₄N₃OH₆)₃·10H₂O. According to the precise description of the cytosine – cytosinium dimer at 210 K, from charge density results, we have established that the main structural change was the suppression on the inversion center between the two molecules of cytosine. This is this not-reversible dynamic cytosine - cytosinium conversion, which is responsible of the additional non-indexed peaks at room temperature after a cooling. Besides this explanation concerning the title compound, Table 9, could be used as a guide for assignation of the proton localization in cytosine-cytosinium dimer, in the case where the hydrogen is not clearly detectable from the diffraction study.

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Supporting Information Available

Tables S1 and S2 are presented as Supplementary material. This information is available free of charge via the Internet at <http://pubs.acs.org>.

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theoretical deformation densities. The static density deformation was determined with the STATDENS program [39]. A quantitative investigation of the bonding was performed by means of a topological analysis according to AIM theory [40]. The definition and the classification of the chemical bond in the topological analysis of ρ are based on the existence of a (3,-1) critical point (CP) between neighboring atoms and the properties of charge density at this point. An important function of the electron density is Laplacian and can be obtained from the second derivatives of the total electron density. The topological analysis of the electron density has been carried out using the NEWPROP program [41, 42]. Pertinent refinement parameters (crystal 3. Number of reflexions used in SHELX, all data: 28 757, unique ($F_0 > 4\sigma F_0$): 20 587; Spherical refinement R(%): 2.88, R_w (%): 3.27, GOF: 1.28, parameters: 28, ; Number of reflexions used in MOLLY($I > 3\sigma(I)$): 18 394 ; Multipole refinement R(%): 2.32 / R_w (%): 2,34 / GOF: 0,91 / parameters: 505).

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Figure captions

Figure 1. a) step 1, 300 K, b) step 3, 300 K c) step 6, 180 K and d) step 7, 100 K (see table 1 for the step conditions).

Figure 2. Cytosine - cytosinium interactions at 210 K: a) between cytosinium A and $A^{(i)}$ and b) between cytosinium B and the cytosine C. $(i) = -x, -y+1, -z+2$

Figure 3. Static model deformation density map. a) Plane containing the A-A dimer and b) plane containing the B-C.

Figure 4. 3D isodensity surface (0.007 eA^{-3}) of the cytosine – cytosinium dimer colored by the electrostatic potential values. Scale in $\text{e}\cdot\text{\AA}^{-1}$ for theoretical plot and unique scale for experimental plots. a) experimental EP values on B cytosinium; b) experimental EP values on A cytosinium; c) experimental EP values on C cytosine; d) theoretical EP values on B-C dimer surface; e) experimental EP values on A-A dimer surface, the symmetry center is indicated with a black circle; f) and g) experimental EP values on B-C dimer surface (two side views). The first H...O distance in the B-C dimer is reported (in blue for B cytosinium, in green for C cytosine).

Figure 5. Cytosine - cytosinium interactions at 100 K: a) to f) six different dimers. In contrast of the structure at 210 K, the third character of the label of the atoms have been omitted and the numbering of the cytosine is indicated in the center of each molecules (from 1 to 12).

Table captions

Table 1. Experimental diffraction steps performed on single crystal of $\text{Na}_3 [\text{V}_{10}\text{O}_{28}] (\text{C}_4\text{N}_3\text{OH}_5)_3 (\text{C}_4\text{N}_3\text{OH}_6)_3 \cdot 10 \text{H}_2\text{O}$.

Table 2. X-ray diffraction experiments and data reduction details on crystal 2.

Table 3. Selected bond lengths (\AA) with estimated standard deviations in parentheses and comparison with the values retrieved in the CSD [20]

Table 4. Selected distances $d_{(\text{H-A})}$ (\AA) of the hydrogen bonds in the cytosine-decavanadate. Estimated standard deviations are given in parentheses. (i) = $-x,-y+1,-z+2$, (iii) = $x,+y,+z+1$, (iv) = $-x+2,-y,-z+1$, (vi) = $x+1,-y,-z+2$, (vii) = $-x+1,-y+1,-z+1$, (viii) = $x,+y,+z-1$

Table 5. Results of the topological analysis of charge density at (3,-1) bond critical point (BCP) of the molecules of cytosines. d_1 and d_2 are the BCP-atom distance (\AA), $\rho(r_c)$ is the electron density at the BCP ($\text{e}\text{\AA}^{-3}$), $\nabla^2 \rho(r_{\text{cp}})$ is Laplacian at the BCP ($\text{e}\text{\AA}^{-5}$), $\square e$ is ellipticity, $\lambda_1, \lambda_2, \lambda_3$, are eigenvalues at the BCP and H is the total energy density (au).

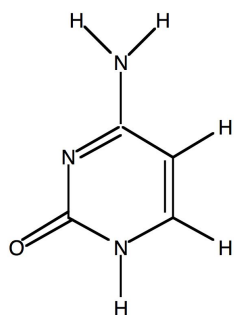
Table 6. Results of the topological analysis of charge density at (3,-1) bond critical point (BCP) of cytosine molecules.

Table 7. Experimental and theoretical atomic net charge (e) and κ values of cytosine molecules. Estimated standard deviations are given in parentheses. ¹ full multipole refinement from [52], ² monopole refinement from [52].

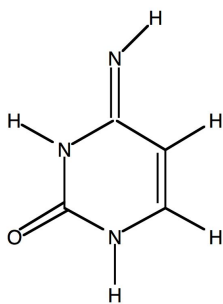
Table 8. Simplified crystal packing description of the dimers. The atoms used for the originated assignment of the cytosine linked to O2 or C5 are in bold. The final assignment (obtained from Table 9) is also reported in the last line according to the nX(Y) code. n = label of the cytosine molecule (from 1 to 12), X = final cytosine assignment (A, B or C), Y = originated cytosine assignment (A, B or C). The code is in bold when the originated and the final assignment is different.

Table 9. Pertinent bond angles (C2-N3-C4, N3-C4-C5, N4-C4-C5) (in °) used for discrimination between cytosine and cytosinium. From literature in the first part of the table, from the structure at 100 K for the second part. Δ is the bond angles difference in the dimer. The summation of the absolute values of these differences is also reported. The final assignment corresponding to a change from the high structure temperature is indicated in bold. ACITEM [21g], CYTZNC [21b], KESVAG [21b], METSI [56], QOCTUX [21d], GYTYEN [21c], LIWSOA [21j], 300 K (this work), 210 K (this work and [28]).

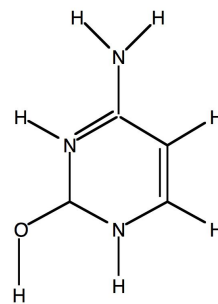
Scheme 1



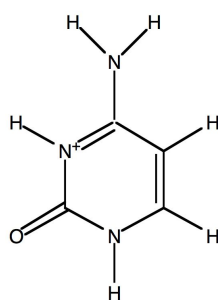
cytosine



aminohydroxo



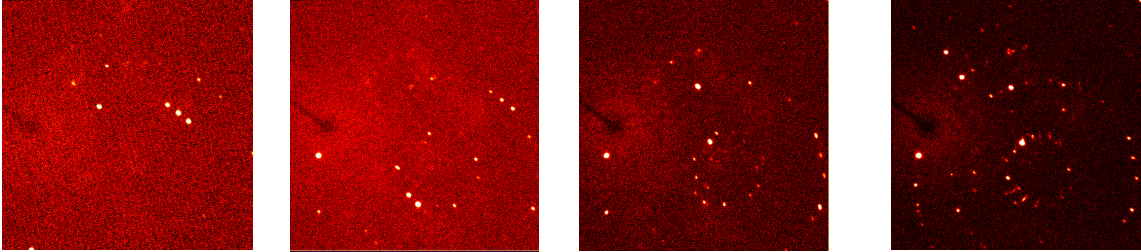
iminooxo



cytosinium

Figures

Figure 1



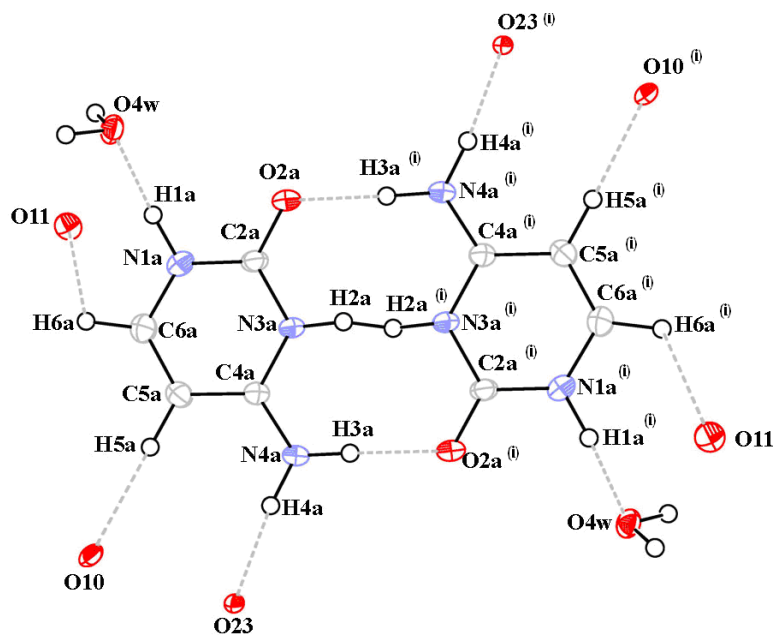
a)

b)

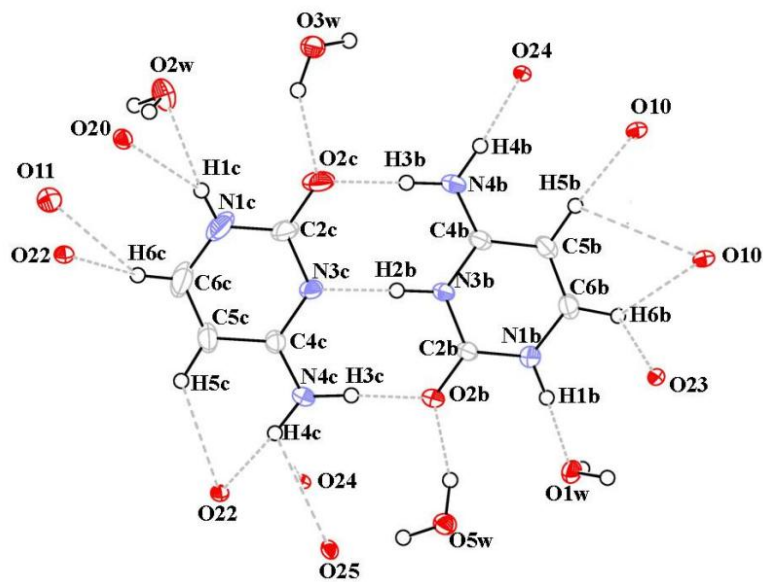
c)

d)

Figure 2

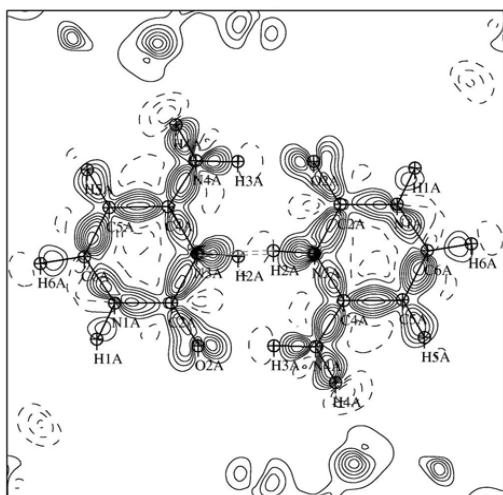


a)

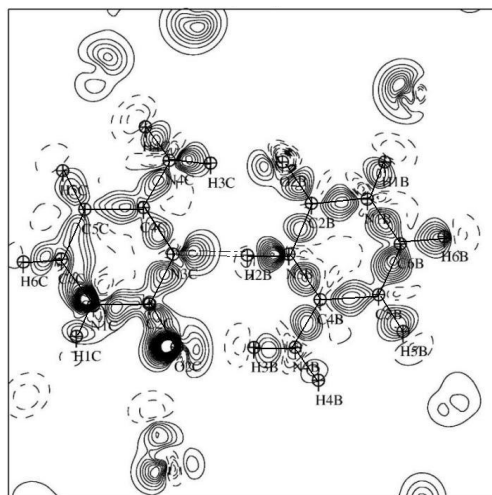


b)

Figure 3



a)



b)

Figure 4

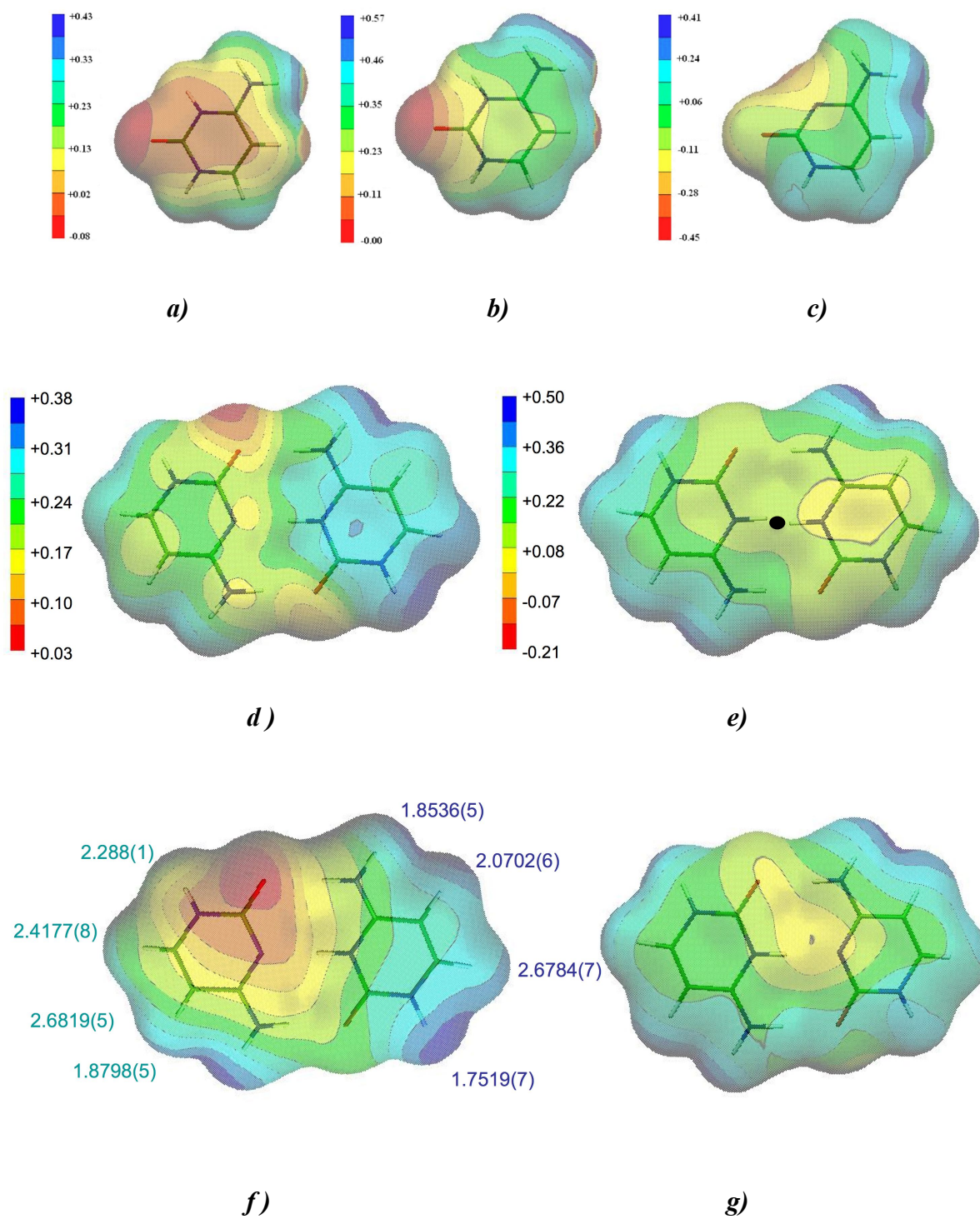


Figure 5

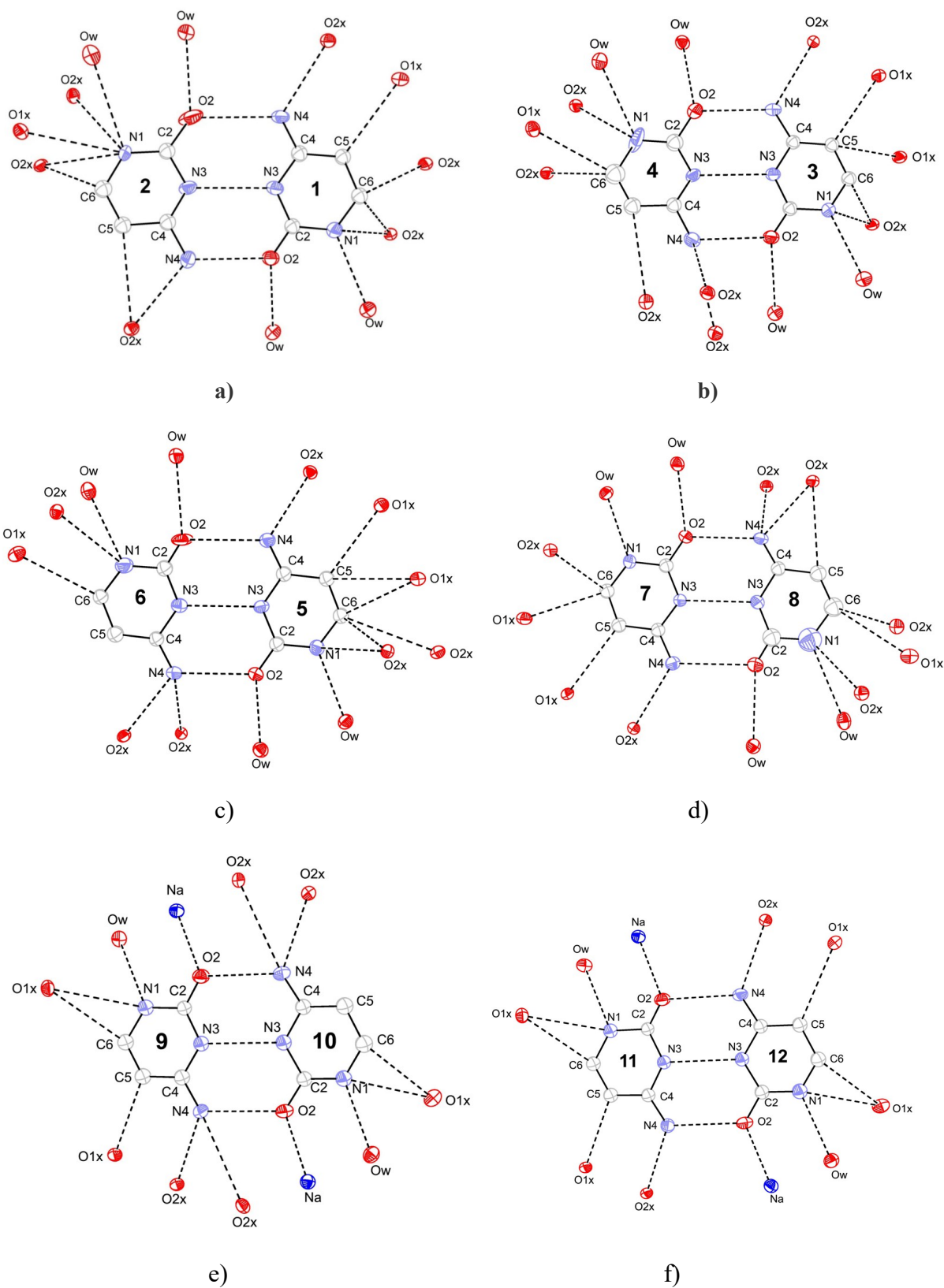


Table 1.

	Crystal	Actions	Observations
Step 1	Crystal 1	orientation matrix at room temperature	$P\bar{1}$, all peaks indexed
Step 2	Crystal 1	ultra freezing at 100 K	
Step 3	Crystal 1	orientation matrix at 100 K	$P\bar{1}$, some peaks non-indexed
Step 4	Crystal 2	orientation matrix at room temperature	$P\bar{1}$, all peaks indexed
Step 5	Crystal 2	data collection at room temperature	See Table 2
Step 6	Crystal 2	slow cooling (10 K/h), orientation matrix every 30 K	$P\bar{1}$, all peaks indexed until 210 K included, at lower temperature additional peaks non-indexed
Step 7	Crystal 2	orientation matrix at 100 K	$P1$ (unit cell x 2) all peaks indexed
Step 8	Crystal 2	data collection at 100 K	See Table 2
Step 9	Crystal 2	slow heating (10 K/h), orientation matrix every 30 K	$P\bar{1}$, some peaks non-indexed
Step 10	Crystal 2	slow cooling (10 K/h), orientation matrix every 30 K	$P\bar{1}$, some peaks non-indexed
Step 11	Crystal 2	data collection (high resolution) at 210 K	See Table 2
Step 12	Crystal 3	orientation matrix at room temperature	$P\bar{1}$, all peaks indexed
Step 13	Crystal 3	slow cooling (10 K/h), orientation matrix every 30 K, until 210 K	$P\bar{1}$, all peaks indexed
Step 13	Crystal 3	data collection (high resolution) at 210 K	See [29, 37]

Table 2

Chemical formula	Na ₃ [V ₁₀ O ₂₈](C ₄ N ₃ OH ₅) ₃ (C ₄ N ₃ OH ₆) ₃ · 10H ₂ O			
Temperature (K)	Air, 298 K	Nitrogen, 210 K	Nitrogen, 100 K	
Wavelength (Å)	0.71073			
Crystal size: (mm) x (mm) x (mm)	0.67 x 0.16 x 0.41			
a (Å)	9.9273(2)	9.8547(1)	12.2144(4)	
b (Å)	12.2332(3)	12.2177(1)	14.9129(5)	
c (Å)	13.1147(3)	13.0854(1)	17.6603(3)	
α (°)	85.968(1)	85.984(1)	106.636(1)	
β (°)	80.082(1)	80.099(1)	103.412(1)	
γ (°)	70.795(6)	70.928(1)	98.719(1)	
Volume(Å ³)	1481.41(6)	1466.68(1)	2914.99(6)	
[sinθ/λ] _{max} (Å ⁻¹)	0.87	1.00	0.87	
Scan width (°) / scan method	0.2 / Δω scan	0.15 / Δω scan	0.2 / Δω scan	
Crystal-to-detector distance (d) (cm)	4.075	4.075	4.075	
Total frames / time per frame (s)	7 200 / 5 to 10	43 200 / 3 to 20	5 400 / 10 to 30	
Absorption correction	SADABS	SADABS	SADABS	
Total measured reflexions	53 299	269 448	77 543	
R _{int} (F ²) after SADABS	0.028	0.031	0.056	
Cell setting, space group	Triclinic, P $\bar{1}$	Triclinic, P $\bar{1}$	Triclinic, P $\bar{1}$	Triclinic, P1
Z	1	1	1	2
Program / Number of reflexions used: all data / unique	SHELX / 15392 / 11198 / (F0 > 4sigF0)	MOLLY // 16983 (I > 3σ(I))	SHELX / 15155 / 13510 / (F0 > 4sigF0)	SHELX / 45601 / 25198 (F0 > 4sigF0)
Refinement type: R(%) / R _w (%) / GOF / parameters	Spherical / 5.41 / 8.52 / 1.030 / 556	Multipole / 3.63 / 3.46 / 1.20 / 505	Spherical / 4.73 / 9.78 / 1.240 / 552	Spherical / 10.02 / 13.18 / 1.097 / 1498
Chemical formula	Na ₃ [V ₁₀ O ₂₈](C ₄ N ₃ OH ₅) ₃ (C ₄ N ₃ OH ₆) ₃ · 10H ₂ O			
Temperature (K)	Air, 298 K	Nitrogen, 210 K	Nitrogen, 100 K	

Wavelength (Å)	0.71073			
Crystal size: (mm) x (mm) x (mm)	0.67 x 0.16 x 0.41			
a (Å)	9.9273(2)	9.8547(1)	12.2144(4)	
b (Å)	12.2332(3)	12.2177(1)	14.9129(5)	
c (Å)	13.1147(3)	13.0854(1)	17.6603(3)	
α (°)	85.968(1)	85.984(1)	106.636(1)	
β (°)	80.082(1)	80.099(1)	103.412(1)	
γ (°)	70.795(6)	70.928(1)	98.719(1)	
Volume(Å ³)	1481.41(6)	1466.68(1)	2914.99(6)	
$[\sin\theta/\lambda]_{\max}$ (Å ⁻¹)	0.87	1.00	0.87	
Scan width (°) / scan method	0.2 / $\Delta\omega$ scan	0.15 / $\Delta\omega$ scan	0.2 / $\Delta\omega$ scan	
Crystal-to-detector distance (d) (cm)	4.075	4.075	4.075	
Total frames / time per frame (s)	7 200 / 5 to 10	43 200 / 3 to 20	5 400 / 10 to 30	
Absorption correction	SADABS	SADABS	SADABS	
Total measured reflexions	53 299	269 448	77 543	
R _{int} (F ²) after SADABS	0.028	0.031	0.056	
Cell setting, space group	Triclinic, P $\bar{1}$	Triclinic, P $\bar{1}$	Triclinic, P $\bar{1}$	Triclinic, P1
Z	1	1	1	2
Program / Number of reflexions used: all data / unique	SHELX / 15392 / 11198 / (F0 > 4sigF0)	MOLLY / 16983 (I > 3 σ (I))	SHELX / 15155 / 13510 / (F0 > 4sigF0)	SHELX / 45601 / 25198 (F0 > 4sigF0)
Refinement type: R(%) / R _w (%) / GOF / parameters	Spherical / 5.41 / 8.52 / 1.030 / 556	Multipole / 3.63 / 3.46 / 1.20 / 505	Spherical / 4.73 / 9.78 / 1.240 / 552	Spherical / 10.02 / 13.18 / 1.097 / 1498

Table 3

bond	Experimental data at 210 K	Literature data, CSD		Experimental data at 210 K	Literature data, CSD	
		Average data	<min, max>		Average data	<min, max>
	Cytosinium B			Cytosine C		
C4-N4	1.322(1)	1.312	1.296-1.327	1.317(1)	1.329	1.320-1.342
C4-C5	1.428(1)	1.417	1.387-1.430	1.427(11)	1.426	1.411-1.440
C5-C6	1.359(1)	1.343	1.329-1.354	1.351(2)	1.340	1.332-1.353
C6-N1	1.358(1)	1.357	1.344-1.372	1.385(2)	1.356	1.353-1.361
N1-C2	1.363(1)	1.365	1.336-1.381	1.362(2)	1.372	1.365-1.382
C2-N3	1.365(1)	1.376	1.360-1.394	1.357(1)	1.351	1.332-1.364
N3-C4	1.354(1)	1.354	1.341-1.362	1.351(1)	1.345	1.337-1.353
C2-O2	1.240(1)	1.220	1.196-1.246	1.254(2)	1.255	1.241-1.274

Table 4

D-H	A	d_(H-A) Cytosinium A	A	d_(H-A) Cytosinium B	A	d_(H-A) Cytosine C
C5-H5	O10 ⁽ⁱ⁾	2.2430(7)	O10 ^(vi)	2.0702(6)	O22 ⁽ⁱⁱⁱ⁾	2.6819(5)
C6-H6	O11 ^(vii)	2.8154(7)	O10	2.6784(7)	O11 ^(iv)	2.4177(8)
N1-H1	O4 _w	1.7006(7)	O1 _w	1.7519(7)	O2 _w	2.288(1)
N4-H4	O23 ⁽ⁱ⁾	1.8678(6)	O24 ^(viii)	1.8536(5)	O22 ⁽ⁱⁱⁱ⁾	1.8798(5)
Total		8,6268		8,3541		9,2674

Table 5.

Type of bond	Bond	d_1	d_2	$\rho(r_c)$	$\nabla^2 \rho(r_c)$ ρ	ϵ	λ_1	λ_2	λ_3	H
C-C										
	C5a-C4a	0.712	0.718	2.00	-18.02	0.34	-16.66	-12.46	11.09	-0.44
	C5a-C6a	0.633	0.728	2.34	-22.57	0.42	-19.69	-13.86	10.99	-0.57
	C5b-C4b	0.711	0.718	1.98	-18.50	0.30	-15.83	-12.14	9.47	-0.44
	C5b-C6b	0.520	0.839	2.16	-19.29	0.32	-1.63	-1.56	10.09	-0.50
	C5c-C4c	0.668	0.760	1.78	-12.01	0.32	-13.39	-10.16	11.54	-0.35
	C5c-C6c	0.661	0.691	2.01	-12.34	0.43	-14.97	-1.52	13.16	-0.42
C-N										
	C2a-N1a	0.575	0.798	2.12	-15.22	0.32	-18.37	-13.92	17.08	-0.47
	C4a-N3a	0.617	0.735	2.20	-14.90	0.21	-17.78	-14.74	17.62	-0.50
	C6a-N1a	0.538	0.820	2.26	-21.05	0.07	-18.59	-17.46	14.99	-0.53
	C2a-N3a	0.616	0.757	2.29	-16.87	0.24	-19.69	-15.86	18.69	-0.53
	C4a-N4a	0.55	0.776	2.47	-24.50	0.20	-21.62	-18.05	15.17	-0.62
	C2b-N3b	0.559	0.806	2.13	-22.53	0.08	-17.38	-16.07	10.92	-0.60
	C2b-N1b	0.535	0.829	2.30	-25.85	0.14	-20.07	-17.59	11.80	-0.57
	C6b-N1b	0.575	0.784	2.33	-23.28	0.13	-19.57	-17.30	13.59	-0.57
	C4b-N3b	0.577	0.777	2.34	-24.26	0.20	-20.33	-16.95	13.02	-0.57
	C4b-N4b	0.558	0.766	2.57	-29.96	0.24	-23.46	-18.98	12.48	-0.68
	C4c-N3c	0.542	0.811	2.11	-15.09	0.25	-16.93	-13.56	15.40	-0.46
	C2c-N1c	0.534	0.840	2.16	-13.73	0.15	-8.30	-7.98	40.94	-0.48
		C2c-N3c	0.546	0.812	2.18	-20.05	0.05	-19.08	-28.21	17.24
	C4c-N4c	0.516	0.802	2.26	-17.15	0.21	-18.78	-15.51	17.14	-0.52
C-O										
	C2a-O2a	0.482	0.760	2.74	-23.89	0.18	-27.09	-22.90	26.10	-0.72
	C2b-O2b	0.480	0.760	2.84	-31.98	0.13	-27.03	-23.99	19.03	-0.79
	C2c-O2c	0.537	0.736	2.58	-14.84	0.04	-22.19	-21.23	28.57	-0.63
H-C										
	H5a-C5a	0.380	0.704	1.80	-17.66	0.15	-17.27	-15.07	14.68	-0.38
	H6a-C6a	0.403	0.677	1.93	-22.50	0.02	-18.09	-17.79	13.48	-0.43
	H6b-C6b	0.445	0.638	1.74	-17.09	0.13	-8.30	-7.98	40.94	-0.36
	H5b-C5b	0.449	0.633	1.86	-19.01	0.13	-2.64	-2.43	14.95	-0.40
	H5c-C5c	0.308	0.778	1.58	-16.37	0.09	-2.09	-2.08	13.04	-0.31
	H6c-C6c	0.202	0.728	1.69	-22.21	0.16	-23.16	-19.98	20.92	-0.36
H-N										
	H4a-N4a									
	H3a-N4a	0.194	0.815	1.81	-27.55	0.05	-28.05	-26.63	27.14	-0.42
	H1a-N1a	0.240	0.768	2.15	-33.69	0.00	-29.37	-28.59	24.28	-0.54
	H2b-N3b	0.158	0.856	1.49	-16.75	0.04	-24.17	-23.33	30.75	-0.29
	H1b-N1b	0.242	0.768	2.06	-31.24	0.05	-28.10	-26.83	23.69	-0.50
	H3b-N4b	0.213	0.795	2.21	-42.33	0.04	-33.50	-32.19	23.36	-0.59
	H4b-N4b	0.240	0.770	2.24	-37.97	0.04	-30.95	-29.89	22.86	-0.59
	H3c-N4c	0.199	0.814	1.91	-27.93	0.05	-9.55	-9.21	43.39	-0.45
	H4c-N4c	0.201	0.807	1.97	-41.13	0.02	-31.33	-30.61	20.81	-0.51
		H1c-N1c	0.162	0.690	3.00	-80.55	0.05	-9.55	-9.21	43.39

Table 6.

		C-C	C-N	C-O
		<min, max>	<min, max>	<min, max>
$\rho(r_{cp})(e\text{\AA}^{-3})$	This work	<1.78 , 2.33>	<2.11 , 2.57>	<2.58 , 2.84>
	Exp. data [50]	<2.09 , 2.39>	<2.36 , 2.51>	2.95
	Theo. 6-31G** study [50]	<2.02 , 2.28>	<2.20 , 2.39>	2.57
$\nabla^2\rho(r_{cp})(e\text{\AA}^{-5})$	This work	<-12.01 , -22.57>	<-13.73 , -29.96>	<-14.84 , -31.98>
	Exp. data [50]	<-17.46 , -22.27>	<-22.23 , -25.54>	-39.77
	Theo. 6-31G** study [50]	<-17.50 , -21.57>	<-20.92 , -26.62>	-19.67

Table 7.

atom	Experimental									Theoretical 4-31G**	
	Cytosinium A		Cytosinium B		Cytosine C		[49]	[52] ¹	[52] ²		[48]
	κ	charge	κ	charge	κ	charge	charge	charge	charge		charge
C2	1.09(1)	0.49(6)	1.05(1)	0.72(6)	1.06(2)	0.52(7)	0.16	0.12(4)	0.50(3)	0.23	
C4	1.051(7)	0.40(6)	1.036(7)	0.49(6)	1.030(7)	0.32(6)	0.13	-0.10(5)	0.29(3)	0.17	
C5	1.051(7)	-0.38(6)	1.036(7)	-0.17(6)	1.030(7)	-0.22(6)	-0.92	0.07(3)	-0.10(3)	-0.18	
C6	1.051(7)	0.10(6)	1.036(7)	0.15(6)	1.030(7)	-0.08(7)	0.67	0.09(5)	0.17(4)	0.83	
O2	1.011(5)	-0.23(6)	0.948(5)	-0.46(6)	0.942(6)	-0.51(8)	-0.31	-0.37(3)	-0.53(2)	-0.44	
N1	1.000(5)	-0.51(6)	0.967(5)	-0.40(6)	0.974(7)	-0.26(8)	-0.67	-0.39(5)	-0.39(4)	-0.66	
N3	1.000(5)	-0.51(6)	0.967(5)	-0.63(6)	0.974(7)	-0.87(8)	-0.23	-0.20(4)	-0.49(2)	0.27	
N4	1.000(5)	-0.47(6)	0.967(5)	-0.49(6)	0.974(7)	-0.45(7)	-0.91	0.08(5)	-0.14(5)	-1.40	
H1	1.160	0.35(4)	1.160	0.39(4)	1.160	0.17(5)	0.13			0.16	
H2	1.160	0.43(4)	1.160	0.41(4)							
H3	1.160	0.34(4)	1.160	0.32(4)	1.160	0.24(4)	0.11	0.11(3)	0.22(4)	0.15	
H4	1.160	0.35(4)	1.160	0.36(4)	1.160	0.48(4)	0.11	0.07(3)	0.28(2)	0.13	
H5	1.160	0.20(4)	1.160	0.14(4)	1.160	0.25(4)	0.02	-0.02(3)	0.05(2)	0.04	
H6	1.160	0.14(4)	1.160	0.17(4)	1.160	0.42(4)	0.06	0.11(3)	0.05(3)	0.07	

Table 8

	Structure at 210 K			Structure at 100 K											
	A	B	C	1	2	3	4	5	6	7	8	9	10	11	12
O2	Na	Ow	Ow	Ow	Ow	Ow	Ow	Ow	Ow	Ow	Ow	Na	Na	Na	Na
N1	Ow	Ow	Ow O2x	Ow O2x	O1x O2x Ow	O2x Ow	Ow O2x	Ow O2x	Ow O2x	Ow O2x	Ow O2x	Ow O1x	O1x Ow	Ow O1x	O1x Ow
C6	O1x	O2x O1x	O1x O2x	O2x O2x	O2x	O2x	O1x O2x	O1x O2x	O1x	O1x O2x	O1x O2x	O1x	O1x	O1x	O1x
C5	O1x	O1x O1x	O2x	O1x	O2x	O1x O1x	O2x	O1x O1x		O1x	O2x	O1x		O1x	O1x
N4	O2x	O2x	O2x O2x	O2x	O2x	O2x	O2x O2x	O1x O2x	O2x O2x	O2x	O2x O2x	O2x O2x	O2x O2x	O2x O2x	O2x O2x
Cytosine type (from the crystal packing)				1B	2C	3B	4C	5B	6C	7B	8C	9A	10A	11A	12A
Final assignation: cytosine type at 100 K				1B(B)	2C(C)	3B(B)	4C(C)	5A(B)	6A(C)	7C(B)	8B(C)	9A(A)	10A(A)	11B(A)	12C(A)

Table 9

	C2-N3-C4	N3-C4-C5	N4-C4-C5	C2-N3-C4	N3-C4-C5	N4-C4-C5	$\Delta(C2N3C3)$	$\Delta(N3C4C5)$	$\Delta(N4C4C5)$	$\sum_i \Delta_i $	Assignment
ACITEM	124.34	117.38	123.82	119.24	121.34	116.99	5.1	-3.96	6.83	15.89	B-C
CYTZNC	124.48	117.03	123.64	118.8	122.64	119.94	5.68	-5.61	3.7	14.99	B-C
KEVSAG	123.87	118.58	123.96	118.73	123.19	120.27	5.14	-4.61	3.69	13.44	B-C
MECTSI	124.64	117.29	124.86	120.02	120.99	121.98	4.62	-3.7	2.88	11.2	B-C
QOCTUX	123.28	118.53	123.64	119.78	121.04	121.59	3.5	-2.51	2.05	8.06	B-C
300 K	122.00	118.87	122.74	121.47	120.27	120.94	0.53	-1.4	1.8	3.73	B-C
210 K	122.15	119.13	122.23	121.88	120.5	120.56	0.27	-1.37	1.67	3.31	B-C
GITYEN	121.49	120.85	120.42	121.12	119.27	122.04	0.37	1.58	-1.62	3.57	A-A
LIWSOA	120.76	121.02	120.91	122.09	119.57	121.82	-1.33	1.45	-0.91	3.69	A-A
300 K	121.91	119.76	121.96	121.91	119.76	121.96	0	0	0	0	A-A
210 K	121.78	119.95	121.73	121.78	119.95	121.73	0	0	0	0	A-A
1 & 2	119.94	121.22	121.62	120.05	124.32	115.02	-0.11	-3.1	6.6	9.81	B-C
3 & 4	122.17	119.35	123.83	119.13	125.32	115.19	3.04	-5.97	8.64	17.65	B-C
5 & 6	122.98	118.01	122.43	123.05	117.95	123.81	-0.07	0.06	-1.38	1.51	A-A
7 & 8	120.4	120.14	121.54	126.36	114.47	126.67	-5.96	5.67	-5.13	16.76	C-B
9 & 10	121.66	120.4	121.28	119.94	120.68	121.85	1.72	-0.28	-0.57	2.57	A-A
11 & 12	125.59	117.67	122.88	119.94	120.68	121.85	5.65	-3.01	1.03	9.69	B-C