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On the accuracy of intermolecular interactions and charge transfer: the case of TTF-CA

Pilar García^a, Slimane Dahaoui^a, Claudine Katan^b, Mohamed Souhassou^a and Claude Lecomte^{a†}

5 High-resolution X-ray diffraction experiments and state of the art density functional theory calculations have been performed. The validity of the atoms in molecules approach is tested for the *neutral-ionic* transition of TTF-CA which involves a transfer of less than one electron between the donor and acceptor molecules. Foremost, crystallographical data have been
10 reassessed along the temperature-induced *neutral-ionic* phase transition undergone by this charge transfer complex. Based on accurate X-ray structures at 105 and 15 K, topological analysis of both DFT and the experimental multipolar electron densities allowed detailed characterization of intra- and interstack intermolecular interactions. Direct quantification of
15 the intermolecular charge transfer and the dipole moment are discussed.

1. Introduction

Molecular and hybrid organic-inorganic materials will gradually replace the harder inorganic materials. The efficiency of some of these materials is linked to the cooperativity concept which is widely used among the molecular materials
20 community. However, cooperativity is still not understood on physical and chemical grounds; it needs thorough investigations of intermolecular interactions (hydrogen bonds, van der Waals contacts, hydrophobicity, charge transfer...) which is a very big challenge for both experimental and theoretical methods. This paper will concentrate on the characterization of intermolecular contacts and on the estimation
25 of the molecular and atomic charges. Intermolecular interactions are usually experimentally characterized by crystallographic or Raman-IR methods: conventional crystallography gives interatomic distances which should be compared to the sum of the so called "ionic" or "van der Waals" radii which are defined from statistical analysis of experimental interatomic distances. IR or Raman methods
30 correlate intermolecular interactions to the variation of typical phonon or molecular group modes but the assumption are also empirical and only qualitative. Therefore one has to go a step further by modelling accurately the charge density in the intermolecular regions: this will give a direct estimation of the molecular charges. This can be achieved either by theoretical methods (Ab initio, DFT) or/and accurate
35 high resolution crystallography.^{1,2} Once a model of the continuous charge density is obtained, then a method has to be chosen for analysing the interactions and

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accurately estimate the atomic group or molecular charges. This method has to be used on experimental and theoretical charge densities.

The charge concept of an atom is not based on physically measurable quantities because it is subjected to atom partitioning. In fact, it is impossible to determine unambiguously the atomic boundary inside which an atomic charge is located: atomic charges depend on the form and the size of the integration volume and on the atomic orbitals used (Ab Initio Methods). Hirshfeld has proposed in 1977³ the stockholder scheme using an atomic weighting function

$$w_j(\vec{r}) = \rho_j^{sph}(\vec{r}) / \rho_{pro}^{sph}(\vec{r}) \quad (1)$$

where ρ^{sph} is the electron density of the independent atom (ρ_i^{sph}) or of the promolecule (ρ_{pro}^{sph}). When multiplied by the modelled density $\rho(\vec{r})$ the function gives an estimation of the charge density of the j^{th} atom of the unit cell

$$\rho_j(\vec{r}) = \rho(\vec{r}) \cdot w_j(\vec{r}) \quad (2)$$

The limiting atomic volume is therefore given by a threshold of w_j . Such atoms have therefore some overlap. Coppens⁴ has also proposed a method based on the Wigner-Seitz method which defines an atomic polyhedron which boundaries are defined by

$$(\vec{r} - \vec{r}_A) \frac{\vec{R}_{AB}}{R_A} = (\vec{r} - \vec{r}_B) \frac{\vec{R}_{AB}}{R_B} \quad (3)$$

R_{AB} being the internuclear distance and R_A and R_B are "a priori" atomic radii. This method which was used to estimate the charge transfer in the TTF-TCNQ material relies on the van der Waals radii. In conclusion all these partitioning applied to both experimental and theoretical models are based on "a priori" assumptions. On the other hand, Mulliken partitioning⁵ is only used with theoretical electron density whereas the κ -P_v refinement method⁶ is widely used by X-ray diffractionists only : it assumes an expanded/contracted spherical pseudoatom density

$$\rho_j(\vec{r}) = \rho_j^{core}(\vec{r}) + P_j^{val} \kappa_j^3 \rho_j^{val}(\kappa_j \vec{r}) \quad (4)$$

where ρ^{core} and ρ^{val} are HF electron densities of the free atoms. Other methods to estimate charges are potential fitted charges^{7,8}. We have shown that these fitted charges are very close to those obtained directly by κ -P_v refinement.

Bader⁹ proposed a topological analysis of the total electron density based on quantum mechanics: the whole molecular system in the crystal is represented by a sum of non overlapping atomic basins: their boundary surface is limited by zero flux surfaces

$$\nabla \rho(\vec{r}) \cdot \vec{n}(\vec{r}) = 0 \quad (5)$$

Integration of charge, dipole and multipole moments inside the atomic basins can be performed. This last method can be applied to both experimental and theoretical data

and has the advantage to characterise intermolecular interactions¹⁰ by bond paths⁷⁰ and charge density at the critical points ($\nabla\rho(\vec{r}) = 0$).

The present discussion proposes to address the accuracy of such an approach on the basis of up to date experimental and theoretical results obtained for a model compound. The tetrathiafulvalene-p-chloranil (TTF-CA; $C_6S_4H_4 \cdot C_6Cl_4O_2$) was chosen because it shows small charge transfer (less than 1e) via intermolecular⁷⁵ interactions that depends on various thermodynamic parameters (T, P, $h\nu$). The TTF-CA complex is built from donor (D=TTF) and acceptor (A=CA) –D-A-D-A-D-A stacks and undergoes a first-order neutral-ionic phase transition (NIT) under temperature variation at a critical temperature of about 80K^{11,12} as well as under pressure variation at a critical pressure of about 11kbar.^{12,13,14} N to I or I to N⁸⁰ conversion can also be induced by photoirradiation as was shown by Koshihara et al.¹⁵ allowing for time-resolved crystallography of such a transformation at the picosecond time scale.¹⁶

According to vibrational spectroscopy and absorption spectra experiments¹⁷, the charge transfer (CT) q_{CT} has been estimated to increase from 0.3e in the N phase to⁸⁵ 0.7 in the I phase. It has been shown that this NIT comes along with a structural phase transition characterized by the loss of inversion symmetry (P21/n to Pn space group) and a dimerization along the stacking axis.¹⁸

Before getting onto the main purpose of this work, crystallographical data need to be reassessed along the temperature-induced *neutral-ionic* phase transition⁹⁰ undergone by TTF-CA, to ensure appropriate starting point. Based on the outcoming X-ray structures obtained at 105 and 15 K, intra- and intermolecular interactions are discussed on the ground of the topology of both experimental and theoretical total electron densities. Corresponding D (TTF) and A (CA) single crystals have also been investigated to provide benchmarks. Quantification of the⁹⁵ intermolecular interaction, charge transfer and the dipole moments through the topological analysis of the electron density will be discussed.

2. Experimental section

2.1 Experimental details

¹⁰⁰ TTF and CA powders were purchased from Lancaster and purified. Single crystals of TTF were obtained by sublimation method. Chloranil crystals were obtained in acetonitrile solution. After a few days of slow evaporation, prismatic yellow single crystals appear in the solution. Single crystals of the TTF-CA complex were grown by sublimation of the component materials in a vacuum-sealed¹⁰⁵ Pyrex glass tube, which was placed at 79 °C in an electric furnace for several days. Dark prismatic-like crystals were obtained.

The temperature evolution of the unit cell parameters was studied using a home-made helium bath cryostat¹⁹ specially designed for a four-circle diffractometer equipped with a 2D detector. The¹¹⁰ sample is mounted on a two-circle goniostat placed in the Helium tail of the cryostat. This goniostat is magnetically coupled to a master magnet fit on the arm of the diffractometer which allows the orientation of the sample inside the cryostat. The cell parameters were determined by the analysis of a set of 84 and 42 images, when cooling and warming, respectively, repeated from 110K to 50K, with an accuracy of 0.02K.

¹¹⁵

For the structures and charge density determinations, three-dimensional diffraction data were collected on Enraf-Nonius Kappa and Oxford-Xcalibur-Sapphire2 CCD-based diffractometers with MoK α radiation and a nominal crystal-to-detector distance of 40 mm. The crystal was cooled from room temperature to 105K with an Oxford Cryosystems N₂ open flow cryostat²⁰ and to 15K with an Oxford Diffraction Helijet open-flow He gas cryosystem.

Integration of frames and data reduction were performed with DENZO²¹ and CrysAlis Red²² programs. The multiple integrated reflections were averaged using SORTAV²³ adapted to area detector data. Internal agreement factors for all data are given in Table 1. Only reflections having I \geq 3 σ (I) were used in the structure and electron density least-squares refinements.

At 105K, TTF-CA crystallizes in the monoclinic system, space group P2₁/n (systematic absences: h0l, h+1 \neq 2n; 0k0, k \neq 2n). At 15K, diffraction intensities of (0k0) with k=2n+1 do exist, while reflections (h0l) with h+l=2n+1 remain systematically absent in agreement with the loss of the 2₁ screw axis, the space group of the ionic phase is then Pn. In order to compare both structures, the coordinates x and z of the Cl19 atom were fixed at high temperature (HT) structure values and only the y coordinate was refined for the low temperature (LT) data. Sorted data were merged in the m point group, and anomalous dispersion corrections were performed as described in Souhassou et al.²⁴ Experimental details are reported in Table 1.

Molecular structures of TTF-CA were solved by direct methods and subsequent Fourier analysis, then, refined, with a full-matrix least-squares refinement (independent atom model IAM, SHELXL-97).²⁵ Hydrogen atoms were first constrained to positions contained from neutron diffraction experiments.¹⁸ The final refined parameters were imported into the MOLLY program to model the aspherical electron density.²⁶ After convergence of the multipolar refinement, H atoms were allowed to relax as their aspherical electron density features are taken into account in the multipolar refinement, leading to unbiased H positions. Recently we have shown that this method gives best C-H bond distances close within 0.03 Å to distances usually found from neutron diffraction.²⁷

It is very important to note that, although the molecules do not lie anymore on inversion centres in the LT phase, molecular geometries are only slightly deformed, remaining almost centrosymmetric as will be discussed later. This feature induces strong correlations between charge density parameters of pseudo symmetry related atoms. In order to avoid these correlations, the following approach was carried out: first all atoms positions and thermal parameters were refined together without any constraints or restraints until convergence, and then the multipole parameters (all κ , P_v and P_m) of atoms related by an inversion centre in the N phase were constrained to be equal. This procedure leads to excellent statistical indexes as R factors are smaller than 1.3%. The maxima and the minima in the residual density maps do not exceed 0.10e Å⁻³. The estimate of the average error in the experimental electron density maps is 0.05 e.Å⁻³, as calculated by^{28,29}

$$\langle \sigma_{res}^2 \rangle^{1/2} = \frac{2}{V} \left[\sum_H (K^{-1}|F_{obs}| - |F_m|)^2 \right]^{1/2} \quad (6)$$

This error is largely smaller in the intermolecular region.

Aspherical modelling of the experimental X-ray structure factors provides the total electron density distribution $\rho(\mathbf{r})$ which is the main observable in the “Quantum Theory of Atoms in Molecules” developed by Bader.³⁰ The topology of the total charge density allows a unique partition of the density into atomic contributions, by defining atomic basin as the volume limited by zero electron density gradient flux. Topological properties at critical points (CP), as electron density, $\rho(r_{CP})$; Laplacian, $\nabla^2\rho(r_{CP})$; electron density curvatures, λ_1 , λ_2 , λ_3 ; local kinetic $G(r_{CP})$ and potential $V(r_{CP})$, energy densities can then be used to describe the molecular structure and to characterize bonding and intermolecular interactions. Topological analysis of experimental electron density was performed analytically using the NEWPROP program.³¹

2.2 Computational details

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Electronic structure calculations were performed in the framework of density functional theory (DFT) using the local density approximation (LDA) parametrization by Perdew and Zunger³², Becke's gradient correction to exchange energy³³ and Perdew's gradient correction to correlation energy.³⁴ Calculations were carried out with the projector-augmented wave (PAW) method³⁵ which uses augmented plane waves to describe the full wave functions and densities without shape approximation. The core electrons were described within the frozen core approximation. The plane wave cutoff for wave functions and densities were fixed respectively to 90 and 360 Ry to ensure required precision for topological analysis. Atomic coordinates were obtained from experimental X-ray data obtained after complete multipolar refinement. CA and TTF isolated molecule calculations were conducted using tetragonal supercells (12.5 x 12.5 x 7.5 Å for TTF and 12 x 12 x 7 Å for CA) including electrostatic decoupling from periodic images.³⁶ For TTF-CA, crystal calculations were performed with three k-points along \mathbf{a}^* .³⁷ Topological analysis of theoretical densities was achieved with the InteGriTy software package using a grid spacing smaller than 0.044 Å.³⁸

Compared to indirect methods, previous DFT investigation of CT complexes found that charge transfer (q_{CT}) is overestimated in the HT phase^{37,39}, certainly due to the approximation used for the exchange and correlation potential that suffers from the well-known problem of the self-interaction error and to a much lesser extent from the missing of the dynamical part of the electron-electron interactions in any local or semi-local approaches. One of the major effects of the too rapid decay of the LDA (or even GGA) exchange-correlation potential is the strong charge delocalization, that leads to a systematic overestimation of q_{CT} in the HT phase for instance.⁴⁰

3. Results and discussion

3.1 Thermal behaviour of the cell parameters

The evolution of the cell parameters as a function of temperature gives a clear evidence of the phase transition at $T_c = 81\text{K}$: \mathbf{a} and \mathbf{c} cell parameters display a slope anomaly, while \mathbf{b} presents an abrupt jump; in all cases the contraction rate decreases below T_c .⁴¹ Fig. 1a shows that on cooling β decreases from 99.06(2)° to 98.68(2)° between 82 and 81K. However, when warming the increase of the β angle occurs between 82 and 83.5K. This hysteresis of about 1K is also observed on the \mathbf{b} parameter (Fig. 1b) and confirms the observed evolution of the intensity of the (0 3 0) reflection.⁴¹ Thus all the cell parameters are sensible to the phase transition, reflecting the 3D character of the temperature induced transition which will be studied using AIM method.

3.2 Structural properties: Phase transition

The molecular structures of TTF-CA at 105 and 15K with atom labelling are reported in Fig. 2. In the high temperature phase (HT, $T > T_c$) the TTF-CA space group is $P2_1/n$ (monoclinic), and the TTF and CA molecules form mixed stacks along \mathbf{a} axis (Fig. 2a). The TTF and CA molecules lie both on inversion centres respectively at (0 1/2 0) and (1/2 1/2 0) and all atoms are in general crystallographic positions. Since D and A lie on inversion centres, the asymmetric unit is made of half a TTF molecule and half a CA molecule. Along the stack, the TTF axis, corresponding to the central C=C bond, forms an angle of 105.9(5)° with the CA O...O axis. As mentioned by Mayerle et al.⁴² and in agreement with the general trend observed in TTF-XA (X=F, Cl, Br)^{42,43}, the best overlap conditions between the highest occupied molecular orbital (HOMO) of D and the lowest unoccupied molecular orbital (LUMO) of A are fulfilled when the angle between the O...O axis of the XA molecule is almost orthogonal to that of the TTF molecule. An edge-on view of the molecules shows several features of interest regarding the molecular configurations and packing. The distance between the TTF and CA molecular centres in the stack is about half the \mathbf{a} parameter (i.e. 3.615(2) Å). However, the mean planes of the TTF and CA molecules are not parallel. The dihedral angle is 1.6(5)°. This angle induces a close contact along the stack between a sulphur atom S4 and the C15-C17 bond, which has an important contribution to the valence band.³⁷ This short contact (i.e. 3.325(2) Å) favours the intermolecular interactions needed for the CT.

In the low temperature phase (LT, $T < T_c$), the crystal space group of TTF-CA becomes Pn . While keeping the same lattice translational symmetry, the organic stack gets dimerized. The asymmetric unit is made of one TTF and one CA molecules, whose centroids are shifted from their HT phase positions, mainly along \mathbf{a} axis, e.g., the stacking axis: -0.08 Å and 0.1 Å for TTF and CA

230 respectively (Fig. 2b). TTF and CA form pairs of D and A characterized by different intra- (d_{intra})
and inter- (d_{inter}) pair distances, defined as the distances between the molecular centres ($d_{\text{intra}} =$
3.416(4) Å and $d_{\text{inter}} = 3.785(4)$ Å). However, as discussed in the dimerization process of
235 tetrathiafulvalene-*p*-bromanil (TTF-BA; $C_6H_4S_4 \cdot C_6Br_4O_2$)⁴³, it is more relevant to compare the
closest intermolecular contacts: the shortest atom to atom contact in the stack connects a carbon
atom of the central C=C bond of TTF to a CA carbon bonded to an oxygen. Due to the dimerization,
two different distances are observed: an intradimer C9...C21 distance (3.136(4) Å) and an interdimer
C2...C15 (3.531(4) Å), while in the HT phase both distances are equivalent by crystal symmetry
(3.370 Å). Other close contacts originate from the overlap between CA's π electrons and the p_z
orbital of one of TTF's sulphur atoms, defining S...C-C contacts. Two non equivalent contacts also
240 appear: the intrapair S11...C23-C15 is reinforced and the corresponding distance reduced to 3.158(4)
Å, while the interpair S4...C21-C17 distance increases to 3.503(4) Å. However, relative orientation
of D and A along the stack are comparable to that of the HT phase, the central C=C bond of TTF
forms an angle of 105.7(1)° with the O...O axes of CA, and the dihedral angle between the mean
planes of D and A is 1.1(1)°.

245 Intramolecular geometries of TTF and CA slightly change when cooling below T_c . In this way, the
multipolar refinement shows an increase of the TTF C=C bond lengths. This effect is larger for the
central C=C bond, from 1.3678(7) Å at 105K to 1.3862(7) Å at 15K. An inverse effect at the limit of
the experimental accuracy occurs for the C-S bonds, with distances decreasing from 1.7494(4) and
250 1.7500(4) Å at 105K to 1.7429(8), 1.7440(9), 1.7439(9) and 1.7499(9) Å at 15K. This evolution of
the C=C and C-S bonds is related to the change of q_{CT} in the complex. The HOMO of TTF is π -
antibonding with respect to the C-S bonds and π -bonding with respect to the C=C bonds, thus, when
 q_{CT} increases, C-S bond lengths decrease, while C=C increase.⁴⁴

255 In the CA molecule all bond lengths are modified due to the phase transition, i.e. by the increase of
 q_{CT} . The single C-C bond lengths decrease, while C=C double bonds increase. However the effect is
larger for C=O and C-Cl bond lengths. The C=O double bonds increase from 1.2227(4) Å at 105K
to 1.237(1) Å for both C=O bonds at 15K. The C-Cl bond lengths also increase from 1.7062(4) and
1.7037(4) to 1.7234(9), 1.7190(9), 1.7176(9) and 1.7162(9) Å. Hence, as the LUMO of CA is π
260 antibonding for both C=O and C-Cl, these bonds increase when q_{CT} becomes higher.

Finally, the U_{ij} anisotropic thermal parameters at 15K are roughly half than those at 105K. This
usual behaviour, also observed in TTF-BA⁴³, largely contrasts with the trends observed in other CT
complexes, such as DM-TTF-CA, where some anisotropic temperature factors along the stacking
265 axis increase below 65K, suggesting an order-disorder type of phase transition.⁴⁵

We can conclude that the observed geometrical changes of TTF and CA are due to the increase of
the q_{CT} between D and A in the complex. But, contrary to the neutron LT structure¹⁸, where the
bond lengths in half a TTF molecule are reported to be 1-2% longer than in the other half, this work
270 shows that the NIT does not induce such large differences. In fact, in the LT phase both TTF and
CA moieties remain almost centrosymmetric. Our results are in agreement with recent synchrotron
results⁴⁶ that show bond lengths differences of the order of the estimated errors (0.004 Å).
Therefore, such small differences cannot characterize the phase transition. One reason for the
discrepancies between neutron and X-Ray structures arises from the small observation/parameter
275 ratio that allowed only for isotropic estimation of Debye Waller factors in the neutron case whereas
a full multipolar refinement was performed on X-ray data.

280 3.3 Charge transfer and volumes

CT is at the origin of the very interesting electronic, magnetic and optical properties of the CT
complexes, and therefore its accurate determination is essential to understand, classify and control
them. However, as discussed on the introductory paragraph, it is not a well-defined quantity and
285 there are several experimental and theoretical methods to calculate it. For TTF-CA, it has been
estimated from the CT absorption band⁴⁷, from vibrational spectroscopy^{17,48} and by empirical
relations between q_{CT} and D or A geometry.⁴⁹ All these methods lead per se to indirect estimations
of q_{CT} . In order to avoid both transferability problems and at the same time to allow for more direct
determinations of q_{CT} , these determinations should be obtained from intrinsic properties, e.g., the
290 electronic structure of the CT complex. These has already been achieved for a couple of NIT

compounds on the basis of DFT calculations^{37,39}, in particular using "Quantum Theory of Atoms in Molecules".³⁰ For TTF-CA, integration over atomic basins leads to q_{CT} of 0.48 and 0.64 e respectively at 300 and 40K. Clearly, q_{CT} is overestimated in the HT phase due to the approximation used for the exchange and correlation (see section 2.2).

295 As discussed above and observed in several TTF-based donor molecules⁵⁰, oxidation of donor molecules produces variations in the central C=C and C-S bond lengths. First principle studies of TTF⁴⁴ and CA⁵¹ isolated molecules in different oxidation states have established the linear effects of molecular ionicity on bond lengths and several vibrational frequencies. In particular, C=O and C-Cl bonds of chloranil and central C=C and C-S bonds in TTF are shown to be particularly sensible to molecular ionicity. These dependences have been related to the contributions of the bonds to the LUMO of CA and HOMO of TTF. However, it has been stressed that these linear relationships may fail due to intermolecular interactions. Moreover, concerning frequency, electron-intramolecular-vibration interactions should not be disregarded.⁵² Examination of the Cambridge Structural Data Base has allowed Umland and co-workers⁴⁹ to establish linear relationships between q_{CT} and several functions (ratio and difference) of these C=C and C-S TTF bonds. This approach was used to estimate q_{CT} in TTF-CA¹⁸ from experimental bond lengths: values of 0.4 and 0.8 e were deduced at 300 and 40K respectively.

310 Starting from our more accurate structural data, the estimates are $q_{CT}=0.40$ e at 105K and $q_{CT}=0.65$ e at 15K. However, if we consider the values of the bond lengths of TTF and CA reported for different oxidation states^{44,51}, charge transfer variations (Δq_{CT}) from 105K to 15K of 0.4 and 0.5 e are obtained using C=O and C-Cl bond lengths respectively. From the TTF C=C and C-S bonds, Δq_{CT} of 0.5 and 0.15e are obtained respectively. When using bond lengths obtained for crystals of pure TTF or CA as reference ($q_{CT} = 0$) and the slopes determined by DFT calculations,^{44,51} results become worse. Indeed, C-Cl bonds lead to $q_{CT} \approx -0.1$ and $q_{CT} \approx 0.4$ e, C-S to $q_{CT} \approx 0.8$ and $q_{CT} \approx 0.9$ e whereas the central C=C bond of TTF to $q_{CT} \approx 0.7$ and $q_{CT} \approx 1.1$ e, respectively at 105 and 15K. Thus, using such approaches to determine q_{CT} is clearly hazardous, and it gives only qualitative trends for Δq_{CT} . From all these linear relationships Δq_{CT} ranges from 0.2 to 0.5e at the NIT. Therefore a direct estimation is needed, based on intrinsic properties, like the electronic structure of the complex. Experimentally, the charge transfer can be obtained from X-ray diffraction data in two ways. First, a refinement in reciprocal space with a modified spherical valence shell (κ -P, formalism⁶), according to equation 4, allows an estimation of the net atomic charges. A similar method was proposed by Espinosa et al.⁵³ to estimate the charge transfer in bis(thiodimethylene)-tetrathiafulvalene tetracyanoquinodimethane (BTDMTTF-TCNQ, $C_{10}H_8S_6^+ \cdot C_{12}H_4N_4^-$) using a P_{lm} multipolar refinement, the atomic net charges can be obtained as:

$$q_j = N_j - P_j^{val} \quad (7)$$

The obtained charge transfer for BTDMTTF-TCNQ (0.75e) has to be compared with the value derived from the diffuse scattering experiments (0.5e).⁵⁴

Both κ -P, and P_{lm} refinements have been performed in TTF-CA using electroneutrality constraint, i.e. the total charge of the unit cell is zero. Thus, these two approaches lead respectively to charge transfers of 0.14 ± 0.05 and 0.06 ± 0.05 e at 105K, while the values increase to 0.67 ± 0.03 and 0.65 ± 0.03 e at 15K. There are not significant differences between the results obtained which show the consistency of the X-ray method.

335 Finally, the topology of the charge density provides another tool to determine the charge transfer, as the sum of the atomic topological charges obtained by integration of the experimental/theoretical charge density over the atomic basins of all atoms of each molecule leads to the molecular charge and thus to q_{CT} . These molecular charges and basin volumes obtained are listed in Table 2 and are compared to those obtained for the CA and TTF single crystals. Within TTF-CA at 105K, the experimental charges of CA and TTF are respectively $-0.22(2)$ and $0.20(2)$ e, increasing to $-0.71(2)$ and $+0.77(2)$ e at 15K. Consequently, the experimental q_{CT} is 0.21 ± 0.02 at HT and 0.74 ± 0.02 at LT, and $\Delta q_{CT} = 0.53$ e. The same approach using the theoretical charge density leads to $q_{CT} \approx 0.64$ e for the LT phase. How can we gauge our approach? First, our values are in pretty good agreement with the κ -P_v and P_{lm} derived charges. Secondly, considering that the unit cell contains two TTF and two CA molecules, a comparison between the volume filled by the atomic basins and the unit cell volume allows for gauging the accuracy of the methods. Summation of the atomic volumes (Table 2) reproduces the cell volume within less than 0.3%, for both temperatures at experimental and

theoretical levels. This nicely demonstrates the ability of such approaches to tackle such small changes, especially regarding atomic quantities. Correlatively, the net integrated charge of the unit cell is also very close to zero.

As the temperature decreases from 105K to 15K, the TTF volume decreases from ≈ 1.5 -2.5 %, while the CA's volume remains almost constant. In fact, as CT increases, TTF loses electron density in favour of CA and volume reduction follows. Moreover, from experiments on TTF-based CT complexes⁵⁵ and ab-initio calculations⁴⁴ it is well known that the geometry of TTF is responsive to environmental effects and phase transitions. On the other hand, the CA electron density increases with CT while lattice parameters shrink, thus maintaining constant the CA molecular volume. Additionally, CA peripheral atoms contain a large number of core electrons that prevents large volume variations compared to TTF. This is particularly clear when volumes obtained within a TTF-CA crystal are compared to those obtained on pure TTF and CA single crystals (Table 2).

3.4 Dipole moment

The calculation in the crystal axis system of the dipole moment for the dimer in the ionic phase can be performed, considering both contributions: the net atomic charges and the atomic dipoles obtained once the topological analysis has been completed. Obviously, the net charge contribution is mainly along the stacking axis **a**: (-12.585, -0.086, 0.155) D. But the contribution of the atomic dipole moments, (16.293, -7.027, 0.457) D, is large, almost lying in the (**a**,**b**) plane reflecting not only the π - π interactions between D and A (along **a**), but also interstack interactions that are dominated by hydrogen bonds as will be discussed in section 3.6.2.

Following dimerization, a ferroelectric dipole moment $\Delta \mu = \mu_{\text{dimer}} - \mu_{\text{interdimer}}$ appears. Values of 0.36 D in the stack and 0.72 D in the unit cell are found. The corresponding dipolar energy,

$Wd = (\Delta \mu)^2 / 4\pi\epsilon_0 V$,⁵⁶ where V is the volume occupied by the dipole moment, is 6.63×10^{-21} J. This small value compared to $k_B T_{N-1}$ ($Wd / k_B = 4.8$ K) shows that long-range dipolar interactions are not the driving force for the NIT, in agreement with previous findings.¹⁸

3.5 Covalent bonds

A quantitative comparison of the covalent bonds strength can be derived from their topological properties. For each type of covalent bond, similar features are found irrespective of molecular environment, e.g. in TTF-CA, TTF or CA crystals or for calculated isolated molecules. Accumulation of electron density is particularly visible in double bonds, the highest density being observed within the C=O bond. The Laplacian of the density also reveals the oxygen lone pairs and the non bonding valence shells of both chlorine and sulphur atoms. Interestingly, while other covalent bonds show almost equal λ_1 and λ_2 curvatures at CPs, indicative of cylindrical symmetry perpendicularly to the bonds, the λ_1 and λ_2 values obtained for all C=C bonds differ significantly (about 30%).

The NIT affects clearly all bond critical points as exemplified by the C2=C9 bond, for which the distance increases from 1.3678(7) to 1.3862(7) Å, the density decreases from 2.16 to 2.05 $\text{e}\text{\AA}^{-3}$ and $\nabla^2 \rho(r_{\text{CP}})$ changes from 21 $\text{e}\text{\AA}^{-5}$ to 18 $\text{e}\text{\AA}^{-5}$ when decreasing the temperature from 105 to 15K. Again, this is directly related to the bonding character of the C=C bonds in the HOMO of TTF. Both C-S and C-Cl bonds are less responsive to CT variation. For C-S, the average experimental $\nabla^2 \rho(r_{\text{CP}})$ and $\rho(r_{\text{CP}})$ vary from -7.2 $\text{e}\text{\AA}^{-5}$, 1.33 $\text{e}\text{\AA}^{-3}$ at 105K to -6.2 $\text{e}\text{\AA}^{-5}$, 1.31 $\text{e}\text{\AA}^{-3}$ at 15K. These values are in good agreement with the literature.⁵⁷ On average, the C-Cl bond CPs lie at 0.76 Å and 0.96 Å from the C atom and the Cl atom respectively. This is directly related to their respective atomic volumes. Moreover, the valence shell charge concentration of the Cl atom is strongly polarized toward the nearby C atom (bonded charge concentration). In the direction of such a polar bond, there is a lack of charge accumulation near the more electronegative atom. This behaviour is also characteristic of C-O and C-F bonds.⁵⁸ During the NIT, C-Cl bonds lengths increase from 1.7045(4) to 1.7191(9) Å in average. $\nabla^2 \rho(r_{\text{CP}})$ and $\rho(r_{\text{CP}})$ change from -4.8 $\text{e}\text{\AA}^{-5}$, 1.40 $\text{e}\text{\AA}^{-3}$ at HT to -4.1 $\text{e}\text{\AA}^{-5}$, 1.32 $\text{e}\text{\AA}^{-3}$ at LT.

The overall agreement between theoretical and experimental results is quite good for all bond CP characteristics. Among the differences, theory gives systematically slightly higher $\rho(r_{\text{CP}})$ than that deduced from experiment and lower curvatures along the bond path (λ_3) for all bonds except C=O. The picture becomes more tricky for C=O bonds: calculated values of $\nabla^2 \rho(r_{\text{CP}})$ show pronounced deviation from experimental ones, especially for λ_3 : experimental/theoretical values of $\nabla^2 \rho(r_{\text{CP}})$ are

-28/3 eÅ⁻⁵ at 105K and -19/-1 in average at 15K. Such a discrepancy between experiment and theory has already been reported in the literature, and explained as originated by the different nature of the radial functions used in DFT and multipolar refinements.⁵⁹ It should also be underlined that, because of the fast variation of the density along the C=O bond path, computations of λ_3 and $\nabla^2\rho(r_{CP})$ are not straightforward, especially when working with densities on grids.³⁸ However, experimental and theoretical $\rho(r_{CP})$ are in excellent agreement: 2.91/2.71 at HT and 2.72/2.63 eÅ⁻³ at LT.

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3.6 Intermolecular Interactions

Up to now, in the TTF-CA family, crystallographic investigations have characterized the intermolecular interactions only by distances between interacting atoms. Accurate electron density studies may go a step further, using topological analysis to describe H bonds and van der Waals interactions by energetical quantities at the intermolecular CPs associated with the interaction. For example, Espinosa et al.^{10, 60} have shown exponential relations between the charge density at CP and the potential and kinetic energy densities at CP. Such analysis needs very accurate data and a software able to calculate properties at the CP (where the density seldom reaches 0.1 eÅ⁻³) with a good precision. Moreover, the corresponding gradient change is very small (flat density) makes the determination of the bond path difficult. This is even more difficult when van der Waals interactions are concerned. The experimental studies described here are accurate enough to characterize the topological properties at CP within a relative error estimated to a maximum of 20%. The most difficult problem is to locate accurately the CP. When the density is obtained from analytical model like multipole refinement, an analytical calculation of the derivatives of $\rho(\vec{r})$ for estimating the topological properties may be performed;³¹ in that way the errors are limited only to the experimental error on the charge density. The topological analysis of the DFT density has been done on a grid³⁸ the positions and the properties of the CPS depend on the grid size and to low extended on the interpolation procedure when the charge density is flat.

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3.6.1. Intrastack Interactions

Starting from the published structures at 300 and 40 K¹⁸, Oison et al.³⁷ have conducted a thorough ab-initio study of TTF-CA. Despite differences in molecular geometries discussed in sect. 3.2, our experimental and theoretical findings on intermolecular interactions are consistent with their results. In particular, topological analysis of the charge density of the HT phase shows the existence of (3,-1) CPs lying between TTF and CA in a plane parallel to the molecules. According to their Laplacian, charge density and potential energy density, only the most important contacts are reported in Table 3 and discussed here. The strongest contact between D and A molecules is clearly visible in Fig. 3a. At HT, a 0.06 eÅ⁻³ isocontour links the two molecules via the S4...C15-C17 contact. This is directly related to the HOMO-LUMO overlap that shows up in the valence band.³⁷ The positive Laplacian (0.59/0.56 eÅ⁻⁵) and the charge density (0.06/0.06 eÅ⁻³) at the CP are characteristic of a closed-shell interaction (Fig. 2a). The agreement between experimental and theoretical results is quite good, but discrepancies are observed in the CP locations and ellipticities as discussed previously. Experimentally, the CP is closer to the C17 atom than to C15, whereas theoretically C15 and C17 distances to the CP are similar. Both experimental and theoretical ellipticities are quite large. This is an indication that the difference in CP location arises from the flatness of the charge density distribution. In the same way, the central C=C bond of TTF is connected to C15 in the experimental density and to C17 in the theoretical one, again with quite large ellipticities. Moreover, both contacts involve the same carbon atoms of CA, namely C17 and C15. Consequently, it becomes very risky to deduce whether these contacts link individual atoms or bridge over bonds. Increased theoretical and experimental precision is required to come to a conclusion for such low and flat charge density regions.

Other (3,-1) CPs are found between the S5 atom and both chlorine atoms of the asymmetric unit. These C119...S5 and C120...S5 contacts exhibit similar characteristics with experimental positive Laplacian of about 0.43 eÅ⁻⁵, a small charge density of 0.04-0.05 eÅ⁻³, and a potential energy density of about -8.1 kJ/mol. CPs in Cl...S contacts are not very frequent in literature. Compared to S...S contacts reported in L-cystine^{57a}, Laplacian, charge density and potential energy density are stronger.

465 In the LT phase, the creation of pairs of DA dimers clearly shows up (Fig. 3b) with an increase of
intradimer charge density. Due to the dimerization process, different inter- and intra dimer CP are
observed (Fig. 2b, Table 3). The S...C-C intra dimer contact remains the dominant one, with a 81%
increase in its strength as measured by $V(r)$. In the dimer pair the strength of the C...C-C and
470 S5...Cl20 increases, while that of S5...Cl19 is slightly reduced. All inter-dimer contacts are
systematically reduced and the observed reduction of CP characteristics, such as potential energy
density, demonstrates their ability to investigate that tiny structural changes. The overall agreement
between experimental and theoretical results is quite good.

475 3.6.2 Interchain Interactions

As *c* cell parameter is twice the length of the two other ones, the stacks are too far along *c* to allow
any atom to atom interactions. Consequently, contacts occur between stacks related by translation
along *b* axis, and by the *n* glide plane symmetry which remains in the ionic phase ($P2_1/n$ to Pn).
480 Interestingly, (3,-1) CPs are found not only on C-H...O, Cl...H and Cl...S contacts between D and A
molecules, but also for S...H (D-D) and Cl...Cl (A-A) ones, as shown in Fig. 4a and in Fig. 4b; the
topological characteristics of the strongest interstack interactions for both HT and LT phases are
given in Table 4.

485 At first observed by the temperature-dependence of the thermal expansion tensor⁶¹, and as discussed
by structure analysis¹⁸ and DFT calculations^{37,62}, interchain hydrogen bonds are the dominant
interaction in TTF-CA for both inter and intra stack interactions. In both HT and LT phases each CA
oxygen atom is linked to two TTF molecules through two different hydrogen bonds. First, as shown
in Figures 4a and 4b, the oxygen atoms accept the strongest bond interaction: O18...H6-C1 and
490 O24...H13-C8 its symmetry equivalent in the HT phase, are almost linear in the [1, 2, 0] direction
and are reinforced in the LT phase, but asymmetrically, O18...H6 being the strongest. In the other
lone pair direction the O18...H14-C10 and O24...H7-C3 interactions are bent and therefore weaker.
They are also strengthened in the ionic phase as their potential energy density increases by more
than 40%. This is related to the contraction of the crystal which is the largest in the plane containing
495 this C-H...O network.⁶¹

The fact that the H bonds are the dominant contacts of all the intra- and interstack interactions
contrasts with the first idea of the quasi-one-dimensionality of TTF-CA. The CT was for a long time
considered to happen as a chain along *a* axis, with weak or no interchain coupling. In this way, the
500 electrostatic interaction between the stacks was assumed to be much smaller than the intrastack one.
However, as discussed above, the temperature evolution of the cell parameters does not agree with
this assumption. Furthermore, as reported by Kawamoto and co-workers⁶³, if the charge distribution
on atoms is considered, the electrostatic energy between TTF and CA aligned along the [1, 2, 0]
direction is attractive and superior to the repulsive one between CA molecules in the *b* direction; It
505 is in clear agreement with the above analysis of the dipole moment. On the other hand, it is well
established that valence and conduction bands have a clear one-dimensional character.³⁷ Thus, the
dimensionality of TTF-CA depends strongly on the property of interest.

In TTF-BA, the estimation of q_{CT} by the Umland empirical method⁴⁹ seems to show that
510 dimerization does not imply CT variation because this complex is already ionic at room
temperature.⁴³ Can we find a rational explanation by looking at the molecular packing and the
corresponding intermolecular interactions? In TTF-BA, (space group $P-1$ to $P1$), the D and A
molecules alternate along *a* and *b* crystallographic axes, the longest axes of the D and A molecules
of the two different stacks being quasiparallel. This geometry leads to bent H bonds, with distances
515 similar to those observed in TTF-CA and therefore of equivalent strength. The main difference
between the two CT complexes comes from the halogen-halogen (Cl-X1...X2-C2) interactions.
Depending on the geometric parameters two types of halogen-halogen interactions have been
defined.^{64,65} In the first one, called head-on contact, both C-X bonds form a nearly collinear
arrangement, with contacts angles θ_1 (C1-X1...X2) and θ_2 (X1...X2-C2) around $160 \pm 10^\circ$. These
520 contacts may arise when a particular crystallographic operation relates the two halogen atoms; the
interaction is repulsive as evidenced by the electrostatic potential.⁶⁶ In the second type, the
interaction is attractive and is called side-on or polarization-induced contact, one of the angles is
 $170 \pm 10^\circ$ and the other $90 \pm 10^\circ$.^{64,65} The preference to form one of these two types of halogen-
halogen contacts is related to the polarizability of the halogen atoms. In TTF-BA attractive type II
525 interactions occur along the [1 1 1] and [1, 1, -1] directions with Br...Br distances of 3.566(3) at

100K and 3.540(3) Å at 15K, much smaller than the sum of van der Waals radii (3.70Å). In TTF-CA, Cl...Cl interactions are repulsive type I interactions with distances of about 3.470(3)Å close to the sum of van der Waals radii (3.50Å).

In conclusion, attractive interactions lead to X...X distances smaller than van der Waals contacts, whereas the repulsive Cl...Cl interactions do not allow such shortening. Topological parameters of the Cl...Cl contacts do not change very much upon the phase transition, as shown in Table 4.

4. Conclusion and perspectives

Accurate X-ray measurement and electron density modelling is a necessary tool to directly investigate intermolecular interactions and tiny charge transfers. For the TTF-CA crystal, the topological analysis of the intermolecular interactions using theoretical DFT and experimental electron densities show that C-H...O interstack contacts are the strongest interactions and are reinforced in the LT phase, whereas the strength of the Cl...Cl interactions, which may be repulsive in that case, do not change when crossing the transition. The role of the latter is still not clear and needs further discussion as for example electrostatic potential calculations.

This study also demonstrates that agreement between theory and experiment is not fully satisfactory and requires more accurate calculations and modelling: theory is not able to reproduce the charge transfer occurring during the NIT and both experiment and theory disagree in the position of the critical points of some important intermolecular interactions as Cl...Cl for example. More critical is the disagreement for finding the C2=C2...C15 (or C17) interaction pathway: both methods find a (3,-1) CP with similar characteristics but experiment would prefer a C2...C15 interaction when theory finds a C2...C17 interaction. Therefore the present accuracy does not allow to discriminate if this interaction links individual atoms (and which atoms) or bridges over bonds.

By comparison with previous work⁴³, halogen... halogen interactions need a better description in order to know when and why these interactions are attractive or repulsive and at which level they compete with C-H...O interactions during the phase transition process. In order to rank these interactions charge density and related electrostatic potential of TTF-XA (X = F, Br, I) are underway using high resolution X-ray diffraction: as an example Fig. 5 shows the contribution of the Cl atom to the electrostatic potential of CA in the ionic phase; clearly two negative lobes appear which should be related with the attractive or repulsive behaviour of X...X interactions. Although the potential distribution is the same than that observed in the vicinity of the terminal halide ligand of CH₃Cl⁶⁶, where the electrostatic energy of the minima is -80 kJ/mol; in TTF-CA -25 and -148 kJ/mol are respectively obtained at 105 and 15K. Therefore, together with the TTF, CA and TTF-CA charge densities, we will have a large panorama of the interactions as a function of the charges, multipole moments, polarizabilities of the halogen atoms and of the geometry and environment of the contacts.

It has been shown that the charge transfer is not an unambiguously defined quantity that can be directly (charge density modelling) or indirectly estimated from molecular geometry or IR spectra. Although D and A molecular geometries are sensible to the charge transfer, its estimation by empirical models is only qualitative: examination of different bond lengths leads to different q_{CT} values. Similarly, the variation of q_{CT} is supposed to be proportional to the frequency of typical phonon modes (e.g., C=O stretching modes of CA). Moreover, these methods present inevitable ambiguities: molecular geometries and frequencies of single TTF and CA molecules are used as references, because those of TTF-CA with $q_{CT}=0$ and $q_{CT}=1e$ are not known.

To our opinion the best way to determine the charges remains in the analysis of the experimental or theoretical charge density. The case of TTF-CA complex has been chosen because the CT is less than one electron shared by 26 atoms. We have shown that both κ -P_v and P_{lm} models give similar and realistic results for the charge transfer. However, they can not be compared with theoretical results, except if we transform the charge density into the theoretical structure factors. A direct comparison is possible with a topological analysis of the charge density as described by Bader that allows for a direct accurate estimation of the charge transfer occurring during the NIT (in the experimental case, $\Delta q_{CT}=0.53e$).

In conclusion these results show that Kappa, multipole and topological methods applied to X-ray data lead to statistically similar results respectively, $\Delta q_C = 0.53, 0.59$ and $0.53e$. This raises the following question; why do we obtain such a good agreement as the definitions of the κ - P_v and topological charges are clearly different: κ - P_v (or multipolar) charges are integrated from reciprocal space calculations on spherical volumes obtained from expanded or contracted HF free infinite atoms whereas the topological charges are integrated inside very non spherical atomic basins of limited volumes. It evidently leads to very different atomic net charges as shown in Table 5 which compares them in the LT phase.

An explanation may be the following: in TTF CA the charge which we consider is summed up over the whole TTF or CA molecule and the different shapes of atoms therefore has no real importance because the charge is integrated over the whole molecular volume. Therefore, as there is an electroneutrality constraint in the refinement process and as we have shown that the total charge summed over all atomic basins of the unit cell is only -0.02 and $0.06e$ at $105K$ and $15K$, instead of zero (Table 3), the small differences between the values obtained from the two methods comes only from the periphery of the molecules which electron density is small. Therefore as both charge partitioning agree it strengthens the X-ray approach but this conclusion drawn in this particular case needs to be discussed and confirmed by other charge estimations on other selected charge transfer materials or ionic species.

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References

- P. Coppens, X-Ray Charge Densities and Chemical Bonding, 1997, Oxford University Press UK
V. G. Tsirelson and R. P. Ozerov Electron Density and Bonding in Crystals, 1996, Inst. Of
Physics Publishing Philadelphia USA.
F.L. Hirshfeld, *Theor. Chem. Acta*, 1977, **44**, 129.
P. Coppens, *Phys. Rev. Lett.*, 1975, **35**, 98.
R. S. Mulliken, *J. Chem. Phys.*, 1955, **23**, 1833.
P. Coppens, T. N. Guru Row, P. Leung, E. D. Stevens, P. J. Becker and Y. W. Yang, *Acta Cryst.*,
1979, **A35**, 63.
N. Bouhaida, N. E. Ghermani, C. Lecomte and A. Thalal, *Acta Cryst.*, 1999, **A55**, 729.
(a) . G. Ángyán and C. Chipot, *Int. J. Quantum Chem.*, 1994, **52**, 17; (b) J. G. Ángyán, C.
Chipot, F. Dehez, C. Hättig, G. Jansen and C. Millot, *J. Comp. Chem.*, 2003, **24**, 997.
R. F. W. Bader, *J. Chem. Phys.*, 1980, **73**, 2871.
Espinosa, J. Molins and C. Lecomte *Chem. Phys. Lett.*, 1998, **285**, 170.
J. B. Torrance, A. Girlando, J. J. Mayerle, J. I. Crowley, V. Y. Lee, P. Batail and S. J. Laplaca,
Phys. Rev. Lett., 1981, **47**, 1747.
Y. Tokura, T. Koda, T. Mitani and G. Saito, *Solid State Communications*, 1982, **43**, 757.
J. B. Torrance, J. E. Vazquez, J. J. Mayerle and V. Y. Lee, *Phys. Rev. Lett.*, 1981, **46**, 253.
A. Girlando and A. Painelli, *Phys. Rev.*, 1986, **B34**, 2131.
S. Koshihara, Y. Takahashi, H. Sakai, Y. Tokura and T. Luty, *J. Phys. Chem.*, 1999, **B103**, 2592.
E. Collet, M. H. Lemée-Cailleau, M. Buron-LeCointe, H. Cailleau, M. Wulff, S. Koshihara, M.
Meyer, L. Toupet, P. Rabiller and S. Techert, *Science*, 2003, **300**, 612.
A. Girlando, R. Bozio, C. Pecile and J. B. Torrance, *Phys. Rev.*, 1982, **B26**, 2306.
M. L. LeCointe, M. H. Lemée-Cailleau, H. Cailleau, B. Toudic, L. Toupet, G. Heger, F. Moussa,
P. Scheiss, K. H. Kraft and N. Karl, *Phys. Rev.*, 1995, **B51**, 3374.
R. Argoud, P. Fertey, P. Bordet and J. Reymann, *Acta Cryst.*, 2000, **A56**, s221.

- J. Cosier and A. M. Glazer, *J. Appl. Cryst.*, 1986, **19**, 105.
Z. Otwinowski, W. Minor, DENZO-SMN, Methods in Enzymology, Vol. 276, Macromolecular
640 Crystallography, Part A, edited by C. W.
CrysAlis, Oxford Diffraction Poland Sp., CCD Data Reduction GUI, version 1.171.27p5 beta
Blessing, R. H. *Crystallogr. Rev.* **1987**, *1*, 3. (New version 1998).
M. Souhassou, E. Espinosa, C. Lecomte, R. H. Blessing, *Acta Cryst.*, 1995, **B51**, 661.
G. M. Sheldrick, SHELXL97. Programs for Crystal Structure Analysis (Release 97-2),
645 University of Göttingen, Germany (1997).
N. Hansen and P. Coppens, *Acta Cryst.*, 1978, **A34**, 909.
C. Jelsch, V. Pichon-Pesme, B. Bartosz, B. Guillot and C. Lecomte, to be published.
D. W. J. Cruikshank, *Acta Cryst.*, 1949, **2**, 65.
B. Rees, *Acta Cryst.*, 1976, **A32**, 483.
650 R. Bader, Atoms in Molecules: A Quantum Theory (Oxford University Press: New York, 1994).
M. Souhassou, R. H. Blessing, *J. Appl. Cryst.*, 1999, **32**, 210.
J. P. Perdew and A. Zunger, *Phys. Rev.*, 1981, **B23**, 5048.
A. D. Becke, *J. Chem. Phys.* 1992, **96**, 2155.
J. P. Perdew, *Phys. Rev. B*, 1986, **33**, 8822.
655 P. E. Blöchl, *Phys. Rev. B*, 1994, **50**, 17953.
P. E. Blöchl, *J. Chem. Phys.*, 1995, **103**, 7422.
V. Oison, C. Katan, P. Rabiller, M. Souhassou, C. Koenig, *Phys. Rev. B*, 2003, **67**, 035120.
C. Katan, P. Rabiller, C. Lecomte, M. Guezo, V. Oison, and M. Souhassou, *J. Appl. Cryst.* 2003,
36, 65.
660 V. Oison, P. Rabiller, C. Katan, *J. Phys. Chem.*, 2004, **A108**, 11049.
(a) R. Merkle, A. Savin and H. Preuss, *J. Chem. Phys.*, 1992, **97**, 9216; (b) E. Ruiz, D. R.
Salahub and A. Vela, *J. Phys. Chem.*, 1996, **100**, 12265; (c) Y. K. Zhang and W. T. Yang, *J.*
Chem. Phys., 1998, **109**, 2604.
M. Buron-Le Cointe, M. H. Lemee-Cailleau, H. Cailleau, B. Toudic, A. Moreac, F. Moussa, C.
665 Ayache and N. Karl, *Phys. Rev.*, 2003, **B68**, 064103.
J. J. Mayerle, J. B. Torrance and J. I. Crowley, *Acta Cryst.*, 1979, **B35**, 2988.
P. Garcia, S. Dahaoui, P. Fertey, E. Wenger and C. Lecomte, *Phys. Rev.*, 2005, **B72**, 104115.
C. Katan; *J. Phys. Chem.*, 1999, **A103**, 1407.
Y. Nogami, M. Taoda, K. Oshima, S. Aoki, T. Nakayama and A. Miura, *Synth. Met.*, 1995, **70**,
670 1219.
A. Tomita and S. Koshihara, 2006, in " First France-Japan Advanced School on Chemistry and
Physics of Molecular Materials".
C. S. Jacobsen and J. B. Torrance, *J. Chem. Phys.*, 1983, **78**, 112.
A. Girlando, I. Zanon, R. Bozio and C. Pecile, *J. Chem. Phys.*, 1978, **68**, 22.
675 T. C. Umland, S. Allie, T. Kuhlmann and P. Coppens, *J. Phys. Chem.*, 1988, **92**, 6456.
P. Guionneau, C. J. Kepert, G. Bravic, D. Chasseau, M. R. Truter, M. Kurmoo and P. Day,
Synth. Meta., 1997, **86**, 1973.
C. Katan, P. E. Blochl, P. Margl and C. Koeniq, *Phys. Rev.*, 1996, **B53**, 12112.
C. Pecile, A. Painelli and A. Girlando, *Mol. Cryst. Liq. Cryst.*, 1989, **171**, 69.
680 E. Espinosa, E. Molins and C. Lecomte, *Phys. Rev. B*, 1997, **56**, 1820.
C. Rovira, J. Tarrés, J. Llorca, E. Molins, J. Veciana, S. Yang, D. O. Cowan, C. Garrigou-
Lagrange, J. Amiel, P. Delhaes, E. Canadell and J. P. Pouget, *Phys. Rev.*, 1995, **B52**, 8747.
(a) J. L. Segura and N. Martin, *Angew. Chem. Int. Ed.*, 2001, **40**, 1372.; (b) M. R. Bryce, *Adv.*
Mater., 1999, **1**, 11.
685 J. Lajzerowicz and J. Legrand, *Phys. Rev.*, 1978, **B17**, 1438.
(a) S. Dahaoui, V. Pichon-Pesme, J. A. K. Howard and C. Lecomte, *J. Phys. Chem. A.*, 1999,
103, 6240; (b) S. Pillet, M. Souhassou, Y. Pontillon, A. Caneschi, D. Gatteschi and C. Lecomte,
New J. Chem., 2001, **1**, 131; (c) F. Benabicha, V. Pichon-Pesme, C. Jelsch, C. Lecomte and A.
Khmou, *Acta Cryst.*, 2000, **B56**, 155; R. Guillot, N. Muzet, S. Dahaoui, C. Lecomte and C.
690 Jelsch, *Acta Cryst.*, 2001, **B57**, 567.
A. Bachn D. Lentz and P. Luger, *J. Phys. Chem.*, 2001, **A105**, 7405.
A. Volkov, and P. Coppens, *Acta Cryst.*, 2001, **A57**, 395.
E. Espinosa, M. Souhassou, H. Lachekar and C. Lecomte, *Acta Cryst.*, 1999, **B55**, 563.
P. Batail, S. J. LaPlaca, J. J. Mayerle and J. B. Torrance, *J. Am. Chem. Soc.*, 1981, **103**, 951.
695 V. Oison, C. Katan and C. Koenig, *J. Phys. Chem.*, 2001, **A105**, 4300.
T. Kawamoto, T. Iizuka-Sakano, Y. Shimoi and S. Abe, *Phys. Rev.*, 2001, **B64**, 205107.

T. Sakurai, M. Sundaralingam, and G. A. Jeffrey, *Acta Cryst.*, 1963, **16**, 354.
N. Ramasubbu, R. Parthasarathy and P. Murray-Rust, *J. Am. Chem. Soc.*, 1986, **108**, 4308.
L. Brammer, E. A. Bruton, and P. Sherwood, *Cryst. Growth Des.*, 2001, **1**, 277.

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Figures

Fig. 1a: Temperature evolution of the beta angle ($^{\circ}$). Cooling (circles) and warming (triangles). The standard error estimated on statistical ground is 0.02° .

Fig. 1b: Temperature evolution of the **b** parameter (\AA). Cooling (circles) and warming (triangles). The standard error estimated on statistical ground is 0.001\AA .

Fig. 2a: Experimental critical points associated to the strongest intrastack contacts in TTF-CA at 105K and shortest distance (C2...C15).

Fig. 2b: Experimental critical points associated to the strongest intrastack contacts in TTF-CA at 15K and shortest intra and interdimer distances (C9...C21 and C2...C15 respectively).

Fig. 3a: Experimental total charge density at 105K. Isocontours of 0.065 e/\AA^3 .

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Fig. 3b: Experimental total charge density at 15K. Isocontours of 0.065 e/\AA^3 : the dimer is clearly identified.

Fig. 4a: Experimental critical points associated to the strongest interstack interactions at 15K.

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Fig. 4b: Experimental critical points around CA molecule associated to the strongest interstack interactions at 15K.

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Fig. 5: Contribution of the chlorine atom to the electrostatic potential at 15K. Contours intervals are 0.02 e/\AA .

Table 1

Table 1 Experimental details

	TTF	CA	TTF-CA	TTF-CA
Formula	C ₆ S ₄ H ₄	C ₆ Cl ₄ O ₂	C ₁₂ S ₄ Cl ₄ O ₂ H ₄	C ₁₂ S ₄ Cl ₄ O ₂ H ₄
M	817.33	491.72	900.38	900.38
T(K)	100	100	105	15
Crystal system	Triclinic	Monoclinic	Monoclinic	Monoclinic
Space group	P -1	P2 ₁ /c	P2 ₁ /n	Pn
Unit cell dimensions	a= 8.3101(3) Å b=12.7800(3) Å c= 7.9582(1) Å α= 99.627(3)° β = 95.906(3)° γ= 101.083(3)°	a= 8.574(1) Å b= 5.690(1) Å c= 8.700(1) Å β=105.937(8) °	a= 7.2297(5) Å b= 7.5933(3) Å c=14.4980(9) Å β=99.15(3)°	a= 7.1999(9) Å b= 7.5556(6) Å c=14.479(1) Å β=98.511(8)°
Volume (Å ³)	817.33	403.43	786.2 (2)	779.02(13)
Z	4	2	2	2
Reflections collected/ independent	141108 / 16513	85699 / 5445	168470 / 9237	75177 / 14634
μ(Mo-Kα) (mm ⁻¹), Resolution (Å ⁻¹)	1.09, 1.10	1.41, 1.19	1.28, 1.14	1.29, 1.16
R _{int}	0.0741	0.0269	0.0364	0.0266
R (I>3σ(I)) ^a	0.0142	0.0167	0.0141	0.0138

^a At the end of multipolar refinement.

Table 2

Table 2 Topological molecular volumes (Å³) and charges (e), compared to those obtained with κ-P_v and multipolar refinement. For TTF* and CA*, data are obtained from experiments on pure TTF and CA single crystals. Other results are relative to molecules in single crystals of TTF-CA at either 105 or 15 K.

	TTF*	TTF ^{105K}	TTF ^{15K}	CA*	CA ^{105K}	CA ^{15K}
V _{exp}	205.12	191.24	186.49	204.92	201.52	201.69
V _{theo}		189.92	187.07		202.82	202.27
q _{exp}	0.01	0.20	0.77	-0.02	-0.22	-0.71
q _{theo}		0.54	0.63		-0.50	-0.64
q _{exp. κ-P_v refinement}		0.14	0.67		-0.14	-0.67
q _{exp. mult refinement}		0.06	0.65		-0.06	-0.65

Table 3

Table 3 Topological characteristics of the electron density at selected (3,-1) intrachain CPs. The λ_1 , λ_2 , λ_3 ($\text{e } \text{\AA}^{-5}$) are the curvatures of the three principal axes of the Hessian matrix, $\nabla^2 \rho$ ($\text{e } \text{\AA}^{-5}$) is the Laplacian of the electron density, ρ ($\text{e } \text{\AA}^{-3}$) is the electron density ε is the ellipticity value and $V(r)$ (kJ/mol) is the potential energy density. The values d_1 , d_2 and d_3 (\AA) represent the distances from the CP to the attractors 1, 2 and 3. First line: experimental electron density, second line: theoretical electron density.

105K	d1	d2	d3	$\nabla^2 \rho$	ρ	ε	$V(r)$	λ_3	λ_2	λ_1
S4... C17-C15	1.816	1.628	1.802	0.59	0.06	1.53	-11.1	0.77	-0.05	-0.13
	1.818	1.692	1.688	0.56	0.06	5.66	-11.0	0.72	-0.02	-0.14
C15... C2-C2A ^a	1.633	1.742	2.194	0.45	0.04	6.68	-7.0	0.51	-0.01	-0.05
C17... C2-C2A ^b	1.713	1.809	1.856	0.43	0.04	7.52	-6.8	0.51	-0.008	-0.01
	1.843	1.851		0.42	0.05	0.44	-8.1	0.57	-0.06	-0.09
C119 ...S5	1.817	1.875		0.45	0.04	0.24	-7.2	0.61	-0.06	-0.08
	1.797	1.840		0.44	0.05	1.22	-8.2	0.58	-0.04	-0.10
C120 ...S5	1.795	1.833		0.52	0.04	0.29	-8.5	0.67	-0.06	-0.08
15K										
Intradimer										
S11... C23-C15	1.730	1.518	1.811	0.74	0.10	0.42	-20.2	1.19	-0.18	-0.26
	1.725	1.601	1.625	0.72	0.08	4.60	-17.0	0.99	-0.04	-0.23
C17... C2-C9 ^a	1.548	1.710	1.872	0.69	0.08	2.23	-15.6	0.93	-0.06	-0.18
C17-C21 ...C9 ^b	1.617	1.571	1.846	0.59	0.06	4.85	-12.0	0.75	-0.02	-0.14
	1.718	1.845		0.52	0.06	1.61	-10.5	0.69	-0.05	-0.12
C120...S5	1.762	1.792		0.62	0.05	0.52	-10.0	0.78	-0.06	-0.09
	1.838	1.832		0.41	0.04	0.69	-6.7	0.56	-0.06	-0.10
C119...S5	1.801	1.848		0.50	0.04	0.51	-8.1	0.66	-0.06	-0.09
Interdimer										
S4... C21-C17	1.855	1.738	1.967	0.38	0.04	1.85	-6.4	0.50	-0.03	-0.09
	1.884	1.875	1.718	0.44	0.04	3.99	-7.4	0.55	-0.02	-0.09
C125...S12	1.819	1.890		0.43	0.04	0.88	-6.8	0.59	-0.06	-0.11
	1.828	1.879		0.43	0.04	0.08	-6.8	0.59	-0.08	-0.08
C126...S12	1.820	1.839		0.44	0.04	2.12	-6.5	0.59	-0.04	-0.11
	1.809	1.845		0.49	0.04	0.21	-7.9	0.64	-0.07	-0.08

^a From experimental electron density. ^b From theoretical electron density. The differences in location of the CP are discussed in the body of the paper.

Table 4

Table 4 Topological characteristics of the electron density at selected (3,-1) interchain CPs. Same comments as Table 3.

105K	d1	d2	$\nabla^2 \rho$	ρ	ϵ	V(r)	λ_3	λ_2	λ_1
H6...O18	0.881	1.396	1.21	0.06	0.05	-16.7	1.70	-0.24	-0.25
	0.878	1.404	1.08	0.09	0.02	-20.2	1.75	-0.33	-0.33
H7...O18	1.000	1.443	0.90	0.06	0.14	-13.9	1.26	-0.17	-0.19
	0.971	1.453	0.88	0.06	0.04	-15.0	1.34	-0.22	-0.23
Cl20...S4	1.722	1.759	0.49	0.06	0.28	-10.2	0.71	-0.10	-0.12
	1.757	1.723	0.54	0.04	0.08	-8.5	0.72	-0.08	-0.09
Cl20...Cl19	1.732	1.734	0.50	0.05	0.10	-8.8	0.70	-0.09	-0.10
	1.743	1.738	0.56	0.04	0.14	-8.6	0.77	-0.09	-0.11
15K									
Short									
H13...O24	0.827	1.343	1.47	0.10	0.03	-26.8	2.20	-0.36	-0.37
	0.813	1.346	1.36	0.11	0.01	-29.0	2.32	-0.48	-0.48
H14...O18	0.901	1.387	1.28	0.09	0.07	-22.9	1.89	-0.3	-0.32
	0.888	1.389	1.17	0.09	0.06	-21.9	1.86	-0.33	-0.35
Cl20...S11	1.704	1.703	0.44	0.06	0.44	-9.7	0.75	-0.12	-0.18
	1.721	1.683	0.62	0.05	0.06	-10.0	0.85	-0.11	-0.12
Cl26...Cl19	1.770	1.727	0.51	0.05	0.14	-8.9	0.73	-0.11	-0.12
	1.747	1.753	0.54	0.04	0.15	-8.3	0.75	-0.09	-0.11
Long									
H6...O18	0.858	1.373	1.28	0.09	0.02	-22.9	1.90	-0.30	-0.31
	0.857	1.389	1.16	0.09	0.03	-22.0	1.91	-0.37	-0.38
H7...O24	0.969	1.405	1.12	0.08	0.09	-19.5	1.64	-0.25	-0.27
	0.935	1.425	1.02	0.07	0.08	-16.7	1.57	-0.26	-0.28
Cl26...S4	1.725	1.740	0.41	0.06	0.23	-9.5	0.71	-0.14	-0.17
	1.751	1.711	0.56	0.04	0.12	-8.9	0.75	-0.08	-0.09
Cl20...Cl25	1.849	1.632	0.56	0.04	1.34	-8.0	0.62	-0.02	-0.04
	1.745	1.733	0.56	0.04	0.14	-8.7	0.78	-0.10	-0.11

Table 5

Table 5 Comparison between κ - P_v , multipolar and topological net charges (e) of atoms of TTF-CA at 15K, statistical errors are 0.02e.

TTF Atom	C1/C8	C2/C9	C3/C10	S4/S11	S5/S12	H6/H13	H7/H14
κ - P_v	-0.26	0.13	-0.24	0.20	0.30	0.14	0.06
Multipolar model	-0.59	-0.23	-0.50	0.38	0.55	0.38	0.34
Topological exp	-0.53	-0.23	-0.29	0.38	0.50	0.31	0.26
Topological theo	-0.14	-0.30	-0.13	0.28	0.33	0.14	0.13
CA Atom	C15/C21	C16/C22	C17/C23	O18/O24	CH9/C125	C120/C126	
κ - P_v	0.22	0.23	0.14	-0.42	-0.27	-0.23	
Multipolar	0.32	-0.04	0.20	-0.44	-0.11	-0.25	
Topological exp	0.92	0.05	0.30	-1.09	-0.19	-0.34	
Topological theo	0.95	0.08	0.08	-1.11	-0.16	-0.16	

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