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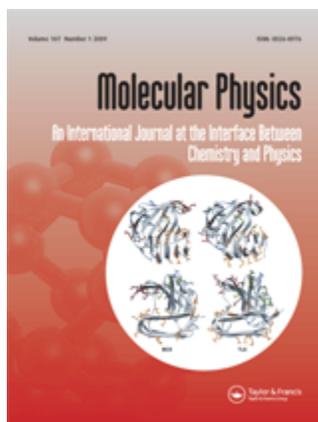
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# Empirical LCAO parameters for $\pi$ molecular orbitals in planar organic molecules

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## Abstract

We present a novel parametrization within a simplified LCAO model (a type of Hückel model) for the description of  $\pi$  molecular orbitals in organic molecules containing  $\pi$ -bonds between carbon, nitrogen, or oxygen atoms with  $sp^2$  hybridization. It is shown that the model is quite accurate in predicting the energy of the highest occupied  $\pi$  orbital and the first  $\pi$ - $\pi^*$  transition energy for a large set of organic compounds. We provide four empirical parameter values for the diagonal matrix elements of the LCAO description, corresponding to atoms of carbon, nitrogen with one  $p_z$  electron, nitrogen with two  $p_z$  electrons, and oxygen. The bond-length dependent formula of Harrison (proportional to  $1/d^2$ ) is used for the non-diagonal matrix elements between neighboring atoms. The predictions of our calculations have been tested against available experimental results in more than sixty organic molecules, including benzene and its derivatives, polyacenes, aromatic hydrocarbons of various geometries, polyenes, ketones, aldehydes, azabenzenes, nucleic acid bases and others. The comparison is rather successful, taking into account the small number of parameters and the simplicity of the LCAO method, involving only  $p_z$  atomic-like orbitals, which leads even to analytical calculations in some cases.

## 1. Introduction

Theoretical and experimental efforts for the determination of the electronic structure of organic molecules started as soon as quantum mechanics was established as the fundamental theory for the microscopic description of matter. These efforts, except for the evaluation of the energy eigenvalues of the electronic states, were concerned also with other aspects, like for example the determination of the symmetry of each electronic state, the assignment of electronic transitions (e.g. singlet-singlet or singlet-triplet transitions, Rydberg transitions,  $\pi$ - $\pi^*$  transitions etc.), and the calculation of the oscillator strength of the transitions. Apart from basic knowledge and the numerous applications of planar organic molecules containing atoms with  $sp^2$  hybridization, the  $\pi$  molecular electronic structure of such compounds is involved in several biological functions. For example, we mention chlorophyll in photosynthesis, the retinal molecule involved in vision or in photon-driven ion pumps like bacteriorhodopsin [1], and many molecules with photobiological functions such as vitamin A, vitamin D precursors, carotene etc., containing polyene chromophores [2]. Also new organic semiconductors based in pentacene and other hydrocarbon molecules have attached an enormous interest regarding their use in molecular electronics [3].

Experimental investigations of the electronic structure of organic molecules started very early, by performing absorption measurements. The ultraviolet absorption spectra of eighteen pyridines and purines [4], fourteen ethylenic hydrocarbon molecules [5], and 1,3 cyclohexadiene [6] have been measured already in 30s. Later Platt and Klevens presented the spectra of several alkylbenzenes [7], some ethylenes and acetylenes [8], and seventeen polycyclic aromatic hydrocarbons composed of fused benzene rings [9] (e.g. phenanthrene and chrysene). At the same period spectra were taken from naphthalene and biphenyl derivatives [10], m- and o-disubstituted benzene derivatives [11], and mono-substituted

1  
2  
3 and p-disubstituted benzene derivatives [12]. Absorption measurements continued with the  
4  
5 same intensity the following decades [13-15]. During the last thirty years, several new  
6  
7 methods emerged for the measurement of the electronic structure of molecules. Some of  
8  
9 them are experimentally easier from conventional absorption spectroscopy and may be able  
10  
11 to probe optically forbidden transitions. For instance singlet-triplet transitions can be easily  
12  
13 assigned by such methods. Particularly, the electron impact method has been applied in  
14  
15 1,3,5 hexatriene [16], resonant enhanced multiphoton ionization in pyrrole, N-methyl  
16  
17 pyrrole, and furan [17], electron scattering spectroscopy in propene [18] and isobutene  
18  
19 [19], and cavity ring-down spectroscopy in 1,3 butadiene [20].  
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25 Early theoretical efforts to describe  $\pi$  molecular structure have been done by Hückel  
26  
27 in 30s [21] (for a recent review on Hückel theory and its aspects see Ref. [22] and also  
28  
29 references therein for important contributions). Platt predicted the first two electronic  
30  
31 transitions of sixteen conjugated molecules by using the LCAO method [23] and also tried  
32  
33 to summarize and justify general laws that govern electronic spectra [24]. Another  
34  
35 theoretical attempt at the same period was done by Pariser and Parr who predicted the first  
36  
37 main visible or ultraviolet absorption bands of benzene and ethylene [25], butadiene,  
38  
39 pyridine, pyrimidine, pyrazine, and s-triazine [26], and several polyacenes [27]. Their  
40  
41 semi-empirical theory, known as PPP theory, was based on antisymmetrized products of  
42  
43 molecular orbitals, obtained using the LCAO approximation. Other models came out in the  
44  
45 following decades, like for example the CNDO/S2 spectroscopic model that was applied  
46  
47 for the description of the electronic excitation spectra of polyacenes (naphthalene to  
48  
49 pentacene), providing results in good agreement with experimental data [28]. In the last  
50  
51 two decades theoretical efforts were focused on more accurate calculations from first  
52  
53 principles. Such methods have been applied in many organic molecules, like for example in  
54  
55 benzene [29, 30], azabenzenes [29], heptacene [30], naphthalene, anthracene, tetracene and  
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3 hexacene [30, 31], pentacene [30, 31, 32], pyrrole [33, 34], furan [34], butadiene and  
4  
5 hexatriene [35], and cyclic ketones and thioketones [36].  
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7

8 Although methods from first principles some times -depending on the used basis set  
9  
10 or the method itself- are not so accurate for all orbitals (especially for the unoccupied  
11  
12 ones), in general they can provide very successful predictions of the electronic structure.  
13  
14 Therefore, methods from first principles are of extreme importance for the interpretation of  
15  
16 molecular electronic spectra. However, as also happens in the experimental observations, a  
17  
18 computationally demanding first principles calculation may not give particular insights at  
19  
20 the underlying mechanism responsible for the obtained result. On the contrary, much  
21  
22 simpler semi-empirical methods, usually containing a few parameters, even though less  
23  
24 accurate, may be in the position to provide a more fundamental understanding of the  
25  
26 electronic structure and its dependence on the physical properties of the system. Excellent  
27  
28 demonstrations of these ideas are provided by the impressive work of Harrison [37], who  
29  
30 was able to account for various properties (ranging from dielectric, to conducting, elastic  
31  
32 etc.) of different categories of solids using such a simple approach, and by Streitwieser  
33  
34 [38], who summarized early efforts along these lines regarding molecular properties.  
35  
36 Inspired by these works, the aim of this article is the evaluation of the electronic structure  
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38 of  $\pi$  molecular orbitals, by using the linear combination of atomic orbitals (LCAO) method  
39  
40 including only  $p_z$ -type orbitals (like in Hückel theory) and a minimal unified set of  
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42 parameters for describing a relatively large number of planar organic molecules.  
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## 53 **2. Theoretical method**

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55 Atoms in planar organic molecules with  $sp^2$  hybridization have their  $p_z$  atomic orbitals  
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57 perpendicular to the molecular plane. The electrons that occupy these orbitals are  
58  
59 eventually delocalized. The LCAO method provides a very simple way to calculate  $\pi$   
60

molecular orbitals, which approximately describe these delocalized electrons. In this approximation the corresponding molecular wavefunction  $\psi(\vec{r})$  is a linear combination of the  $p_z$  atomic orbitals from each atom, or preferably, as it is used in the present work, of atomic-like orbitals  $p$  which resemble the  $p_z$  atomic orbitals:

$$\psi(\vec{r}) = \sum_{i=1}^N c_i p_i(\vec{r}) . \quad (1)$$

The summation index,  $i$ , runs among the  $N$  atoms of the molecule, which contribute  $p_z$  electrons in  $\pi$  bonds. Here we ignore all other orbitals (including the  $sp^2$  hybrids) and consider only the Hamiltonian in the subspace of  $p$  orbitals.

Multiplying the Schrödinger equation,

$$\hat{H} \psi = E \psi , \quad (2)$$

with the conjugate atomic-like orbital  $p_j^*(\vec{r})$  and integrating, we obtain the linear system

$$\sum_{i=1}^N [(H_{ji} - E\delta_{ji})c_i] = 0 , \quad \text{for } j=1,2,\dots,N . \quad (3)$$

Here we have assumed orthogonality of the  $p$  orbitals located in different atoms (this can be achieved by a proper choice of the atomic-like orbitals), i.e.,  $\int p_j^*(\vec{r}) p_i(\vec{r}) d^3r = \delta_{ji}$ , where  $\delta$  is the delta of Kronecker, otherwise the corresponding overlap integral should be included in Eq. (3). The Hamiltonian matrix elements  $H_{ji}$  are given by

$$H_{ji} = \int p_j^*(\vec{r}) \hat{H} p_i(\vec{r}) d^3r . \quad (4)$$

Thus, in this approximation we obtain the coefficients,  $c_i$ , which provide the  $\pi$  molecular orbitals through Eq. (1), and the corresponding energy eigenvalues  $E$  by numerical diagonalization of the Hamiltonian matrix, as it can be seen from Eq. (3). The only information needed in this approach is the values of the matrix elements,  $H_{ij}$ . Since these

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matrix elements are empirically obtained, the explicit knowledge of the atomic-like  $p$  orbitals, constituting the orthogonal basis of the expansion in Eq. (1), is not necessary.

Regarding the diagonal matrix elements,  $H_{ii}$  (known also as Coulomb integrals or on-site energies), depending on the atom in which the index  $i$  is referred, we use the values  $\varepsilon_C = -6.7$  eV for carbon,  $\varepsilon_{N_2} = -7.9$  eV for nitrogen with one electron in the  $p_z$  atomic orbital (i.e. with coordination number 2),  $\varepsilon_{N_3} = -10.9$  eV for nitrogen with two electrons in the  $p_z$  atomic orbital (i.e. with coordination number 3), and  $\varepsilon_O = -11.8$  eV for oxygen. We arrived at these empirical values after a series of simulations of the electronic structure of various organic molecules. Initially we tried to use the ionization energies of the elements C (-11.26 eV), O (-13.62 eV), N (-14.53 eV), as it is usually chosen for the diagonal matrix elements. However, for all the molecules examined, using the ionization energies of the elements led to large deviations from the experimentally known molecular ionization energies of the highest occupied  $\pi$  orbital. Therefore, we used  $\varepsilon_C$  as a free parameter and tried to fit the vertical ionization energy of many cyclic and non-cyclic hydrocarbons. It turned out that the value  $\varepsilon_C = -6.7$  eV, results in good agreement (within less than 13% deviation) between the calculated and the experimental  $\pi$  ionization energies of the investigated molecules. At this point we mention that benzene's  $\pi$  structure can be analytically solved within our approach. From the experimental result of the ionization energy of benzene, one can determine the value of  $\varepsilon_C$  by matching the experimental result with the analytically obtained HOMO energy. Using Eq. (5) (see below) for the interatomic matrix elements in benzene, this procedure yields the value -6.8 eV for  $\varepsilon_C$ . The analytically derived result almost coincides with the empirical value of  $\varepsilon_C = -6.7$  eV used in our work. Next, in order to determine the values of the other on-site energies we fixed  $\varepsilon_C = -6.7$  eV and examined organic molecules containing nitrogen and oxygen atoms,

with emphasis in the nucleic acids' bases, considering  $\varepsilon_{N_2}$ ,  $\varepsilon_{N_3}$ , and  $\varepsilon_O$  as free parameters. A comparison between the calculated energy of the highest occupied  $\pi$  orbital and the corresponding experimental value of vertical ionization energy led to the above mentioned values of  $\varepsilon_{N_2}$ ,  $\varepsilon_{N_3}$ , and  $\varepsilon_O$ . We note that different LCAO diagonal energies for two types of nitrogen atoms, distinguished by their coordination numbers as  $N_2$  and  $N_3$ , i.e. with one or two electrons in the  $p_z$  atomic orbital, have also been used in the literature [38,39].

The nondiagonal (i.e. interatomic) matrix elements  $H_{ij}$  (known also as resonance integrals) are zero if the indices  $i$  and  $j$  refer to atoms without a direct bond between them, while for neighboring bonded atoms we use the expression proposed by Harrison [37]:

$$H_{ij} = V_{pp\pi} = -0.63 \frac{\hbar^2}{md_{ij}^2}, \text{ for } i, j \text{ referring to neighboring bonded atoms. (5)}$$

Here,  $m$  is the electron mass and  $d_{ij}$  is the length of the bond between the atoms  $i, j$ . Harrison' formula is universal and applies to corresponding matrix elements between different elements. The proportionality of the matrix elements to  $1/d^2$  is not valid for arbitrary distances, but only at distances near to the equilibrium interatomic distances in matter. We remark that this expression for  $H_{ij}$  describes the matrix elements between adjacent  $p_z$ -type orbitals under the hypothesis that their overlap is ignored. This is consistent with our previously mentioned assumption in deriving Eq. (3). Harrison's interatomic matrix elements are very popular among physicists because, as we already mentioned, they can successfully describe a large variety of properties of materials within a simple LCAO approximation [37]. Such a dependence of the interatomic matrix elements (proportional to  $1/d^2$ ) has not been used by chemists in the application of LCAO in molecules, where the rather more complicated Wolfsberg-Helmholz expression [40] is

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3 widely applied. However, since the interatomic distances are similar in molecules and  
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5 solids, one expects that Harrison's matrix elements (5) can be also applied in molecules.  
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7 Their advantage, compared to the well-known Wolfsberg-Helmholz interatomic matrix  
8  
9 elements, is that they are considerably simpler and readily applied when the interatomic  
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11 distance (bond length)  $d$  is known. Further, as we show in this work, they can be rather  
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13 successfully used for estimating the energy of the highest occupied  $\pi$  orbital and the first  $\pi$ -  
14  
15  $\pi^*$  transition in a large number of organic molecules. The geometries and the interatomic  
16  
17 distances  $d_{ij}$  in all the theoretically investigated molecules in this work, apart from the  
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19 nucleic acids bases, are obtained from the NIST website [41].  
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25 We stress that the term LCAO, as used in this work, it does not imply that the  
26  
27 values of Coulomb and resonance integrals have been estimated through the use of atomic  
28  
29 orbitals. Instead, these values have been empirically chosen through comparisons with  
30  
31 various experimental data (molecular ionization energies have been used for the on-site  
32  
33 energies and solid-state band structure features for the interatomic matrix elements). Our  
34  
35 LCAO method with the above values of diagonal and nondiagonal matrix elements can be  
36  
37 considered as a type of Hückel model with explicit bond-length dependence of the  
38  
39 resonance integrals.  
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43  
44 The energy eigenvalues obtained from the numerical diagonalization of the  
45  
46 Hamiltonian matrix correspond to the electronic spectrum of  $\pi$  molecular orbitals. Then the  
47  
48 occupied and unoccupied  $\pi$  orbitals of the organic molecule can be found by counting all  
49  
50  $p_z$  electrons contributed by the atoms of the molecule and arrange them successively in  
51  
52 couples of different spin in accordance to Pauli principle. Additionally the  $\pi$ - $\pi^*$  transitions  
53  
54 can be obtained. In some molecules like benzene and polyacenes the HOMO-LUMO gap  
55  
56 (i.e. the energy gap between the highest occupied molecular orbital and the lowest  
57  
58 unoccupied molecular orbital) corresponds to a  $\pi$ - $\pi^*$  transition. However, this is not always  
59  
60

the case, as we discuss later, for example in the case of polyenes or some heterocyclic organic compounds, where the HOMO-LUMO gap is not a  $\pi$ - $\pi^*$  transition.

### 3. Results and Discussion

#### 3.1 Benzene, polyacenes, and aromatic hydrocarbons of various geometries

In Table I we present our results for benzene, polyacenes and a number of aromatic hydrocarbons with many rings and various architectures. In particular, we show the calculated ionization energy of the highest occupied  $\pi$  molecular orbital ( $\pi$  HOMO),  $I\pi_{th}$ , the corresponding experimental value,  $I\pi_{exp}$ , and the respective % deviation,

$$\Delta I\pi = \frac{I\pi_{th} - I\pi_{exp}}{I\pi_{exp}} \times 100, \quad (6)$$

for each organic molecule of the table. Furthermore the calculated energy of the first excited  $\pi^*$  orbital ( $\pi$  LUMO),  $L\pi_{th}$ , is displayed along with the resulting theoretical  $\pi$ - $\pi^*$  energy gap,  $\pi$ - $\pi^*_{th}$ , the experimental one,  $\pi$ - $\pi^*_{exp}$ , and the corresponding % deviation,

$$\Delta\pi\text{-}\pi^* = \frac{(\pi\text{-}\pi^*_{th}) - (\pi\text{-}\pi^*_{exp})}{\pi\text{-}\pi^*_{exp}} \times 100 \quad (7).$$

The experimental values in this and the following tables correspond to vertical ionizations or excitations (i.e. without a change in the structure of the molecule). It must be mentioned that the first  $\pi$ - $\pi^*$  transition in this kind of molecules of Table I corresponds to the HOMO-LUMO gap.

In respect to the ionization energy (the energy that must be given for the evacuation of an electron from the highest occupied  $\pi$  molecular orbital), the LCAO predicted results are in very good agreement with the experimental data. The biggest deviation (12.5%) is in

1  
2  
3 hexacene and the smaller ones in benzene and naphthalene. For larger polyacenes  
4  
5 (anthracene and tetracene) the relative error is between 5-7.5%, while it exceeds 10% for  
6  
7 even larger polyacenes (pentacene and hexacene). In most aromatic hydrocarbons of Table  
8  
9 I with nonlinear architecture the deviation is larger than 8%, but less than 12% (apart from  
10  
11 benzo[p]hexaphene-naphtho(2',3':1,2)pentacene, where it is 12.1%).  
12  
13  
14

15  
16 Regarding the  $\pi$ - $\pi^*$  energy gap, the deviations are larger and in certain cases are  
17  
18 more than 40%. In particular in pentacene, hexacene and benzo[p]hexaphene-  
19  
20 naphtho(2',3':1,2)pentacene the relative errors are 43-45%, 44% and 42%, respectively. In  
21  
22 other six molecules the deviations are between 30-40%, while in the remaining fourteen  
23  
24 from the organic compounds of Table I where experimental data are available, the relative  
25  
26 error is below 30%. The lowest deviation for the  $\pi$ - $\pi^*$  energy gap is found in benzene. In  
27  
28 **general our results underestimate the  $\pi$ - $\pi^*$  gaps (apart from the case of benzene).** It must be  
29  
30 mentioned here that the energy gap decreases for polyacenes as the number of benzene  
31  
32 rings increases. This is due to a wider splitting of the energy states resulting from the  $p_z$   
33  
34 atomic orbitals, as the number of atoms increases. This experimentally verified trend is  
35  
36 captured from our theoretical calculations, even though the corresponding deviations  
37  
38 increase for larger polyacenes.  
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45 A comparison of the calculated first  $\pi$ - $\pi^*$  transition between the used model and  
46  
47 methods from first principles can be made for some molecules of Table I. Particularly for  
48  
49 benzene, the LCAO predicts a HOMO-LUMO gap of 5.0 eV and several first principles  
50  
51 methods 4.84 eV [42], 5.14 eV [29], and 5.24-5.28 eV [30]. For the polyacenes first  
52  
53 principles calculations are more accurate. For example in naphthalene the LCAO gives a  
54  
55 value of 3.2 eV and various methods from first principles predict 4.09-5.27 eV [31], 4.38-  
56  
57 4.88 eV [43], and 4.27 eV [30], which are closer to the experimental value of 3.9-4.0 eV,  
58  
59 except for the extreme values of 4.88 eV and 5.27 eV. Furthermore, a comparison between  
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2  
3 LCAO predictions, experimental data, and first principles calculations presented in Refs.  
4  
5 [30], [31], and [43] for anthracene, tetracene, pentacene, and hexacene confirms this  
6  
7 conclusion.  
8  
9

### 10 11 12 13 *3.2 Polyenes*

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16  
17 The results for several polyenes are shown in Table II. In most of these molecules the  $\pi$   
18  
19 molecular orbitals can be obtained readily analytically in our approach, since there are only  
20  
21 a couple of  $sp^2$  hybridized atoms forming  $\pi$ -bonds. Here, the maximum deviation between  
22  
23 the predicted and experimental  $\pi$ HOMO ionization energy is around 11-12%. The highest  
24  
25 deviations 11.4-12% and 9.5-11.2% are found in 2-butene,2,3-dimethyl and ethylene,  
26  
27 respectively. These are the only deviations among the investigated molecules exceeding  
28  
29 10%, while the others are no more than 8%.  
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34  
35 The theoretical  $\pi$ - $\pi^*$  energy gap exhibits larger deviations from the experimental  
36  
37 value, as happened in the previous subsection. Again there is a systematic underestimate of  
38  
39 the  $\pi$ - $\pi^*$  transition. The lowest deviation is around 19% in 1-propene,2-methyl (isobutene)  
40  
41 and the highest 42%-43% in 1,3,5-hexatriene. Here these higher deviations may result from  
42  
43 the existence of Rydberg states, which interpolate energetically in between the  $\pi$  states.  
44  
45 Note that the actual HOMO-LUMO gap in these molecules is not a  $\pi$ - $\pi^*$  transition, but a  
46  
47 transition from a  $\pi$  state to a Rydberg state. For a better prediction of the electronic spectra  
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49 of these molecules within the LCAO approximation, the higher-energy atomic states can be  
50  
51 included in the atomic orbital expansion of the molecular wavefunction. In the expansion  
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53 used here, for simplicity, the  $\pi$  molecular orbitals were considered as isolated (far  
54  
55 energetically) in respect to the other orbitals and only the  $p_z$  atomic states were included  
56  
57  
58  
59  
60 in the LCAO method. It must be also mentioned that for ethylene there is a dispute on

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2  
3 whether the  $\pi\text{-}\pi^*_{\text{exp}}$  value shown in Table II corresponds to a vertical transition or to a  
4  
5  
6 twisted configuration of the molecule. The latter hypothesis is supported by a number of  
7  
8 theoretical studies of increasing accuracy, which have led to a final estimate of about 8.0  
9  
10 eV for the vertical transition energy [71,72,73].  
11

12  
13 The exclusion of Rydberg states from our consideration explains also why first  
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15 principles methods may provide a better agreement with the observed values when applied  
16  
17 in ethylene, 1,3-butadiene, and 1,3,5-hexatriene. Especially for ethylene such methods  
18  
19 estimate the first  $\pi\text{-}\pi^*$  transitions in the region of 7.97-8.54 eV [74], which is closer to the  
20  
21 experimental value than the present LCAO estimation of 5.5 eV. In 1,3-butadiene the  
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23 experimental value is 5.9 eV and first principles calculations predict 6.12-8.54 eV [74]. By  
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25 using the LCAO, the prediction is 3.6 eV which is less accurate than 6.12 eV but no worse  
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27 than 8.54 eV. For 1,3,5-hexatriene first principles evaluations range between 5.01-7.36 eV  
28  
29 [74]. Again the upper limit of this region is no more accurate than the simple LCAO  
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31 calculation (2.8 eV).  
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### 39 *3.3 Benzene derivatives and azulene*

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42 In the Table III below we report our calculations for some benzene derivatives and for  
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44 azulene. Once more, the evaluated ionization energies of the highest occupied  $\pi$  molecular  
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46 orbital are pretty close to the corresponding experimental values. The highest relative errors  
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48 appear in tetralin and p-xylene (8.2-9.5% and 7.0-9.5% respectively), while in most of the  
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50 remaining cases the relative errors are below 5%.  
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55  
56 Regarding the first  $\pi\text{-}\pi^*$  transition, little information could be found. Particularly  
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58 we were able to find results only for toluene, styrene, p-xylene and azulene. In toluene the  
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60 prediction is close to the experimental value (the relative error is 6%). In p-xylene and

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3 styrene the deviation between theoretical and experimental data is 11% and 18%,  
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5 respectively, while in azulene the relative error is around 28%.  
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### 10 3.4 Organic compounds containing nitrogen atoms 11

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14  
15 Table IV presents the same quantities with the previous tables for cyclic heteroatomic  
16  
17 organic compounds containing nitrogen atoms. The results for the  $\pi$ HOMO ionization  
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19 energies show a relatively good agreement with the experimental values. Three among the  
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21 seven molecules of this table show deviations larger than 10%. The biggest deviation is  
22  
23 17.5% in pyrimidine and the next one is 14.5% in 1,3,5-triazine (or s-triazine). It must be  
24  
25 mentioned here that in some of the molecules of Table IV the HOMO is not a  $\pi$  molecular  
26  
27 orbital, but an antibonding n orbital, originating from the nitrogen atoms contained in these  
28  
29 molecules. Generally, the highest deviations appear in those molecules. Many efforts have  
30  
31 been done in order to clarify whether the HOMO of these molecules is a n or a  $\pi$  state. It  
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33 seems that the HOMO orbital is ordered as a n state in 1,3,5-triazine [103], pyrimidine, and  
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35 pyridazine [54].  
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42 Regarding the  $\pi$ - $\pi^*$  energy gaps, small deviations are obtained using the simple  
43  
44 LCAO method in this family of molecules. For azabenzenes (1,3,5-triazine, pyrimidine,  
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46 pyridazine, pyridine) the deviations do not exceed 6%. Results obtained from first  
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48 principles calculations in azabenzenes are also similar to the experimental ones. Such first  
49  
50 principles predictions are 5.33-5.80 eV for 1,3,5-triazine, 4.93-5.44 eV for pyrimidine,  
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52 4.86-5.31 eV for pyridazine, and 4.84-5.22 eV for pyridine [29]. We see that for  
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54 azabenzenes, the LCAO method, although simple, is almost as accurate as methods from  
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56 first principles for the prediction of the first  $\pi$ - $\pi^*$  transition.  
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For pyrrole, despite the fact that various types of experiments and theoretical investigations have been devoted to its study, a detailed assignment of many transitions has not been achieved, yet. Earlier experiments indicated that the first  $\pi$ - $\pi^*$  transition is at 5.22 eV [15], but later calculations proposed that the specific transition is of different nature [33b], or others defined the first  $\pi$ - $\pi^*$  transition in a higher energy region [33a, 34]. Information about the  $\pi$ - $\pi^*$  transitions of 1H-imidazole and 1H-imidazole,2-methyl could not be found.

To demonstrate the beauty of such a simple schemes, like the one used here, we mention that in the case of 1,3,5-triazine (or s-triazine) the LCAO  $\pi$  molecular electronic structure can be readily obtained analytically, because of its high symmetry; the original  $6 \times 6$  matrix that has to be diagonalized ends up to a  $2 \times 2$  matrix by virtue of the Bloch theorem. Note that due to the symmetry, all interatomic distances are identical in this molecule leading to identical non-diagonal matrix elements  $V_{pp\pi}$ . The analytical derivation of s-triazine's  $\pi$  structure provides a physical insight on the origin of the differences observed in its  $\pi$ - $\pi^*$  transition and ionization energy as compared for example with those of benzene. The  $\pi$ HOMO and  $\pi$ LUMO energies of s-triazine are given by

$$E_{HOMO}^{s\text{-triaz.}} = \varepsilon_{C-N} - \sqrt{(V_{C-N})^2 + (V_{pp\pi}^{s\text{-triaz.}})^2}, \quad E_{LUMO}^{s\text{-triaz.}} = \varepsilon_{C-N} + \sqrt{(V_{C-N})^2 + (V_{pp\pi}^{s\text{-triaz.}})^2},$$

where  $\varepsilon_{C-N} = \frac{\varepsilon_C + \varepsilon_{N_2}}{2} = -7.3$  eV is the mean position between the  $\varepsilon_C$  and  $\varepsilon_{N_2}$  on-site

energies, and  $V_{C-N} = \frac{\varepsilon_C - \varepsilon_{N_2}}{2} = 0.6$  eV is one half of their difference.  $V_{pp\pi}^{s\text{-triaz.}} \approx -2.63$  eV is

the interatomic matrix element obtained from Eq. (5), when substituting s-triazine's interatomic distances  $d_{ij}^{s\text{-triaz.}} \approx 1.35$  Å [41]. The analytically obtained  $\pi$ - $\pi^*$  gap is

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$2\sqrt{(V_{C-N})^2 + (V_{pp\pi}^{s-triaz.})^2} \approx 5.4$  eV, while the ionization energy is  $|E_{HOMO}^{s-triaz.}| \approx 10.0$  eV. The

corresponding quantities for benzene are

$$E_{HOMO}^{benz.} = \varepsilon_C - |V_{pp\pi}^{benz.}|, \quad E_{LUMO}^{benz.} = \varepsilon_C + |V_{pp\pi}^{benz.}|,$$

where  $V_{pp\pi}^{benz.} \approx -2.5$  eV is the interatomic matrix element of Eq. (5) for benzene. The  $\pi$ - $\pi^*$

gap of benzene is  $2|V_{pp\pi}^{benz.}| \approx 5.0$  eV and its ionization energy is  $|E_{HOMO}^{benz.}| \approx 9.2$  eV.

Therefore, we see that 1,3,5-triazine is expected to exhibit a larger  $\pi$ - $\pi^*$  transition than

benzene, due to the smaller interatomic distances in s-triazine (leading to larger resonance

integrals  $V_{pp\pi}$ ), while there is also a smaller contribution arising from the difference

between carbon and nitrogen on-site energies (the term  $V_{C-N}$ ). The ionization energy of s-

triazine is larger than that of benzene, mainly due to the fact that nitrogen has a lower on-

site energy than carbon (compare the term  $\varepsilon_{C-N}$  in  $E_{HOMO}^{s-triaz.}$  with  $\varepsilon_C$  in  $E_{HOMO}^{benz.}$ ). There is

also a smaller contribution to the increased ionization energy of s-triazine resulting from

the same reasons that lead to a larger  $\pi$ - $\pi^*$  gap (see the terms  $\sqrt{(V_{C-N})^2 + (V_{pp\pi}^{s-triaz.})^2}$  and

$|V_{pp\pi}^{benz.}|$  in  $E_{HOMO}^{s-triaz.}$  and  $E_{HOMO}^{benz.}$ , respectively).

### 3.5 Organic compounds containing oxygen atoms

Table V includes the theoretical results and the corresponding experimental data for cyclic and linear heteroatomic organic compounds containing oxygen atoms. For the non-cyclic molecules of this table the LCAO results for  $\pi$  orbitals can be readily obtained analytically because only two atoms are involved in  $\pi$ -bonding.

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4 Regarding the  $\pi$ HOMO ionization energies of the depicted molecules the  
5 theoretical results do not differ more than 15% from the experimental values. The highest  
6 deviation is 14.5% in 2-pentanone, while for four molecules of Table V the deviation  
7 between the theoretical prediction and the experimental result does not exceed 10%.  
8 Similarly to the previous subsection 3.4, many molecules of Table V have an n orbital as  
9 HOMO. Specifically, this is known to be the case in acetone [109], acetaldehyde [110], 2-  
10 pentanone [111], and p-benzoquinone [112].

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21 Looking at the  $\pi$ - $\pi^*$  energy gaps, larger deviations are obtained and our results  
22 underestimate the experimental values. In acetaldehyde there is a relatively small deviation  
23 of around 9%. The relative errors in p-benzoquinone and 2,4-cyclopentadiene-1-one, which  
24 are the highest obtained in respect to all tables in this work, are about 49% and 52%,  
25 respectively. A possible reason for which the LCAO method fails to predict the first  $\pi$ - $\pi^*$   
26 transition in 2,4-cyclopentadiene-1-one, is the strong interaction of  $\pi$  orbitals with near  
27 degenerate states [36], which are ignored in the present treatment. Regarding p-  
28 benzoquinone two n states are found between the highest  $\pi$  and the lowest  $\pi^*$  orbitals, and  
29 therefore the simple LCAO method used here is not able to give accurate results [113]. We  
30 mention at this point that the first  $\pi$ - $\pi^*$  transition is optically forbidden in this molecule  
31 [113]. For acetone, the first  $\pi$ - $\pi^*$  transition has not yet been clarified, due to the coupling of  
32  $\pi$  orbitals with Rydberg or n states [114]. Information could not be found for the molecules  
33 of 2-pentanone and 2,4-cyclohexadien-1-one,6-methylene.

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51 To give another example of physical insight gained by analytical relations, like in  
52 the previous subsection, we discuss the cases of acetaldehyde and acetone and compare  
53 them with propene and 1propene,2methyl, respectively. Regarding the  $\pi$  structure,  
54 acetaldehyde and acetone (propene and 1propene,2methyl) can be treated similarly, since  
55 the  $\pi$  system contains only the C=O (C=C) bond, i.e. just two atoms. Moreover the bond  
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lengths are almost the same in both molecules,  $d_{C=O}^{acetald.} \approx d_{C=O}^{acetone} \approx 1.21 \text{ \AA}$  ( $d_{C=C}^{propene} \approx d_{C=C}^{1propene,2meth.} \approx 1.33 \text{ \AA}$ ) [41], leading to identical  $\pi$  electronic structures within our method. The  $2 \times 2$  Hamiltonian matrix can be trivially diagonalized in these cases. The  $\pi$ HOMO and  $\pi$ LUMO energies of acetone and acetaldehyde are

$$E_{HOMO}^{acet.} = \varepsilon_{C-O} - \sqrt{(V_{C-O})^2 + (V_{pp\pi}^{acet.})^2}, \quad E_{LUMO}^{acet.} = \varepsilon_{C-O} + \sqrt{(V_{C-O})^2 + (V_{pp\pi}^{acet.})^2},$$

where  $\varepsilon_{C-O} = \frac{\varepsilon_C + \varepsilon_O}{2} = -9.25 \text{ eV}$ ,  $V_{C-O} = \frac{\varepsilon_C - \varepsilon_O}{2} = 2.55 \text{ eV}$ , and

$$V_{pp\pi}^{acet.} = -0.63 \frac{\hbar^2}{m(d_{C=O}^{acet.})^2} \approx -3.3 \text{ eV}. \text{ The } \pi\text{-}\pi^* \text{ transition is } 2\sqrt{(V_{C-O})^2 + (V_{pp\pi}^{acet.})^2} \approx 8.3 \text{ eV}$$

and the  $\pi$  ionization energy is  $|E_{HOMO}^{acet.}| \approx 13.4 \text{ eV}$ . The results for propene and 1propene,2methyl are

$$E_{HOMO}^{prop.} = \varepsilon_C - |V_{pp\pi}^{prop.}|, \quad E_{LUMO}^{prop.} = \varepsilon_C + |V_{pp\pi}^{prop.}|,$$

where  $V_{pp\pi}^{prop.} = -0.63 \frac{\hbar^2}{m(d_{C=C}^{prop.})^2} \approx -2.7 \text{ eV}$ . The  $\pi\text{-}\pi^*$  gap is  $2|V_{pp\pi}^{prop.}| \approx 5.4 \text{ eV}$  and the

ionization energy  $|E_{HOMO}^{prop.}| \approx 9.4 \text{ eV}$ . These analytical results are similar to those presented in the previous subsection. They show for example that acetaldehyde (or acetone) has a much larger  $\pi\text{-}\pi^*$  transition than propene (or 1propene,2methyl) due to significant contributions from both the smaller bond length (i.e. the larger  $V_{pp\pi}$  value) and the very different on-site energies of  $\varepsilon_C$  and  $\varepsilon_O$  (the  $V_{C-O}$  term). Regarding the ionization energies of these molecules, similar conclusions can be drawn here, as in the comparison between s-triazine and benzene examined in subsection 3.4. However because the quantities  $|\varepsilon_{C-O} - \varepsilon_C|$ ,  $V_{C-O}$ , and  $|V_{pp\pi}^{acet.} - V_{pp\pi}^{prop.}|$ , are much larger here than the corresponding ones,  $|\varepsilon_{C-N} - \varepsilon_C|$ ,  $V_{C-N}$ , and  $|V_{pp\pi}^{s-triaz.} - V_{pp\pi}^{benz.}|$ , of the previous subsection, more important

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3 differences in the  $\pi$ - $\pi^*$  transitions and ionization energies are expected now, as compared  
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6 to those between 1,3,5-triazine and benzene. This indeed is confirmed by the experiment.  
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### 10 3.6 Nucleic acids bases

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15 Molecules of biological interest like the DNA and RNA bases are included in Table VI.  
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17 Starting the discussion from the ionization energies of the highest occupied  $\pi$  orbital, we  
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19 notice that good agreement with the experiment is obtained. The larger deviation is found  
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21 in uracil (5.3-7.2%). In adenine and thymine the relative errors are smaller (2.4-3.5% and  
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23 0-2.2%, respectively), while in guanine and cytosine the present theoretical results are  
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25 identical or almost identical with the observed values.  
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29 Concerning the energy of the first  $\pi$ - $\pi^*$  transition, relatively small deviations are  
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31 obtained. In adenine and guanine appear the larger relative errors, 16-21% and 12-16%,  
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33 respectively. The lowest deviation is in cytosine, where the present LCAO prediction  
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35 coincides with some of the experimental observations. Comparison between the  
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37 experimental data and results from several first principles methods for all molecules of  
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39 Table VI shows that the simple method used here is not much worse than the latter  
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41 theoretical methods. In the case of adenine, the energy of the first  $\pi$ - $\pi^*$  transition, as  
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43 predicted by several methods from first principles, is found in the range of 4.97-5.13 eV  
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45 [122, 123, 124]. This overestimates the experimental value (4.5-4.8 eV) and it is only  
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47 slightly better than the underestimated value predicted by our simple LCAO method (3.8  
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49 eV). Similar is the situation for guanine, where first principles methods evaluate the first  $\pi$ -  
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51  $\pi^*$  energy in the region 4.76-4.96 eV [122, 125], overestimating the experimental result  
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53 (4.3-4.5 eV), while the LCAO prediction is 3.8 eV. Regarding the molecule of cytosine the  
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55 results from first principles methods are 4.39-4.71 eV [126, 127, 128], which are in very  
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3 good agreement with the experiment (4.5-4.7 eV), as it is also the case for the accurate  
4 LCAO prediction (4.5 eV). Finally for the molecules of thymine and uracil methods from  
5 first principles predict the first  $\pi$ - $\pi^*$  transition in a region of 4.75-5.17 eV [125, 129, 130]  
6 and 4.82-5.44 eV [129, 130, 131, 132], respectively. The lower values in these regions are  
7 in better agreement with the experimental results, comparing to the LCAO method, but this  
8 is not the case for the higher predicted energies. It must be mentioned also that for these  
9 two molecules first principles calculations predict that the HOMO-LUMO transition is a  $n$ -  
10  $\pi^*$  transition. This does not seem to agree with the earlier established general acceptance  
11 that in all these five DNA and RNA bases the HOMO and LUMO are  $\pi$  orbitals [133,134].  
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#### 27 4. Conclusions

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30 Using the simplest form of the LCAO approximation, which takes into account only  $p_z$   
31 atomic-like orbitals, and a minimal set of unified parameters (four for the diagonal matrix  
32 elements corresponding to atoms of carbon ( $\varepsilon_C = -6.7$  eV), nitrogen with coordination  
33 number 2 ( $\varepsilon_{N_2} = -7.9$  eV), nitrogen with coordination number 3 ( $\varepsilon_{N_3} = -10.9$  eV), and  
34 oxygen ( $\varepsilon_O = -11.8$  eV), and the interatomic matrix elements between first neighboring  
35  $p_z$ -type atomic orbitals, Eq. (5), proposed by Harrison [37]), the  $\pi$  molecular electronic  
36 structure of more than sixty planar organic molecules with  $sp^2$  hybridized atoms has been  
37 evaluated. The energies of the  $\pi$ HOMO states and the  $\pi$ HOMO- $\pi$ LUMO gaps have been  
38 compared with experimental data.  
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53 The choice of the values of these four empirical parameters has been obtained  
54 through optimization in respect to the ionization energy of the highest occupied  $\pi$   
55 molecular orbital. In particular, the value of  $\varepsilon_C$  has been obtained first, by considering the  
56 molecules presented in Tables I-III (46 molecules totally). The resulting optimized value is  
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3 almost the same as that obtained analytically considering the  $\pi$ HOMO of benzene. Then,  
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5 keeping fixed this value of  $\varepsilon_C$ , the remaining values of  $\varepsilon_{N_2}$ ,  $\varepsilon_{N_3}$ , and  $\varepsilon_O$  have been  
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8 obtained considering the molecules presented in Tables IV-VI (19 molecules totally), with  
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10 emphasis on the four DNA bases of Table VI. We mention that such an optimization is not  
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12 a demanding computational process due to the simplicity of the method.  
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16 Our theoretical calculations predict the experimental value of the  $\pi$ HOMO energy  
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18 of sixty five organic compounds with a relative error of less than 15% in all cases  
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20 examined, except for pyrimidine, where the deviation is 17.5%. Regarding the first  $\pi$ - $\pi^*$   
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22 transition, the deviations from the experimental observations are larger, but no more than  
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24 about 50%, although there is not any adjustable parameter in this case (we emphasize that  
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26 the optimization for obtaining the empirical parameters did not include the  $\pi$ - $\pi^*$  gaps). In  
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28 forty five from the investigated molecules, experimental data for the first  $\pi$ - $\pi^*$  transition  
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30 were available. In most of the cases our results underestimate the  $\pi$ - $\pi^*$  gaps. In one case the  
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32 deviation from the theoretical prediction is around 52%, in other five cases the deviation is  
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34 between 40-50%, while in the other cases the relative error is below 40%.  
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40 Taking into account the simplicity of the method in respect to more accurate first  
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42 principles calculations, the minimal number of parameters used, and the complicated actual  
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44 electronic structure in some of the investigated cases (appearance of both  $\pi$  and Rydberg or  
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46 n states in the vicinity of HOMO and LUMO), it seems that the LCAO approach presented  
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48 here provides a relatively accurate tool for a quick and easy derivation of theoretical  
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50 estimates concerning the  $\pi$  electronic structure (at least the  $\pi$ HOMO and  $\pi$ LUMO states of  
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52 interest) for planar organic molecules, which apart from carbon and hydrogen may also  
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54 include nitrogen and/or oxygen atoms. This model computationally requires just a trivial  
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56 diagonalization and can be easily used from no-specialists in hard theoretical or numerical  
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58 calculations. We mention also that some cases can be treated even analytically within our  
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3 approach. Compared to earlier Hückel approaches [38], our method offers improved  
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5 predictions, at least for the molecular properties examined here.  
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8 The rather accurate description of the HOMO and LUMO energies of DNA bases  
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10 (the relative error is no more than 3.5% for HOMO orbitals and does not exceed 21% for  
11  
12 the HOMO-LUMO transitions) suggests that the obtained bases' wavefunctions can be  
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14 used for estimating interbase coupling parameters (using appropriate atomic matrix  
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16 elements [150] in a Slater-Koster type of coupling [151]), which are relevant for hole or  
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18 electron transfer between DNA bases [152]. Such parameters can be used in  
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20 phenomenological tight-binding descriptions of charge transfer along DNA [153]. Another  
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22 problem of biological interest which is related with the nature of  $\pi$  molecular orbitals in  
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24 planar organic groups and could be possibly investigated within our approach [154],  
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26 concerns the lowest excited state of flavin in the FADH<sup>-</sup> cofactor of the enzyme photolyase,  
27  
28 which is involved in radiative DNA damage repair [155].  
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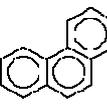
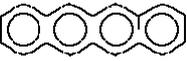
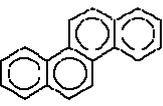
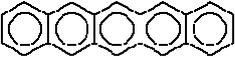
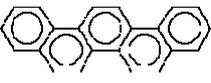
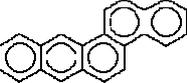
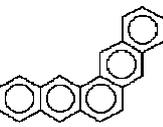
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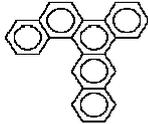
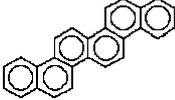
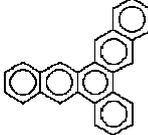
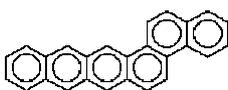
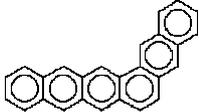
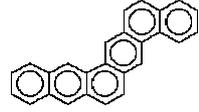
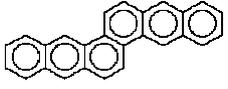
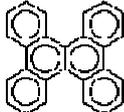
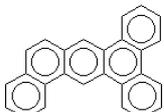
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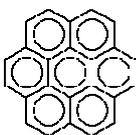
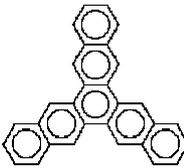
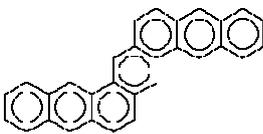
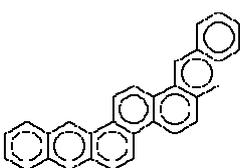
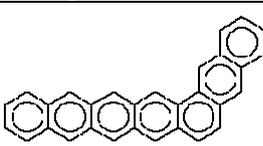
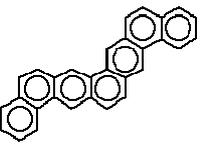
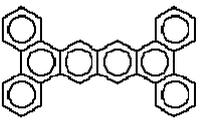
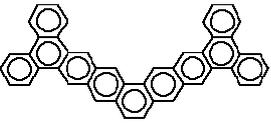
**Table I.** Benzene, polyacenes, and aromatic hydrocarbons of various geometries. The first column depicts the organic molecule. The second and the third column present the experimental and theoretical, respectively, ionization energy of the highest occupied  $\pi$  molecular orbital. The fourth column has the % relative error between the experimental and theoretical values. In the fifth column the evaluated energy of the first excited  $\pi^*$  orbital is shown. The sixth and the seventh columns include the theoretical and the experimental  $\pi$ - $\pi^*$  energy gap, respectively. The last column presents the % deviation between the calculated and experimental  $\pi$ - $\pi^*$  gaps.

Organic molecule	$I\pi_{th}$ (eV)	$I\pi_{exp}$ (eV)	$\Delta I\pi$ (%)	$L\pi_{th}$ (eV)	$\pi$ - $\pi^*_{th}$ (eV)	$\pi$ - $\pi^*_{exp}$ (eV)	$\Delta \pi$ - $\pi^*$ (%)
 Benzene	9.2	9.2-9.3 <sup>[44-55]</sup>	(-) 0-1.1	-4.2	5.0	4.7 <sup>[56]</sup> -4.9 <sup>[57]</sup>	(+) 2-6
 Naphthalene	8.3	8.1-8.3 <sup>[46,49,50,53,58-61]</sup>	(+) 0-2.5	-5.1	3.2	3.9 <sup>[56]</sup> -4.0 <sup>[57]</sup>	(-) 18-20
 Anthracene	7.8	7.4 <sup>[49,50,60,62,63]</sup>	(+) 5.4	-5.6	2.2	3.3 <sup>[56,57]</sup>	(-) 33
 Phenanthrene	8.3	7.9 <sup>[49,50,62,64-66]</sup>	(+) 5.1	-5.1	3.2	3.5 <sup>[9]</sup>	(-) 9
 Naphthacene (or Tetracene)	7.5	7.0 <sup>[49,67]</sup>	(+) 7.1	-5.9	1.6	2.6 <sup>[56,57]</sup>	(-) 38
 Chrysene	8.1	7.6 <sup>[49,50,67]</sup>	(+) 6.6	-5.3	2.8	3.4 <sup>[56]</sup>	(-) 18
 Pentacene	7.3	6.6 <sup>[49]</sup>	(+) 10.6	-6.1	1.2	2.1 <sup>[57]</sup> -2.2 <sup>[56]</sup>	(-) 43-45
 Picene	8.1	7.5 <sup>[49,50]</sup>	(+) 8.0	-5.3	2.8	3.3 <sup>[68]</sup>	(-) 15
 3,4-benzotetraphene	7.8	7.2 <sup>[49]</sup>	(+) 8.3	-5.6	2.2	3.3 <sup>[68]</sup>	(-) 33
 Pentaphene	7.9	7.3 <sup>[49,50]</sup>	(+) 8.2	-5.5	2.4	-	-

(Table I, continued)

Organic molecule	$I\pi_{th}$ (eV)	$I\pi_{exp}$ (eV)	$\Delta I\pi$ (%)	$L\pi_{th}$ (eV)	$\pi-\pi^*_{th}$ (eV)	$\pi-\pi^*_{exp}$ (eV)	$\Delta \pi-\pi^*$ (%)
 Hexacene	7.2	6.4 <sup>[49]</sup>	(+) 12.5	-6.2	1.0	1.8 <sup>[69]</sup>	(-) 44
 Naphtho[2,3-g]chrysene	7.9	7.2 <sup>[49]</sup>	(+) 9.7	-5.5	2.4	3.1-3.3 <sup>[68]</sup>	(-) 23-27
 Benzo[c]picene-fulminene	8.0	7.2 <sup>[50]</sup>	(+) 11.1	-5.4	2.6	3.2 <sup>[68]</sup>	(-) 19
 6,7benzopentaphene- Benzo[h]pentaphene	8.0	7.4 <sup>[49]</sup>	(+) 8.1	-5.4	2.6	3.2 <sup>[68]</sup>	(-) 19
 Naphtho[2,1a]naphthacene- naphtho(2',1':1,2)tetracene	7.6	6.8 <sup>[68]</sup>	(+) 11.8	-5.9	1.7	2.7 <sup>[68]</sup>	(-) 37
 Hexaphene	7.6	6.9 <sup>[68]</sup>	(+) 10.1	-5.8	1.8	2.7 <sup>[68]</sup>	(-) 33
 3,4benzopentaphene- Benzo[c]pentaphene	7.9	7.2 <sup>[68]</sup>	(+) 9.7	-5.5	2.4	3.0 <sup>[68]</sup>	(-) 20
 Dibenzo[b,k]chrysene- anthraceno(2',1':1,2)anthracene	7.7	7.0 <sup>[68]</sup>	(+) 10.0	-5.7	2.0	3.0 <sup>[68]</sup>	(-) 33
 Dibenzo[g,p]chrysene	8.0	7.2 <sup>[49]</sup>	(+) 11.1	-5.4	2.6	-	-
 Tribenz[a,c,h]Anthracene	8.1	7.4 <sup>[49,50]</sup>	(+) 9.5	-5.3	2.8	-	-

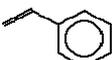
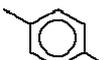
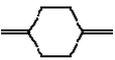
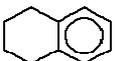
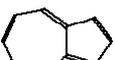
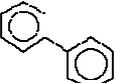
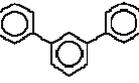
(Table I, continued)

Organic molecule	$I\pi_{th}$ (eV)	$I\pi_{exp}$ (eV)	$\Delta I\pi$ (%)	$L\pi_{th}$ (eV)	$\pi-\pi^*_{th}$ (eV)	$\pi-\pi^*_{exp}$ (eV)	$\Delta \pi-\pi^*$ (%)
 Coronene	8.1	7.3 <sup>[50,70]</sup>	(+) 11.0	-5.3	2.8	3.0 <sup>[56]</sup>	(-) 7
 Trinaphthylene-naphtho (2',3':6,7)Pentaphene	8.1	7.4 <sup>[68]</sup>	(+) 9.5	-5.3	2.8	3.1 <sup>[68]</sup>	(-) 10
 Naphtho[2,3c]Pentaphene- Naphtha(2',3':3,4)pentaphene	7.8	7.0 <sup>[68]</sup>	(+) 11.4	-5.6	2.2	2.9 <sup>[68]</sup>	(-) 24
 Dibenzo[b,n]picene-2,3:8,9- dibenzopicene	7.8	7.2 <sup>[68]</sup>	(+) 8.3	-5.6	2.2	2.9 <sup>[68]</sup>	(-) 24
 Benzo[p]hexaphene-naphtho (2',3':1,2)pentacene	7.4	6.6 <sup>[68]</sup>	(+) 12.1	-6.0	1.4	2.4 <sup>[68]</sup>	(-) 42
 Dibenzo[c,m]pentaphene	7.8	7.1 <sup>[49,50]</sup>	(+) 9.9	-5.6	2.2	-	-
 Tetrabenzo[a,c,j,l]naphthacene	7.8	7.0 <sup>[50]</sup>	(+) 11.4	-5.6	2.2	-	-
 Tetrabenzo[a,c,g,s]heptaphene	7.7	6.9 <sup>[49]</sup>	(+) 11.6	-5.7	2.0	2.6 <sup>[68]</sup>	(-) 23

**Table II.** Polyenes. The columns represent the same quantities as in Table I.

Organic molecule	$I\pi_{th}$ (eV)	$I\pi_{exp}$ (eV)	$\Delta I\pi$ (%)	$L\pi_{th}$ (eV)	$\pi-\pi^*_{th}$ (eV)	$\pi-\pi^*_{exp}$ (eV)	$\Delta\pi-\pi^*$ (%)
 Ethylene	9.5	10.5-10.7 <sup>[47, 75-78]</sup>	(-) 9.5-11.2	-4.0	5.5	7.6 <sup>[75]</sup> -7.7 <sup>[79]</sup>	(-) 28-29
 Propene	9.4	9.7-10.2 <sup>[47, 75-77, 80, 81]</sup>	(-) 3.1-7.8	-4.0	5.4	7.2 <sup>[82]</sup>	(-) 25
 1-propene, 2-methyl	9.4	9.4-9.5 <sup>[80, 83]</sup>	(-) 0-1.1	-4.0	5.4	6.7 <sup>[84]</sup>	(-) 19
 2-butene	9.4	9.1 <sup>[45]</sup>	(+) 3.3	-4.0	5.4	-	-
 2-butene, 2,3-dimethyl	9.3	8.3-10.5 <sup>[45, 85]</sup>	(±) 11.4-12.0	-4.1	5.2	-	-
 1,3-butadiene	8.5	9.0 <sup>[86]</sup> -9.1 <sup>[75]</sup>	(-) 5.6-6.6	-4.9	3.6	5.9 <sup>[87, 88]</sup>	(-) 39
 1,3,5-hexatriene	8.1	8.3 <sup>[75]</sup>	(-) 2.4	-5.3	2.8	4.8 <sup>[56]</sup> -4.9 <sup>[16, 89]</sup>	(-) 42-43

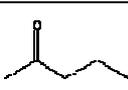
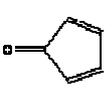
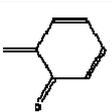
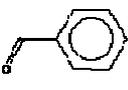
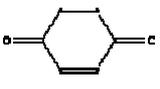
**Table III.** Benzene derivatives and azulene. The columns show the same quantities as in Table I.

Organic molecule	$I\pi_{th}$ (eV)	$I\pi_{exp}$ (eV)	$\Delta I\pi$ (%)	$L\pi_{th}$ (eV)	$\pi-\pi^*_{th}$ (eV)	$\pi-\pi^*_{exp}$ (eV)	$\Delta\pi-\pi^*$ (%)
 Toluene	9.2	8.8-9.0 <sup>[44,47,48,51,52,90]</sup>	(+) 2-4.5	-4.2	5.0	4.7 <sup>[91]</sup>	(+) 6
 Styrene	8.5	8.5-8.6 <sup>[44,53,61,92]</sup>	(-) 0-1.2	-4.9	3.6	4.4 <sup>[91,93]</sup>	(-) 18
 p-xylene	9.2	8.4-8.6 <sup>[44,52,94]</sup>	(+) 7.0-9.5	-4.2	5.0	4.5 <sup>[91]</sup>	(+) 11
 Benzene, 1,2-dimethyl	9.2	8.6-8.8 <sup>[44,52,59,95]</sup>	(+) 4.5-7.0	-4.2	5.0	-	!
 Benzene, 1,4-diethyl	8.2	8.4 <sup>[96]</sup>	(-) 2.4	-5.2	3.0	-	!
 1,4-cyclohexadiene, 3,6-bis(methylene)-	7.7	7.9 <sup>[97]</sup>	(-) 2.5	-5.7	2.0	-	!
 Tetralin	9.2	8.4 <sup>[95]</sup> -8.5 <sup>[98]</sup>	(+) 8.2-9.5	-4.2	5.0	-	!
 1-ethylnaphthalene	8.1	7.9-8.0 <sup>[99]</sup>	(+) 1.3-2.5	-5.3	2.8	-	!
 Azulene	7.9	7.4 <sup>[59,100]</sup>	(+) 6.8	-5.6	2.3	1.8 <sup>[101]</sup>	(+) 28
 Biphenyl	8.5	8.3 <sup>[65]</sup> - 8.4 <sup>[64]</sup>	(+) 1.2-2.4	-4.9	3.6	-	!
 m-terphenyl	8.4	8.1 <sup>[102]</sup>	(+) 3.7	-5.0	3.4	-	!

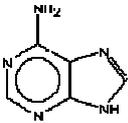
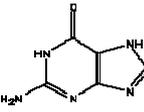
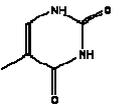
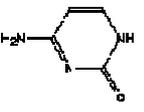
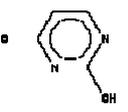
**Table IV.** Organic compounds containing nitrogen atoms. The columns represent the same quantities as in Table I.

Organic molecule	$I\pi_{th}$ (eV)	$I\pi_{exp}$ (eV)	$\Delta I\pi$ (%)	$L\pi_{th}$ (eV)	$\pi-\pi^*_{th}$ (eV)	$\pi-\pi^*_{exp}$ (eV)	$\Delta \pi-\pi^*$ (%)
 1,3,5-triazine	10.0	11.7 <sup>[103]</sup>	(-) 14.5	-4.6	5.4	5.6 <sup>[56]</sup> -5.7 <sup>[104]</sup>	(-) 4-5
 Pyrimidine	9.4	11.4 <sup>[54]</sup>	(-) 17.5	-4.6	4.8	5.1 <sup>[56,104]</sup>	(-) 6
 Pyridazine	9.4	10.5 <sup>[54]</sup>	(-) 10.5	-4.6	4.8	4.9 <sup>[105]</sup> -5.0 <sup>[56,105]</sup>	(-) 2-4
 Pyridine	9.2	9.7 <sup>[54]</sup> -9.8 <sup>[106]</sup>	(-) 5.2-6.1	-4.5	4.7	4.8 <sup>[56]</sup> -5.0 <sup>[104]</sup>	(-) 2-6
 Pyrrole	8.3	8.2 <sup>[107]</sup>	(+) 1.2	-3.8	4.5	-	-
 1H-imidazole	8.4	8.8 <sup>[107]</sup> -9.0 <sup>[108]</sup>	(-) 4.5-6.7	-3.8	4.6	-	-
 1H-imidazole,2methyl	8.4	8.5 <sup>[108]</sup>	(-) 1.2	-3.9	4.5	-	-

**Table V.** Organic compounds containing oxygen atoms. The columns represent the same quantities as in Table I.

Organic molecule	$I\pi_{th}$ (eV)	$I\pi_{exp}$ (eV)	$\Delta I\pi$ (%)	$L\pi_{th}$ (eV)	$\pi-\pi^*_{th}$ (eV)	$\pi-\pi^*_{exp}$ (eV)	$\Delta \pi-\pi^*$ (%)
 Acetone	13.4	12.6 <sup>[109]</sup>	(+) 6.3	-5.1	8.3	-	-
 Acetaldehyde	13.4	13.2 <sup>[110]</sup>	(+) 1.5	-5.1	8.3	9.1 <sup>[115]</sup>	(-) 9
 2-pentanone	13.4	11.7 <sup>[111]</sup>	(+) 14.5	-5.1	8.3	-	-
 2,4-cyclopentadiene-1-one	8.5	9.5 <sup>[116]</sup>	(-) 10.5	-6.9	1.6	3.3 <sup>[117]</sup>	(-) 52
 2,4-cyclohexadien-1-one, 6-methylene	8.9	8.8 <sup>[118]</sup>	(+) 1.1	-6.3	2.6	-	-
 Benzaldehyde	9.2	9.6 <sup>[58]</sup> -9.8 <sup>[119]</sup>	(-) 4.2-6.1	-5.7	3.5	4.4 <sup>[120]</sup>	(-) 20
 p-benzoquinone	9.4	10.9 <sup>[112]</sup>	(-) 13.8	-7.3	2.1	4.1 <sup>[121]</sup>	(-) 49

**Table VI.** Nucleic acids bases. The columns represent the same quantities as in Table I.

Organic molecule	$I\pi_{th}$ (eV)	$I\pi_{exp}$ (eV)	$\Delta I\pi$ (%)	$L\pi_{th}$ (eV)	$\pi-\pi^*_{th}$ (eV)	$\pi-\pi^*_{exp}$ (eV)	$\Delta \pi-\pi^*$ (%)
 Adenine	8.2	8.4 <sup>[133]</sup> -8.5 <sup>[134]</sup>	(-) 2.4-3.5	-4.4	3.8	4.5-4.8 [135-141]	(-) 16-21
 Guanine	8.2	8.2-8.3 <sup>[133-134]</sup>	(-) 0-1.2	-4.4	3.8	4.3-4.5 [135-137, 139, 142]	(-) 12-16
 Thymine	9.0	9.0-9.2 <sup>[133,143,144]</sup>	(-) 0-2.2	-4.8	4.2	4.6-4.7 [135,136,138-141, 145]	(-) 9-11
 Cytosine	8.9	8.9 <sup>[133]</sup>	0	-4.4	4.5	4.5-4.7 [136,137,139,140, 146-148]	(-) 0-4
 Uracil	9.0	9.5-9.7 <sup>[133,143,144,149]</sup>	(-) 5.3-7.2	-4.8	4.2	4.8 <sup>[137]</sup>	(-) 13