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((7008 words))

Catalase-like activity of bovine met-hemoglobin: interaction with the pseudo-catalytic peroxidation of anthracene traces in aqueous medium

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Keywords: Enzyme catalysis, hemoglobin, bioremediation process, PAH, promiscuous activity

Abbreviations:

- ABTS 2,2'-azino-bis(3-ethylbenzthiazoline-6-sulphonic acid
- ANT Anthracene
- CYP Cytochrome P450
- Hb Hemoglobin
- HRP Horeseradish peroxidase
- Met-Hb Met-hemoglobin
- NADPH Nicotinamide adenine dinucleotide phosphate
- PAH Polycyclique Aromatic Hydrocarbons
- UFLC Ultra fast liquid chromatography

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3 Hemoglobin is a member of hemoprotein superfamily whose main role is to transport
4 O_2 in vertebrate organisms. It has also two known promiscuous enzymatic activities,
5 namely peroxidase and oxygenase. Here we show for the first time that bovine
6 hemoglobin also presents a catalase-like activity characterized by a V_{max} of 344
7 $\mu M/min$, a K_M of 24 mM and a k_{cat} equal to 115 min^{-1} . For high anthracene and
8 hemoglobin concentrations and low hydrogen peroxide concentrations, this activity
9 inhibits the expected oxidation of anthracene, which occurs through a peroxidase-like
10 mechanism. Anthracene belongs to the polycyclic aromatic hydrocarbons family
11 whose members are carcinogenic and persistent pollutants found in industrial waste
12 waters. Our results show that anthracene oxidation by hemoglobine and hydrogen
13 peroxide follows a typical bi-bi ping-pong mechanism with a V_{max} equal to 0.250
14 $\mu M/min$, $K_{M(H_2O_2)}$ of 80 μM , $K_{M(ANT)}$ of 1.1 μM and k_{cat} of 0.17 min^{-1} . The oxidation of
15 anthracene is shown to be pseudo-catalytic because an excess of hemoglobin and
16 hydrogen peroxide is required to make PAH completely disappear. Thus bovine
17 hemoglobin presents, in different degrees, all the catalytic activities of the
18 hemoprotein group which makes it a very interesting protein for biotechnological
19 processes and with which one can study structure-activity relationships.
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1. Introduction

Hemoglobin (Hb) is a well-known protein. It is a heterotetramer that consists of two alpha and two beta subunits (α_2 / β_2), each one containing the prosthetic group protoporphyrin IX [1]. This peculiar structure allows Hb to bind, transport and release oxygen in all vertebrate organisms. Depending on the iron atom oxidation state, different types of Hb can be found, notably oxy-, deoxy- and met-hemoglobin [2, 3]. The met-hemoglobin (met-Hb), produced *in vivo* through an autoxidation process, is characterized by a ferric (Fe^{III}) state of the hemin and is unable to bind oxygen.

Hb belongs to the hemoprotein super family which includes transport proteins such as globin, responsible for O_2 transport and storage, cytochrome, which is responsible for electron transport and nitrophorine, a protein responsible for nitric oxide transport [4] and also three catalytic proteins (enzyme) classes named heme-thiolates, peroxidases and catalases.

- Heme-thiolate (EC 1.14) is the recommended collective name for the class of proteins that includes the cytochrome P450 enzymes (CYP). The most common reaction catalyzed by these enzymes is monooxygenation, defined as the insertion of one atom of oxygen into a substrate from molecular oxygen, with the concomitant release of water [5, 6, 7, 8] (Figure 1).
- Peroxidases (EC 1.11.1) use hydrogen peroxide (H_2O_2) as an electron acceptor to catalyze oxidative reactions [9, 10, 11, 12, 13] (Figure 1).
- Catalases, which belong to a sub group of peroxidase (EC 1.11.1.6), catalyze the disproportionation reaction of H_2O_2 . In this case, H_2O_2 acts both as an electron acceptor and an electron donor at the same time. The reaction yields molecular oxygen (O_2) and water [14] (Figure 1).

The hemoprotein family is a very interesting group of proteins for structure-activity relationship studies because the active center, iron porphyrin, is the same for all of the members, and the different activities arise from structural and conformational properties of the protein [15].

Most of the hemoprotein family members exhibit multiple types of activity. In 1994, Sun *et al.* pointed out that chloroperoxidase from *Caldariomyces fumago* (EC 1.11.1.10) had a catalase-like activity [16]. In 2001, Hernandez-Ruiz *et al.* demonstrated the same phenomenon for the horseradish peroxidase (HRP; EC 1.11.1.7) [17]. On the other hand catalases also show peroxidase activities [18]. In 1979, a new group of hemoprotein was discovered and named catalase-peroxidase (KatGs) [19]. KatGs are found in prokaryotes and fungi. They display a predominant catalase activity and a weak peroxidase activity with a broad specificity [20]. Nevertheless, on the basis of their primary sequence KatGs have been recognized to be part of the class I peroxidase super family [21]. All these elements tend to show that different hemoprotein activities may coexist, and is highly dependent on their tridimensional structure and on experimental conditions.

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3 Although the principal function of Hb is to transport oxygen, it has been shown
4 that this protein has other functions [22], such as pseudo-enzymatic activities [23, 24,
5 25]. Oxyhemoglobin, in the presence of CYP reductase and the nicotinamide adenine
6 dinucleotide phosphate (NADPH) cofactor, hydroxylates aromatic substrates such as
7 aniline through a CYP-like oxygenase mechanism [26, 27, 28]. In the presence of
8 H₂O₂, met-hemoglobin, can catalyze oxygenation of various substrates like 2,2'-
9 azino-bis(3-ethylbenzthiazoline-6-sulphonic acid (ABTS) and polycyclic aromatic
10 hydrocarbons (PAH) through a peroxidase-like mechanism [29, 30, 31]. All these
11 different activities of Hb are assumed to be the result of the heme ability to activate
12 and react with oxygen species.
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18 In the present paper, we report for the first time that Hb, in addition of its
19 oxygenase- and peroxidase-like activities, also presents a catalase-like activity. This
20 oxygen-producing activity has been pointed out during the study of anthracene (ANT)
21 oxidation through the peroxidase pathway of Hb. Anthracene belongs to the PAH
22 family which are carcinogenic and mutagenic pollutants found in waste waters from
23 petroleum refineries as traces (ng to µg/L) [32]. Bovine met-hemoglobin is a cheap
24 material that comes from the food industry and is a very promising candidate for
25 bioremediation of PAH in water. Here we present a detailed study of anthracene
26 oxidation mechanism by met-hemoglobin and H₂O₂ revealing that in aqueous
27 medium ANT oxidation is a "pseudo-catalytic" reaction. Furthermore, we show that to
28 successfully remove PAH from water, care should be taken for high Hb and ANT
29 concentrations low H₂O₂ concentrations otherwise the peroxidase activity of Hb is
30 inhibited by its newly demonstrated catalase activity. Kinetic parameters and
31 optimum pH have been determined for both anthracene oxidation (peroxidase
32 activity) and oxygen production (catalase activity).
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39 2. Materials and methods

40 41 Chemicals

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43 Anthracene (ANT) and anthraquinone were purchased from Sigma-Aldrich
44 (99% purity), acetonitrile was of HPLC grade (SDS). Bovine met-hemoglobin (met-
45 Hb) was obtained from Sigma-Aldrich as a lyophilized powder (ref H2625). This Hb
46 was used without any purification and considered as a pure preparation with a
47 molecular weight of 64 500 Da. The purity of the preparation has been checked using
48 the SDS-page technique (12% acrylamide) using standard protocol and coomassie
49 blue staining. No traces of other protein has been detected. H₂O₂ was obtained as a
50 35% solution from Sigma-Aldrich. Chemicals for the buffers were reagent grade
51 (Fluka). The ionic strength of the buffers was 0.05 M. For solutions of pH = 3, 4 and
52 5, a sodium citrate buffer was used (C₆H₈O₇ / C₆H₇O₇Na). For buffers of pH = 5.5, 6, 7
53 and 8, a phosphate buffer was used (NaH₂PO₄ / Na₂HPO₄). All solutions were
54 prepared using de-ionized water from a milli-Q purification system from Millipore.
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Anthracene oxidation study

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3 ANT oxidation studies were conducted in a 100 ml reaction mixture containing
4 from 0.3 to 15 μM Hb, 0.3 to 3 μM ANT, 1% acetonitrile (v/v) and 50 mM phosphate
5 buffer at pH = 5. ANT oxidation was started by adding various concentration of H_2O_2
6 (from 0.075 to 3 mM). The reaction progress was followed by ultrafast liquid
7 chromatography (UFLC). ANT disappearance was calculated from the decreasing
8 area of chromatogram peaks between reaction time $t = 0$ and $t = x$. For initial rates
9 determination, reactions were stopped at $t_x = 5$ min and the total disappearance
10 measurement reactions were stopped at $t_x = 15$ min. Reactions were performed in
11 flasks protected by aluminum foil to avoid ANT photo-oxidation cross-reaction. To
12 ensure reproducible results, after reaction and before UFLC analysis, samples have
13 been dilute in 50% (v/v) acetonitrile. The dilution step is necessary to stop the
14 reaction and allows also a good stability to the samples before UFLC analysis. All
15 reactions were repeated three times. Blanks without H_2O_2 and without Hb have also
16 been realized. Oxidation products were identified by GC-MS after a continuous
17 extraction of the media with 200 ml of methylene chloride. The organic phase was
18 dried with Na_2SO_4 and concentrated to 1 ml. Anthraquinone production has been
19 quantified thanks to a calibration curve done with the commercial standard
20 compound.
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28 UFLC analyses

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31 UFLC system used is a Shimadzu instrument, which was equipped with a
32 Supelco reversed phase column C18-PAH (50 mm length 4.6 mm internal diameter
33 and 3 μM particles size). The chromatographic apparatus is composed of two pumps
34 LC-20AD, an automatic sampler SIL-20AHT, a diode array detector SPD-M20A, a
35 column oven CTO-20A, and a communication bus module CBM-20A.
36
37 Chromatograms were monitored with LabPower Shimadzu software. The separation
38 method consists of a gradient between solvent A (50/50 (v/v) water/acetonitrile) and
39 solvent B (acetonitrile) starting from 0% of B over 0.5 minutes, then increasing B up
40 to 75% from 0.5 to 3 minutes, maintaining 75% of B from 3 to 4.2 minutes, increasing
41 B up to 100% from 4.2 to 4.5 and maintaining it to 5 minutes. Column oven is
42 maintained at 40 $^\circ\text{C}$ during the analysis. The retention time of ANT in this condition is
43 2.3 min. Peak integration has been done with the maximum wave length adsorption
44 of the compound (251 nm).
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49 GC-MS analyses

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51 GC-MS analyses were performed with a Shimadzu GC-2010 coupled to a
52 Shimadzu MS-2010 and equipped with a Supelco SPB-5MS capillary column (30 m x
53 0,25mm). Ionization was carried out by electronic impact (70 eV). The temperature of
54 the ion source was 200 $^\circ\text{C}$ and the temperature interface was 280 $^\circ\text{C}$. The oven
55 temperature program starts from 100 $^\circ\text{C}$ during 5 minutes, then increasing 5 $^\circ\text{C}/\text{min}$ to
56 280 $^\circ\text{C}$ and remaining at 280 $^\circ\text{C}$ for 11 minutes. 1 μl of the extracted and
57 concentrated sample was injected with a split equal to 10 and detection in the mass
58 spectrometer was done in scan mode between 50 and 400 (m/z). In these conditions,
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3 the retention times of anthracene and anthraquinone are 15 and 17 minutes
4 respectively. Compound identification has been done with the NIST mass spectral
5 database. The principal ion masses (m/z) for anthraquinone were 76, 152, 180, 208
6 and 178 for anthracene
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9 Oxygen production measurement

10 Oxygen production measurement was performed with an oxygen electrode
11 CellOx 325 coupled to an Inolab 730 oxymeter, both from WTW Company. The
12 measurement was started by addition of H_2O_2 under soft stirring. The reaction
13 volume was fixed at 100 ml. Oxygen production rates were measured in the first
14 minute of the reaction. The effect of pH and of different H_2O_2 and Hb concentration
15 has been studied. Reactions were repeated three times and blanks without H_2O_2 and
16 without Hb have been performed.
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22 Kinetics parameters determination

23 The kinetic parameters of the catalase-like activity and anthracene oxidation
24 activity of Hb have been determined graphically by plotting experimental data
25 obtained by measuring the initial rates of activity (V_i) toward the substrates (H_2O_2 ,
26 ANT) concentration (S). The Lineweaver-Burk representation, $1/V_i = f(1/S)$, allows
27 the determination of the different kinetic parameters: Michaelis constant (K_M),
28 maximum rate activity (V_m) and indirectly the catalytic constant (k_{cat}).
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33 3. Results and discussion

34 Anthracene oxidation through the peroxidase activity of Hb

35 To determine optimal pH reaction conditions, the peroxidase activity of Hb
36 ($0.15 \mu M$) was measured between pH values of 3.0 and 8.0 in the presence of $150 \mu M$
37 of H_2O_2 and $0.3 \mu M$ of ANT (Figure 2). The results show a narrow pH range for
38 peroxidase activity, with an optimum pH value between 5.0 and 6.0. At a pH of 5.0,
39 the maximum peroxidase activity corresponds to $76 \pm 5\%$ of ANT elimination. At pH
40 4.0 and 7.0, the remaining activity is less than 40% and 20% of the maximum activity,
41 respectively. In the following studies, pH was maintained at 5.0.
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48 In order to be as close as possible to waste water specificities, the challenge
49 of the present work was to use very low co-solvent (acetonitrile) addition, $\sim 1\%$ (v/v),
50 and low ANT concentration ($0.3\text{--}3 \mu M$). Previously, Vazquez-Duhalt *et al.* had already
51 noticed that the biocatalytic oxidation of PAH by Hb and H_2O_2 is strongly dependent
52 on organic solvent concentration with an optimum of 15% (v/v) acetonitrile [29].
53 According to this study, decreasing the acetonitrile concentration leads to a serious
54 decrease of activity. For example, in 10% (v/v) acetonitrile, the relative specific
55 activity of Hb falls to 60%. This may be due either to the fact that, for organic co-
56 solvent concentrations below 10%, the diffusion of PAH to the protein active site is
57 slow, or that amount of co-solvent was not sufficient to dissolve the PAH in water
58 since Vazquez-Duhalt *et al.* used an initial PAH concentration of $30 \mu M$. It should be
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3 noted that this concentrations 100- to 1000-fold higher than the one present in waste
4 water and which represents the goal of our study. As PAH compounds are very
5 hydrophobic, they exhibit a low solubility in water. For example, the ANT solubility in
6 pure water is 7.3 μM [33]. Furthermore, H_2O_2 is a suicide substrate which inactivates
7 the protein at the same time that it is used for the catalytic reaction [34]. One could
8 think that, for low acetonitrile content, the inactivation of the protein is faster than the
9 PAH oxidation because of a better H_2O_2 diffusion.
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14 As a result, we show that total ANT disappearance, for a concentration in
15 water of 0.3 μM in a medium containing 1% (v/v) acetonitrile, is possible if a
16 stoichiometric amount or an excess of Hb with respect to ANT is used. Nearly total
17 disappearance of ANT was reached when using a stoichiometric quantity or a two-
18 fold molar excess of Hb ($94 \pm 5\%$ and $97 \pm 5\%$, respectively) (Figure 3A).
19 Noteworthy, a very large molar excess of H_2O_2 was necessary to reach good levels
20 of ANT oxidation. Indeed an excess of at least 1000, with respect to Hb
21 concentration, was needed to reach 90% of ANT oxidation (Figure 3B). Anthracene
22 oxidation, in these conditions, can be qualified of "pseudo-catalytic" because, even if
23 Hb accelerates the reaction (allowing oxidation of ANT in the presence of H_2O_2),
24 excesses of Hb and peroxide are required for a complete disappearance of ANT.
25 One could think that the sub-stoichiometric behavior of the reaction limits the
26 potential application of the process at an industrial scale but, as bovine met-Hb is a
27 waste product from the food industry, it doesn't. In fact, abattoirs are an unlimited
28 source of cheap bovine met-Hb. Moreover, the total disappearance of ANT under our
29 conditions is reached in almost 15 minutes, which is much faster than usual catalytic
30 process used to remove PAH. Methods such as oxidation catalyzed by porphyrin [35]
31 or inorganic mineral oxide catalyst [36] with H_2O_2 or UV take at least several hours to
32 reach their maximal efficiency.
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41 Concerning ANT oxidation products, a small peak has been observed by the
42 UFLC technique corresponding to 35 % of the ANT that had disappeared. It was
43 identified as anthraquinone by GC-MS The other remaining 65% is supposed to be
44 covalently linked to Hb as previously shown in the literature for other PAHs [37,38]. In
45 the present study, we use the term 'disappear' to qualify the reaction of Hb and H_2O_2
46 over ANT, because only a part of the disappearing substrate has been clearly
47 identified as an oxidized product, and further studies should be necessary to
48 determine exactly the state of ANT linked to the protein and its localization on the
49 protein. This could be performed by using anthracene labeled by radioactive ^{14}C
50 followed by the entire enzymatic digestion of the reacted hemoglobin to identify on
51 which part of the hemoglobin the anthracene is attached, and also by analyzing
52 oxidized PAH released after total hydrolysis of Hb as shown by Melikian *et al* [39].
53 Most probably the ANT is under a semi-oxidized form as no adsorption of ANT and
54 anthraquinone onto the protein has been observed while running blank reactions
55 without H_2O_2 . Considering this fact, the term 'oxidation' can also be employed to
56 qualify the reaction of ANT with Hb and H_2O_2 It should be noted that if the cross-
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3 linking of the ANT oxidation product to the protein is verified, it is also an advantage
4 for an industrial process, because a major part of the oxidation product is trapped
5 within the protein. If Hb is immobilized on a support, a recovery of the oxidized-
6 HAP/protein system is possible as well as its elimination by simple calcination
7 process. Even if the oxidized forms of PAH are more biodegradable, they are also
8 more carcinogenic [40] and as such, this global elimination process has to be
9 considered.
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14 In order to understand the mechanisms of ANT oxidation by Hb and to
15 measure kinetic parameters, initial rates of the reaction (V_i) have been assayed for
16 different ANT concentrations (0.3, 0.6, 1, 2, 3 μM), various H_2O_2 concentrations
17 (0.075, 0.15, 0.3, 0.75 and 1.5 mM) and fix Hb concentration (1.5 μM). The
18 Lineweaver-Burk representations, $1/V_i = f(1/[\text{ANT}])$ and $1/V_i = f(1/[\text{H}_2\text{O}_2])$ (Figure 4)
19 show parallel lines for low ANT and high H_2O_2 concentration, which means that the
20 oxidation of ANT by Hb fits a bi-bi ping-pong mechanism. This mechanistic model is
21 typical of the peroxidase group [41,42]. First, H_2O_2 binds to the porphyrin iron to be
22 activated, and then ANT reacts with the activated oxygen species. Nevertheless, for
23 high ANT and low H_2O_2 concentrations, the lines are no longer parallel, showing a
24 slowing of the ANT disappearance rates. It is not easy to observe in Figure 4
25 because rates are very close to one another, due to the tight differences between the
26 various ANT concentrations studied, but for $\text{H}_2\text{O}_2 \leq 0.15$ mM and $\text{ANT} \geq 2$ μM the
27 rates of reaction decrease.
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34 Kinetic parameters determined for anthracene oxidation were $V_{\text{max}} = 0.25 \pm$
35 0.01 $\mu\text{M}/\text{min}$, $K_{\text{M}(\text{H}_2\text{O}_2)} = 80$ μM , $K_{\text{M}(\text{ANT})} = 1.1$ μM and $k_{\text{cat}} = 0.170 \pm 0.005$ min^{-1} . The
36 very tight K_{M} for anthracene is surprising, but it should be noted that it has not been
37 determined under saturation conditions. Indeed due to its poor solubility in an
38 aqueous medium, the maximum concentration of anthracene was 3 μM far below Hb
39 saturation ($[\text{Hb}] = 1.5$ μM). Moreover, the Hb concentrations required to accurately
40 measure oxidation rates should be high compared to the ANT concentrations. Thus,
41 it was not feasible to work with an excess of anthracene in a medium containing only
42 1% (v/v) of acetonitrile. Nevertheless, our data shows that the system works almost
43 as if it was saturated with anthracene. Indeed, with saturating H_2O_2 concentrations,
44 the initial rates over 1 μM ANT, increased just slightly (Figure 4). Considering this
45 fact, our results showed that the reaction was slow ($V_{\text{max}} = 0.25 \pm 0.01$ $\mu\text{M}/\text{min}$) but it
46 was compensated by the good affinity of ANT for Hb ($K_{\text{M}(\text{ANT})} = 1.1$ μM). The
47 observed rate (0.17 min^{-1}), in our study, is below the reported rate (0.4 min^{-1}) in the
48 previous work done by Vazquez-Duhalt *et al* in 1995 and 2000 [29, 31]. Firstly, this
49 difference can be explained by the co-solvent concentration. In our study, 1 %
50 acetonitrile was used in contrast to 15 % in the Vazquez-Duhalt work. More
51 acetonitrile in the media can increase the ANT solubility with an impact on the kinetic
52 parameters. On the other hand, Vazquez-Duhalt *et al* worked with human Hb which
53 can have a slight different reactivity. But even if both proteins react in exactly the
54 same way, the purification process to obtain the proteins can vary and bovine met-
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Hb, even if it is as pure as human Hb, can be less active. Whatever it is, our approach is more suited to a possible industrial waste water biotechnological treatment process because of the reactions conditions (low co-solvent and PAH concentrations) and because bovine met-Hb is surely more suitable than human Hb for a biotechnological process.

Working with a fixed and low H_2O_2 concentration (150 μM), variable Hb concentrations (0.15, 0.3, 1.5, 3, 15 μM) and high anthracene concentration (3 μM) (Figure 5), in the zone where ANT oxidation rates were slow down (Figure 4), showed that this inhibition is not due to a typical substrate excess inhibition. On one hand, Hb peroxidase relative activity, (in terms of total anthracene disappearance) increased from 63 to 100% between 0.2 and 1.5 μM Hb. On the other hand, the same relative activity falls down from 100 to 58% between 1.5 and 15 μM Hb. To summarize, considering results from the graph 4-A and B, discussed above, the inhibition of ANT oxidation takes place for $[\text{Hb}]/[\text{H}_2\text{O}_2] \leq 100$ and $[\text{H}_2\text{O}_2]/[\text{ANT}] \leq 75$. These results define an inhibition zone which should be taken into account for a likely industrial application for PAH removal in aqueous media. Nevertheless, most of the time, reaction conditions are outside of this inhibition zone because a large excess of peroxide compared to Hb is used and the PAH concentration in waste water is very low (nM).

These results suggested that H_2O_2 was consumed through an alternative pathway. Additionally, for $[\text{Hb}] = 15 \mu\text{M}$ and $[\text{H}_2\text{O}_2] = 15 \text{ mM}$, small gas bubbles appeared and foam was forming. This observation strongly suggested that Hb produced a gas from H_2O_2 . The sole gas that could potentially be produced by hemoproteins is O_2 , coming from the dismutation of H_2O_2 through a catalase-like mechanism. During this type of activity H_2O_2 plays both an oxidative and a reductive role to form molecular oxygen [14] (Figure 1). Nevertheless, this reaction has never been observed for Hb.

Catalase-like activity of Hb

In order to study the potential catalase-like activity of Hb and to ensure reliable results, following experiments have been performed in absence of ANT. A high concentration of H_2O_2 was used in order to have significant O_2 production rates. O_2 emission was measured by oxymetry.

Firstly, optimal pH reaction condition was determined by measuring the catalase activity between pH = 3.0 and 8.0 with Hb 3 μM in the presence of 5 mM H_2O_2 (Figure 2). Results show a wide pH range for O_2 emission. The optimum pH activity is between 5.0 <pH< 7.0, but even at pH = 8.0, more than 50% of Hb catalase-like activity remains. The catalase activity of Hb is less sensitive to pH compared to its peroxidase activity. It should be noted that at a pH of 8, only the

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3 catalase activity occurs, but there is no pH value at which one can work with only the
4 peroxidase activity of bovine met-Hb. The following experiments were run at pH =
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8 For $[H_2O_2] = 1.5$ mM, increasing the Hb concentrations (1.5, 3, 15, 30 μ M) led
9 to an effective production of O_2 with regards to Hb concentration. Whereas the
10 solubility of O_2 in water is only 8 mg/l, an emission of 22 mg of O_2 per liter has been
11 measured in 10 minutes with $[Hb] = 30$ μ M (Figure 6). The O_2 production rates
12 evolved linearly ($r^2 = 0.9971$, Figure 6) clearly demonstrating the catalytic nature of
13 this catalase-like activity. Initial O_2 production rates have been determined for $[H_2O_2]$
14 = 5 mM and various Hb concentrations (1.5, 3, 5, 7.5, 10, 12.5, 15 μ M). For a fixed
15 H_2O_2 concentration, more hemoglobin gives more oxygen production.
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19 The Lineweaver-Burk representation of O_2 production (Figure 7), measured at
20 a fix Hb concentration (3 μ M) versus increasing H_2O_2 concentrations (1, 5, 10, 25, 50
21 mM), allowed us to determine the kinetic constants of Hb catalase-like activity.
22 Results give $V_{max} = 344 \pm 24$ μ M/min, $K_M = 24.4 \pm 5.2$ mM and $k_{cat} = 115 \pm 8$ min^{-1} . It
23 shows that Hb catalase-like activity is fast but the affinity of H_2O_2 for Hb is weak. In
24 term of comparison, Hernandez-Ruiz *et al.* have calculated the catalytic parameters
25 for the catalase-like activity of horseradish Peroxidase and found: $K_M = 4$ mM and k_{cat}
26 = 110 min^{-1} [17]. The kinetic parameters of the catalase-like activity of
27 chloroperoxidase from *C. fumago* measured by Sun *et al.* are $K_M = 3.3$ mM and $k_{cat} =$
28 $54\ 000$ min^{-1} [16]. The catalase activity of Hb is faster but in the same range of order
29 of HRP, but much slower than chloroperoxidase. However, the affinity of H_2O_2 is less
30 for Hb than for HRP and chloroperoxidase, which are equal to each other.
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36 An explanation for this catalase activity could be the presence of a
37 contaminating catalase in the met-hemoglobin preparation. However this hypothesis
38 can be easily contradicted since the kinetic parameters of Hb catalase activity are
39 very different from those found for real catalases, such as catalase from *Micrococcus*
40 *lysodeikticus* [43, 44], featuring a K_M of 1100 mM for H_2O_2 , which is almost 46-fold
41 higher than the one calculated for Hb (25 mM). Additionally, the Hb catalase activity
42 presents a saturation kinetic under steady-state conditions, which is not the case of
43 the bacterial catalase (Figure 7). These kinetic elements are sufficient to show that
44 Hb preparation is not contaminated by catalase. Two different batches of Hb were
45 used with two commercial purities (Sigma-Aldrich ref 2625 and ref 2500), and the
46 same catalase activity was noticed. Moreover, a sodium dodecyl sulfate poly-
47 acrylamide gel electrophoresis has been performed with Hb and no traces of other
48 proteins could be detected, showing that this activity doesn't come from some
49 contaminating residual enzymatic activity (Figure 8). Nevertheless, as this
50 hemoglobin catalase-like activity has never been described before, it should be
51 verified, in further studies, if it is only in the case of bovine met-Hb or if it also exists
52 for human and other Hb proteins.
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59 The Hb catalase activity described here can be defined as a promiscuous
60 catalytic activity [45] since it is a side reaction catalyzed by the wild-type Hb which is
distinctly different from the main function known for Hb. This "promiscuous" property

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3 is thought to be an evolution-related process [46] allowing organisms to survive under
4 changing conditions [47]. Thus it could be assumed that this Hb catalase activity is
5 the legacy of natural evolution, from a common proteic ancestor of hemoproteins
6 which has differentiated progressively one from the other, creating several different
7 functions from a common active site. Few structural mutations induced during the
8 time could have given to some hemoproteins either an enzymatic activity like
9 peroxidases, catalases or oxygenases, or the ability to transport small molecules to
10 others. The presence of significant or weak promiscuous activities widens the
11 biotechnological applicability of Hb and makes the hemoprotein superfamily a very
12 interesting group for structure-activities relationship studies. Some researchers have
13 already underlined common structural elements between these different classes of
14 enzymes [48, 49] and further studies, including studies of Hb, should yield important
15 information to understand how the same global structure can lead to different
16 activities. For example, Jakopitsch *et. al.* have shown that some distal amino acids
17 such as Asp152 and Trp122 are essential for the catalase activity, but not for the
18 peroxidase activity of the bifunctional KatGs [50, 51]. Bovine met-Hb displays a totally
19 different heme distal environment, which could lead to other conclusions concerning
20 the origin of catalase activity.

21
22 In addition, it is possible that met-hemoglobin, which cannot bind O₂, could act *in*
23 *vivo* as a protective protein against H₂O₂ accumulation. Indeed, hydrogen peroxide is
24 constantly produced from oxygen active species in erythrocytes [52, 53]. Thus it
25 could be transformed to O₂ through met-hemoglobin catalase activity. On this
26 support, it can be supposed that slight mutation modifying this catalase-like activity of
27 hemoglobin could have important consequences *in vivo*, like, for example, onto the
28 lipid and lipoprotein peroxidation [54, 55, 56].

29
30 Concerning PAH removal from industrial waste water, when using the peroxidase
31 activity of Hb, the inhibition of the ANT disappearance rates can be eliminated by
32 avoiding the inhibition zone described above, but the catalase-like activity will always
33 run in parallel of the peroxidase activity, consuming some H₂O₂ molecules.

4. Concluding remarks

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The ANT oxidation by the peroxidase-like activity of hemoglobin proceeds through a bi-bi ping-pong mechanism. In aqueous medium, anthracene oxidation rates and turnover numbers are very low, probably due to the limited diffusion of anthracene to the active site of hemoglobin. It can be qualified as a pseudo-catalytic system. In aqueous media, total disappearance of anthracene traces can be performed within 15 minutes but with high amount of Hb. This requirement does not prohibit industrial applications because bovine met-Hb is a waste from food industry with a low production cost.

Bovine met-Hb presents a catalase-like activity described here for the first time. This activity interfered with the peroxidase-like activity especially for $[Hb]/[H_2O_2] \leq 100$ and $[H_2O_2]/[ANT] \leq 75$.

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3 Thus, hemoglobin is a good example of accidental catalytic promiscuity as
4 recently defined [44]. Despite its principal role is to transport oxygen, it also
5 possesses low peroxidase-, oxygenase- and catalase-like activities.
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12 ***Acknowledgment***

13
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15 The authors are grateful for the financial support from the TOTAL France S.A
16 company and to Naseem Ramsahye for the English revision of the manuscript.
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Figures legend:

Figure 1. Catalytic cycles of hemoproteins (the square is representing the heme group): 1- CYP Monooxygenation cycle: the insertion of one atom of oxygen into a substrate RH from molecular oxygen with the concomitant release of water through a complex pathway needing electrons and protons transfer before obtaining the active compound $\text{Fe}^{\text{IV}}=\text{O}^+$ which reacts with RH to gives ROH (this pathway can be quenched by the presence of carbon monoxide); 2- In the peroxidase pathway, H_2O_2 oxidizes directly Fe^{III} into the active compound $\text{Fe}^{\text{IV}}=\text{O}^+$ which then reacts with RH; 3- The catalase pathway: the active compound $\text{Fe}^{\text{IV}}=\text{O}^+$, formed by the action of a first H_2O_2 molecule, can react with a second H_2O_2 molecule and lead to the release of O_2 and H_2O .

Figure 2. Relative hemoglobin pseudo-peroxidase and catalase activity versus pH. For peroxidase-like activity, the initial anthracene concentration was 0.3 μM , 1% acetonitrile (v/v), $[\text{H}_2\text{O}_2] = 150 \mu\text{M}$ and $[\text{Hb}] = 0.15 \mu\text{M}$. The maximum activity, in terms of total ANT disappearance, was $76 \pm 5\%$ at $\text{pH}=5$. Analyses were done after complete reaction (15 minutes) by UFLC. For catalase-like activity, $[\text{Hb}] = 3 \mu\text{M}$, $[\text{H}_2\text{O}_2] = 5 \text{ mM}$. Analyses were done with an oxymeter during the first minutes after addition of H_2O_2 . The maximum activity, in terms of initial O_2 production rate, was 62 $\mu\text{M}/\text{min}$ at $\text{pH}=6$.

Figure 3. A : Total anthracene disappearance versus hemoglobin concentration. The initial anthracene concentration was 0.3 μM in phosphate buffer 50 mM, $\text{pH} = 5$, 1% acetonitrile (v/v). A molar excess $[\text{H}_2\text{O}_2]/[\text{Hb}] = 1000$ was used for all reactions. Analyses were made after complete reaction (15 minutes) by UFLC. B: Relative activity of Hb on ANT elimination at variable $[\text{H}_2\text{O}_2]/[\text{Hb}]$ ratios. The initial anthracene concentration was 0.3 μM in phosphate buffer 50 mM, $\text{pH} = 5$, 1% acetonitrile (v/v). Hb concentration was 0.15 μM . Maximum activity, in terms of total ANT disappearance was $85 \pm 5\%$ for $[\text{H}_2\text{O}_2]/[\text{Hb}] = 5000$. Analyses were made after complete reaction (15 minutes) by UFLC

Figure 4. Lineweaver-Burk plot of initial anthracene disappearance rates, at variable H_2O_2 concentration (A) and at various anthracene concentrations (B). Reactions were conducted in phosphate buffer 50 mM $\text{pH} = 5$, with different H_2O_2 concentrations (0.075, 0.15, 0.3, 0.75, 1.5 mM), various ANT concentrations (0.3, 0.6, 1, 2, 3 μM) and a fixed Hb concentration (1.5 μM). Analyses were done after 5 minutes reaction by UFLC.

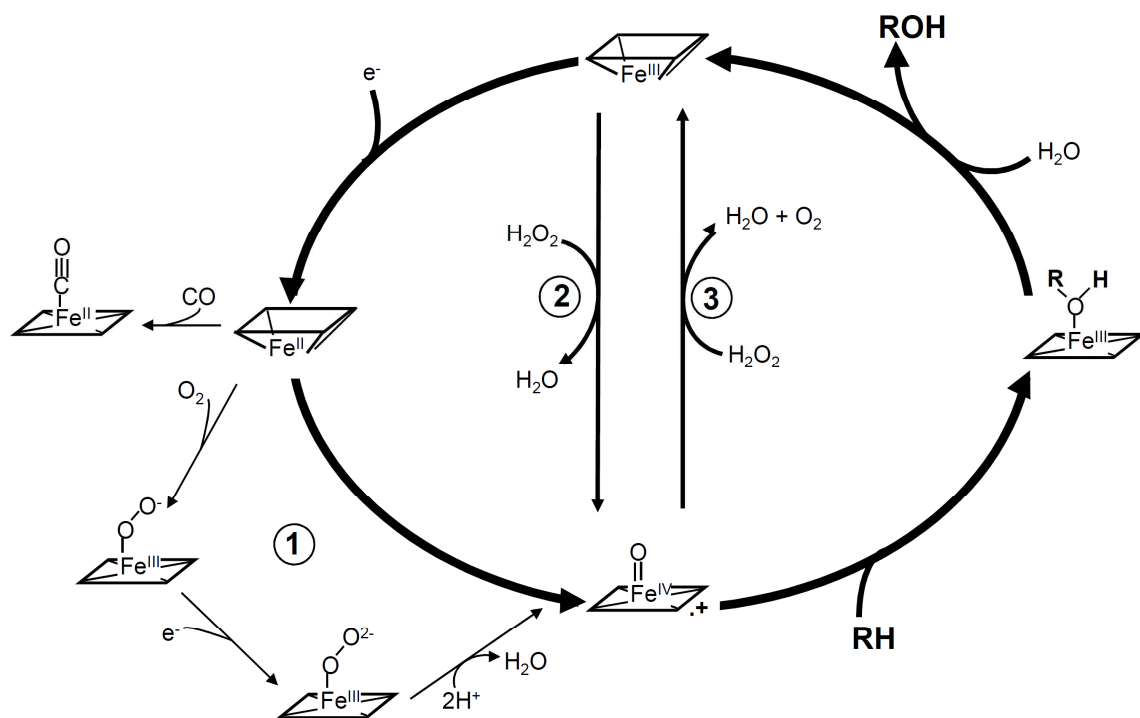
Figure 5. Relative peroxidase-like activity of Hb on anthracene oxidation at low H_2O_2 concentration (150 μM) in function of Hb concentrations. Reactions were done in phosphate buffer 50 mM, $\text{pH} = 5$, $[\text{ANT}] = 3 \mu\text{M}$. The maximum activity, in term of total disappearance, was $57 \pm 5\%$ with $[\text{Hb}] = 1.5 \mu\text{M}$. Analyses were done after complete reaction (15 min) by UFLC

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Figure 6. A: O₂ emission (mg/L) for [H₂O₂] = 1.5 mM and various Hb concentration (0, 1.5, 3, 15, 30 μM). Measurements were done with an oxymeter. B: Linear initial O₂ production rates in function of Hb concentration (1.5, 3, 5, 7.5, 10, 12.5, 15 μM). H₂O₂ concentration was kept fixed for all reactions (5 mM). Reactions were done in a phosphate buffer 50 mM, pH=6. Measurements were done with an oxymeter in the first minutes after H₂O₂ addition.

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Figure 7. A: Initial O₂ production rates by Hb in function of [H₂O₂] showing the saturation kinetic of Hb catalase activity. Reactions were done in phosphate buffer 50 mM, pH = 6, [Hb] = 3 μM. Measurements were done with an oxymeter during the first minutes after H₂O₂ addition. B: Lineweaver-Burk plot of initial O₂ production rates.

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Figure 8. SDS-page gel photo. Left: standards with various molecular weights. Phosphorylase b (97400 Da), bovine serum albumin (66200 Da), ovalbumin (45000 Da), carbonic anhydrase (31000 Da), soybean trypsin inhibitor (21500 Da) and lysozyme (14400 Da). Right: Hb (10g/L).

Fig 1



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Fig 2

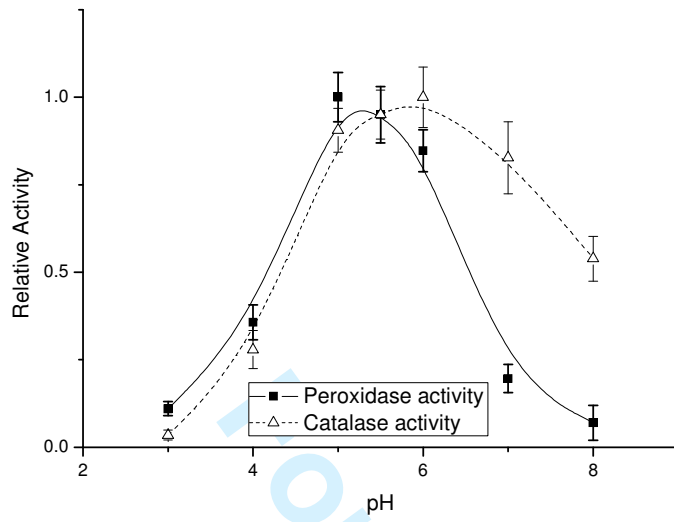
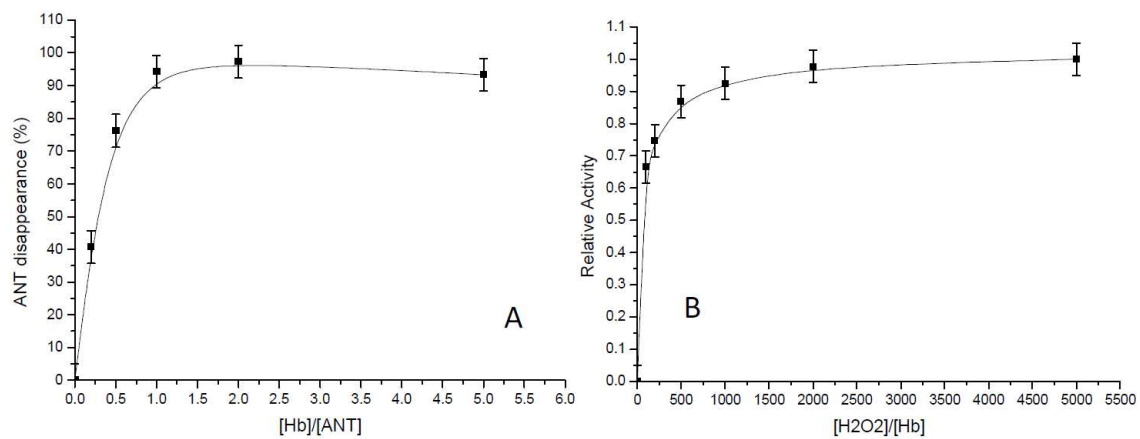


Fig 3 A and B



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Fig 4 A and B

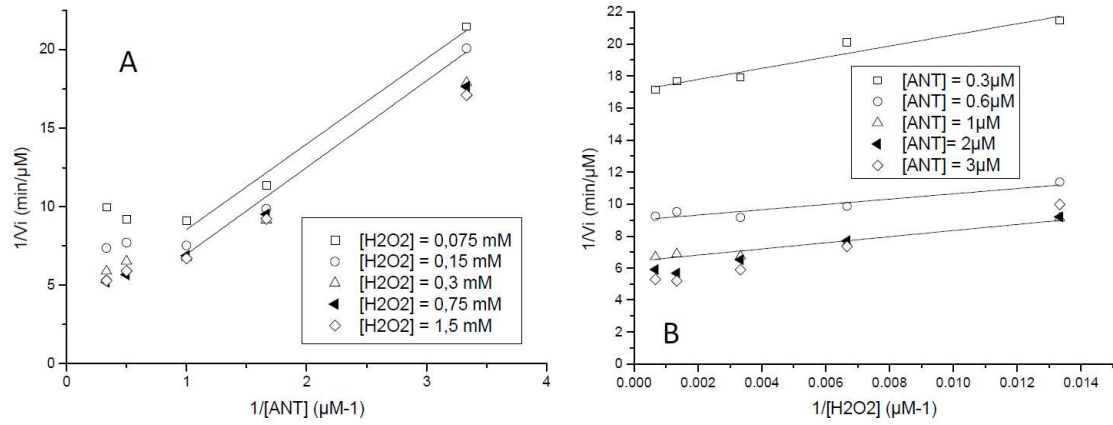


Fig 5

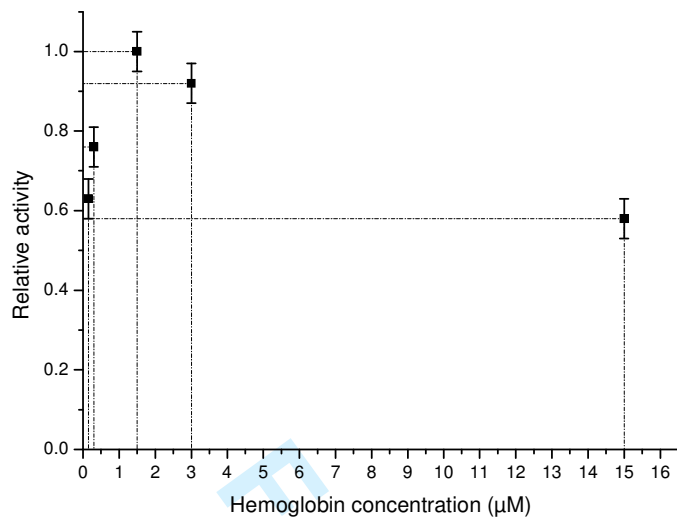


Fig 6 A and B

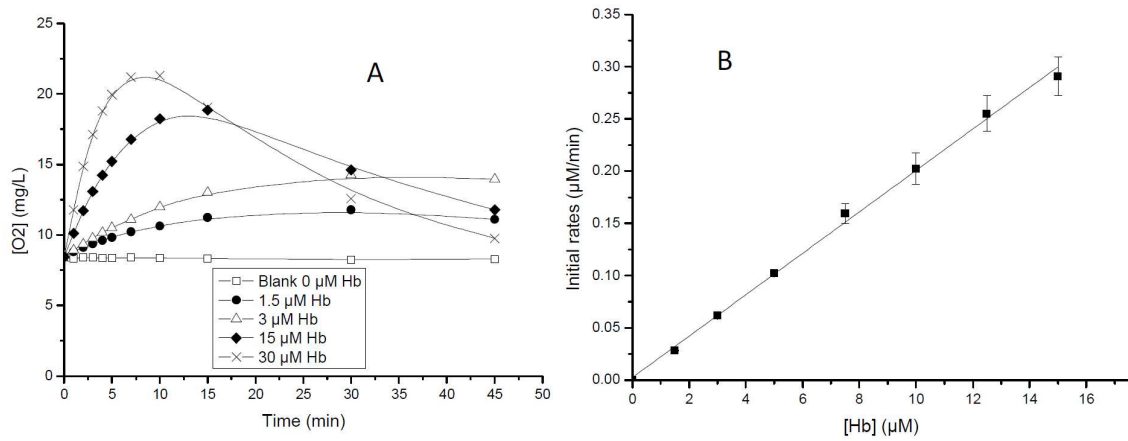


Figure 7 A and B

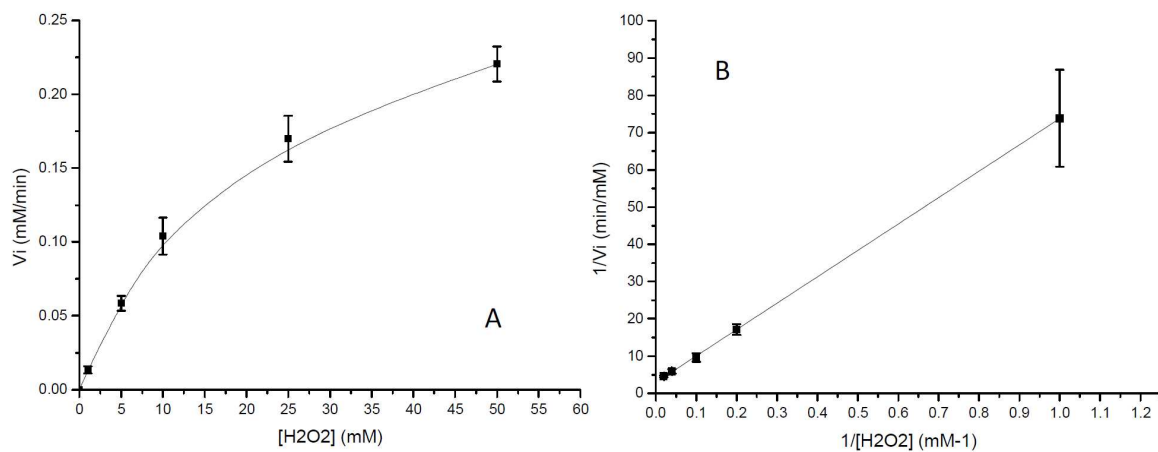


Fig 8

