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Bidentate ligation of magnesium by 1,2-dimethoxyethane in the gas phase

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Abstract

The 1:1 Mg \cdots 1,2-dimethoxyethane (Mg-DXE) complexes are studied experimentally and theoretically. They are generated by a laser ablation source in a supersonic expansion. They are studied spectroscopically by Resonance Two-Photon Ionization (R2PI). DFT/B3LYP and *ab initio* calculations using the MOLPRO quantum chemistry package are performed to document their ground and excited states in a series of geometry ranging from monodentate to bidentate ligation of Mg by the O atoms of DXE. An absorption band is observed in the 27800-30500 cm^{-1} range, which, thanks to the calculations, is attributed to the bidentate complex. The structure of the band is discussed in terms of the excitation of electronic states of the complex correlating adiabatically to the 3s3p ^1P and 3s4s ^1S states of Mg at large separation between Mg and DXE.

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I. INTRODUCTION

Crown ethers are important compounds in coordination chemistry for their ability to bind specific metal cations. Their cyclic structure formed by the association of n $-\text{CH}_2 - \text{CH}_2 - \text{O}-$ units puts the n oxygen atoms in proper position to bind a metal cation, which is *maintained* at the centre of the crown.¹ This corresponds to a polydentate ligation since the metal ion is bound to several anchoring sites of the ligand molecule. A simpler variant of polydentate interaction is found with 1,2-dimethoxyethane (DXE), $\text{CH}_3 - \text{O} - \text{CH}_2 - \text{CH}_2 - \text{O} - \text{CH}_3$, which also contains the $-\text{CH}_2 - \text{CH}_2 - \text{O}-$ repeating unit (note that the name 1,2-dimethoxyethane is quite often shortened as DME. We nevertheless prefer DXE since it makes no confusion with dimethyl ether). DXE is less rigid than the closed ring of crown ethers and can exist as various rotamers, one of which has the two O atoms in position to provide a bidentate interaction with a metal ion. It has the conformation labeled *tgt* (*t* for *trans* and *g* for *gauche*) with respect to the O-C, C-C and C-O bonds (Fig. 1 shows the *tgt* and *ttt* rotamers and the bond axes, which define the *trans* and *gauche* isomerizations). DXE is often used as a benchmark for studying polydentate ligation of ions.^{2,3} Its less pronounced rigidity will appear as an advantage in the present work, which deals with the ligation of neutral metal atoms.

FIGURE 1

Gas phase studies allow one to form and to study spectroscopically, molecular assemblies, which are free of external perturbations. The purpose then, is to investigate intrinsic properties of these assemblies very directly, and to perform extensive comparisons with theoretical approaches. This is exemplified for DXE-alkali ion²⁻⁵ and DXE-transition metal ion^{6,7} complexes. A refined analysis of the $\text{M}^+ - \text{O}$ bonds in the DXE-alkali ion complexes showed that they are non-covalent and determined by electrostatic charge-dipole, polarization and partial charge transfer interactions. The charge transfer is significant for the light alkali ions, Li^+ and Na^+ , with a transfer of $0.3 e^-$ and $0.1 e^-$, respectively.^{2,3} The strength of the bidentate interaction of the alkali ion with the two O atoms of DXE reduces the O-C-C-O dihedral angle from 73.8° in the *tgt* rotamer of free DXE down to 47.6° in the

$\text{Li}^+(\text{DXE})$ complex (note that these values depend on the level of the calculation by a few degrees). The effect of reducing the dihedral angle is smaller for the larger cations, in line with the larger $\text{M}^+ - \text{O}$ bond length when M^+ goes along the $\text{Li}^+ \rightarrow \text{Cs}^+$ sequence.³

Much less attention has been given to neutral systems although the $\text{M} - \text{DXE}$ interaction is quite important from a fundamental point of view. As opposed to metal ions, neutral metal atoms do not exhibit strong bonds with DXE in the ground state. There, the interaction reflects a subtle balance between several contributions, none of which is dominant, in contrast with the situation encountered above with the charge-dipole interaction in $\text{M}^+ - \text{DXE}$. A work by Yeh and Su reports the formation of the $\text{Na} - \text{DXE}$ complex in a flow generator and its photoionization study. The experimental work is complemented by MP2 *ab initio* and density functional calculations.⁸ It was found that the most stable conformer of the $\text{Na} - \text{DXE}$ complex has a cyclic structure, *i.e.* Na has a bidentate ligation with DXE. Yang and coworkers also investigated such systems in a work where $\text{Li}, \text{Na} - \text{DXE}$ complexes were produced by the laser vaporization/supersonic beam technique and studied experimentally by ZEKE spectroscopy.⁹ Again, DFT and MP2 calculations complemented the experimental work. The ZEKE spectroscopy was unsuccessful on the $\text{Na} - \text{DXE}$ complex and only adiabatic ionization could be documented, in qualitative agreement to those reported.⁸ In contrast, the ZEKE spectroscopy of the $\text{Li} - \text{DXE}$ complex yields information on the vibrational modes of the ground state ionized complex, which were fully assigned. Finally, the calculation on the $\text{Li} - \text{DXE}$ complex showed that the bidentate ligation affects essentially the O-C-C-O dihedral angle of DXE, which is reduced to 49.0° , a value, which is comparable to that encountered above in the ionized complex $\text{Li}^+(\text{DXE})$.⁹

The ligation of neutral alkaline earth atoms by DXE offers a further degree of complexity in the ligation mechanism, by the presence of two valence electrons in the atom. In particular the ligation might behave differently whether the alkaline earth atom is in the ground state or electronically excited, because the ground state has a closed shell structure, whereas it is not the case in electronically excited states.

The present work is a first attempt to address this question by documenting the $\text{Mg} - \text{DXE}$ complex, in a combined experimental and theoretical approach where the

ground and the first electronically excited states of the complex are investigated. A motivation is to examine how the ligation of Mg in various electronic states affects the flexible O-C-C-O dihedral angle of the DXE molecule. The experimental work is performed using the laser vaporization/supersonic beam technique combined with the Resonance Two-Photon Ionization technique. It provides spectroscopic information on the Mg – DXE complex. The theoretical work associates B3LYP/6-311++G** calculations of various Mg – DXE geometries in the ground state, ranging from the *tgt* to the *ttt* rotamers, with an *ab initio* approach providing vertical excitation energies at these geometries. This enables to explore the effect of ligation on the magnesium spectroscopy when the ligation switches from bidentate to monodentate. In the latter calculation, the electronic problem is treated by the Internally Contracted Multi-Reference Configuration Interaction method (IC-MRCI) using the pseudo potential approach to reduce the electron problem to that of the valence electrons.

II. EXPERIMENTAL AND COMPUTATIONAL TECHNIQUES

A. Experimental

The experimental apparatus combines a laser ablation system, nanosecond ionization lasers and a linear time-of-flight mass spectrometer (TOF-MS). Briefly, magnesium atoms are evaporated by the laser ablation within the beam source where they are mixed with ligand molecules. The supersonic expansion, which follows, generates the desired Mg – DXE complexes among other species. All of them are carried in the beam after collimation. They are ionized using the lasers and the resulting ions are mass analyzed using TOF-MS. More extensive details are given now concerning the various components of the apparatus and the operating procedure.

FIGURE 2

1. *Beam source*

Cold clusters are generated using the laser evaporation source schemed in Fig. 2. It is composed of a piezoelectric pulsed valve, a nozzle of 0.5 mm diameter followed by a channel, 1 mm in diameter. A rotating rod of magnesium (6 mm in diameter, 80 mm long) is ablated by the 532 nm doubled output of a Nd-YAG laser, in the configuration of Smalley.¹⁰ A mixture of 1% DXE in helium is premixed with 10% argon in a cylinder. It is passed through the nozzle and entrains the Mg vapor created by the laser ablation in a 10 mm long channel. The beam is free to expand at the tip of the channel. It produces cold Mg – DXE complexes, but as usual in laser evaporation sources, several other species are present in the beam, including oxides coming from an oxygen pollution of the rod and various types of mixed clusters. The operating conditions are carefully adjusted to optimize the formation of the Mg-DXE complex (see Sec. III A). The pressure in the 10 mm channel turns out to be an important parameter with this respect. It is controlled by changing the backing pressure of the gas line feeding the pulsed valve and by adjusting the duration of the valve opening. The molecular beam is collimated by a 2 mm skimmer before entering the chamber where the spectroscopic studies are performed.

2. *Ionization lasers*

A pulsed tuneable dye laser (LambdaPhysik LPD 3000) is pumped by the 532 nm output of a nanosecond Nd-YAG laser. It is used for the Resonance Two-Photon Ionization (R2PI) study of the complexes. The species present in the molecular beam are characterized by single photon ionization in the range 207-240 nm (6.05-5.2 eV) by frequency doubling a dye laser delivering tunable light between 414 and 480 nm. An excimer laser (GAM Laser EX5) operating at 157 nm is used also to ionize these species with a photon energy close to 8 eV.

3. *Clocking*

The time delays between the pulsed components of the beam source, lasers and detection tools represent essential parameters to control. This is achieved *via* a digital delay generator

controlled by a quartz resonator (designed and build at Laboratoire de PhotoPhysique Moléculaire - LPPM - Université Paris Sud). The time sequence is the following: *i*) the two Nd-YAG lasers are charged at first. Then, *ii*), the ablation laser is fired. This defines the time origin t_0 . *iii*), the pulsed valve is opened at time t_1 . The time delay $t_1 - t_0$ is varied by 1 μs steps. It is adjusted to ablate magnesium within the leading edge of the gas pulse. The tuning range of $t_1 - t_0$ where complexes are observed is limited to 10 μs . Finally, *iv*), the ionization laser is fired at a time t_2 when the complexes arrive in the observation zone. The $t_2 - t_0$ time delay is also controlled by 1 μs steps.

B. Computational

Two kinds of calculations are performed. First, a series of geometries of the Mg-DXE complex is generated using the GAUSSIAN 98 package.¹¹ It samples the O-C-C-O dihedral angle θ of DXE between -180° and 180° , hence exploring geometrical conformation where the DXE molecule switches between the *tgt* and the *ttt* rotamers. Second, vertical excitation energies to the first four excited singlet states of the complex are calculated, at each geometries, using the MOLPRO quantum chemistry package.¹² A limited exploration of the first four triplet states is also performed.

The GAUSSIAN calculations are performed according to the density functional (DFT) Becke 3-parameter Lee, Yang and Parr (B3LYP) hybrid method in an all-electron treatment, which uses the 6-311++G** basis set. The *modredundant* option is used to scan the O-C-C-O dihedral angle θ of DXE between -180° and 180° by 10° steps, while optimizing all the other coordinates.

The MOLPRO calculations are aimed at providing accurate excitation energies of the complex in an energy range, which overlaps the $27800\text{-}30500\text{ cm}^{-1}$ range, where an absorption bands is found experimentally. This corresponds to documenting the singlet states of Mg-DXE, which correlate adiabatically to the $3s3p\ ^1P$ and $3s4s\ ^1S$ states of Mg at large separation between Mg and DXE. The technique successfully applied in¹⁴ to describe the MgH molecule, is used again. Accordingly, the core electrons of Mg are described by

a $[\text{Mg}^{2+}]$ effective core potential (ECP), complemented by a core-polarization potential (CPP). The parameters defining the ECP and the CPP are given in,¹⁴ as the uncontracted [10s,11p,9d] Gaussian basis set, which describes the valence electrons of Mg. These choices provide a reliable prediction of the magnesium energy levels, up to the $3s3d^1D$ state at 46403 cm^{-1} , as discussed.¹⁴ In particular, the prediction is within 340 cm^{-1} for the $3s3p^1P$ and $3s4s^1S$ states and the ionization energy, which are relevant here (see Tab. I). Triplet states, which are relevant for the discussion below are also reported in the table. The lowest triplet energy is very satisfactory and differs by only 219 cm^{-1} from the experiment, whereas the upper state is predicted within 657 cm^{-1} , only.

The core electrons of the oxygen ($[\text{O}^{6+}]$) and carbon ($[\text{C}^{4+}]$) atoms in DXE are also described by effective core potentials. Those of Ref.¹⁵ are used together with the corresponding Gaussian basis set for the valence electrons. Finally the electrons of the H atoms in DXE are described by the cc-pVDZ basis set.¹⁶

The electronic problem is treated in the MOLPRO calculations with the same spirit as in our former calculations on the $\text{Ca} - \text{CH}_3\text{F}$ complex.¹⁷ A RHF calculation of the Mg^+ -DXE ion is performed first. A state-averaged MCSCF calculation of the neutral $\text{Mg} - \text{DXE}$ complex follows, which reads the orbitals from the RHF calculation. The average includes the 5 lowest singlet states (S_0 to S_4). They correlate adiabatically to the $3s^2^1S$, $3s3p^1P$, and $3s4s^1S$ states of magnesium at large $\text{Mg} - \text{DXE}$ separation. The active space of the MCSCF calculation consists of two electrons in 10 orbitals (asymptotically correlating to the 3s, 3p, 4s and 3d orbitals of Mg). This provides the reference states of the Internally Contracted Multi Reference Configuration Interaction (IC-MRCI) calculation,¹⁸ which provides the results presented below. Due to the state averaging, which includes states of fairly large energy, the ground state is not extremely accurate in this calculation. We preferred a separate MCSCF+IC-MRCI calculation, which includes this sole state. Then, a reduced active space of 2 electrons in 5 orbitals is sufficient.

TABLE I

III. RESULTS

A. Characterization of the beam

Mg-DXE complexes are easily formed and detected by a one-photon ionization at 212 nm. The mass spectrum shown in Fig. 3 is typical of those observed under soft expansion conditions, which are achieved by complying with three factors: *i*) minimum opening time of the pulsed valve; *ii*) lowest possible backing pressure in the foreline and *iii*) minimum concentration of DXE in the backing gas. This ensures that the 1:1 Mg – DXE complex is the dominant species among Mg clusters as observed in Fig. 3.

FIGURE 3

B. Resonant Two-Photon Ionization (R2PI)

The R2PI study of the 1:1 Mg – DXE complex is performed with the laser tuned in the 25000-33000 cm^{-1} energy range (3.1-4.1 eV). The complex is found to have absorption bands within this range, between 27800 and 30500 cm^{-1} , enabling the R2PI scheme. A first photon is absorbed by the complex to an electronic state, which is sufficiently long lived to be ionized by a second photon of same wavelength. The large ion signal observed at $m/z = 114$ corresponds to the 1:1 $\text{Mg}^+ - \text{DXE}$ ion with no indication of fragmentation to smaller masses or broadening of the mass peak.

The $\text{Mg}^+ - \text{DXE}$ ion signal is obtained by integrating the mass peak at the relevant mass at the output of the TOF-MS. It is shown in Fig. 4 as a function of the frequency of the R2PI laser. The three curves in the figure are recorded when tuning the TOF-MS to 114, 115 and 116 amu, which correspond to the mass of the complex when formed from the $^{24,25,26}\text{Mg}$ isotopes, respectively. The three spectra have the same shape. They have a weakly structured wing between 27800 cm^{-1} and 28380 cm^{-1} , which is followed by strong structures. The latter belong apparently to a long vibrational progression, but although

the spacing of the structures appears regular, their intensity is very irregular. A shoulder of weaker intensity appears above 29500 cm^{-1} . It is weakly and irregularly structured.

FIGURE 4

The vibrational progression displayed in Fig. 4 does not seem to be shifted, within the experimental uncertainties, when switching from an isotope of magnesium to the other. Given the width of the vibrational bands, $40\text{ to }60\text{ cm}^{-1}$, if existing, the isotope shifts must be smaller than $10\text{ to }20\text{ cm}^{-1}$. This small value is an indication that Mg is not directly involved in the vibrational mode corresponding to the progression observed in Fig. 4 as it would be the case, for instance, in the Mg-DXE stretching mode. Fig. 5 shows the expected isotope shift for this mode, assuming a vibrational frequency $\omega_e = 110\text{ cm}^{-1}$, which corresponds approximately to the separation between the large amplitude structures observed in Fig. 4. This figure is built from the expression below, which describes the isotopic shift as a mass effect in the anharmonic oscillator approximation:¹⁹

$$\omega_e^{i'} - \omega_e^i = (\rho - 1)\omega_e^i \quad (1)$$

with

$$\rho = \sqrt{\frac{\mu^i}{\mu^{i'}}} \quad (2)$$

In these expressions, i and i' distinguish the two isotopes, μ^i, i' being their reduced masses and ω_e^i, i' the corresponding vibrational frequencies. The predicted effect amounts to more than 20 cm^{-1} when the vibrational quantum number v is larger than $v=6$. Since the v -dependent isotopic shift of the experimental observation falls below this value, magnesium is presumably not involved directly in the Mg-DXE deformation mode, which produces the large amplitude progression observed in Fig. 4. A deformation mode, which is not strongly coupled to the movement of Mg, is for instance the torsion of the DXE moiety.

FIGURE 5

C. Calculations

Fig. 6 shows a cut of the ground state energy surface S_0 of the Mg-DXE complex along the O-C-C-O dihedral angle. The full curve is provided by the DFT/B3LYP calculation and the dashed curve by the MCSCF+IC-MRCI calculation. As seen in Sec II B, the potential energy of the complex is optimized in the DFT/B3LYP calculation for all the coordinates at a fixed dihedral angle, θ . The same calculation is then repeated at different dihedral angles such that $-180^\circ < \theta < 180^\circ$ in order to explore the ligation of Mg by DXE. Two equivalent absolute energy minima are found along the DFT/B3LYP curve in the figure when the O-C-C-O dihedral angle is equal to $\pm 60^\circ$. They correspond to the DXE moiety in the *gauche* configuration with respect to the C-C central axis. Two equivalent secondary minima are also observed with the dihedral angle equal to $\pm 180^\circ$. Their energy is 110 cm^{-1} higher than that of the absolute minima. The height of barrier separating the absolute and secondary minima is 880 cm^{-1} , whereas that between the two equivalent absolute minima at $\pm 60^\circ$ is 1810 cm^{-1} . The DFT/B3LYP curve (full curve) is compared in the figure to that provided by the MCSCF+IC-MRCI calculation (dashed curve). Both are in qualitative agreement. Nevertheless, the equivalent absolute minima in the latter calculation are observed with the O-C-C-O dihedral angle equal to $\pm 48^\circ$, the energy of the secondary minima is 366 cm^{-1} and the barrier heights of 880 and 1810 cm^{-1} in the DFT/B3LYP calculation are changed to 1552 and 1035 cm^{-1} , respectively.

FIGURE 6

The Mg-DXE binding energy, uncorrected for the Basis Set Superposition Error (BSSE), can be estimated by the energy difference between the MCSCF+IC-MRCI calculation for the ground state complex at the absolute minimum and the same calculation

when Mg is put at a very large distance (100 Å). The estimated binding energy is 1580 cm⁻¹.

Fig. 7 shows a cut of the ground state S_0 and of the four lowest singlet excited potential energy surfaces S_1 to S_4 of the Mg-DXE complex provided by the MCSCF+IC-MRCI calculation, as a function of the O-C-C-O dihedral angle. The upper curve of the figure is the RHF energy of the ground state Mg⁺-DXE ion. The four excited curves and the ionic curve document the vertical excitation and the vertical ionization energy of the complex from each point of the ground state curve. Qualitatively, these four states show a similar behavior, yet significantly different from that of the ground state. In particular, the barrier between the equivalent absolute minima at $\pm 48^\circ$ in the ground state does not exist in the excited states, nor in the ion. Instead, a flat minimum extending between -40° and $+40^\circ$ is observed along each curve.

FIGURE 7

The experimental R2PI signal shown in Fig. 4 is reproduced in Fig. 7, assuming that the experimental spectrum corresponds to a vertical excitation from either the absolute or the secondary minimum of the ground state curve. The experimental spectrum overlaps the potential energy curves, when an excitation from the absolute minimum at $\pm 48^\circ$ is assumed, whereas is significantly below the potential energy curves with the other assumption.

We already mentioned a striking observation in Fig. 7: the four excited curves and the ionic curve of the Mg-DXE system have a similar behavior as the O-C-C-O dihedral angle θ of DXE is varied. This finds an easy interpretation in Fig. 8. The latter displays the first excited orbital of the Mg-DXE system provided by the state averaged MCSCF calculation. The large size of the excited orbital, compared to the size of the DXE molecule, lets the underlying Mg⁺ ion almost free to interact with the DXE molecule, with little screening by the excited orbital. The dominant configuration corresponding to the first excited state of the Mg-DXE system has an electron in this orbital, hence creating the condition for a similar variation of the interaction energy of both this excited state of Mg – DXE and the ground state of Mg⁺ – DXE upon torsion of the DXE moiety. Such an interpretation is also

valid for the three other excited states of Mg-DXE, which appear in Fig. 7. The size of the excited orbitals associated with these states is indeed comparable or larger than that of the first excited orbital shown in Fig. 8. In particular, the DXE moiety is very much embedded into the orbital of either dominant 4s or 3p character, when the 3p orbital is perpendicular to the C-C bond of the molecule.

FIGURE 8

The other salient feature in Fig. 7 is the very different behavior of the ground state on the one hand and the excited states on the other, upon the torsion motion of the DXE moiety. The ground state behavior is a consequence of a compensation between a stronger attraction between Mg and DXE when the DXE torsion makes the two oxygen atoms facing Mg and a repulsive interaction between the two oxygen atoms. In contrast in the excited states, the attraction between Mg⁺ ion underneath the excited orbital and the oxygen atoms overwhelms the repulsion between the two O atoms. Hence, the excited potential is essentially flat as long as the two O atoms are in position to interact with the Mg⁺ ion core. Fig. 8 shows that this is still possible when the O-C-C-O dihedral angle θ is 60°. Nevertheless, the energy of the excited states increases slightly, due to the destabilization of the ground state, which makes the distance between Mg and the O atoms to increase. It switches from 2.41 to 2.79 Å when θ is changed from 0° to 60°. At larger values of θ , a single O atom interacts with the Mg⁺ ion core, reducing substantially the stabilization of the excited states. This effect accounts for the substantial enhancement of the excited state energies when $\theta > 60^\circ$. This enhancement is even larger because the ground state itself is slightly destabilized and the distance between Mg and the closest O atom goes up to 3.47 Å at $\theta = 180^\circ$.

IV. DISCUSSION

A. Conformation of the ground state Mg-DXE complex

The conformational landscape of DXE is quite complex since numerous rotamers are formed by rotation of the chemical groups forming the DXE molecule about the three axes shown in Fig. 1. The available information, which is both experimental and theoretical has been reviewed recently.²¹ It indicates that the total population of the *gauche* rotamers with respect to the O-C-C-O torsion, is dominant in the gas phase at 273.15 K over that of the *trans* rotamer. The population ratio between the O-C-C-O *gauche* and *trans* rotamers is ranging between 71:29 and 79:21, with the *gauche* rotamers corresponding to *tgt* and *tgg'* essentially and the *trans* rotamer corresponding to *ttt* (the *gg'* notation indicates that the two *gauche* torsions have opposite orientations). The larger *gauche* fraction is due to the larger statistical weight of the corresponding rotamers, which inverts the fact that they are slightly less stable than the *ttt* rotamer (*tgt* is 108 cm^{-1} less stable than *ttt*).²¹

The *gauche* rotamers of DXE have the two O atoms in proper position for a bidentate ligation of magnesium, whereas the *trans* rotamer is expected to lead to a monodentate ligation. The Mg-DXE complex is formed in the present work by collision between Mg and DXE in the supersonic expansion. Given that the DXE molecule has essentially the *gauche* configuration in the gas phase, we infer that the complex should be of the bidentate type, essentially. This expectation is reinforced by the fact that the bidentate ligation of Mg is favored energetically over the monodentate ligation. Fig. 6 shows indeed that the well at 180° (*ttt* rotamer), which is associated with the monodentate ligation is higher in energy than the well corresponding to the bidentate ligation. Nevertheless, the *tgt* and *ttt* rotamers of free DXE are separated by a barrier of about 900 cm^{-1} (0.11 eV),²¹ which is almost not reduced by the ligation (see Fig. 6). This barrier might not be overcome in the molecular beam source when Mg collides the small fraction of *ttt* DXE molecule, which is present in the gas phase. Hence, a minor fraction of monodentate Mg-DXE complexes may exist in the beam. However, according to the assignment performed in Sec. IV C, this does not seem to be the case.

The fact that the Mg-DXE complex is easily formed in supersonic expansion is consistent with the fairly large binding energy of 1580 cm^{-1} , which is estimated from the MCSCF+IC-MRCI calculation.

B. The $(\text{Mg} - \text{DXE})^+$ molecular ion

Sec. III A shows that the Mg-DXE complex can be ionized by 212 nm (5.85 eV) photons. Given the ionization energy of free magnesium, 7.644 eV,¹³ the ionization energy of the complex is lowered by at least 1.8 eV compared to that of free magnesium. The calculation shown in Fig. 7 is consistent with this observation since the ionization energy of the bidentate complex taken along the dotted-dashed line at 48° is found to be 43082 cm^{-1} (5.34 eV), corresponding to a stabilization of the Mg^+DXE ion by 2.32 eV with respect to the dissociation as $\text{Mg}^+ + \text{DXE}$.

The R2PI spectra of Fig. 4 are extending between 27800 and 30500 cm^{-1} (3.45 to 3.78 eV). Hence, 6.90 to 7.56 eV are available for the ionization. Since the calculated ionization energy of the complex is 5.34 eV, an excess energy of 1.6 to 2.2 eV is available, which is shared between the kinetic energy of the photoelectron and the internal excitation of the molecular ion. This excess energy is not large enough for the ion to dissociate as $\text{Mg}^+ + \text{DXE}$. The fact that no fragment ion is observed in the R2PI experiment, suggests that either no other fragmentation channels is accessible energetically or, if existing, that an activation barrier makes the ion stable.

C. Electronic excitation of Mg-DXE and excited state dynamics

Fig. 7 brings the R2PI signal observed experimentally together with the calculated potential energy curves, according to the assumption that the vertical excitation proceeds either from the absolute minimum of the ground state at $\pm 48^\circ$ (bidentate ligation of Mg by the *tgt* DXE rotamer) or from the secondary minimum at $\pm 180^\circ$ (monodentate ligation by the *ttt* rotamer).

The experimental signal overlaps the calculated curve when the bidentate solvation is assumed, whereas no overlap is observed in the other case. The discrepancy is then *c.a.* 4000 cm^{-1} . This large value can hardly be assigned to the accuracy of the calculation. The agreement between experiment and calculation for the $3s^2\ ^1S \rightarrow 3s3p\ ^1P$ $3s^2\ ^1S \rightarrow 3s4s\ ^1S$ transitions of free Mg is indeed better than 340 cm^{-1} (see Sec. IIB and Tab. I) and these states correlate adiabatically to the molecular states of the Mg-DXE complex shown in Fig. 7. Hence, we consider that the accuracy on the calculated molecular state is large enough to state from Fig. 7, that the R2PI spectrum observed between 27800 and 30500 cm^{-1} is due to the Mg-DXE complex with a bidentate ligation of Mg by DXE. The fact that an intense R2PI signal is observed suggests that one or several of the electronic states of Mg-DXE, which can be accessed resonantly from the ground state bidentate complex, are sufficiently long lived to be ionized efficiently. The purpose now is to identify these states.

The Franck-Condon region of excitation of the complex, corresponding to the bidentate ligation of Mg, with the O-C-C-O dihedral angle equals to 48° , is shown as a dotted-dashed line in Fig. 7. It crosses the excited potential curves at 25492 , 27416 , 28151 and 29863 cm^{-1} as listed in Tab. II. The states, which are crossed at the three lowest energies, *i.e.* S_1 , S_2 and S_3 correlate adiabatically to $\text{Mg}(3s3p\ ^1P)$ at infinite separation between Mg and DXE. The fourth state, S_4 , correlates adiabatically to $\text{Mg}(3s4s\ ^1S)$. Excitation to this state would be symmetry forbidden in free Mg, but the transition is allowed in the complex because of the presence of DXE, which lowers the symmetry of the system and mixes the $4s$ orbital with the $3p$ orbital, which is pointing towards the central C-C bond of DXE. Hence, vertical excitation to each of these levels is possible. They are discussed in the next section as tools to assign the observed features of the R2PI spectrum to the excitation of the Mg-DXE complex to specific molecular energy states.

TABLE II

D. Identification of three bright states

According to the Franck-Condon picture, the R2PI spectrum of the Mg-DXE complex shown in Fig. 7 should feature maxima or shoulders when the photon energy matches a vertical one-photon resonance with excited states of the complex. According to the previous section, such resonances are expected at 25492, 27416, 28151 and 29863 cm^{-1} from the MCSCF+IC-MRCI calculation, corresponding to the vertical excitation of the S_1 , S_2 , S_3 and S_4 states of the complex, respectively.

The expected feature of larger energy at 29863 cm^{-1} falls within the blue shoulder of weak intensity of the absorption spectrum (between 29600 and 30500 cm^{-1}) observed in Fig. 7. We thus assign this shoulder to the excitation of the S_4 state, which correlates adiabatically to $\text{Mg}(3s4s^1\text{S})$.

The expected features at 27416 and 28151 cm^{-1} , fall within 900 cm^{-1} of the structured region of the absorption spectrum (between 28300 and 29600 cm^{-1}). This suggests that the vertical excitation to S_2 and S_3 , which correlate adiabatically to $\text{Mg}(3s3p^1\text{P})$, participate to this region of the R2PI spectrum. Moreover, Fig. 7 shows also that the equilibrium geometry of the S_2 and S_3 states is planar along the O-C-C-O dihedral angle, whereas the vertical excitation from the equilibrium geometry of S_0 forms these states in a *gauche* geometry, with the dihedral angle equal to 48° . Hence, a strong excitation of the torsion mode of the DXE moiety is expected by this excitation. This is consistent with the observation made in Sec. IIIB that the vibrational progression with $\omega_e = 110 \text{ cm}^{-1}$ observed in Fig. 4 does not involve the displacement of the Mg atom, directly. Moreover, the shape of the middle curve correlating adiabatically to $\text{Mg}(3s3p^1\text{P})$ can be adequately fitted by a parabola, consistent with a harmonic torsion oscillator with $\omega_e = 110 \text{ cm}^{-1}$. Nevertheless, the experimental spectrum cannot be fitted satisfactorily by a single progression. This observation is in line with the fact that two different electronic states participate to the R2PI spectrum in this region and perturbs each other.

E. Identification of a dark state

The last expected feature in the R2PI spectrum correspond to the vertical excitation of S_1 , which correlates with $\text{Mg}(3s3p^1P)$. It is predicted at 25492 cm^{-1} and falls 2300 cm^{-1} below the red wing of the R2PI spectrum at 27800 cm^{-1} . The region overlapping 25492 cm^{-1} was explored experimentally and no R2PI signal was observed. Nevertheless, there is no doubt from the calculation that an energy state exists in this region, and that it is accessible through a strongly allowed transition. The transition dipole moment of the $S_1 \leftarrow S_0$ vertical excitation, provided by the MCSCF+IC-MRCI calculation, is indeed 2.23 e a_0 , almost as large as the strongly allowed $3s3p^1P \leftarrow 3s^2^1S$ of free Mg (2.33 e a_0).

We thus infer that the excitation to S_1 does not lead to an observable R2PI signal because a non-adiabatic energy transfer is at play, which evacuates the electronic energy out of this state in a short time compared to the 3 ns duration of the R2PI laser pulse. A possible interpretation could be a bond rearrangement within the S_1 surface, or a strong coupling between this surface and S_0 through a conical intersection. A limited exploration of the S_1 surface was performed, which does not support this assumption. An alternative interpretation is a coupling between the S_1 surface and a triplet surface. Fig. 9 shows a correlation diagram of the first singlet and triplet states of free Mg with those of the Mg-DXE complex. It is built using the energies given in Tabs. I and II and the binding energy, which has been calculated for the ground state complex (1580 cm^{-1}). This diagram brings the important result that only the T_3 state is repulsive toward the Mg+DXE dissociation and from Tab. II we know that this state is 259 cm^{-1} above S_1 . S_1 is thus expected to be strongly coupled to the dissociative T_3 state in the excitation region and results in the formation of $\text{Mg}(3s3p^3P) + \text{DXE}$. The R2PI experiment is performed with a single color, and Fig. 9 shows that a photon exciting the S_1 of the complex from S_0 does not have enough energy to ionize free Mg in the $3s3p^3P$ state.

FIGURE 9

Hence we surmise that S_1 is a dark state because its lifetime is considerably reduced by

the coupling with the T_3 state, which dissociates as $\text{Mg}(3s3p^3P) + \text{DXE}$, the Mg atom, not being detectable with photons of 25492 cm^{-1} .

V. SUMMARY AND CONCLUSIONS

The 1:1 Mg \cdots 1,2-dimethoxyethane (Mg-DXE) complex was explored experimentally. It was generated in a laser ablation source and studied spectroscopically by Resonance Two-Photon Ionization (R2PI). The experiment was complemented by calculations, DFT/B3LYP and *ab initio*, to document the ground state and the four lowest excited electronic states, which correlate adiabatically to the $3s3p^1P$ and $3s4s^1S$ states of Mg at large separation between Mg and DXE, for a series of geometry ranging from monodentate to bidentate ligation of Mg by the oxygen atoms of DXE. A complex absorption band is observed in the R2PI spectrum between 27800 and 30500 cm^{-1} . The combined experiments and calculations show that it must be attributed to the bidentate complex, the bond between the metal and DXE being generated by the interaction of the Mg atom with the two oxygen atoms of DXE. The calculation showed that the ground state bidentate complex is 366 cm^{-1} more stable than the monodentate complex. Moreover, the DXE moiety has a *gauche* geometrical configuration at the optimum geometry of the ground state complex, also adapted to the bidentate ligation. An exploration of the O-C-C-O dihedral angle of DXE showed that the DXE moiety adopts a planar chair geometrical configuration in the excited states, which is also adapted to a bidentate ligation. Because of the geometry difference between the ground and excited states, a well pronounced progression assigned to the torsion mode along the C-C bond dominates the R2PI spectrum. Finally, the quantum calculations suggests that the observed band is actually the superposition of three electronic bands, hence explaining the complex nature of the observed R2PI spectrum.

The present operating conditions did not allow to form $\text{Mg}(\text{DXE})_2$ complexes. Nevertheless, their study would be interesting from a fundamental point of view. The geometry of such 4-ligated complexes could be quite far away from planarity as it is the case when an alkali ion is ligated by two DXE molecules.³ This situation would offer interesting comparisons with the ligation by suitable crown ethers, which are essentially planar.

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²⁰ See EPAPS Document No. xxx for the optimized geometries used in the calculations of the potential energies for the ground state (S_0) and the four lowest excited states (S_1 to S_4). This document may be retrieved via the EPAPS homepage (<http://www.aip.org/pubservs/epaps.html>) or from <ftp.aip.org> in the directory /epaps/. See the EPAPS homepage for more information.

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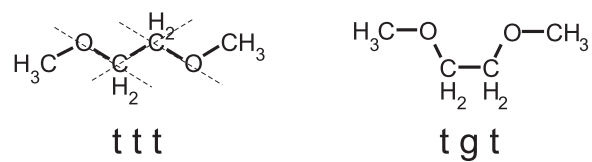


FIG. 1: Scheme of two rotamers of the 1,2-dimethoxyethane (DXE) molecule. The dashed lines on the *ttt* rotamer are the axes, which define the *trans* (*t*) and *gauche* (*g*) conformations.

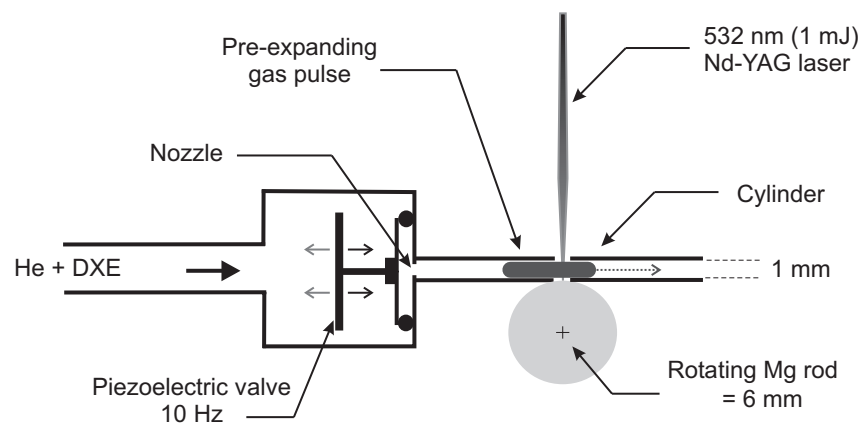


FIG. 2: Outline of the laser ablation-molecular beam experimental setup used to obtain cold coordination complexes.

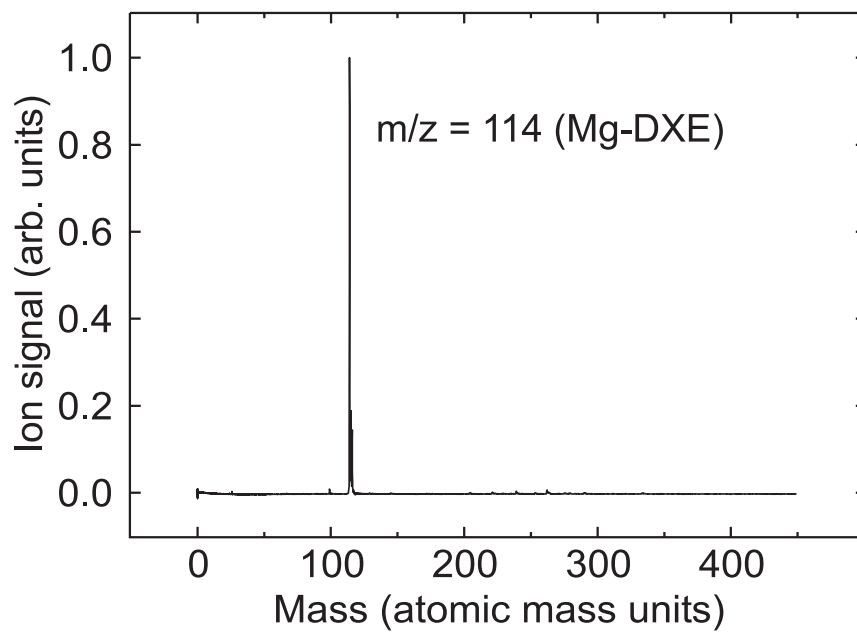


FIG. 3: Mass spectrum observed in a Resonance Two Photon Ionization (1+1) experiment at 345 nm. The expansion conditions are optimized so as to form preferentially the 1:1 Mg-DXE complex as explained in the text.

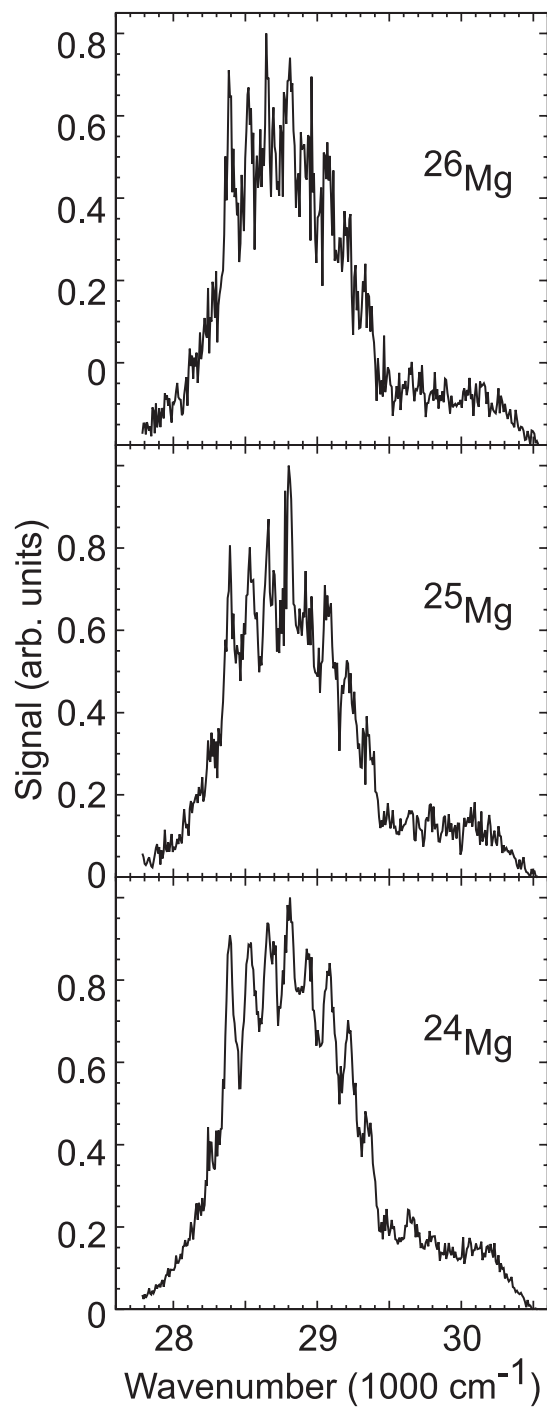


FIG. 4: R2PI spectrum of the 1:1 Mg-DXE complex. The three spectra correspond to the natural isotopes of magnesium, as labeled in the figure.

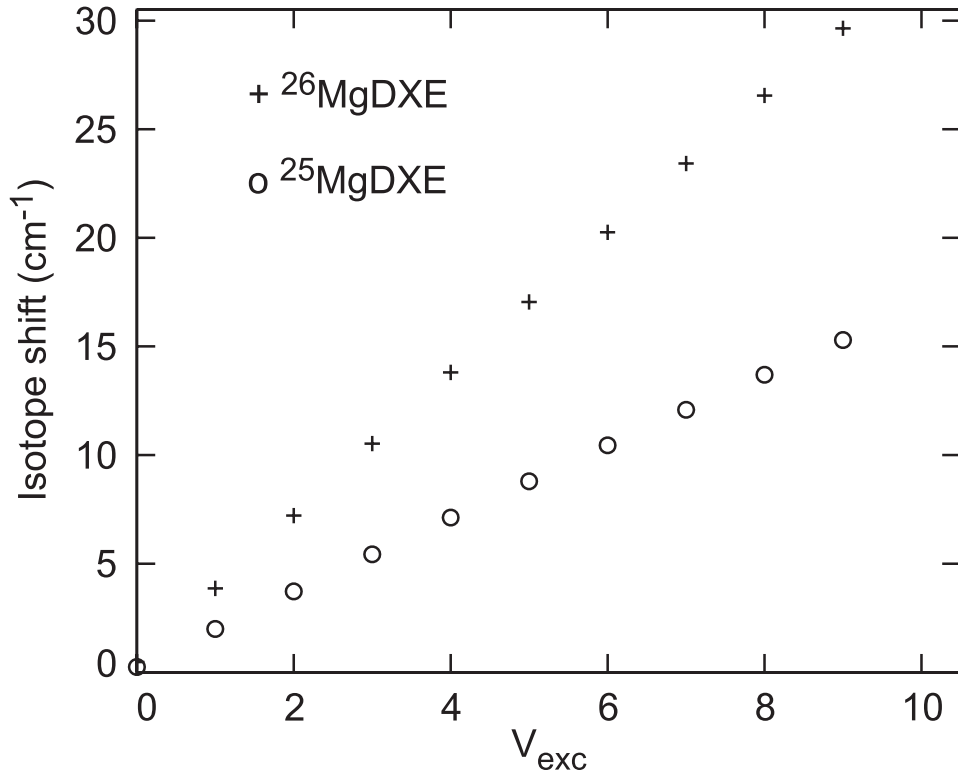


FIG. 5: Simulation of the isotopic shift for the vibrational stretching mode Mg-DXE, assuming a vibrational frequency $\omega_e = 110 \text{ cm}^{-1}$ consistent with the separation between the structures observed in Fig. 4. The isotopic shift is relative to the most abundant isotope, $^{24}\text{Mg} - \text{DXE}$. It is given as a function of the vibrational level of the excited state, v_{exc} , for $^{25}\text{Mg} - \text{DXE}$ (\circ) and $^{26}\text{Mg} - \text{DXE}$ ($+$).

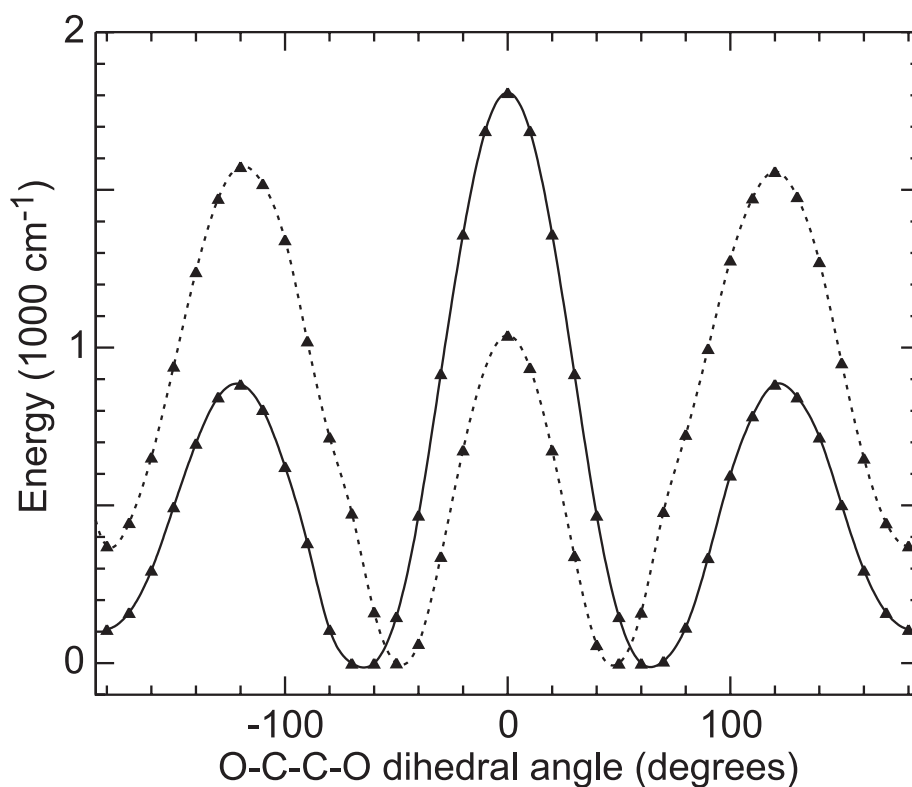


FIG. 6: Ground state potential energy of the Mg-DXE complex as a function of the O-C-C-O dihedral angle. The full and dashed curves are obtained from the DFT/B3LYP/6-311++G** and MCSCF+IC-MRCI calculations, respectively. The curve are spline fits to the calculated points, which are shown as full triangles. The O-C-C-O skeleton of DXE is planar when the O-C-C-O dihedral angle is equal to 0° and $\pm 180^\circ$, having the boat and chair conformations, respectively. See²⁰ for the optimized geometries used in the calculations.

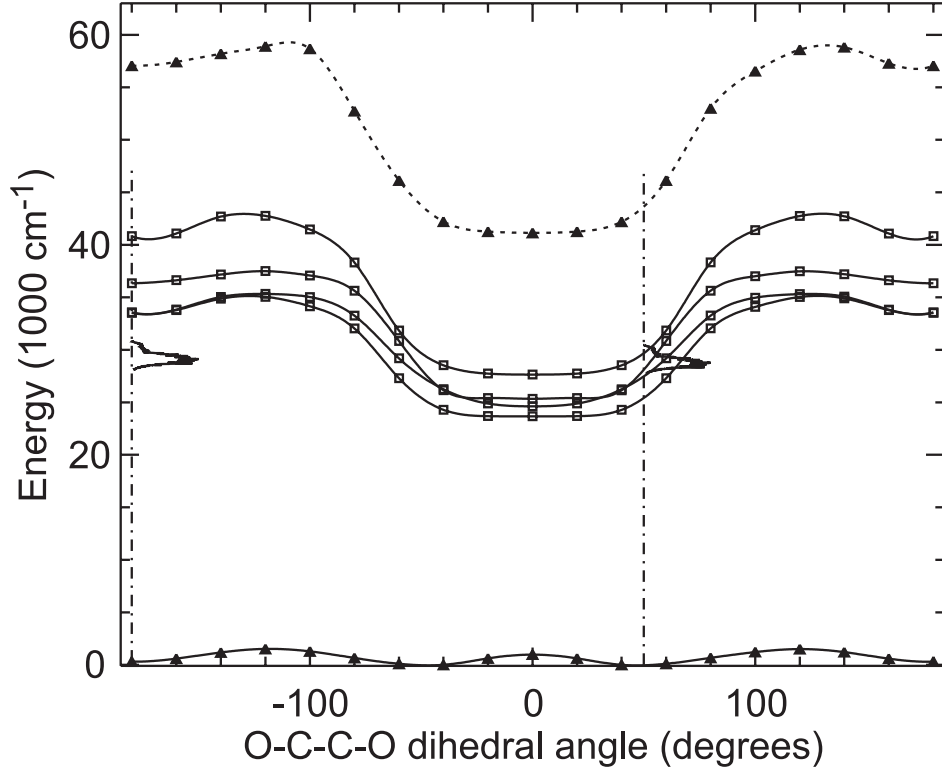


FIG. 7: Potential energies provided by the MCSCF+IC-MRCI calculations for the ground state (S_0) and the four lowest excited states (S_1 to S_4) of the Mg-DXE complex and for the ground state of the Mg^+ -DXE ion as a function of the O-C-C-O dihedral angle. The excited and ionic curves represent the vertical excitation and ionization energies from the ground state. The experimental R2PI signal of Fig. 4 is reproduced, assuming a vertical excitation from either the absolute or the secondary minimum of the ground state curve. The dotted-dashed lines are plotted at the absolute and secondary minima of the ground state curve. Their length, 47183 cm^{-1} (5.85 eV), corresponds to the energy of the photon used to characterize the beam in Sec. III A. See²⁰ for the optimized geometries used in the calculations.

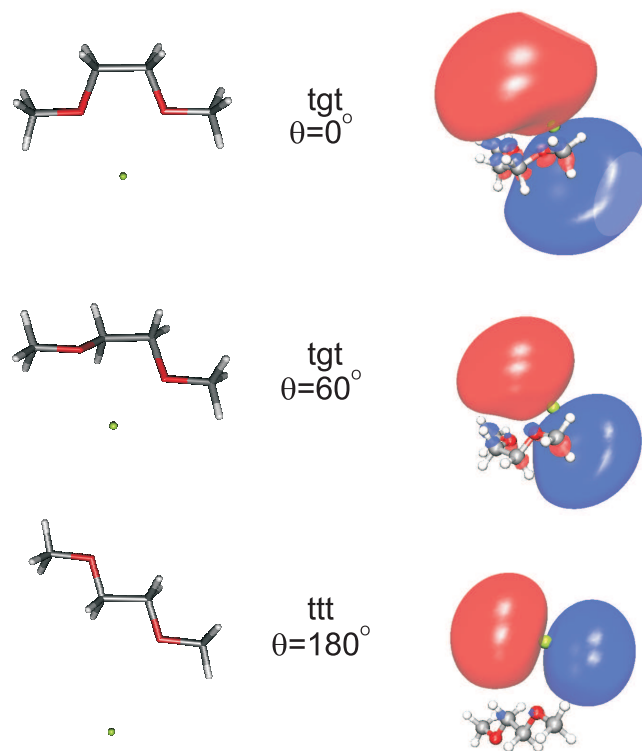


FIG. 8: Output of the Gaussian calculation (left) showing the geometry of the Mg-DXE system for three values of the O-C-C-O dihedral angle θ , 0° , 60° and 180° , from top to bottom, respectively. The Mg atom is shown in green and the O atoms in red. The corresponding first excited orbital provided by the state averaged MCSCF step of the MOLPRO calculation is shown on the right side of the figure, together with the Mg-DXE system.

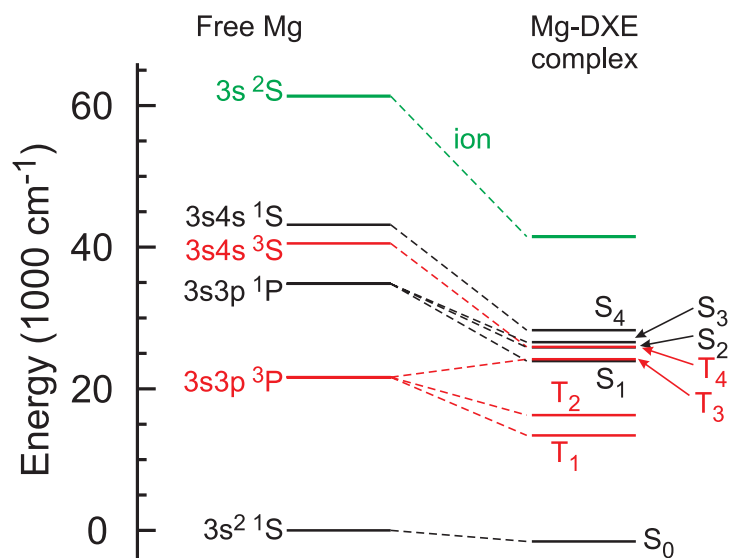


FIG. 9: Correlation diagram between the energy states of free Mg and those of the Mg-DXE complex at the equilibrium geometry of the ground state complex (O-C-C-O dihedral angle equal to 48°). Singlet and triplet states are in black and red, respectively. The correlation between Mg^+ and Mg^+ -DXE is shown in green.

TABLE I: Term symbols and ionization energies of Mg atom. The column labeled Δ is the energy difference between experiment (Exp.) and calculation (Calc.). All energy values are given in cm^{-1} .

Level	Exp. ^a	Calc. ^b	Δ
Mg			
$3s^2 \ ^1S$	0	0	
$3s3p \ ^3P$	21891	21620	271
$3s3p \ ^1P$	35051	34832	219
$3s4s \ ^3S$	41197	40540	657
$3s4s \ ^1S$	43503	43162	341
Mg ⁺			
$3s^2 \ ^2S$	61669	61330	339

^a Ref.¹³

^b This work, IC-MRCI calculations

TABLE II: Vertical excitation and ionization energies of Mg-DXE. The energies (cm^{-1}) are given for each state at 4 geometries of the DXE moiety, each characterized by the dihedral angle θ . All energies are referred to the ground state in each geometry.

Level	$\theta = 0^\circ$	$\theta = 48^\circ$	$\theta = 60^\circ$	$\theta = 180^\circ$
Mg-DXE				
S ₀	0	0	0	0
S ₁	22685	25492	27190	33202
S ₂	23637	27416	29077	33255
S ₃	24354	28151	30730	36024
S ₄	26665	29863	31729	40501
Mg ⁺ – DXE				
T ₁	13574	14983	15994	20111
T ₂	16693	17858	18580	20394
T ₃	23180	25751	25893	24021
T ₄	24431	27498	29930	38356
D ₀	40274	43082	45474	56643