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Double Structural Transition in Hybrid Material MIL-53 upon Hydrocarbon Adsorption: The Thermodynamics Behind the Scenes

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In this paper, we rationalize the thermodynamics behind the guest-induced structural transitions of hybrid material MIL-53, showing that the existence of a double transition depends on the *relative* guest affinities for the two phases, rather than absolute adsorption enthalpies. Based on a simple model, we interpret recent experimental data of alkane adsorption and predict transition pressures.

Porous metal–organic frameworks (MOFs) are a topical class of materials that display an extremely large range of crystal structures and host–guest properties, potentially giving them a major impact in adsorption, separation and storage of strategic gases (H₂, CO₂, CH₄, ...). A growing number of these materials show exceptional guest-responsive behaviors upon gas adsorption, due to the flexibility of their organic–inorganic frameworks. This includes examples of progressive swelling or contraction (also called breathing), pore deformation, and amorphous-to-crystal and crystal-to-crystal structural transitions.¹ The MIL-53 materials family,² a particularly eye-catching case of the last category, has attracted a lot of attention due to its large flexibility and the occurrence of a double structural transition upon adsorption of some gases (CO₂, H₂O, C₂H₆, ...) but not others (H₂, CH₄).³ It was also reported very recently⁴ that liquid phase adsorption of *para*-xylene does induce a structural transition, while *ortho*- and *meta*-xylene do not. The studies performed so far, both experimentally and by molecular simulation, mainly focused on structural characterization and energetics (by calorimetry, forcefield-based calculations and DFT). However, the current depiction of these guest-induced structural transitions is lacking a general thermodynamic interpretation of all the results obtained so far; Llewellyn et al. indeed highlighted the necessity of understanding “the thermodynamic conditions for the host-guest interactions which allow the breathing to happen”.³

We recently developed a generic thermodynamic framework for the understanding of guest-induced structural transitions in flexible nanoporous materials such as MOFs,⁵ by use of the osmotic pseudo-ensemble. For a material that has two possible framework structures and where gas adsorption follows type I isotherms⁶, we proposed a full taxonomy of possible guest-induced structural transitions. This classification relies only on a few key parameters, such as the free energy difference, ΔF , between the (empty) host structures, their pore volumes, $V_p^{(i)}$, and the adsorption affinities for the guest, K_i . This method also allows us to calculate one of these parameters when the pressures of structural transitions are known. The robustness of the method was demonstrated on systems exhibiting such contrasting behaviors as “breathing” and “gate opening”. In

the case of MIL-53 (Al), in particular, we used the available CO₂ adsorption isotherm to calculate a free energy difference between the empty large pore (**lp**, Fig 1S) and narrow pore (**np**, Fig 2S) forms of $\Delta F \approx 2.5$ kJ/mol. This value was then used to successfully predict the position of the low-pressure **lp**→**np** transition, confirmed by calorimetry.⁷ This rather small value of 2.5 kJ/mol is of the order of kT at room temperature, explaining the bistability of MIL-53 (Al). Indeed, all guest-responsive hybrid materials we studied so far exhibit free energy differences between host structures in the range of 2–5 kJ/mol. This is in sharp contrast with the hypothesis of Llewellyn et al. that ΔF should be much greater than kT ,³ but not incompatible with the vision of Zhang and Chen that the energy *barrier* between different structures has to be much larger than kT for materials where flexibility is thermodynamically controlled.⁸

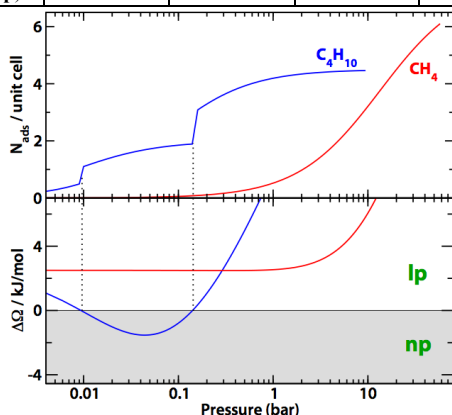
We now turn to the issue of how the presence or absence of guest-induced structural transitions can be related to the properties of the guest molecule. New experimental data on the adsorption of various species in the pores of MIL-53 has been published very recently, both in gas phase (for C_nH_{2n+2}, $n \leq 4$)³ and liquid phase (for *o*-, *m*- and *p*-xylene).⁴ In case of materials such as MIL-53, where the large-pore (**lp**) form is intrinsically more stable than the narrow-pore (**np**) form, our taxonomy predicts either the occurrence of two structural transitions upon gas adsorption, or the absence of any transition. This is determined by a balance between intrinsic stability of the crystal structures, adsorption affinities and accessible volume. In the case of alkanes adsorption in MIL-53, adsorption isotherms from Ref 3 show that the variation of pore volume (and adsorbed quantities at saturation) for both phases is small and the main factor is the change in adsorption affinities for different guests (see Table 1S). Within the limits of this model, the presence or absence of structural transitions for a given adsorbate is unambiguously determined by one factor only: the ratio K_{np}/K_{lp} of adsorption affinities in the two structures. A double guest-induced structural transition happens if and only if K_{lp}/K_{np} is large enough so that the following inequality holds:⁵

$$(V_p^{(np)} - V_p^{(lp)}) \ln \left(\frac{K_{np}/K_{lp} \times V_p^{(lp)} - V_p^{(np)}}{V_p^{(lp)} - V_p^{(np)}} \right) + V_p^{(lp)} \ln \frac{K_{np}}{K_{lp}} \geq \frac{\Delta F}{\rho RT} \quad (1)$$

If, on the contrary, K_{lp}/K_{np} is too small, there will be no structural transition at all. This conclusion that the existence of a double structural transition depends on the *relative* affinities of the adsorbate in both structures, is at variance with the view that the double transition is dictated solely by the affinity for the **lp** form (with a critical adsorption enthalpy of ~ 20 kJ/mol).³ Our findings are directly supported by very recent experiments showing that, although affinities

Table 1: Predicted pressures for the **lp**→**np** and **np**→**lp** transitions.

Guest	CO ₂	C ₂ H ₆	C ₃ H ₈	C ₄ H ₁₀
$P(\text{lp} \rightarrow \text{np})$	0.3 bar	0.17 bar	30 mbar	9.0 mbar
$P(\text{np} \rightarrow \text{lp})$	5 bar	3.3 bar	0.45 bar	0.15 bar

**Figure 1:** Upper panel: adsorption isotherms of CH₄ (in red) and C₄H₁₀ (in blue) in MIL-53 (Cr), in a Langmuir model. Lower panel: difference in osmotic potential between **lp** and **np** phases, as a function of pressure. Vertical dotted lines correspond to C₄H₁₀-induced structural transition.

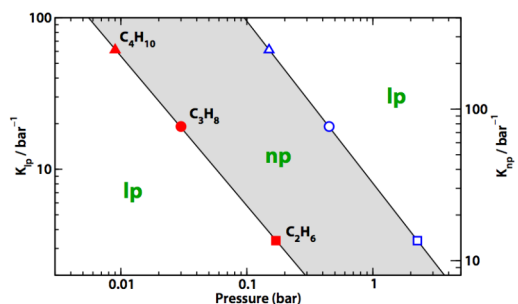
of xylene isomers for MIL-53 (Al) follow the order *ortho* > *meta* > *para*, only *para*-xylene induces a double structural transition upon adsorption.⁴

By way of illustration, we show in Figure 1 the adsorption isotherms of C₄H₁₀ and CH₄ in MIL-53 (Cr) at 303 K, using for each structure (**lp** and **np**) Langmuir isotherms fitted on experimental data³ (the parameters are detailed in the Supporting information). For C₄H₁₀, $K_{\text{np}}/K_{\text{lp}} \approx 4$ and there are two successive structural transitions ($\Delta\Omega$ changes sign twice); this leads to an isotherm where two steps can be clearly seen in logarithmic pressure scale. Contrarily, for CH₄, $K_{\text{np}}/K_{\text{lp}} \approx 1$ and the **lp** structure is thermodynamically favored throughout the pressure range; the isotherm, being simply that of CH₄ in MIL-53 **lp**, has no step.

Finally, we focus on the linear relationship observed by Llewellyn et al. between the **np**→**lp** transition pressure and the adsorption enthalpies of gases in the **np** and **lp** structures. To understand this, we study the evolution of the pressure at which structural transitions happen in the material as a function of K_{np} and K_{lp} . To relate these adsorption affinities to adsorption enthalpies and entropies, we write:

$$K_i = A \exp(-\Delta H_i / RT) \exp(\Delta S_i / R) \quad (2)$$

We can then show (see Supporting Information) that the pressure of both the **np**→**lp** and **lp**→**np** structural transitions exhibits a linear $\log(P)$ vs. ΔH relation, if the $K_{\text{np}}/K_{\text{lp}}$ ratio is constant. By virtue of Eq 2, this condition translates to $\Delta\Delta H = \Delta H_{\text{np}} - \Delta H_{\text{lp}}$ and $\Delta\Delta S = \Delta S_{\text{np}} - \Delta S_{\text{lp}}$ being constant among a series of adsorbates. While the first equality for alkane adsorption in MIL-53 (Cr) can be verified from the Monte Carlo results presented in Ref 3, we suggest that the second is a consequence of the often observed linear relationship between ΔH_{ads} and ΔS_{ads} for adsorption of alkanes in microporous materials.⁹ We thus conclude that the logarithmic relationship between P and adsorption enthalpies is not a universal feature but a direct consequence of a constant difference between adsorption enthalpies in both forms of MIL-53. Furthermore, we predict that the pressure

**Figure 2:** Existence domains of the **lp** and **np** phases of MIL-53 upon gas adsorption, with guests of different affinities (K_{lp} and K_{np}). Symbols correspond to **lp**→**np** and **np**→**lp** (respectively filled and open) transitions for the C₂ to C₄ alkanes.

of the **lp**→**np** transition follows this same logarithmic relationship, enabling us to predict transition pressures that have not been experimental measured yet. Finally, we can draw a phase diagram (Fig 2) of MIL-53 (Cr) as a function of guest pressure and host-guest affinities, for a series of adsorbates with constant $K_{\text{lp}}/K_{\text{np}}$ ratio. We show on this phase diagram the points corresponding to the phase transitions induced by C₂H₆, C₃H₈ and C₄H₁₀. It is noteworthy that the pressure domain of the **np** phase, which appears to have a constant width in the logarithmic scale of Fig 2, actually gets smaller for adsorbates with larger affinities for the solid. Figures 3S and 4S show additional 2D and 3D cuts in the phase diagram as described by our model.

In conclusion, we have elucidated and rationalized the thermodynamics of the “breathing” phenomenon of MIL-53, highlighting the case of hydrocarbon adsorption. We demonstrated that the existence of the breathing is determined by the relative affinities of the sorbate for the **lp** and **np** phases, and proposed an explanation for the linear $\log(P)$ vs. ΔH_{ads} relation observed by Llewellyn et al.³

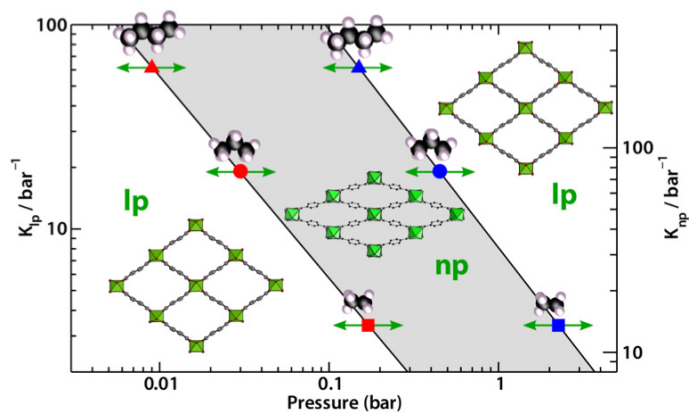
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Supporting Information Available: Calculation details, structures of MIL-53 (Cr) in **lp** and **np** form and 2D representations of the solutions of Eq 1. This material is available free of charge via the Internet at <http://pubs.acs.org>.

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Abstract

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