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**The chemistry  
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# The chemistry influencing ODEs in the Polar Boundary Layer in spring: a model study

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Title Page

Abstract

Introduction

Conclusions

References

Tables

Figures

◀

▶

◀

▶

Back

Close

Full Screen / Esc

Printer-friendly Version

Interactive Discussion



## Abstract

Near-total depletions of ozone have been observed in the Arctic spring since the mid 1980s. The autocatalytic cycles involving reactive halogens are now recognized to be of main importance for Ozone Depletion Events (ODEs) in the Polar Boundary Layer (PBL). We present sensitivity studies using the model MISTRA in the box-model mode on the influence of chemical species on these ozone depletion processes. In order to test the sensitivity of the chemistry under polar conditions, we compared base runs undergoing fluxes of either Br<sub>2</sub>, BrCl, or Cl<sub>2</sub> to induce ozone depletions, with similar runs including a modification of the chemical conditions. The role of HCHO, H<sub>2</sub>O<sub>2</sub>, DMS, Cl<sub>2</sub>, C<sub>2</sub>H<sub>4</sub>, C<sub>2</sub>H<sub>6</sub>, HONO, NO<sub>2</sub>, and RONO<sub>2</sub> was investigated. Cases with elevated mixing ratios of HCHO, H<sub>2</sub>O<sub>2</sub>, DMS, Cl<sub>2</sub>, and HONO induced a shift in bromine speciation from Br/BrO to HOBr/HBr, while high mixing ratios of C<sub>2</sub>H<sub>6</sub> induced a shift from HOBr/HBr to Br/BrO. Cases with elevated mixing ratios of HONO, NO<sub>2</sub>, and RONO<sub>2</sub> induced a shift to BrNO<sub>2</sub>/BrONO<sub>2</sub>. The shifts from Br/BrO to HOBr/HBr accelerated the aerosol debromination, but also increased the total amount of deposited bromine at the surface (mainly via increased deposition of HOBr). These shifts to HOBr/HBr also hindered the BrO self-reaction. In these cases, the ozone depletion was slowed down, where increases in H<sub>2</sub>O<sub>2</sub> and HONO had the greatest effect. The tests with increased mixing ratios of C<sub>2</sub>H<sub>4</sub> highlighted the decrease in HO<sub>x</sub> which reduced the production of HOBr from bromine radicals. In addition, the direct reaction of C<sub>2</sub>H<sub>4</sub> with bromine atoms led to less available reactive bromine. The aerosol debromination was therefore strongly reduced. Ozone levels were highly affected by the chemistry of C<sub>2</sub>H<sub>4</sub>. Cl<sub>2</sub>-induced ozone depletions were found unrealistic compared to field measurements due to the rapid production of CH<sub>3</sub>O<sub>2</sub>, HO<sub>x</sub>, and ROOH which rapidly convert reactive chlorine to HCl in a “chlorine counter-cycle”. This counter-cycle efficiently reduces the concentration of reactive halogens in the boundary layer. Depending on the relative bromine and chlorine mixing ratios, the production of CH<sub>3</sub>O<sub>2</sub>, HO<sub>x</sub>, and ROOH from the counter-cycle can significantly affect the bromine chemistry. Therefore, the pres-

## The chemistry influencing ODEs

M. Piot and R. von Glasow

Title Page

Abstract

Introduction

Conclusions

References

Tables

Figures

◀

▶

◀

▶

Back

Close

Full Screen / Esc

Printer-friendly Version

Interactive Discussion



ence of both bromine and chlorine in the air may unexpectedly lead to a slow down in ozone destruction. For all  $\text{NO}_y$  species studied ( $\text{HONO}$ ,  $\text{NO}_2$ ,  $\text{RONO}_2$ ) the chemistry is characterized by an increased bromine deposition on snow reducing the amount of reactive bromine in the air. Ozone is less depleted under conditions of high mixing ratios of  $\text{NO}_x$ . The production of  $\text{HNO}_3$  led to the acid displacement of  $\text{HCl}$ , and the release of chlorine out of salt aerosols ( $\text{Cl}_2$  or  $\text{BrCl}$ ) increased.

## 1 Introduction

Since the first reports of dramatic decreases of ozone (so-called ODEs) in the Arctic during spring (Bottenheim et al., 1986; Oltmans and Komhyr, 1986; Barrie et al., 1988), intensive efforts have been made to better understand the processes involved in the observed ozone loss. These events with low mixing ratios of ozone were correlated with high concentrations of filterable bromine (f-Br) (Barrie et al., 1988, 1989; Bottenheim et al., 1990). Later, BrO was observed with Long-Path DOAS (LP-DOAS, see Hausmann and Platt, 1994). Lehrer et al. (1997) showed a striking correlation between f-Br and BrO. During the ALERT2000 campaign,  $\text{Br}_2$  and BrCl were measured with mixing ratios as high as 30–35  $\text{pmol mol}^{-1}$  in April, while  $\text{Cl}_2$  was not observed above its detection limit of about 2  $\text{pmol mol}^{-1}$  (Foster et al., 2001; Spicer et al., 2002). Similarly, results from the chemical amplification method used by Perner et al. (1999) showed mixing ratios of  $\text{ClO}_x$  (Cl+ClO) not exceeding 2  $\text{pmol mol}^{-1}$  in spring. Only the measurements by Tuckermann et al. (1997) in Spitzbergen in 1995 indicated ClO mixing ratios up to 21  $\text{pmol mol}^{-1}$ . These values were not observed, however, in the follow-up campaign in 1996. During the TOPSE aircraft program Ridley et al. (2003) showed that ozone depletions were actually widespread in the Arctic region. Indeed, Zeng et al. (2003) estimated that about 20% of the Arctic regions were influenced by persistent near-surface ODEs in spring. This is in accordance with remote sensing data from satellites measuring column BrO (Richter et al., 1998; Wagner and Platt, 1998). For more details on polar ODEs, see Simpson et al. (2007b).

## The chemistry influencing ODEs

M. Piot and R. von Glasow

Title Page

Abstract

Introduction

Conclusions

References

Tables

Figures

◀

▶

◀

▶

Back

Close

Full Screen / Esc

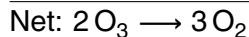
Printer-friendly Version

Interactive Discussion



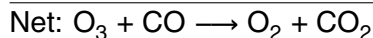
Under ODE conditions halogen catalytic reaction cycles are responsible for the depletion of O<sub>3</sub>. Halogens (X, Y=Br, Cl, I) directly destroy ozone via three main cycles (see, e.g., [von Glasow and Crutzen, 2007](#); [Simpson et al., 2007b](#)). Rate coefficients between iodine and ozone are highest, but we do not focus on iodine chemistry in this paper. Bromine is the most abundant and therefore, the most efficient halogen species for the ozone destruction:

Cycle I:



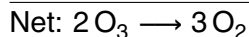
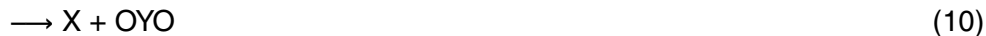
This cycle I is the fastest ozone-depleting reaction cycle for X=Br.

Cycle II:



Reaction (5) is very fast and represents a main pathway for the production of HOX.

Cycle III:



**The chemistry  
influencing ODEs**

M. Piot and R. von  
Glasow

Title Page

Abstract

Introduction

Conclusions

References

Tables

Figures

◀

▶

◀

▶

Back

Close

Full Screen / Esc

Printer-friendly Version

Interactive Discussion



The interhalogen Reactions (8) to (10) have similar effects on ozone to those of bromine itself.

If enough CH<sub>3</sub>O<sub>2</sub> is present in the air mass, the following reaction might become a substantial pathway to convert BrO into other reactive halogen species (Aranda et al., 1997):



HO<sub>2</sub> impacts the speciation of bromine species via reaction with BrO (see Reaction 5). In addition, it directly reacts with Br atoms to form HBr:



Reaction (5) is the most efficient one for the modification of the bromine speciation, followed by Reactions (11–12) (with an order of magnitude less efficient). The reaction rates of Reaction (13) remain usually small. However, it may become important when the ratio [Br]:[BrO] increases (e.g., when ozone depletion is nearly complete).

In order to sustain a significant amount of reactive halogens in the gas phase, additional mechanisms involving the liquid and solid phases (aerosols, ice crystals/snow) are necessary to activate halides and recycle less-reactive gas phase halogens. In the early 1990s Fan and Jacob (1992) and McConnell et al. (1992) suggested the following important heterogeneous reaction path for the liberation of Br<sub>2</sub> and BrCl from sea salt (based on data from Eigen and Kustin, 1962), involving HOBr:



The proportional release of Br<sub>2</sub>/BrCl from Reaction (14) was experimentally studied by Fickert et al. (1999). The required acidity may be supplied by strong acids, such as man-made HNO<sub>3</sub> and H<sub>2</sub>SO<sub>4</sub> (see also Mozurkewich, 1995; Tang and McConnell, 1996; Vogt et al., 1996). Reaction (14), releasing two bromine atoms in the gas phase from the uptake of HOBr is called the “bromine explosion” (Platt and Lehrer, 1996). The

**The chemistry  
influencing ODEs**M. Piot and R. von  
Glasow

Title Page

Abstract

Introduction

Conclusions

References

Tables

Figures

◀

▶

◀

▶

Back

Close

Full Screen / Esc

Printer-friendly Version

Interactive Discussion



**The chemistry  
influencing ODEs**M. Piot and R. von  
Glasow

Title Page

Abstract

Introduction

Conclusions

References

Tables

Figures

◀

▶

◀

▶

Back

Close

Full Screen / Esc

Printer-friendly Version

Interactive Discussion



source of halogens in the PBL remains unclear. Recently, efforts to locate the source focused on freshly-new sea ice and frost flowers (Martin et al., 1995; Rankin et al., 2000; Kaleschke et al., 2004; Dominé et al., 2005), as they provide adequate surfaces, called “brine” (Richardson and Keller, 1966), containing highly concentrated halides. Simpson et al. (2007a) found a high correlation between fresh sea ice and high levels of BrO.

Box models (Fan and Jacob, 1992; McConnell et al., 1992; Tang and McConnell, 1996; Sander et al., 1997; Michalowski et al., 2000; Evans et al., 2003) and one-dimensional models (Lehrer et al., 2004; Piot and von Glasow, 2007) have investigated the reaction cycles and have increased our understanding of the halogen/ozone chemical processes. Piot and von Glasow (2007) modeled meteorological and chemical processes that may influence the occurrence of an ODE with the snow surface acting as an efficient recycling surface. They also investigated the relevance of observed levels of non-halogen species on the halogen/ozone chemistry. However, the relative importance of these relevant species in the Arctic on the ozone chemistry has to be better understood.

Numerous concentration and flux measurements of compounds in the Arctic have been reported in the literature, but the variability of these compounds in relation to the ozone/halogen chemistry needs a clearer investigation to better understand the occurrence of an ODE. In this paper, we examined the potential effects of the presence of several chemical species in the PBL on the ozone level. Several modifications (in mixing ratios or fluxes) compared to conditions in base runs are investigated for HCHO, H<sub>2</sub>O<sub>2</sub>, DMS, Cl<sub>2</sub>, C<sub>2</sub>H<sub>4</sub>, C<sub>2</sub>H<sub>6</sub>, HONO, NO<sub>2</sub>, and RONO<sub>2</sub>. Implications for the chemistry in the PBL are discussed in Sects. 4.3 to 4.9.

## 2 Model description and setup

We used the model MISTRA which was initially designed to study the physics and chemistry of the marine boundary layer (von Glasow et al., 2002a,b; von Glasow and

Crutzen, 2004) but was recently extended and applied also to Arctic conditions (Piot and von Glasow, 2007). MISTRA is a one-dimensional model which can also be used as a box model which is what we did for this study. In this study, we used the box model in the Lagrangian mode.

We chose the solar conditions (relevant for the calculation of photolysis frequencies) for Alert, Canada (82.5°N, 62.3°W) as in early spring (solar declination of +7°). We chose a boundary layer height of 300 m, as often observed during spring in the high Arctic (Hopper and Hart, 1994; Hopper et al., 1998; Ridley et al., 2003). The relevant meteorological parameters in the box-model mode are the temperature, relative humidity, and the particle size distribution. The uptake of gases by aerosols is temperature dependent (Schwartz, 1986) because the molecular speed, the gas phase diffusivity, and the reactive uptake depend on the temperature. We want to stress, however, that the temperature dependencies of the uptake parameters are based on estimates as found in the literature.

In the sensitivity studies we present in this paper we distinguish an air mass influenced by the presence of sea water (coastal conditions) from an aged air mass over snow-covered areas (background conditions). The air mass composition is modified by the presence of sea water (see Table 1). Sea salt aerosols (SSA) are only produced in runs under coastal conditions. They provide an additional source of potentially releasable bromine (Reaction 14) and represent an important medium for recycling less reactive bromine compounds.

SSAs are produced by bursting bubbles at the sea surface (Woodcock, 1953; Prupacher and Klett, 1997). We used the parameterization from Monahan et al. (1986) to estimate the flux of these particles (with wind speed  $u=5\text{ m s}^{-1}$ ). The initial composition of SSAs is listed in Table 2. According to measurements made by Koop et al. (2000) SSAs and sulfate aerosols are liquid under the conditions of our model runs.

The very stiff chemical differential equation system is solved with a Rosenbrock third-order integrator using automatically adjusted timestep ( $\Delta t=10^{-10}$  to 10 s) (Sandu et al., 1996).

**The chemistry  
influencing ODEs**M. Piot and R. von  
Glasow

Title Page

Abstract

Introduction

Conclusions

References

Tables

Figures

◀

▶

◀

▶

Back

Close

Full Screen / Esc

Printer-friendly Version

Interactive Discussion





**The chemistry  
influencing ODEs**M. Piot and R. von  
Glasow

Title Page

Abstract

Introduction

Conclusions

References

Tables

Figures

◀

▶

◀

▶

Back

Close

Full Screen / Esc

Printer-friendly Version

Interactive Discussion



All chemical reactions in gas and aqueous phase, equilibria and transfer reactions are calculated as one coupled system using the kinetic preprocessor KPP which allows rapid change of the chemical mechanism (Damian et al., 2002). The temperature, humidity and particle size distribution are updated every 10 s. Photolysis frequencies are calculated online (Landgraf and Crutzen, 1998) every 10-min timestep with a four-stream model.

We used a two-day, one-dimensional model run to calculate the photolysis rates at the lowermost layer (5 m), used in the initialization of the box model. SSAs are produced during this “spinup” run. The box-model sensitivity runs last four days.

A diurnal variation of the temperature ( $\Delta T=2$  K) is prescribed, inducing a temperature range from  $T=243.5$  to  $245.5$  K in the model (see Jobson et al., 1994; Hopper et al., 1994; Ridley et al., 2003).

The initial mixing ratios for gas phase species ( $\xi$ ) are based on observations made at Alert, in April (Table 3). MISTRA includes a comprehensive set of gas phase reactions as well as chemical reactions in aerosol particles focusing on halogen species. Exchanges between the two phases are also taken into account. The set of used reactions is similar to von Glasow et al. (2002b), but updated with data from the IUPAC compilation (February 2006, available from <http://www.iupac-kinetic.ch.cam.ac.uk/>).

The chemical reaction mechanism in the model has been updated with additional relevant species for the Arctic environment: alkyl nitrate  $\text{RONO}_2$  (with the alkyl group  $\text{R}=\text{C}_n\text{H}_{2n+1}$ ), bromoform  $\text{CHBr}_3$  and methyl bromide  $\text{CH}_3\text{Br}$ . The model includes 169 gas phase reactions (H-O-S-C-N-Br-Cl), as well as 150 aqueous phase reactions, 60 phase exchange reactions, 13 heterogeneous reactions and 21 equilibria for both sulfate and sea salt aerosols. The updated set of reactions is available as electronic attachment to this paper. <http://www.atmos-chem-phys-discuss.net/8/7391/2008/acpd-8-7391-2008-supplement.pdf>

All particles above their deliquescence humidity (70% and 75% of relative humidity for sulfate and sea salt aerosols, respectively) are treated as aqueous solutions. The same holds for particles that had been activated or have been released as droplets

above their crystallisation humidity (40% and 42%, respectively). A detailed description of the aerosol characteristics can be found in [von Glasow et al. \(2002b\)](#).

Under coastal conditions, we prescribed surface fluxes of dimethylsulfide (DMS) and ammonia (NH<sub>3</sub>), as well as mixing ratios of biogenic bromine (see Table 1).

### 5 3 Chemical reaction cycles

This section is meant to give an introduction to reaction cycles that will be discussed in the remainder of this paper. For a more thorough discussion please refer to, e.g., [Seinfeld and Pandis \(1998\)](#) or [Finlayson-Pitts and Pitts \(1999\)](#).

#### 3.1 HCHO and HO<sub>x</sub> chemistry

10 Formaldehyde (HCHO) is an important source of oxidizing free radicals (HO<sub>x</sub>) for the PBL, and has therefore received considerable interest in the last decade ([Barrie et al., 1994](#); [Sumner and Shepson, 1999](#); [Hutterli et al., 2002](#); [Jacobi et al., 2002](#); [Ridley et al., 2003](#); [Jacobi et al., 2004](#)). Measurements in the air, firn air, and in snow revealed variations of HCHO determined by several processes. Its uptake/release between the snow  
15 and adjacent firn air is temperature-dependent. Additionally, photochemical reactions and ventilation of the firn air to the above layers play an important role in the production of HCHO from the snowpack. For more details on the production of HCHO the reader is referred to [Simpson et al. \(2007b\)](#). Its main loss reactions are:



If present in sufficient amounts, HCHO may modify the bromine chemistry via:



Title Page

Abstract

Introduction

Conclusions

References

Tables

Figures

◀

▶

◀

▶

Back

Close

Full Screen / Esc

Printer-friendly Version

Interactive Discussion





The reaction between HCHO and chlorine is discussed in Sect. 3.3. The chemistry of formaldehyde is investigated in Sect. 4.3.

Hydrogen peroxide (H<sub>2</sub>O<sub>2</sub>) in the Arctic is found in high concentrations in snow. Snow-air fluxes have been measured at several sites in polar regions (Führer et al., 1996; Hutterli et al., 2001; Jacobi et al., 2002; Hutterli et al., 2004). It is reversibly deposited to the snow due to nonlinear processes occurring between the atmosphere and the snow (Conklin et al., 1993; Neftel et al., 1995). H<sub>2</sub>O<sub>2</sub> has a significant impact on the lifetime of trace gases as it constitutes a large potential source for gas phase oxidants (HO<sub>x</sub>) which contribute to the atmospheric oxidizing capacity. H<sub>2</sub>O<sub>2</sub> is destroyed in Reactions (20–22) (producing HO<sub>x</sub>) and is produced in Reaction (23).



The chemistry of H<sub>2</sub>O<sub>2</sub> is discussed in Sect. 4.4.

### 3.2 DMS chemistry

Charlson et al. (1987) suggested a role of dimethylsulfide (DMS) in regulating climate. This effect, if present, remains ill-quantified though. Toumi (1994), based on kinetic data from Barnes et al. (1991), was the first to suggest that BrO might play an important role in the oxidation of DMS. von Glasow and Crutzen (2004) discussed uncertainties in the oxidation of DMS and pointed out that the net effect of DMS oxidation products, considering BrO as oxidant, on clouds might be contrary to those suggested by Charlson et al. (1987), namely a decrease of cloud albedo instead of an increase.

**The chemistry  
influencing ODEs**

M. Piot and R. von  
Glasow

Title Page

Abstract

Introduction

Conclusions

References

Tables

Figures

◀

▶

◀

▶

Back

Close

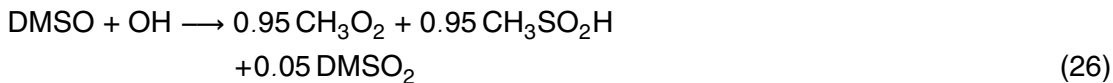
Full Screen / Esc

Printer-friendly Version

Interactive Discussion



The following are the most important halogen-DMS reactions:



The interactions between DMS and bromine are discussed in Sect. 4.5.

### 3.3 Chlorine chemistry

The chemistry of chlorine has been extensively studied. Keene et al. (1999) provided an inventory for sources of reactive chlorine. Jobson et al. (1994) were the first to indirectly measure concentrations of chlorine in the Arctic troposphere.

However, low chlorine levels observed in the Arctic indicated its minor role in ozone depletions (Perner et al., 1999; Foster et al., 2001; Spicer et al., 2002). Nevertheless, chlorine may considerably modify hydrocarbons and radical budgets (Jobson et al., 1994; Ariya et al., 1998, 1999; Ramacher et al., 1999): the destruction of ozone by chlorine chemistry is described in Sect. 1. Additionally, if substantial levels of chlorine are present the following reactions may become important:



where ROOH is a hydroperoxide (see, e.g., Jacob, 2000; Frey et al., 2006) with  $\text{R}=\text{C}_n\text{H}_{2n+1}$ . Chlorine chemistry is investigated in Sect. 4.6.

Title Page

Abstract

Introduction

Conclusions

References

Tables

Figures

◀

▶

◀

▶

Back

Close

Full Screen / Esc

Printer-friendly Version

Interactive Discussion



### 3.4 C<sub>2</sub>H<sub>4</sub> chemistry

Measurements of fluxes of alkenes from the snow have shown the presence of local sources for ethene (C<sub>2</sub>H<sub>4</sub>) in the Arctic. [Bottenheim et al. \(2002a\)](#) measured atmospheric mixing ratios as high as 100 pmol mol<sup>-1</sup> in early spring with a slow decrease with season. In our set of reactions, three reactions characterize the chemistry of C<sub>2</sub>H<sub>4</sub>:



The rate coefficients for Reactions (32) and (33) are about 200 and 2000 times higher than for Reaction (34), respectively. The reactions involving C<sub>2</sub>H<sub>5</sub>O<sub>2</sub> and H<sub>2</sub>C(OH)CH<sub>2</sub>OO will not be further detailed in this paper as their reaction rates remain very small under polar conditions. In MISTRA the possible formation of organically-bound bromine compounds as reaction product of Reaction (34) ([Keil and Shepson, 2006](#)) is not taken into account. Other reactions with heavier alkenes are not included in the model. We investigated the effect of these three reactions in Sect. 4.7.

### 3.5 C<sub>2</sub>H<sub>6</sub> chemistry

Alkanes are commonly present in the Arctic due to transport from the source regions (Eurasia and northern America mostly, see [AMAP report, 1998](#)). In April, ethane (C<sub>2</sub>H<sub>6</sub>) is usually observed in the surface air at about 1.5–2.5 nmol mol<sup>-1</sup> with a gradual decline with season ([Jobson et al., 1994](#); [Ariya et al., 1999](#); [Bottenheim et al., 2002a](#); [Evans et al., 2003](#)). In our model C<sub>2</sub>H<sub>6</sub> chemistry includes the following reactions:



Title Page

Abstract

Introduction

Conclusions

References

Tables

Figures

◀

▶

◀

▶

Back

Close

Full Screen / Esc

Printer-friendly Version

Interactive Discussion



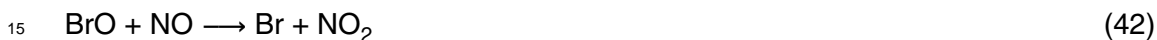
C<sub>2</sub>H<sub>6</sub> does not directly react with bromine. As for alkenes, heavier alkane compounds are not included in the model. The sensitivity studies on the influence of C<sub>2</sub>H<sub>6</sub> on the ozone/halogen chemistry are presented in Sect. 4.8.

### 3.6 NO<sub>x</sub> chemistry

5 NO<sub>x</sub> species are photochemically produced in the snowpack (Honrath et al., 1999, 2000a,b; Zhou et al., 2001; Beine et al., 2002; Jacobi et al., 2004). The major NO<sub>x</sub> reactions are listed below.



With sufficiently high NO<sub>x</sub> concentrations, the concentrations of bromine radicals may be altered via:



After their formation, bromine nitrate and nitrite may decompose back or photolyze to release reactive bromine:



## The chemistry influencing ODEs

M. Piot and R. von Glasow

Title Page

Abstract

Introduction

Conclusions

References

Tables

Figures

◀

▶

◀

▶

Back

Close

Full Screen / Esc

Printer-friendly Version

Interactive Discussion



BrONO<sub>2</sub> may also hydrolyse on liquid surfaces via:



Also, [Sander et al. \(1999\)](#) suggested that BrONO<sub>2</sub> may heterogeneously react with aerosol halides without required acidity (based on experimental observations from [Behnke et al., 1997](#)):



However, in the Arctic, the release of reactive bromine under background NO<sub>x</sub> levels via this reaction is relatively small (100 times less efficient than Reaction 14, due to a small rate coefficient for Reaction 47).

The chemistry of nitrous acid (HONO) plays an important role for the complete nitrogen cycle in the troposphere ([Perner and Platt, 1979](#); [Heikes and Thompson, 1983](#)). Measurements in the Arctic during spring indicated intensive photochemical productions in the snowpack constituting a major source of HONO for the boundary layer ([Zhou et al., 2001](#); [Beine et al., 2003](#); [Amoroso et al., 2006](#)). The only relevant sink for HONO is its photolysis producing both highly reactive OH and NO molecules:



RONO<sub>2</sub> is likely to contribute to the NO<sub>x</sub> budget ([Brasseur et al., 1999](#)). For a detailed description of the production and loss of RONO<sub>2</sub>, see [Carter and Atkinson \(1985\)](#). In the Arctic RONO<sub>2</sub> mainly reacts via the following reactions:



The chemistry of the NO<sub>y</sub> species (NO<sub>x</sub>, HONO, and RONO<sub>2</sub>) is investigated in Sect. 4.9.

Title Page

Abstract

Introduction

Conclusions

References

Tables

Figures

◀

▶

◀

▶

Back

Close

Full Screen / Esc

Printer-friendly Version

Interactive Discussion



## 4 Discussion of the sensitivity studies

### 4.1 Overview

We prescribed a source of gas phase halogens to reproduce ozone depletions with observed time scales. In this paper, for the sake of simplicity, we did not intend to realistically reproduce the source of bromine in the model. For a thorough investigation of the source of bromine, the reader is referred to the recent paper from [Piot and von Glasow \(2007\)](#). We quantified the required fluxes of halogen in accordance with their effects on ozone. Then, we associated the resulting ozone depletions with the classification proposed by [Ridley et al. \(2003\)](#) to define the type of ODE. Regarding ozone mixing ratios, we define a partial ODE (PODE) as  $4 < \xi_{\text{Ozone}} \leq 20 \text{ nmol mol}^{-1}$  and a major ODE (MODE) as  $\xi_{\text{Ozone}} \leq 4 \text{ nmol mol}^{-1}$ . We distinguished major ODEs developing within one day from major ODEs developing within four days (called M1 and M4). Partial ODEs developing within 4 days are named P4. Depletions of ozone are reproduced via different sources of halogens. In this study, we analyzed  $\text{Br}_2$ ,  $\text{Cl}_2$  and  $\text{BrCl}$  as the potential sources. Note that no diurnal variation of fluxes is taken into account in this study. The applied source of halogens is prescribed as a constant flux (see Table 4).

Excessive fluxes of chlorine are discussed in Sect. 4.6.

The runs including a prescribed source of halogens and initialized as shown in Table 3 represent our “base runs”. Figure 2 illustrates the M1, M4 and P4 ODEs caused by the prescribed  $\text{Br}_2$  fluxes (Table 4). The names of our sensitivity runs include the source of halogens ( $\text{Br}_2$ ,  $\text{BrCl}$ , or  $\text{Cl}_2$ ) and additionally any change compared to the base run. The run  $\text{Br}_2\text{-M4}$  is one of our base runs with a major ozone depletion developing within four days caused by a  $\text{Br}_2$  flux. The run  $\text{Br}_2\text{-M4-DMS}=0.20 \text{ nmol mol}^{-1}$  is the same run with a different value for the initial mixing ratio of DMS than in the base run.

Our sensitivity studies consist of modifying the amount of a species compared to the amount in the base runs. We want to stress that only relevant cases with infor-

### The chemistry influencing ODEs

M. Piot and R. von Glasow

Title Page

Abstract

Introduction

Conclusions

References

Tables

Figures

◀

▶

◀

▶

Back

Close

Full Screen / Esc

Printer-friendly Version

Interactive Discussion





**The chemistry  
influencing ODEs**M. Piot and R. von  
Glasow

Title Page

Abstract

Introduction

Conclusions

References

Tables

Figures

◀

▶

◀

▶

Back

Close

Full Screen / Esc

Printer-friendly Version

Interactive Discussion



5 mative features for halogen or ozone concentrations are discussed in this paper. This paper does not intend to reproduce observed conditions, but rather to investigate the potential influence of several species on the halogen/ozone chemistry. We differentiate three different modification types: a change in initial mixing ratio only, in flux only, or in both. The studied values, for mixing ratios, were increased/decreased by an order of magnitude compared to the initial value used in the base runs. In case flux measurements for a species are known, we used an adequate average value for our sensitivity tests (e.g., see Sect. 4.9). When no flux measurement is known at the appropriate location and time of the year, we assumed the fluxes as function of their initial mixing ratio. After evaluation of the potential range of study, lower/upper values for the fluxes in these sensitivity studies are chosen to be 10%/200% of the initial mixing ratio, per day (e.g.,  $\xi_{\text{ethane}}=1.5 \text{ nmol mol}^{-1}$ ; flux=0.15/3.0 nmol mol<sup>-1</sup> day<sup>-1</sup>). We chose 200% for the upper limit in our study as higher values would induce fluxes in majority exceeding any field observation. The value 10% induced small fluxes which do not considerably modify the gas phase mixing ratio.

15 These fluxes remain constant throughout the model simulations. Diurnal variations that may be observed in the field (Sumner and Shepson, 1999; Zhou et al., 2001; Hutterli et al., 2001; Foster et al., 2001; Bottenheim et al., 2002a) are therefore not taken into account.

20 Due to the large number of sensitivity runs we only selected the most relevant runs in this paper. We chose the runs to be discussed with an ODE type (M1, M4, P4), based on the relevance of the applied conditions or their specific chemistry compared to the base run.

## 4.2 Details of base runs

25 In order to comprehensively analyze the sensitivity runs, we first describe the main features of the base runs. In Fig. 2, O<sub>3</sub> undergoing a P4 ODE or a M4 ODE is continuously depleted, with a net slow down in destruction at night due to the absence of reactive halogens. The destruction of O<sub>3</sub> during a M1 is very rapid and occurs within

**The chemistry  
influencing ODEs**M. Piot and R. von  
Glasow

Title Page

Abstract

Introduction

Conclusions

References

Tables

Figures

◀

▶

◀

▶

Back

Close

Full Screen / Esc

Printer-friendly Version

Interactive Discussion



a few hours. In Fig. 3, we show the chemistry of compounds other than  $O_3$  for the two base runs  $Br_2$ -M4 and  $Br_2$ -P4.  $Br_2$ -M1 is omitted for clarity. In addition, the base run  $Br_2$ -M1 will not be investigated in great details in this paper as the required flux is rather unlikely high. Similarly, details of  $BrCl$ - and  $Cl_2$ -induced ODEs will be described in upcoming sections.

In both M4 and P4 cases, total gas phase bromine ( $Br_x$ , Fig. 3a) increases during the three first simulated days due to the constant flux of  $Br_2$ . The production of  $BrO$  via Reaction (1) increases gradually over time.  $BrO$  undergoes a strong diurnal cycle (Fig. 3b) as  $Br$  atoms are quasi absent at night. Sea salt aerosol  $Br^-$  is liberated in both cases via the bromine explosion cycle (Fig. 3c) and less reactive bromine is mostly recycled via reaction in acidic sulfate aerosols (not shown). On the last simulated day bromine compounds evolve differently in cases M4 and P4: in the case P4,  $O_3$  remains above  $11 \text{ nmol mol}^{-1}$ , and the rates for Reaction (1) do not decrease strongly due to the reduction of  $O_3$  mixing ratios. Therefore,  $BrO$  remains at high mixing ratios during the last day ( $\sim 25 \text{ pmol mol}^{-1}$ ). In contrast, in the case M4, the  $O_3$  destruction is nearly complete 12 h before the end of the run. Rates for Reaction (1) drastically decrease and the  $BrO$  production is also reduced (see sharp decrease in Fig. 3b). The decrease of this reaction rate leads to a shift in speciation from  $BrO$  to  $Br$ , and therefore, from  $HOBr$  to  $HBr$  (see Reactions 5 and 13). The reduction of  $HOBr$  and increase of  $HBr$  lead to the re-bromination of the aerosols (see SSA  $Br^-$  in Fig. 3c, on the last day. Sulfate  $Br^-$  not shown).  $HO_x$  mixing ratios on the first day are approximately 1 and  $1.5 \text{ pmol mol}^{-1}$  in cases M4 and P4, respectively, with a rapid decrease to values below  $0.3 \text{ pmol mol}^{-1}$  for the following days. These mixing ratios are similar to measurements obtained by Bloss et al. (2007) in coastal Antarctica.

### 4.3 Formaldehyde influence

We investigated the influence of a constant flux of  $HCHO$  on the development of an ODE. Figure 4 shows the comparison between the base run  $Br_2$ -P4 and the run  $Br_2$ -

**The chemistry  
influencing ODEs**M. Piot and R. von  
Glasow

Title Page

Abstract

Introduction

Conclusions

References

Tables

Figures

◀

▶

◀

▶

Back

Close

Full Screen / Esc

Printer-friendly Version

Interactive Discussion



P4-HCHO where a flux of  $5.0 \times 10^9 \text{ molec cm}^{-2} \text{ s}^{-1}$  is prescribed under background air conditions (for these conditions, see Table 1). This value is similar to the one used in Michalowski et al. (2000), and the resulting HCHO mixing ratios in the gas phase are in good agreement with observations (de Serves, 1994; Sumner and Shepson, 1999; Sumner et al., 2002). As explained in Sect. 3.1 a high concentration of HCHO (Fig. 4b) increases the reaction rates of Reactions (18) and (19) and increases the production of HO<sub>2</sub> via Reactions (16) and (17) (Fig. 4c). Therefore, in this case the halogen speciation is rapidly shifted from X and XO to HOX and HX (also see Reactions 5 and 13). Mixing ratios of HBr, HCl, HOBr, and HOCl rapidly increase (Figs. 4g to i). The uptake of these compounds onto aerosols maintains the rapid aerosol dehalogenation via the bromine explosion process (Reaction 14, see Fig. 4d).

The sea salt aerosol dehalogenation depends on the ratio  $[\text{Br}^-]:[\text{Cl}^-]$  (see Fickert et al., 1999; Adams et al., 2002). Under our conditions, this dehalogenation occurs mostly in the form of a Br<sub>2</sub> liberation in the gas phase, but BrCl and Cl<sub>2</sub> are also produced.

After the near-complete aerosol debromination on the first day (after  $\approx 10$  h), both runs undergo a shift in speciation from bromine radicals to HOBr and HBr (shift stronger in Br<sub>2</sub>-P4-HCHO). This shift reduces the BrO self-reaction (Reactions 2–3) which is an efficient cycle for the release of Br atoms. Furthermore, bromine deposition on snow strongly increases (see accumulated deposition in Fig. 4k), significantly decreasing the total gas phase bromine concentration in Fig. 4j. We calculated that HOBr is responsible for nearly 80% of this increase.

Interestingly, mixing ratios of Br<sub>2</sub> in the gas phase (Fig. 4f) show a decrease compared to the base run. This decrease is a consequence of the large loss of bromine on snow which reduces the bromine loading in the gas phase.

At the end of the model run the amount of highly reactive BrO is reduced by 65% (Fig. 4e). As a result O<sub>3</sub> mixing ratios are  $10 \text{ nmol mol}^{-1}$  higher than in the base run (Fig. 4a). Note that Br<sup>-</sup> in SSA increases during the last simulated day in the base run due to insufficient HOBr in the aqueous phase (see solid black line, Fig. 4d).

**The chemistry  
influencing ODEs**M. Piot and R. von  
Glasow

Title Page

Abstract

Introduction

Conclusions

References

Tables

Figures

◀

▶

◀

▶

Back

Close

Full Screen / Esc

Printer-friendly Version

Interactive Discussion



In summary, a flux of HCHO as prescribed here efficiently reduces the ozone depletion process, as the PODE4 threshold is reached only at the very end of the simulation.

In all studied cases with elevated HCHO mixing ratios, HCHO strongly modified the halogen speciation. However, the conditions of high mixing ratios of HCHO differently impacted the ozone chemistry, depending on the concentration of  $\text{Br}^-$  in sea salt aerosols. With high initial aerosol  $\text{Br}^-$  concentrations, the bromine explosion is accelerated by the presence of more HCHO and  $\text{HO}_x$  (see explanations above). With very low initial  $\text{Br}^-$  concentrations, the chemistry is characterized only by the shift in bromine speciation and a resulting reduction of available highly reactive bromine radicals (not shown).

Under conditions of near-total  $\text{O}_3$  depletion and high concentrations of HCHO in the air (e.g.,  $\text{Br}_2$ -M1 with a constant flux of HCHO), the conversion from bromine radicals to HBr and HOBr rapidly leads to an increase in aerosol  $\text{Br}^-$  concentration and therefore, in its deposition on snow. Mixing ratios of  $\text{Br}_x$  dramatically decrease (not shown). We noted no major difference in the chemistry when we applied this case to coastal air conditions (with the changes listed in Table 1).

Under conditions of  $\text{Cl}_2$ -induced ODEs, the presence of high HCHO fluxes led to no relevant changes in the ozone chemistry (maximum differences of  $1 \text{ nmol mol}^{-1}$ ).

When prescribing a HCHO flux of  $6.0 \times 10^7 \text{ molec cm}^{-2} \text{ s}^{-1}$ , as estimated from Jacobi et al. (2002) (see Piot and von Glasow, 2007), the influence on bromine was negligible. We conclude that this flux of HCHO has insignificant effects on ozone. These sensitivity studies show that higher fluxes of HCHO possibly causing the observed gas phase concentrations are required to significantly impact the ozone chemistry.

#### 4.4 $\text{H}_2\text{O}_2$

In this section we compare the base run  $\text{Br}_2$ -M4 to  $\text{Br}_2$ -M4- $\text{H}_2\text{O}_2=1.5 \times 10^{10} \text{ molec cm}^{-2} \text{ s}^{-1}$  (Fig. 5). To the best of our knowledge, fluxes of  $\text{H}_2\text{O}_2$  at Alert in spring have not yet been measured. Therefore, we used values for

**The chemistry  
influencing ODEs**M. Piot and R. von  
Glasow

Title Page

Abstract

Introduction

Conclusions

References

Tables

Figures

◀

▶

◀

▶

Back

Close

Full Screen / Esc

Printer-friendly Version

Interactive Discussion



this flux as described in Sect. 4.1. Such fluxes out of the snowpack have already been measured by Hutterli et al. (2004), but their location (Summit, Greenland) and the period of measurement (summer) differ significantly from our conditions. This H<sub>2</sub>O<sub>2</sub> flux induces gas phase mixing ratios three to four times higher than observations in the Arctic spring (de Servès, 1994). However, these mixing ratios remain in the range of late spring/summertime measurements (Bales et al., 1995; Hutterli et al., 2001; Jacobi et al., 2002).

The photolysis of H<sub>2</sub>O<sub>2</sub> represents its primary loss pathway (Reaction 20), producing highly reactive OH radicals (Fig. 5d). Reactions (7) and (21) mainly increase the concentration of HO<sub>2</sub> (Fig. 5c). Therefore, as explained in Sect. 4.3, higher reaction rates for Reaction (5) and (13) lead to lower mixing ratios of Br and BrO (Fig. 5f). The higher mixing ratios of reaction products HOBr (Fig. 5e) and HBr maintain an efficient recycling in SSAs (Fig. 5g). Again, the shifted speciation from Br/BrO to HBr/HOBr reduces the efficiency of the BrO self-reaction, and the increase in bromine deposition reduces the total amount of gas phase bromine. Mixing ratios of BrO decrease by up to 10 pmol mol<sup>-1</sup> between day 2 and 3 compared to the base run Br<sub>2</sub>-M4. (Fig. 5f). Consequently, mixing ratios of O<sub>3</sub> are about 9 nmol mol<sup>-1</sup> higher than in the base run after three days. On the last simulated day, O<sub>3</sub> mixing ratios reach the M4 threshold with a 12-hour delay compared to the base run. As the ozone depletion is not complete on this last simulated day, BrO mixing ratios remain high, while they rapidly decrease in the base run Br<sub>2</sub>-M4 (see Fig. 5f). Similarly, the bromine recycling through the aerosol phase remains efficient, keeping Br<sup>-</sup> concentrations low (Fig. 5g), while the base run undergoes a re-bromination (see Sect. 4.2).

This flux of H<sub>2</sub>O<sub>2</sub> substantially affected the ozone destruction. The aerosol debromination remained more efficient than in the base model run. However, the deposition of bromine on snow (more than 2.0 × 10<sup>-7</sup> mol m<sup>-2</sup> compared to 0.4 × 10<sup>-7</sup> for the base run) was higher than the bromine production in the gas phase and it strongly reduced the amount of reactive bromine.

When prescribing a flux of 1.0 × 10<sup>8</sup> molec cm<sup>-2</sup> s<sup>-1</sup>, estimated from Jacobi et al.

(2002) (see [Piot and von Glasow, 2007](#)) for early spring, the ozone chemistry showed no major change. Such a value for the  $\text{H}_2\text{O}_2$  flux does not substantially affect the concentration of  $\text{HO}_x$  in the gas phase. It seems, from these results, that the fluxes of  $\text{H}_2\text{O}_2$  estimated for the high Arctic in spring are not significantly influencing the ozone chemistry. These sensitivity studies showed that only higher fluxes (most probably related to higher temperatures, e.g., in late spring/summer, see [Hutterli et al., 2001](#)) may effectively influence the ozone/halogen chemistry.

#### 4.5 DMS and DMSO “counter-cycle”

To the best of our knowledge, fluxes of DMS have not been measured in the Arctic spring. The value for the flux used in this section is explained in Sect. 4.1. Figure 6 shows a comparison between the base run  $\text{Br}_2\text{-M4}$  and the model run  $\text{Br}_2\text{-M4-DMS}=100\text{ pmol mol}^{-1}$  with a flux of  $4.0 \times 10^9 \text{ molec cm}^{-2} \text{ s}^{-1}$ . We investigated DMS only under coastal conditions as it is produced in the ocean.

The primary effect of high concentrations of DMS on the ozone/halogen chemistry is through Reaction (25):  $\text{BrO}$  oxidizes DMS and produces  $\text{Br}$  radicals which represents an efficient additional recycling pathway for  $\text{BrO}$ . Indeed, this reaction leads to slightly more ozone depletion during the first simulated day (about 0.7%). The reaction product DMSO (Fig. 6c) then reacts with  $\text{OH}$  (Reaction 26) to produce  $\text{CH}_3\text{O}_2$  (Fig. 6d). This represents a key reaction initiating a cycle that we call the DMSO “counter-cycle” (see Fig. 7). An increase in  $\text{CH}_3\text{O}_2$  induces more  $\text{HCHO}$  (Fig. 6e) and  $\text{HO}_2$  which eventually affect the bromine distribution. This leads, again, to a shift in bromine speciation as explained in Sects. 4.3 and 4.4. Note that the increase of  $\text{CH}_3\text{O}_2$  and  $\text{HCHO}$  starts on the first simulated day, while the bromine speciation shift becomes important only on the following day. The relative variation of  $\text{BrO}$  compared to the base run is anti-correlated to the relative variations of  $\text{HOBr}$  and  $\text{HBr}$ . The presence of more  $\text{HOBr}$  and  $\text{HBr}$  in the gas phase (Figs. 6h and i) accelerates the aerosol debromination (see Fig. 6k). However, the deposition of (mostly)  $\text{HOBr}$  represents a great sink for bromine (Fig. 6l).  $\text{Br}_x$  is only slightly affected by the competition between aerosol debromination

### The chemistry influencing ODEs

M. Piot and R. von Glasow

[Title Page](#)[Abstract](#)[Introduction](#)[Conclusions](#)[References](#)[Tables](#)[Figures](#)[◀](#)[▶](#)[◀](#)[▶](#)[Back](#)[Close](#)[Full Screen / Esc](#)[Printer-friendly Version](#)[Interactive Discussion](#)

acceleration enhancing  $\xi_{\text{Br}_x}$  and the increase of bromine deposition decreasing  $\xi_{\text{Br}_x}$  (Fig. 6j). In this run, ozone reaches the M4 threshold with a delay of only two hours.

In summary, contrary to what one might have expected the reaction between DMS and BrO is not dominating the overall effect of DMS on bromine/ozone under the conditions of this run. The production of  $\text{CH}_3\text{O}_2$  via reaction between DMSO and OH becomes rapidly important and leads to less Br radicals. Under conditions of near-total ozone destruction (e.g., run Br<sub>2</sub>-M1 after the first day), the shift in bromine speciation from BrO and Br to mostly HBr leads to a efficient bromination and re-bromination of sulfate and sea salt aerosols, respectively (not shown).

DMS does not react substantially with ClO and the rate coefficient of Reaction (24) is very small. Therefore, DMS does not have an effect on Cl<sub>2</sub>-induced ODEs.

#### 4.6 Cl<sub>2</sub> and “chlorine counter-cycle”

All previous ODEs we investigated were induced by a Br<sub>2</sub> flux. In this section we present results on Cl<sub>2</sub>-induced ODEs (see Table 4) to study the influence of a Cl<sub>2</sub> flux on ozone chemistry. Not surprisingly, the prescribed Cl<sub>2</sub> fluxes required to reduce O<sub>3</sub> within the observed time scales were unrealistically high. Nevertheless, we describe these sensitivity runs as they appear interesting for a better understanding of the chemical cycles.

We present in Fig. 8 the base run Cl<sub>2</sub>-M1. Upon photolysis of Cl<sub>2</sub> on the first simulated day, the two main reactions using chlorine (Reactions 1 and 27) lead to a strong production of ClO and CH<sub>3</sub>O<sub>2</sub> (see Fig. 8). Those two reaction products react together to yield HCHO (Reaction 30). This reaction, increasing HCHO mixing ratios (Fig. 8), also accelerates the reaction rate of Reaction (31). This reaction chain efficiently converts Cl/ClO (see HCl compared to the sharp decreases in Cl/ClO a few hours after each sunrise in Fig. 8). Additionally, ROOH is produced substantially via Reaction (28) due to the high concentrations of CH<sub>3</sub>O<sub>2</sub> and HO<sub>2</sub>. Thus, the reaction rate of Reaction (29) also increases and contributes to the rise of HCl mixing ratios. All these reaction pathways are schematically described in Fig. 9. This figure clearly shows that

### The chemistry influencing ODEs

M. Piot and R. von Glasow

Title Page

Abstract

Introduction

Conclusions

References

Tables

Figures

◀

▶

◀

▶

Back

Close

Full Screen / Esc

Printer-friendly Version

Interactive Discussion



the reaction chain initiated by Reaction (27) efficiently induces a direct (Reactions 27, 29, and 31) or indirect (Reactions 28 and 30) conversion from reactive chlorine to less reactive HCl. HCl drastically increases to reach unrealistically high mixing ratios of several tens of  $\text{nmol mol}^{-1}$ .

5 In summary, Cl and ClO radicals rapidly react to produce compounds ( $\text{HO}_x$ , ROOH or  $\text{CH}_3\text{O}_2$ ) which eventually react back with Cl/ClO to release HCl. We call this negative feedback “chlorine counter-cycle”.

As highlighted in this section, a higher flux of chlorine radicals induces an even stronger “chlorine counter-cycle”, shifting chlorine to HCl. In order to induce a major ODE within one day (M1), it is necessary to prescribe an unrealistically high  $\text{Cl}_2$  flux of  $3.0 \times 10^{11} \text{ molec cm}^{-2} \text{ s}^{-1}$  to keep the reaction between Cl and  $\text{O}_3$  (Reaction 1) efficient. Mixing ratios of ClO (up to  $1 \text{ nmol mol}^{-1}$ ) and Cl (order of  $\text{pmol mol}^{-1}$ ) are inconsistent with measurements made in the Arctic (Tuckermann et al., 1997; Impey et al., 1997; Perner et al., 1999; Boudries and Bottenheim, 2000). We conclude that the presence of this efficient “chlorine counter-cycle”, as explained in this section, makes a  $\text{Cl}_2$ -induced ODE unrealistic.

This counter-cycle cannot initiate in case of  $\text{Br}_2$ -induced ODEs, as there is no reaction between bromine atom and  $\text{CH}_4$ , which is the key reaction to yield  $\text{CH}_3\text{O}_2$  in case of  $\text{Cl}_2$ -induced ODEs.

20 It is noteworthy mentioning the effect of this counter-cycle on  $\text{Br}_2$ -induced ODEs including high concentrations of chlorine. In our sensitivity studies this counter-cycle appeared to have a substantial influence on the eventual ozone mixing ratios in several runs. As an example, Fig. 10 shows the comparison between the base run  $\text{Br}_2$ -M4 and  $\text{Br}_2$ -M4- $\text{Cl}_2=5.0 \times 10^8 \text{ molec cm}^{-2} \text{ s}^{-1}$  under coastal air conditions. In this case Cl concentrations of up to  $2.5 \times 10^{-3} \text{ pmol mol}^{-1}$  released via  $\text{Cl}_2$  photolysis is sufficiently high to “activate” the counter-cycle and therefore, slow down the ozone destruction via reduction of BrO and Br. After the sunrise on the first day, mixing ratios of ClO rapidly rise to a maximum of  $7 \text{ pmol mol}^{-1}$  (Fig. 10b) mostly via Reaction (1). This contributes to the slight decrease in  $\text{O}_3$  on the first day compared to the base run

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## The chemistry influencing ODEs

M. Piot and R. von Glasow

---

[Title Page](#)[Abstract](#)[Introduction](#)[Conclusions](#)[References](#)[Tables](#)[Figures](#)[◀](#)[▶](#)[◀](#)[▶](#)[Back](#)[Close](#)[Full Screen / Esc](#)[Printer-friendly Version](#)[Interactive Discussion](#)



**The chemistry  
influencing ODEs**M. Piot and R. von  
Glasow

(Fig. 10a). As highlighted in this section, the high mixing ratios of Cl and ClO lead to the activation of the chlorine counter-cycle:  $\text{CH}_3\text{O}_2$ ,  $\text{HO}_x$ , and  $\text{ROOH}$  substantially increase (see Figs. 10d and e). Several hours after each sunrise of the model run, chlorine radicals undergo the sharp decrease attributed to the counter-cycle (see Fig. 10b with mean values similar to observations from Perner et al., 1999; Tuckermann et al., 1997) and are converted to unreactive HCl (Fig. 10c). As a consequence, higher mixing ratios of  $\text{CH}_3\text{O}_2$  increase reaction rates of Reactions (11–12), while reaction rates of Reactions (5, 13, 18 and 19) are higher due to increased  $\text{HO}_2$  and HCHO (see Fig. 10e). Again, bromine undergoes a shift in speciation from Br/BrO to HBr/HOBr (Figs. 10f to h).

This case highlights that high concentrations of bromine together with chlorine may lead to unexpected halogen interactions. The eventual effect on ozone depends on the activation state of the chlorine counter-cycle.

The chemistry of BrCl-induced ODEs is similar to that of a  $\text{Br}_2$ -induced ODE including a flux of chlorine. In the P4 ozone destruction, the release of chlorine via the flux of BrCl ( $1.0 \times 10^8 \text{ molec cm}^{-2} \text{ s}^{-1}$ , see Table 4) is too small to “activate” the chlorine counter-cycle. The bromine chemistry in the BrCl-P4 case remains nearly identical to the chemistry described for the base run  $\text{Br}_2$ -P4. Therefore, the required flux for BrCl to induce an equivalent P4 as for  $\text{Br}_2$ -P4 is stoichiometrically similar to the  $\text{Br}_2$  flux. Such a flux of BrCl induces realistic loadings of chlorine in the model:  $\xi_{\text{ClO}} \sim 2 \text{ pmol mol}^{-1}$  (see Perner et al., 1999),  $\xi_{\text{Cl}} \sim 4 \times 10^{-4} \text{ pmol mol}^{-1}$  (see Jobson et al., 1994; Boudries and Bottenheim, 2000), and  $\xi_{\text{Cl}_2} \sim 2 \text{ pmol mol}^{-1}$  (see Foster et al., 2001). The BrCl-M4 base run does show a small influence of the counter-cycle on the bromine chemistry. This can be highlighted by increased concentrations of HBr compared to the concentrations in  $\text{Br}_2$ -M4 (increase of about 40%). Therefore, the BrCl flux required to induce an equivalent M4 ODE as in the  $\text{Br}_2$  case is slightly higher than the stoichiometric ratio with  $\text{Br}_2$  ( $1.9 \times 10^8$  compared to  $2 \times 9.0 \times 10^7 \text{ molec cm}^{-2} \text{ s}^{-1}$ ). In addition, loadings of chlorine in the model do not differ substantially from the BrCl-P4 run. Mixing ratios of BrCl in BrCl-M4 reach  $7 \text{ pmol mol}^{-1}$ , while they have a maximum of  $5 \text{ pmol mol}^{-1}$

Title Page

Abstract

Introduction

Conclusions

References

Tables

Figures

◀

▶

◀

▶

Back

Close

Full Screen / Esc

Printer-friendly Version

Interactive Discussion



**The chemistry  
influencing ODEs**M. Piot and R. von  
Glasow

Title Page

Abstract

Introduction

Conclusions

References

Tables

Figures

◀

▶

◀

▶

Back

Close

Full Screen / Esc

Printer-friendly Version

Interactive Discussion



in BrCl-P4. In the BrCl-M1 base run, the chlorine counter-cycle is clearly activated and it strongly affects the bromine chemistry by shifting reactive bromine to mainly HBr (10 to 20 times higher in BrCl- than Br<sub>2</sub>-induced ODEs, not consistent with predictions from other models, see [Fan and Jacob, 1992](#); [Sander et al., 1997](#); [Lehrer et al., 2004](#)). BrCl mixing ratios reach a maximum of 38 pmol mol<sup>-1</sup> which is a rather high value compared to measurements made by [Foster et al. \(2001\)](#). The required flux of BrCl, compared to Br<sub>2</sub>, is substantially higher than the stoichiometric ratio (3.4×10<sup>9</sup> compared to 2×1.5×10<sup>9</sup> molec cm<sup>-2</sup> s<sup>-1</sup>), showing that the chlorine counter-cycle is activated. These studies show that only the chemistry of the BrCl-P4 and BrCl-M4 runs lies within the range of observations.

#### 4.7 C<sub>2</sub>H<sub>4</sub>

Ethene (C<sub>2</sub>H<sub>4</sub>) chemistry is characterized by the three reactions listed in Sect. 3.4. In this section we compare the base run Br<sub>2</sub>-M4 with Br<sub>2</sub>-M4-C<sub>2</sub>H<sub>4</sub>=0.8 nmol mol<sup>-1</sup> under coastal conditions (Fig. 11). Such a high mixing ratio for C<sub>2</sub>H<sub>4</sub> corresponds to an upper limit, but this value lies within the variability of observations during Arctic spring ([Doskey and Gaffney, 1992](#); [Ariya et al., 1999](#)).

Mixing ratios of OH (Fig. 11c) and Cl (released from SSA; Fig. 11d) decrease, but the Cl chemistry does not affect the ozone/bromine concentration under these conditions ( $\xi_{\text{Cl}}$  too low, but similar to measurements from [Jobson et al., 1994](#); [Boudries and Bottenheim, 2000](#)). The increase in  $\xi_{\text{C}_2\text{H}_4}$  reduces HO<sub>x</sub> lifetime via Reaction (33).

Therefore, this reaction tends to reduce the formation of HOBr from Br/BrO radicals. However, in the case of Br<sub>2</sub>-M4-C<sub>2</sub>H<sub>4</sub>=0.8 nmol mol<sup>-1</sup> Reaction (34) is accelerated (see Fig. 11e). This reaction prevails over Reactions (32) and (33) during the whole model run and increases the formation of HBr. Thus, BrO mixing ratios decrease strongly (Fig. 11i). Clearly, the decrease of HOBr (Fig. 11g) reduces the efficiency of the bromine explosion mechanism and the shift of bromine speciation to HBr and its subsequent uptake to particles leads to a drastic increase in sulfate and sea

**The chemistry  
influencing ODEs**M. Piot and R. von  
Glasow

Title Page

Abstract

Introduction

Conclusions

References

Tables

Figures

◀

▶

◀

▶

Back

Close

Full Screen / Esc

Printer-friendly Version

Interactive Discussion



salt aerosol  $\text{Br}^-$  (Fig. 11h, not shown for sulfate aerosols).  $\text{Br}^-$  concentrations in SSA are up to  $20 \text{ nmol mol}^{-1}$  higher than in the base run  $\text{Br}_2\text{-M4}$ . This scenario shows that the presence of large amounts of  $\text{C}_2\text{H}_4$  notably reduces the efficiency of the bromine explosion. This and the conversion from Br atoms to less reactive HBr lead to less available bromine radicals for the depletion of ozone. The ozone mixing ratio drops only to  $11 \text{ nmol mol}^{-1}$  within four days, which is equivalent to a “P4” event.

However, it is important to note that the  $\text{C}_2\text{H}_4$  destruction reactions may compete with each other, depending on the amount of  $\text{C}_2\text{H}_4$  in the model run. In order to assess the influence of Reaction (33) compared to Reaction (34), we performed additional model runs with  $\xi_{\text{C}_2\text{H}_4} = 0.5$  and  $0.01 \text{ nmol mol}^{-1}$ . In Fig. 12, we compare ozone mixing ratios of the different runs with ozone in (A)  $\text{Br}_2\text{-M4}$  and (B)  $\text{Br}_2\text{-P4}$ . In Fig. 12a, these ratios for  $\text{C}_2\text{H}_4 = 0.8$  and  $\text{C}_2\text{H}_4 = 0.5 \text{ nmol mol}^{-1}$  show that ozone for these runs is less depleted than for the base run, due to the reduction in the bromine explosion cycle, as explained above. Interestingly, however, ozone is also temporarily less depleted than in the base run under conditions of  $\xi_{\text{C}_2\text{H}_4} = 0.01 \text{ nmol mol}^{-1}$  (less  $\text{C}_2\text{H}_4$  than in the base run). With low  $\xi_{\text{C}_2\text{H}_4}$  the reduced rate of Reaction (33) allows more  $\text{HO}_x$  to react with bromine radicals to form HOBr. Similarly, a reduction in the rate of Reaction (34) leads to less HBr. The increase and slight decrease of HOBr and HBr, respectively, induce a moderate acceleration of the SSA debromination compared to the base run. Nevertheless, as highlighted in previous sections, the formation of HOBr reduces  $\text{Br}_x$  via increased deposition at the surface.

As a consequence, less bromine radicals are available compared to the base run. The difference to the base run M4 is highest in the morning of day 3. Later, ozone rapidly decreases and becomes more depleted than in the base run M4. This change in ozone destruction is the result of a large decrease in reaction rate of Reaction (34). At the end of the model run  $\text{Br}_2\text{-M4-C}_2\text{H}_4 = 0.01 \text{ nmol mol}^{-1}$ ,  $\text{C}_2\text{H}_4$  is completely destroyed. Therefore, the concentration of Br atoms increases compared to the base run, which leads to a more efficient ODE on the last day. This specific time span highlights the importance of Reaction (34) in reducing the amount of available Br atoms, although

the rate coefficient for this reaction is small.

Under conditions of a P4 ODE, an increase of  $\xi_{\text{C}_2\text{H}_4}$  compared to  $0.08 \text{ nmol mol}^{-1}$  (base run Br<sub>2</sub>-P4) leads, again, to a decrease in ozone depletion via the reduction of the bromine explosion and the conversion from Br atoms to HBr (see Fig. 12b). The model run with  $\xi_{\text{C}_2\text{H}_4}=0.01 \text{ nmol mol}^{-1}$  displays a similar chemistry as explained for the case Br<sub>2</sub>-M4. However, C<sub>2</sub>H<sub>4</sub> does not drop down to zero at the end of the run as in Br<sub>2</sub>-M4-C<sub>2</sub>H<sub>4</sub>= $0.01 \text{ nmol mol}^{-1}$ . Therefore, Reaction (34) remains efficient during the whole model run. Ozone mixing ratios at the end of the model run (Fig. 12b) remain slightly higher than in the base run P4.

This study highlighted that high concentrations of ethene strongly reduced the bromine explosion and therefore, the ozone depletion. Low concentrations of ethene, however, allow substantial amounts of HO<sub>x</sub> to be present in the atmosphere, which shifts reactive bromine to less reactive HOBr. Such concentrations may then lead to a decrease in ozone destruction as well.

Prescribing a flux of  $1.3 \times 10^8 \text{ molec cm}^{-2} \text{ s}^{-1}$  of C<sub>2</sub>H<sub>4</sub>, as measured by Swanson et al. (2002), showed only a very weak influence on ozone:  $\xi_{\text{O}_3}$  only increased by less than  $1 \text{ nmol mol}^{-1}$ . Fluxes measured in the field seem to have no noticeable influence on the bromine/ozone chemistry, but notice that the resulting C<sub>2</sub>H<sub>4</sub> mixing ratios in the model only increases by  $10 \text{ pmol mol}^{-1}$  due to this prescribed flux. Only higher fluxes of C<sub>2</sub>H<sub>4</sub> may affect the halogen/ozone chemistry.

## 4.8 C<sub>2</sub>H<sub>6</sub>

First, we will discuss sensitivity studies on Br<sub>2</sub>-induced ODEs. The observed variability of C<sub>2</sub>H<sub>6</sub> in the Arctic is relatively small (between 1 and  $4 \text{ nmol mol}^{-1}$ , see Ariya et al., 1999). When investigating realistic values for C<sub>2</sub>H<sub>6</sub> mixing ratios, we noted no relevant change in ozone. For a better understanding of the C<sub>2</sub>H<sub>6</sub> chemistry, we also investigated the effects of higher mixing ratios (an order of magnitude higher than observed values, with  $\xi_{\text{C}_2\text{H}_6}=15 \text{ nmol mol}^{-1}$ , see Table 3).

### The chemistry influencing ODEs

M. Piot and R. von Glasow

Title Page

Abstract

Introduction

Conclusions

References

Tables

Figures

◀

▶

◀

▶

Back

Close

Full Screen / Esc

Printer-friendly Version

Interactive Discussion



**The chemistry  
influencing ODEs**M. Piot and R. von  
Glasow

Title Page

Abstract

Introduction

Conclusions

References

Tables

Figures

◀

▶

◀

▶

Back

Close

Full Screen / Esc

Printer-friendly Version

Interactive Discussion

We compare now the base run Br<sub>2</sub>-M4 with Br<sub>2</sub>-M4-C<sub>2</sub>H<sub>6</sub>=15 nmol mol<sup>-1</sup> under background conditions (Fig. 13). In our model C<sub>2</sub>H<sub>6</sub> only reacts with OH and Cl with a rate coefficient for Reaction (35) approximately 500 times higher than for Reaction (36). Under Br<sub>2</sub>-M4 conditions, however, the concentration of Cl is very low: Reaction (35) is negligible compared to Reaction (36). The removal of OH radicals from the atmosphere by C<sub>2</sub>H<sub>6</sub> (Reaction 36) is an important pathway reducing the concentration of HO<sub>x</sub> (Fig. 13c and d) as well as HCHO (Fig. 13e). As explained in previous sections this decrease in oxidant concentrations reduces the production of HOBr and HBr (see Fig. 13g) and slows the bromine explosion cycle down. In this investigation, the limitation in HO<sub>x</sub> prevails and increases the mixing ratio of BrO by at most 2 pmol mol<sup>-1</sup> (Fig. 13f). The more efficient BrO self-reaction leads to a faster recycling of bromine oxide and a stronger ozone destruction (with a maximum decrease of 2 nmol mol<sup>-1</sup> compared to the base model run, Fig. 13a). ODEs caused by Br<sub>2</sub> emissions are enhanced under high concentrations of ethane. Ethane cleanses the air from high reactive oxidants which hinder the BrO self-reaction. However, relatively high mixing ratios of ethane do not substantially affect the ozone chemistry.

We also investigated Cl<sub>2</sub>-induced ODEs with increased mixing ratios of C<sub>2</sub>H<sub>6</sub>. In all ODE cases (M1, M4 and P4), Reaction (35) prevailed and induced weaker depletions of ozone. C<sub>2</sub>H<sub>6</sub> effectively diminishes the availability of Cl atoms. Therefore, C<sub>2</sub>H<sub>6</sub> directly reduces the chlorine-induced ozone destruction chain. By prescribing 15 nmol mol<sup>-1</sup> of C<sub>2</sub>H<sub>6</sub> compared to 1.5 nmol mol<sup>-1</sup> in the base runs, ozone only reached the M2 threshold compared to Cl<sub>2</sub>-M1, P4 compared to Cl<sub>2</sub>-M4, and no ODE compared to Cl<sub>2</sub>-P4, respectively.

This study shows that ethane is mainly important for chlorine-related chemistry via direct reaction between C<sub>2</sub>H<sub>6</sub> and Cl. Ethane does not influence the bromine chemistry markedly: only reaction with OH radicals slightly reduces HO<sub>x</sub> concentrations.

## 4.9 HONO, NO<sub>2</sub>, and RONO<sub>2</sub>

Field measurements in the Arctic spring have highlighted the production of HONO molecules originating from the snowpack. Zhou et al. (2001) measured hourly-averaged HONO flux in the range of 0–1.0×10<sup>9</sup> molec cm<sup>-2</sup> s<sup>-1</sup> with a mean value of 5.0×10<sup>8</sup> molec cm<sup>-2</sup> s<sup>-1</sup>. In this section we investigate the importance of this mean value for the flux of HONO. Figure 14 shows the comparison between Br<sub>2</sub>-P4 and Br<sub>2</sub>-P4-HONO=5.0×10<sup>8</sup> molec cm<sup>-2</sup> s<sup>-1</sup> under background air conditions. At daytime, photolysis is the dominant loss for HONO (Reaction 48). OH (Fig. 14c) and NO (see Fig. 14e) radicals are rapidly produced.

As described earlier in this paper higher mixing ratios of OH induce a stronger production of HO<sub>2</sub> (mostly via Reaction 7, Fig. 14d). The presence of higher mixing ratios of HO<sub>2</sub> compared to the base run leads to a shift in bromine speciation from Br/BrO to HBr/HOBr (Figs. 14h and i). Again, as described earlier, such HOBr mixing ratios accelerate the aerosol debromination via the bromine explosion (Fig. 14j), but its deposition on snow also substantially reduces the amount of available bromine. Furthermore, the production of HOBr in the case of Br<sub>2</sub>-P4-HONO=5.0×10<sup>8</sup> molec cm<sup>-2</sup> s<sup>-1</sup> induces a slow down in the rapid BrO self-reaction. These reactions contribute to the reduction of available highly reactive bromine (Fig. 14g) for the ozone depletion.

In addition, NO<sub>x</sub> production from the photolysis of HONO accelerates the rates of Reactions (40) to (47). Reaction (42) accelerates the BrO recycling into Br atoms. However, Reactions (40) and (41) rapidly produce BrNO<sub>2</sub> and BrONO<sub>2</sub> during daytime, with BrONO<sub>2</sub> reaching maxima of about 2 pmol mol<sup>-1</sup> (not shown). BrONO<sub>2</sub>, more reactive than BrNO<sub>2</sub>, may heterogeneously react with aerosol surfaces to convert halides into photolysable halogens (Reaction 47). However, the aerosol debromination induced by this reaction pathway is about two orders of magnitude smaller than via Reaction (14). Indeed, BrNO<sub>2</sub> and BrONO<sub>2</sub> rather represent a temporary gas phase reservoir of bromine. In this model run, deposition of BrONO<sub>2</sub> was found important for the loss of bromine on the snow. We calculated that the total bromine deposition

### The chemistry influencing ODEs

M. Piot and R. von Glasow

Title Page

Abstract

Introduction

Conclusions

References

Tables

Figures

◀

▶

◀

▶

Back

Close

Full Screen / Esc

Printer-friendly Version

Interactive Discussion



**The chemistry  
influencing ODEs**M. Piot and R. von  
Glasow

Title Page

Abstract

Introduction

Conclusions

References

Tables

Figures

◀

▶

◀

▶

Back

Close

Full Screen / Esc

Printer-friendly Version

Interactive Discussion



contributed, at the end of the model run, more than 30% to the total amount of bromine in the run compared to 10% for the base run (not shown). Bromine deposition on snow is higher than the bromine release from the bromine explosion.  $\text{Br}_x$  mixing ratios decrease (Fig. 14f) and reduce the availability of reactive bromine (Fig. 14g). Consequently,  $\text{O}_3$  is destroyed less and displays a final mixing ratio of  $16.5 \text{ nmol mol}^{-1}$  after the four simulated days compared to 12.0 for  $\text{Br}_2$ -P4 (Fig. 14a).

Under conditions of the prescribed flux of HONO from the snow, maximum  $\text{NO}_x$  mixing ratios remain near  $2 \text{ pmol mol}^{-1}$ , as opposed to undetectable values in the base run  $\text{Br}_2$ -P4 (see Fig. 14e after the first simulated day). The flux of HONO induced a moderate acid displacement in SSAs (see Robbins et al., 1959). The presence of  $\text{NO}_x$  leads to the production of  $\text{HNO}_3$  throughout the model run (via mainly Reaction 38, Fig. 14k). The uptake of this strong acid in SSAs induces an increase in acidity ( $\text{H}^+$ ) and  $\text{NO}_3^-$  ions. In our model run with a HONO flux, the SSA  $\text{H}^+$  concentration is 5 to 10 times higher than in the base run, which modifies the equilibrium of reaction  $\text{H}^+ + \text{Cl}^- \leftrightarrow \text{HCl}_{\text{aq}}$  to the right. Therefore, the uptake of  $\text{HNO}_3$  leads to the outgassing of HCl.

Gas phase HCl chemistry is, however, driven at daytime by the increased release of  $\text{Cl}_2$  and  $\text{BrCl}$  in the gas phase compared to  $\text{Br}_2$  (Fig. 14l), as the equilibrium for Reaction (14) is also modified.

At night the absence of OH radicals stops the production of  $\text{HNO}_3$  via Reaction (38). Therefore, the acid displacement in SSA is interrupted as well and  $\text{HCl}_{\text{gas}}$  is taken up into aerosols. HCl reaches a maximum of  $16 \text{ pmol mol}^{-1}$  at the end of the model run (Fig. 14l) and SSA chloride displays a net decrease over the model run (not shown).

These model results indicate that concentrations of  $\text{NO}_x$ , under the influence of such a flux of HONO in Arctic regions, may be associated with high concentrations of gas phase chlorine, if chloride is present in aerosols (also see Hara et al., 2002). We encourage field experiments to find evidences for this specific chemistry in order to confirm these results.

Under coastal air conditions, the presence of high mixing ratios of HOBr and  $\text{BrONO}_2$  with a higher number of SSAs containing  $\text{Br}^-$  (due to the SSA production) induces more

aerosol debromination. These conditions imply more  $\text{Br}_x$  in the air during the model run than under background conditions. After four days,  $\text{O}_3$  is about  $2.5 \text{ nmol mol}^{-1}$  more depleted under coastal air than under background air conditions.

We also performed runs with lower/higher initial mixing ratios of  $\text{NO}_2$  (not shown). The related chemistry was similar to that of HONO. Prescribing higher initial mixing ratios of  $\text{NO}_2$  mainly induced a shift in speciation from  $\text{Br}/\text{BrO}$  to  $\text{BrNO}_2/\text{BrONO}_2$ . This represents a reservoir of less reactive bromine, which reduces the amount of available reactive bromine. Additionally, accumulated bromine deposition increases, mostly due to higher deposition rate of  $\text{HOBr}$  and  $\text{BrONO}_2$ . Ozone is less depleted in all studied cases that include more  $\text{NO}_2$  in the air. However, the difference in  $\text{O}_3$  mixing ratios compared to the base runs is rather small (nearly between 1 and  $4 \text{ nmol mol}^{-1}$  for maximum  $\xi_{\text{NO}_2} = 0.2 \text{ nmol mol}^{-1}$ ). Acid displacement occurred in all cases as well.

We also studied the influence of higher mixing ratios of  $\text{RONO}_2$  (not shown).  $\text{RONO}_2$  chemistry is characterized by Reactions (49) to (51). In both  $\text{Br}_2\text{-M4}$  and  $\text{Br}_2\text{-P4}$  cases the rates of the three reactions remained relatively small, but Reaction (51) prevailed. Only few  $\text{pmol mol}^{-1}$  of  $\text{NO}_2$  molecules were produced in the model runs. The related chemistry remained similar to that of  $\text{NO}_2$ . However, large variations of  $\xi_{\text{RONO}_2}$  did not substantially influence  $\text{O}_3$ : when multiplying  $\xi_{\text{RONO}_2}$  by a factor 10 compared to the base value (Table 3), ozone only increased by  $3 \text{ nmol mol}^{-1}$ , in both  $\text{Br}_2\text{-M4}$  and  $\text{Br}_2\text{-P4}$  cases. Acid displacement of  $\text{HCl}$  by  $\text{HNO}_3$  also occurred under such chemistry.

Among HONO,  $\text{NO}_2$  and  $\text{RONO}_2$ , HONO was found the most important species affecting the bromine/ozone chemistry due to the production of both highly reactive OH and NO.  $\text{NO}_2$  and  $\text{RONO}_2$  only showed a limited influence on bromine/ozone. We therefore encourage experimentalists to record HONO concentrations in particular, simultaneously with bromine/ozone.  $\text{RONO}_2$  only weakly affected ozone, but it is a relevant tracer providing information on the concentration of chlorine radicals (see Reaction 51).

**The chemistry  
influencing ODEs**M. Piot and R. von  
Glasow

Title Page

Abstract

Introduction

Conclusions

References

Tables

Figures

◀

▶

◀

▶

Back

Close

Full Screen / Esc

Printer-friendly Version

Interactive Discussion





## 5 Discussion of deposited bromine on snow

In this paper, we did not consider the recycling of bromine from the snow, as we focused on boundary layer chemical reactions influencing bromine/ozone. However, recycling of deposited bromine on snow appears to be an important process for the re-emission of reactive bromine (see Foster et al., 2001; Peterson and Honrath, 2001; Spicer et al., 2002). Moreover, Piot and von Glasow (2007) showed that the deposition/re-emission process is essential for the timing of an ODE.

The model results highlighted here apply only for constant fluxes of halogens. The different processes leading to the re-emission of deposited bromine on snow are not explicitly taken into account. Therefore, temporal/spatial variations of the flux of halogens in our model run are not investigated in this paper. The simulations where bromine deposition on snow represents an important loss pathway would have opposite effects compared to simulations including recycling on snow. In the sensitivity studies presented here, the deposition of bromine on snow only represents a loss of bromine for the studied airmass. For results including recycling from the snow, the reader is referred to Piot and von Glasow (2007). It is important to stress that the chemistry in the PBL drastically changes, whether deposited bromine is recycled as reactive bromine or not.

## 6 Conclusions

The chemistry of HCHO, H<sub>2</sub>O<sub>2</sub>, DMS, Cl<sub>2</sub>, C<sub>2</sub>H<sub>4</sub>, C<sub>2</sub>H<sub>6</sub>, HONO, NO<sub>2</sub>, and RONO<sub>2</sub> was investigated. Their impact on halogen/ozone in the PBL was assessed by the use of the box model MISTRA in the Lagrangian mode. We compared base runs undergoing Br<sub>2</sub>-, Cl<sub>2</sub>-, or BrCl-induced ODEs with similar runs including a modification in flux or mixing ratio of a species. Under conditions of elevated mixing ratios of HCHO, H<sub>2</sub>O<sub>2</sub>, DMS, and Cl<sub>2</sub> halogen radicals are shifted to HOX/HX (X=Br, Cl) due to the production of HO<sub>x</sub> radicals. This shift in speciation increases the bromine explosion cycle in SSAs.

### The chemistry influencing ODEs

M. Piot and R. von Glasow

Title Page

Abstract

Introduction

Conclusions

References

Tables

Figures

⏪

⏩

◀

▶

Back

Close

Full Screen / Esc

Printer-friendly Version

Interactive Discussion



However, the main effect of this shift is rather the reduction in BrO self-reaction associated with the increased deposition on snow (mainly via HOBr deposition). In these cases, the total amount of gas phase bromine is reduced and the ozone depletion process slows down.

5 It was shown in this paper that ODEs caused by fluxes of Cl<sub>2</sub> lead to unrealistic amounts of reactive chlorine compared to field measurements. Reactions involving reactive chlorine also release CH<sub>3</sub>O<sub>2</sub>, HO<sub>x</sub>, and ROOH which react, again, with chlorine radicals to produce HCl. We called this reaction chain “chlorine counter-cycle”. Very large amounts of HCl are produced in order to deplete ozone. We noted that this  
10 counter-cycle may be “activated” rapidly by the presence of sufficient concentrations of chlorine radicals (e.g., model run Br<sub>2</sub>-M4-Cl<sub>2</sub>=5.0×10<sup>8</sup> molec cm<sup>-2</sup>s<sup>-1</sup>). In that case, the activation of the chlorine counter-cycle unexpectedly leads to the reduction of reactive bromine and reduces the ozone depletion. Similar activations of this counter-cycle are also observed for the base run BrCl-M1.

15 The chemistry of C<sub>2</sub>H<sub>6</sub> mainly reduces the concentration of HO<sub>x</sub> in the air, which modifies the bromine speciation from HOBr/HBr to BrO/Br. C<sub>2</sub>H<sub>6</sub> cleanses the air from oxidants. However, the influence of high mixing ratios of C<sub>2</sub>H<sub>6</sub> on ozone is found very weak. C<sub>2</sub>H<sub>6</sub> has stronger effects on ODEs caused by Cl<sub>2</sub> fluxes, as it converts chlorine radicals to HCl rather rapidly.

20 The chemistry of C<sub>2</sub>H<sub>4</sub> clearly affects the concentrations of HO<sub>x</sub> and bromine radicals. Higher mixing ratios of C<sub>2</sub>H<sub>4</sub> compared to base runs show a drastic reduction in bromine explosion efficiency. Less HOBr and more HBr clearly lead to the re-bromination of SSAs and the bromination of sulfate aerosols. The reduction of available gas phase bromine leads to less ozone depletion. Nevertheless, we also noted  
25 that lower mixing ratios of C<sub>2</sub>H<sub>4</sub> may also lead to less ozone depletion. Under such C<sub>2</sub>H<sub>4</sub> conditions, longer HO<sub>x</sub> lifetimes lead to higher concentrations of HOBr than in the base run: the aerosol debromination is stronger. As mentioned for other species, the increase in HOBr leads to increased deposition on snow and a slow down in BrO self-reaction. Under low mixing ratios of C<sub>2</sub>H<sub>4</sub>, Br<sub>x</sub> also decreases compared to the

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## The chemistry influencing ODEs

M. Piot and R. von Glasow

---

[Title Page](#)[Abstract](#)[Introduction](#)[Conclusions](#)[References](#)[Tables](#)[Figures](#)[⏪](#)[⏩](#)[◀](#)[▶](#)[Back](#)[Close](#)[Full Screen / Esc](#)[Printer-friendly Version](#)[Interactive Discussion](#)

base run and the ozone depletion process slows down.

Among HONO, NO<sub>2</sub> and RONO<sub>2</sub>, HONO was found the most influencing nitrogen-containing species. The photolysis of HONO both releases highly reactive OH and NO. Both radicals induce a shift in bromine speciation to produce HOBr/HBr and BrONO<sub>2</sub>/BrNO<sub>2</sub>. The main effect of this shift is the reduction in gas phase reactive bromine via deposition on snow. Ozone is less destroyed. For all three nitrogen-containing species, we noted an acid displacement in SSAs from HNO<sub>3</sub> to HCl and the increased release of chlorine compared to bromine out of SSAs. We found more gas phase chlorine in model runs including high NO<sub>x</sub> than in base model runs.

Again, we want to stress that recycling of deposited halogen in/on snow is not included in this model. Therefore, variations in deposition on snow are not taken into account. Differences in model results may be important if this recycling on snow is included.

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## The chemistry influencing ODEs

M. Piot and R. von Glasow

Title Page

Abstract

Introduction

Conclusions

References

Tables

Figures

◀

▶

◀

▶

Back

Close

Full Screen / Esc

Printer-friendly Version

Interactive Discussion



**The chemistry  
influencing ODEs**M. Piot and R. von  
Glasow

Title Page

Abstract

Introduction

Conclusions

References

Tables

Figures

◀

▶

◀

▶

Back

Close

Full Screen / Esc

Printer-friendly Version

Interactive Discussion

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**The chemistry  
influencing ODEs**M. Piot and R. von  
Glasow

Title Page

Abstract

Introduction

Conclusions

References

Tables

Figures

◀

▶

◀

▶

Back

Close

Full Screen / Esc

Printer-friendly Version

Interactive Discussion



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**The chemistry  
influencing ODEs**M. Piot and R. von  
Glasow

Title Page

Abstract

Introduction

Conclusions

References

Tables

Figures

◀

▶

◀

▶

Back

Close

Full Screen / Esc

Printer-friendly Version

Interactive Discussion

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**The chemistry  
influencing ODEs**M. Piot and R. von  
Glasow

Title Page

Abstract

Introduction

Conclusions

References

Tables

Figures

◀

▶

◀

▶

Back

Close

Full Screen / Esc

Printer-friendly Version

Interactive Discussion



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influencing ODEs**M. Piot and R. von  
Glasow

Title Page

Abstract

Introduction

Conclusions

References

Tables

Figures

◀

▶

◀

▶

Back

Close

Full Screen / Esc

Printer-friendly Version

Interactive Discussion

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**The chemistry  
influencing ODEs**M. Piot and R. von  
Glasow

Title Page

Abstract

Introduction

Conclusions

References

Tables

Figures

◀

▶

◀

▶

Back

Close

Full Screen / Esc

Printer-friendly Version

Interactive Discussion

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**The chemistry  
influencing ODEs**

M. Piot and R. von  
Glasow

Title Page

Abstract

Introduction

Conclusions

References

Tables

Figures

◀

▶

◀

▶

Back

Close

Full Screen / Esc

Printer-friendly Version

Interactive Discussion



**The chemistry  
influencing ODEs**M. Piot and R. von  
Glasow

Title Page

Abstract

Introduction

Conclusions

References

Tables

Figures

◀

▶

◀

▶

Back

Close

Full Screen / Esc

Printer-friendly Version

Interactive Discussion

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**The chemistry  
influencing ODEs**M. Piot and R. von  
Glasow

Title Page

Abstract

Introduction

Conclusions

References

Tables

Figures

◀

▶

◀

▶

Back

Close

Full Screen / Esc

Printer-friendly Version

Interactive Discussion



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**The chemistry  
influencing ODEs**

M. Piot and R. von  
Glasow

Title Page

Abstract

Introduction

Conclusions

References

Tables

Figures

◀

▶

◀

▶

Back

Close

Full Screen / Esc

Printer-friendly Version

Interactive Discussion



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ACPD

8, 7391–7453, 2008

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**The chemistry  
influencing ODEs**

M. Piot and R. von  
Glasow

---

Title Page

Abstract

Introduction

Conclusions

References

Tables

Figures

⏪

⏩

◀

▶

Back

Close

Full Screen / Esc

Printer-friendly Version

Interactive Discussion



## The chemistry influencing ODEs

M. Piot and R. von  
Glasow

**Table 1.** Differences in air composition between background and coastal air conditions.

| Species            | $\xi_{\text{background}}$<br>(nmol mol <sup>-1</sup> ) | $\xi_{\text{coastal}}$<br>(nmol mol <sup>-1</sup> ) |
|--------------------|--|---|
| CH <sub>3</sub> Br | 0.0  | 0.012   |
| CHBr <sub>3</sub>  | 0.0  | 0.006   |
| DMS                | 0.0  | 0.01 <sup>(a)</sup>                                 |
| Fluxes             | molec cm <sup>-2</sup> s <sup>-1</sup>                 | molec cm <sup>-2</sup> s <sup>-1</sup>              |
| DMS                | 0.0  | 2.0 × 10 <sup>9(b)</sup>                            |
| NH <sub>3</sub>    | 0.0  | 4.0 × 10 <sup>8</sup>                               |
| Sea salt prod.     | No   | Yes   |

<sup>(a)</sup>: Ferek et al. (1995)

<sup>(b)</sup>: estimated from Quinn et al. (1990)

Title Page

Abstract

Introduction

Conclusions

References

Tables

Figures

◀

▶

◀

▶

Back

Close

Full Screen / Esc

Printer-friendly Version

Interactive Discussion



**The chemistry  
influencing ODEs**M. Piot and R. von  
Glasow

Title Page

Abstract

Introduction

Conclusions

References

Tables

Figures

I◀

▶I

◀

▶

Back

Close

Full Screen / Esc

Printer-friendly Version

Interactive Discussion

**Table 2.** Initial composition of sea salt aerosols (Jaenicke, 1988; Andrews et al., 2004).

| Species                               | Cl <sup>-</sup> | Na <sup>+</sup> | Mg <sup>2+</sup> | SO <sub>4</sub> <sup>2-</sup> | K <sup>+</sup> | Ca <sup>2+</sup> | HCO <sub>3</sub> <sup>-</sup> | Br <sup>-</sup> | I <sup>-</sup> |
|---------------------------------------|-----------------|-----------------|------------------|-------------------------------|----------------|------------------|-------------------------------|-----------------|----------------|
| Concentration (mmol l <sup>-1</sup> ) | 550             | 470             | 53               | 28                            | 10             | 10               | 2                             | 0.85            | 0.001          |



## The chemistry influencing ODEs

M. Piot and R. von  
Glasow

**Table 3.** Initial mixing ratios ( $\xi$ ) for gas phase species under background conditions.

| Species                       | $\xi$ (nmol mol <sup>-1</sup> ) | Reference  |
|-------------------------------|---------------------------------|--|
| SO <sub>2</sub>               | 0.01                            | Barrie and Hoff (1984), Bottenheim et al. (1990) |
| O <sub>3</sub>                | 40.0                            | Bottenheim et al. (1986), Anlauf et al. (1994)   |
| NO <sub>2</sub>               | 0.02                            | Beine et al. (1997)                              |
| HNO <sub>3</sub>              | 0.05                            | Leaitch et al. (1994), Ridley et al. (2003)      |
| HONO                          | 0.01                            | Li (1994), Zhou et al. (2001)                    |
| RONO <sub>2</sub>             | 0.14                            | Sander et al. (1997)                             |
| PAN                           | 0.3                             | Ridley et al. (2003), Stroud et al. (2003)       |
| CO                            | 150.0                           | Ramacher et al. (1999)                           |
| Methane                       | 1800.0                          | Worthy et al. (1994)                             |
| Ethane                        | 1.5                             | Jobson et al. (1994), Hopper et al. (1994)       |
| Ethene                        | 0.08                            | Bottenheim et al. (2002b)                        |
| HCHO                          | 0.1                             | Barrie et al. (1994), Sumner and Shepson (1999)  |
| H <sub>2</sub> O <sub>2</sub> | 0.3                             | de Serves (1994), Snow et al. (2002)             |

Title Page

Abstract

Introduction

Conclusions

References

Tables

Figures

◀

▶

◀

▶

Back

Close

Full Screen / Esc

Printer-friendly Version

Interactive Discussion



**The chemistry  
influencing ODEs**M. Piot and R. von  
Glasow**Table 4.** Prescribed halogen fluxes in  $\text{molec cm}^{-2} \text{s}^{-1}$ . High fluxes of chlorine are discussed in Sect. 4.6.

|               | Partial ODE-P4<br>( $\text{molec cm}^{-2} \text{s}^{-1}$ ) | Major ODE-M4<br>( $\text{molec cm}^{-2} \text{s}^{-1}$ ) | Major ODE-M1<br>( $\text{molec cm}^{-2} \text{s}^{-1}$ ) |
|---------------|--|--|--|
| $\text{Br}_2$ | $5.0 \times 10^7$  | $9.0 \times 10^7$  | $1.5 \times 10^9$  |
| BrCl          | $1.0 \times 10^8$  | $1.9 \times 10^8$  | $3.4 \times 10^9$  |
| $\text{Cl}_2$ | $3.0 \times 10^9$  | $2.0 \times 10^{10}$                                     | $3.0 \times 10^{11}$                                     |

Title Page

Abstract

Introduction

Conclusions

References

Tables

Figures

I ◀

▶ I

◀

▶

Back

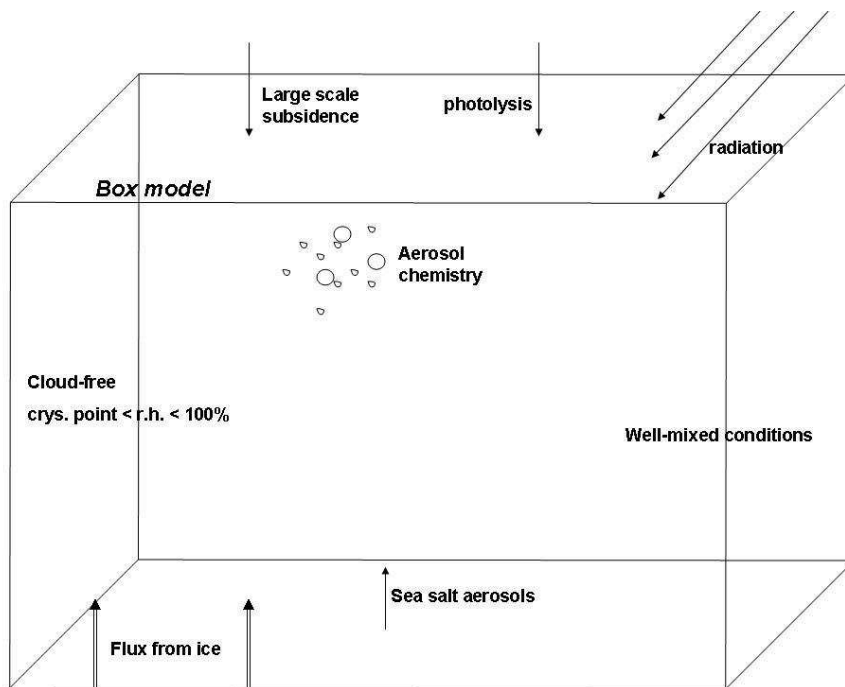
Close

Full Screen / Esc

Printer-friendly Version

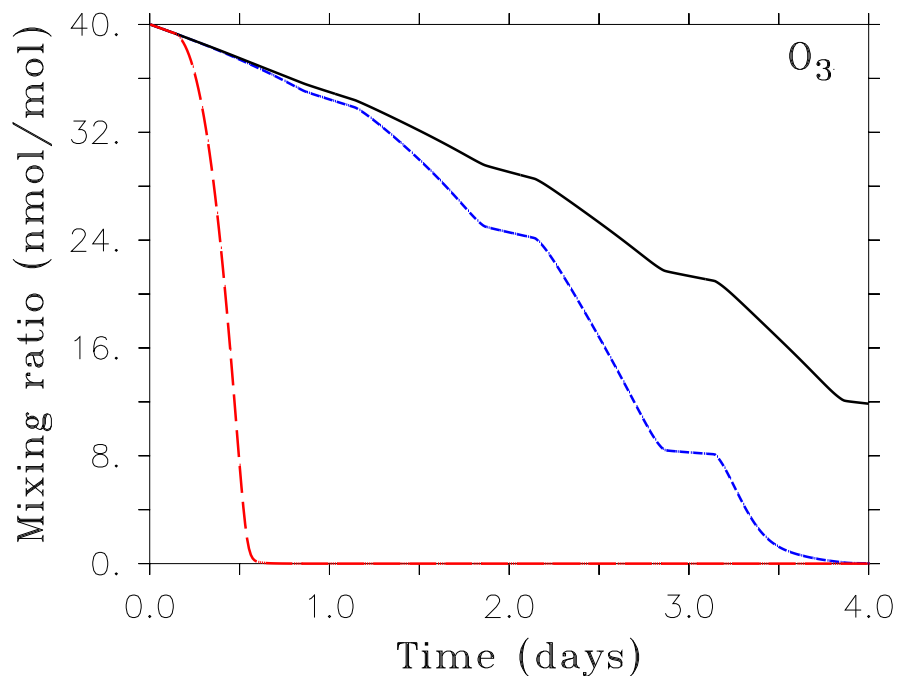
Interactive Discussion



The chemistry  
influencing ODEsM. Piot and R. von  
Glasow

**Fig. 1.** Schematic depiction of the most important processes included in the box model version of MISTRA, applied for Arctic conditions. Fluxes from ice and sea salt production are switched ON/OFF, depending on the air mass type (see text and Table 1).

[Title Page](#)[Abstract](#)[Introduction](#)[Conclusions](#)[References](#)[Tables](#)[Figures](#)[◀](#)[▶](#)[◀](#)[▶](#)[Back](#)[Close](#)[Full Screen / Esc](#)[Printer-friendly Version](#)[Interactive Discussion](#)

The chemistry  
influencing ODEsM. Piot and R. von  
Glasow

**Fig. 2.**  $Br_2$ -induced ozone depletions (M1, M4, and P4, see text). Solid black line:  $F_{Br_2} = 5.0 \times 10^7 \text{ molec cm}^{-2} \text{ s}^{-1}$ ; dashed blue line:  $F_{Br_2} = 9.0 \times 10^7 \text{ molec cm}^{-2} \text{ s}^{-1}$ ; dashed red line:  $F_{Br_2} = 1.5 \times 10^9 \text{ molec cm}^{-2} \text{ s}^{-1}$ . Respective  $Cl_2$ - and  $BrCl$ -induced ODEs are not shown.

Title Page

Abstract

Introduction

Conclusions

References

Tables

Figures

◀

▶

◀

▶

Back

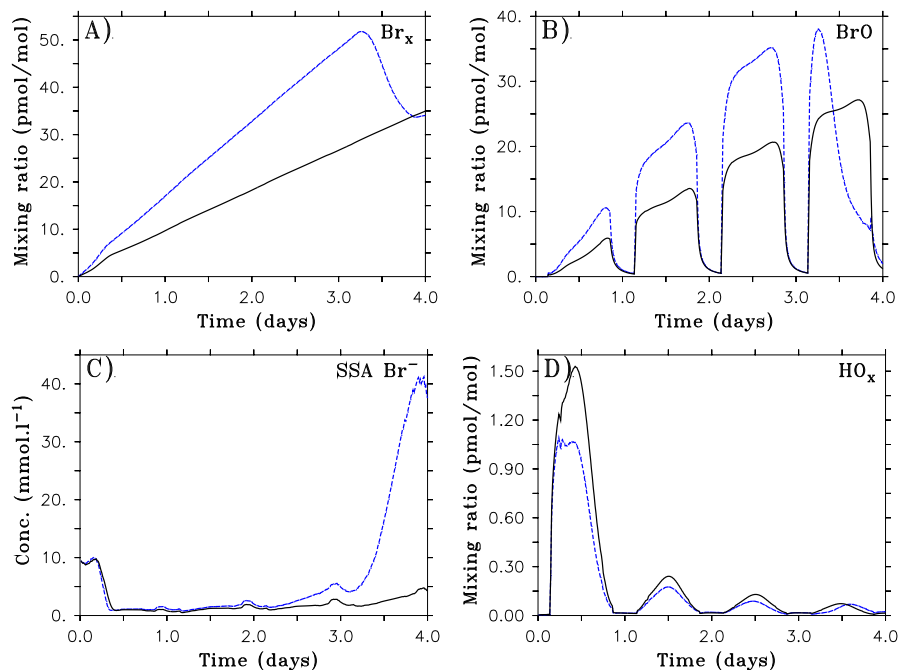
Close

Full Screen / Esc

Printer-friendly Version

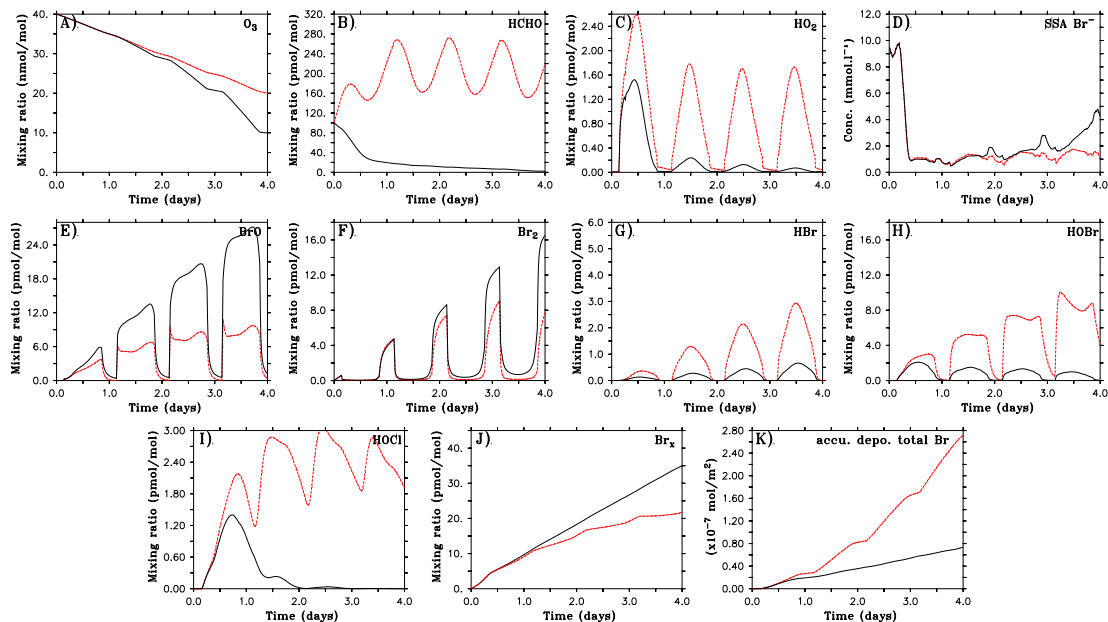
Interactive Discussion



The chemistry  
influencing ODEsM. Piot and R. von  
Glasow

**Fig. 3.** Chemistry of Br<sub>2</sub>-induced ODEs for background air conditions. Solid black line: base run Br<sub>2</sub>-P4; dashed blue line: base run Br<sub>2</sub>-M4. **(A)** total gas phase bromine Br<sub>x</sub>, **(B)** BrO, **(C)** sea salt aerosol Br<sup>-</sup>, **(D)** HO<sub>x</sub> (OH+HO<sub>2</sub>).

[Title Page](#)[Abstract](#)[Introduction](#)[Conclusions](#)[References](#)[Tables](#)[Figures](#)[◀](#)[▶](#)[◀](#)[▶](#)[Back](#)[Close](#)[Full Screen / Esc](#)[Printer-friendly Version](#)[Interactive Discussion](#)

The chemistry  
influencing ODEsM. Piot and R. von  
Glasow

**Fig. 4.** Solid black line: base run  $\text{Br}_2\text{-P4}$ ; dashed red line:  $\text{Br}_2\text{-P4-HCHO}=5.0 \times 10^9$  molec  $\text{cm}^{-2} \text{s}^{-1}$ . Background air conditions. **(A)**  $\text{O}_3$ , **(B)**  $\text{HCHO}$ , **(C)**  $\text{HO}_2$ , **(D)** sea salt aerosol  $\text{Br}^-$ , **(E)**  $\text{BrO}$ , **(F)**  $\text{Br}_2$ , **(G)**  $\text{HBr}$ , **(H)**  $\text{HOBr}$ , **(I)**  $\text{HOCl}$ , **(J)** total bromine  $\text{Br}_x$ , **(K)** accumulated deposition of total gas phase bromine on snow.

Title Page

Abstract

Introduction

Conclusions

References

Tables

Figures

◀

▶

◀

▶

Back

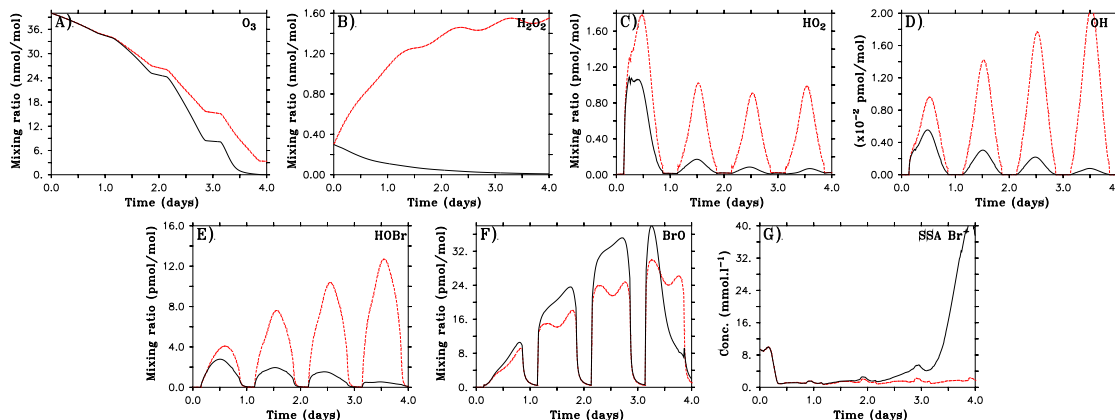
Close

Full Screen / Esc

Printer-friendly Version

Interactive Discussion



The chemistry  
influencing ODEsM. Piot and R. von  
Glasow

**Fig. 5.** Solid black line: base run  $\text{Br}_2\text{-M4}$ ; dashed red line:  $\text{Br}_2\text{-M4}$ - $\text{H}_2\text{O}_2=1.5 \times 10^{10} \text{ molec cm}^{-2} \text{ s}^{-1}$ . Background air conditions. (A)  $\text{O}_3$ , (B)  $\text{H}_2\text{O}_2$ , (C)  $\text{HO}_2$ , (D)  $\text{OH}$ , (E)  $\text{HOBr}$ , (F)  $\text{BrO}$ , (G) sea salt aerosol  $\text{Br}^-$ .

Title Page

Abstract

Introduction

Conclusions

References

Tables

Figures

◀

▶

◀

▶

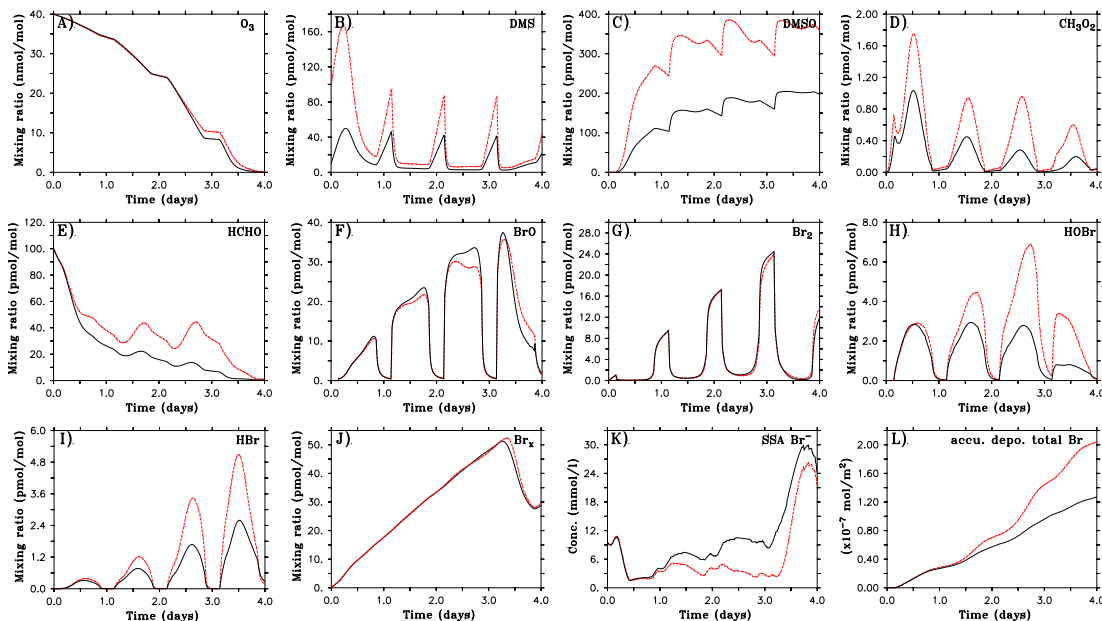
Back

Close

Full Screen / Esc

Printer-friendly Version

Interactive Discussion

The chemistry  
influencing ODEsM. Piot and R. von  
Glasow

**Fig. 6.** Solid black line:  $\text{Br}_2\text{-M4}$ ; dashed red line:  $\text{Br}_2\text{-M4-DMS}=0.1 \text{ nmol mol}^{-1}$  with a flux of  $4.0 \times 10^9 \text{ molec cm}^{-2} \text{ s}^{-1}$ . Coastal air conditions. **(A)**  $\text{O}_3$ , **(B)** DMS, **(C)** DMSO, **(D)**  $\text{CH}_3\text{O}_2$ , **(E)** HCHO, **(F)** BrO, **(G)**  $\text{Br}_2$ , **(H)** HOBr, **(I)** HBr, **(J)** total bromine  $\text{Br}_x$ , **(K)** sea salt aerosol  $\text{Br}^-$ , **(L)** accumulated deposition of total gas phase bromine.

Title Page

Abstract

Introduction

Conclusions

References

Tables

Figures

◀

▶

◀

▶

Back

Close

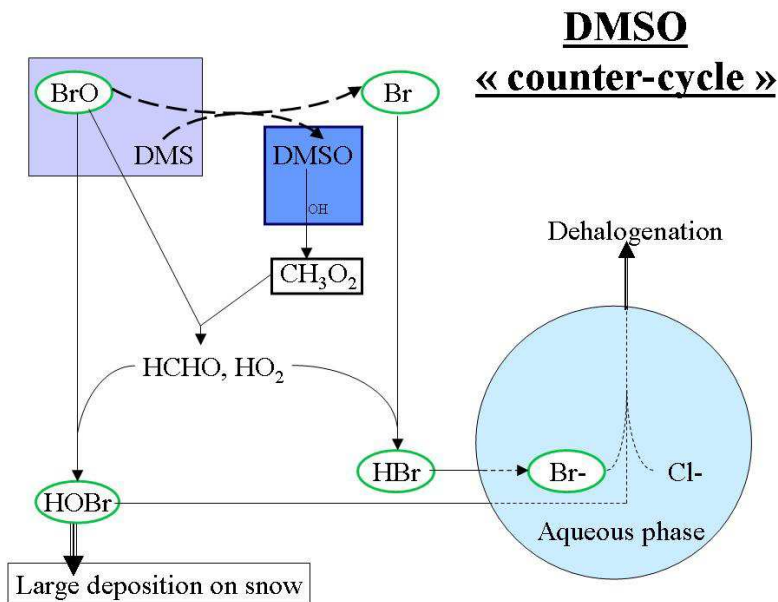
Full Screen / Esc

Printer-friendly Version

Interactive Discussion

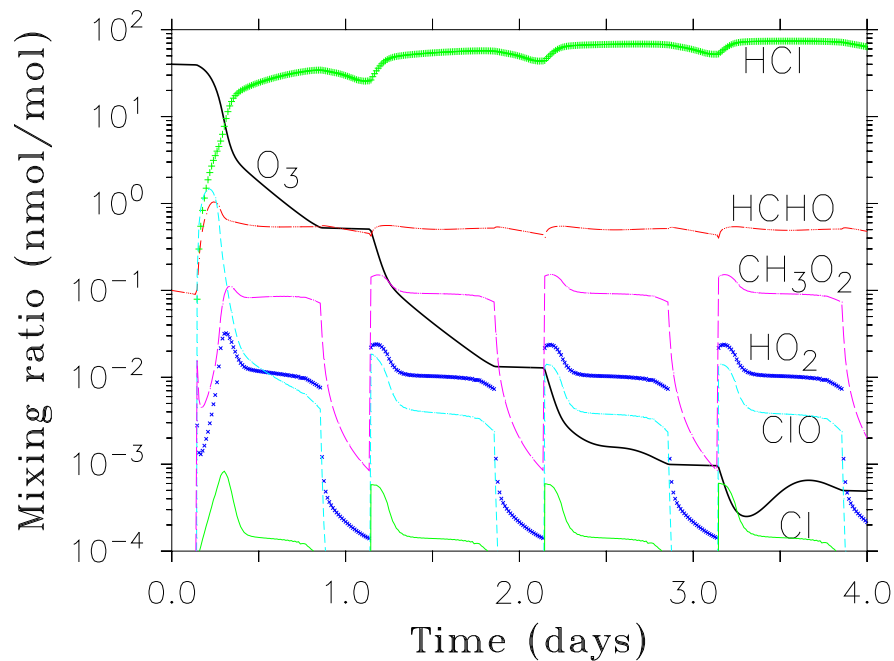




The chemistry  
influencing ODEsM. Piot and R. von  
Glasow

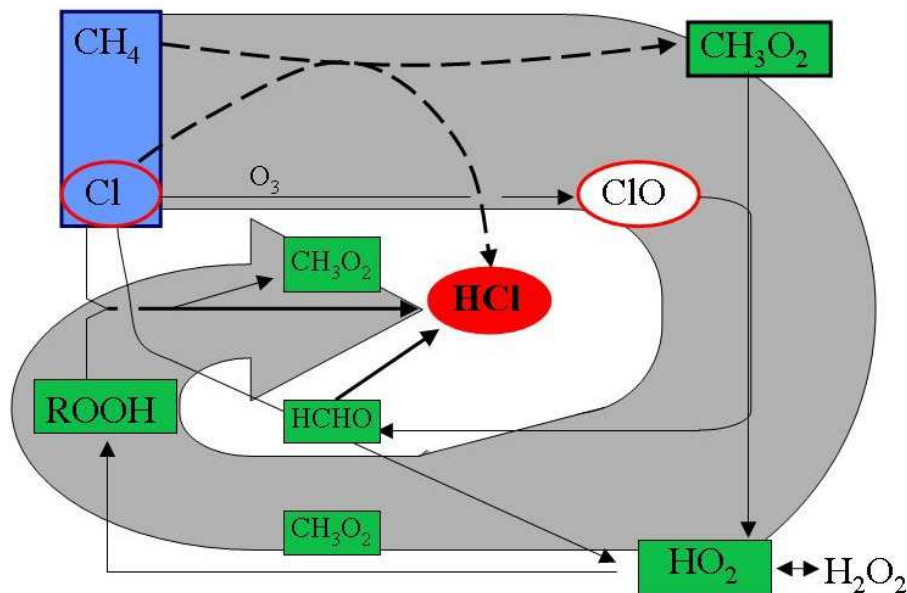
**Fig. 7.** Schematic description of the DMS chemistry and the DMSO “counter-cycle”. The initiating reaction is depicted as bold dashed line. The reaction highlighted in the blue square is the key reaction producing CH<sub>3</sub>O<sub>2</sub>.

[Title Page](#)[Abstract](#)[Introduction](#)[Conclusions](#)[References](#)[Tables](#)[Figures](#)[◀](#)[▶](#)[◀](#)[▶](#)[Back](#)[Close](#)[Full Screen / Esc](#)[Printer-friendly Version](#)[Interactive Discussion](#)

The chemistry  
influencing ODEsM. Piot and R. von  
Glasow

**Fig. 8.** Simulated evolution of gas phase  $O_3$ , HCHO,  $HO_2$ , HCl, ClO, Cl, and  $CH_3O_2$  using the  $Cl_2$ -M1 base run.

[Title Page](#)[Abstract](#)[Introduction](#)[Conclusions](#)[References](#)[Tables](#)[Figures](#)[◀](#)[▶](#)[◀](#)[▶](#)[Back](#)[Close](#)[Full Screen / Esc](#)[Printer-friendly Version](#)[Interactive Discussion](#)

The chemistry  
influencing ODEsM. Piot and R. von  
Glasow

**Fig. 9.** Most important reactions caused by high mixing ratios of gaseous  $\text{Cl}_2$ . The initiating reaction is highlighted by a blue square. Red circles: chlorine species. Green squares: key species for the  $\text{HCl}$  production. The grey arrow in background is to show the overall direction taken by the “chlorine counter-cycle”.

Title Page

Abstract

Introduction

Conclusions

References

Tables

Figures

◀

▶

◀

▶

Back

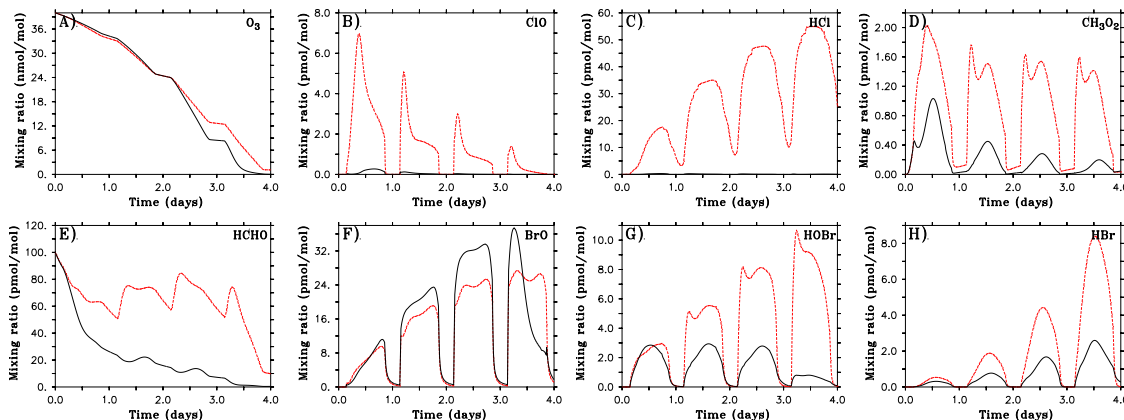
Close

Full Screen / Esc

Printer-friendly Version

Interactive Discussion



The chemistry  
influencing ODEsM. Piot and R. von  
Glasow

**Fig. 10.** Solid black line:  $\text{Br}_2\text{-M4}$ ; Dashed red line:  $\text{Br}_2\text{-M4-Cl}_2=5.0 \times 10^8 \text{ molec cm}^{-2} \text{ s}^{-1}$ . Coastal air conditions. (A)  $\text{O}_3$ , (B)  $\text{ClO}$ , (C)  $\text{HCl}$ , (D)  $\text{CH}_3\text{O}_2$ , (E)  $\text{HCHO}$ , (F)  $\text{BrO}$ , (G)  $\text{HOBr}$ , (H)  $\text{HBr}$ .

Title Page

Abstract

Introduction

Conclusions

References

Tables

Figures

◀

▶

◀

▶

Back

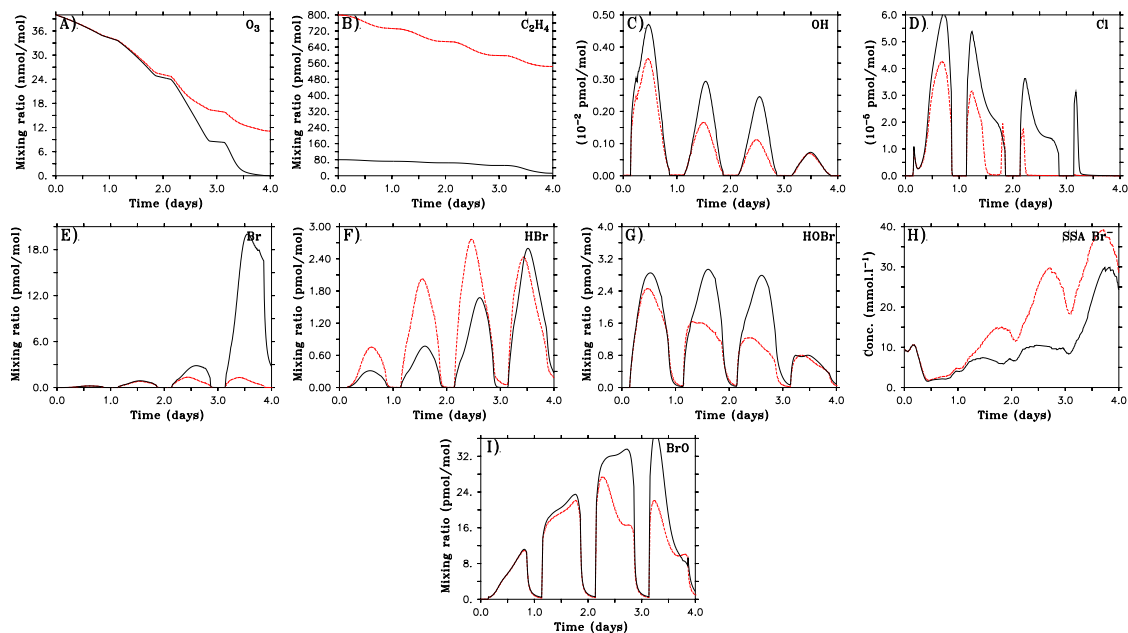
Close

Full Screen / Esc

Printer-friendly Version

Interactive Discussion





**Fig. 11.** Solid black line:  $\text{Br}_2\text{-M4}$ ; dashed red line:  $\text{Br}_2\text{-M4-C}_2\text{H}_4=0.8 \text{ nmol mol}^{-1}$ . Coastal air conditions. (A)  $\text{O}_3$ , (B)  $\text{C}_2\text{H}_4$ , (C) OH, (D) Cl, (E) Br, (F) HBr, (G) HOBr, (H) sea salt aerosol  $\text{Br}^-$ , (I) BrO.

## The chemistry influencing ODEs

M. Piot and R. von Glasow

Title Page

Abstract

Introduction

Conclusions

References

Tables

Figures

◀

▶

◀

▶

Back

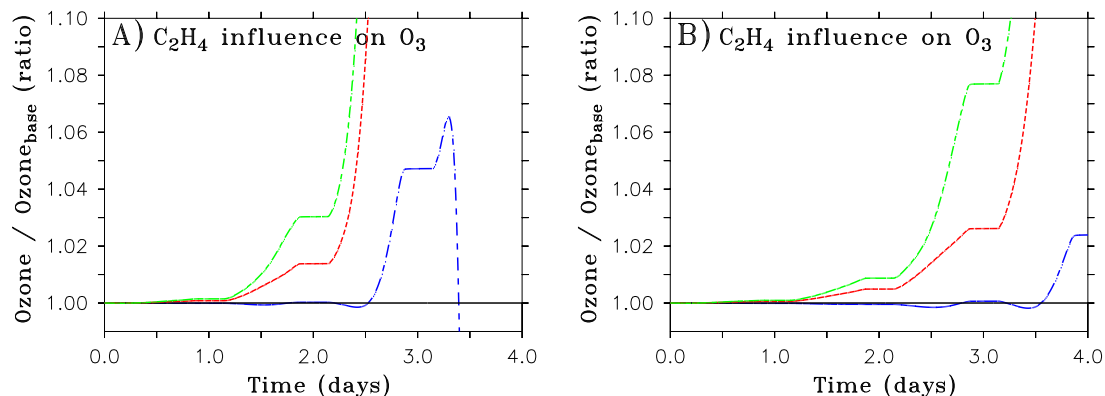
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Full Screen / Esc

Printer-friendly Version

Interactive Discussion



The chemistry  
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**Fig. 12.** Ratio of O<sub>3</sub> in sensitivity runs to the base run for **(A)** M4 and **(B)** P4 ODEs, respectively. Dash-dotted blue line:  $\xi_{\text{C}_2\text{H}_4}=0.01 \text{ nmol mol}^{-1}$ ; dashed red line:  $\xi_{\text{C}_2\text{H}_4}=0.5 \text{ nmol mol}^{-1}$ ; dash-dotted green line:  $\xi_{\text{C}_2\text{H}_4}=0.8 \text{ nmol mol}^{-1}$ . Coastal air conditions.

Title Page

Abstract

Introduction

Conclusions

References

Tables

Figures

◀

▶

◀

▶

Back

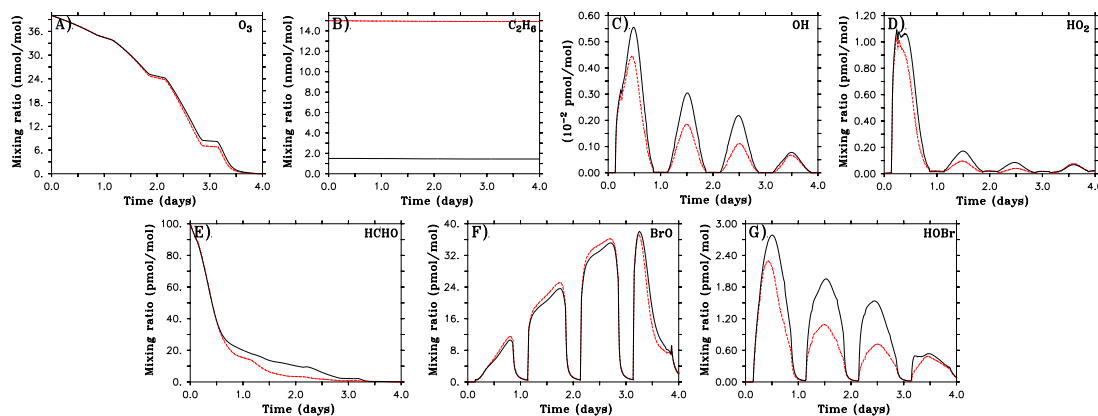
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**Fig. 13.** Solid black line:  $\text{Br}_2\text{-M4}$ ; dashed red line:  $\text{Br}_2\text{-M4-C}_2\text{H}_6=15 \text{ nmol mol}^{-1}$ . Background air conditions. **(A)**  $\text{O}_3$ , **(B)**  $\text{C}_2\text{H}_6$ , **(C)**  $\text{OH}$ , **(D)**  $\text{HO}_2$ , **(E)**  $\text{HCHO}$ , **(F)**  $\text{BrO}$ , **(G)**  $\text{HOBr}$ .

Title Page

Abstract

Introduction

Conclusions

References

Tables

Figures

◀

▶

◀

▶

Back

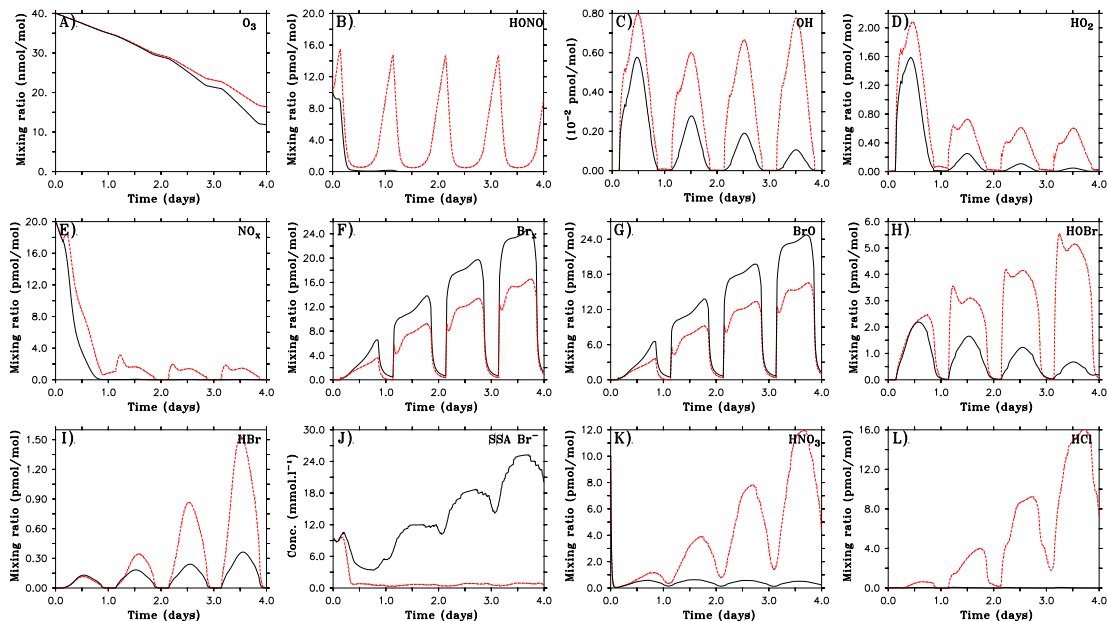
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Full Screen / Esc

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Interactive Discussion



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**Fig. 14.** Solid black line:  $\text{Br}_2\text{-P4}$ ; dashed red line:  $\text{Br}_2\text{-P4-HONO} = 5.0 \times 10^8 \text{ molec cm}^{-2} \text{ s}^{-1}$ . Background air conditions. (A)  $\text{O}_3$ , (B) HONO, (C) OH, (D)  $\text{HO}_2$ , (E)  $\text{NO}_x$ , (F) total bromine  $\text{Br}_x$ , (G)  $\text{BrO}$ , (H)  $\text{HOBr}$ , (I)  $\text{HBr}$ , (J) sea salt aerosol  $\text{Br}^-$ , (K)  $\text{HNO}_3$ , (L)  $\text{HCl}$ .

Title Page

Abstract

Introduction

Conclusions

References

Tables

Figures

◀

▶

◀

▶

Back

Close

Full Screen / Esc

Printer-friendly Version

Interactive Discussion

