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Atmospheric Chemistry and Physics

Gaseous (DMS, MSA, SO_2, H_2SO_4 and DMSO) and particulate (sulfate and methanesulfonate) sulfur species over the northeastern coast of Crete

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Abstract. A detailed study of the levels, the temporal and diurnal variability of the main compounds involved in the biogenic sulfur cycle was carried out in Crete (Eastern Mediterranean) during the Mediterranean Intensive Oxidant Study (MINOS) field experiment in July–August 2001. Intensive measurements of gaseous dimethylsulfide (DMS), dimethylsulfoxide (DMSO), sulfur dioxide (SO_2), sulfuric (H_2SO_4) and methanesulfonic acids (MSA) and particulate sulfate (SO_4^{2-}) and methanesulfonate (MS^-) have been performed during the campaign.

Dimethylsulfide (DMS) levels ranged from 2.9 to 136 pmol·mol^{-1} (mean value of 21.7 pmol·mol^{-1}) and showed a clear diurnal variation with daytime maximum. During nighttime DMS levels fall close or below the detection limit of 2 pmol·mol^{-1}. Concurrent measurements of OH and NO_3 radicals during the campaign indicate that NO_3 levels can explain most of the observed diurnal variation of DMS. Dimethylsulfoxide (DMSO) ranged between 0.02 and 10.1 pmol·mol^{-1} (mean value of 1.7 pmol·mol^{-1}) and presents a diurnal variation similar to that of DMS. SO_2 levels ranged from 220 to 2970 pmol·mol^{-1} (mean value of 1030 pmol·mol^{-1}), while nss-SO_4^{2-} and MS^- ranged from 330 to 7100 pmol·mol^{-1}, (mean value of 1440 pmol·mol^{-1}) and 1.1 to 37.5 pmol·mol^{-1} (mean value of 11.5 pmol·mol^{-1}) respectively.

Of particular interest are the measurements of gaseous MSA and H_2SO_4. MSA ranged from below the detection limit (3×10^4) to 3.7×10^7 molecules cm^{-3}, whereas H_2SO_4 ranged between 1×10^5 and 9.0×10^7 molecules cm^{-3}. The measured H_2SO_4 maxima are among the highest reported in literature and can be attributed to high insolation, absence of precipitation and increased SO_2 levels in the area. From the concurrent SO_2, OH, and H_2SO_4 measurements a sticking coefficient of 0.52±0.28 was calculated for H_2SO_4. From the concurrent MSA, OH, and DMS measurements the yield of gaseous MSA from the OH-initiated oxidation of DMS was calculated to range between 0.1–0.4%. This low MSA yield implies that gaseous MSA levels can not account for the observed MS^- levels. Heterogeneous reactions of DMSO on aerosols should be considered to explain the observed levels of MS^-.

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1 Introduction

Oceanic dimethylsulfide (DMS) emission is proposed to play a key role in climate regulation, through its oxidation products, especially in the remote marine atmosphere (Charlson et al., 1987). Once emitted into the atmosphere DMS can be oxidized by OH, NO_3 and possibly other radical species as Cl atoms and BrO radicals forming a variety of products such as sulfur dioxide (SO_2), methanesulfonic acid (MSA) and dimethylsulfoxide (DMSO) (e.g. Watts et al., 1990; Urbanski et al., 1998; Davis et al., 1998; Arsene et al., 1999 and references therein).

The proposed climatic impact of DMS is decisively determined from H_2SO_4 production contributing to the formation of new hygroscopic sulfate particles. On the other hand, the contribution of MSA, another main oxidation end-product of DMS, to new particle formation is expected to be negligible, at least in the polluted marine atmosphere (Kreidenweis and

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Seinfeld, 1988). Particles can play an important climatic role by scattering solar radiation and acting as cloud condensation nuclei (CCN), thus influencing the earth's albedo and climate (IPCC, 2001).

The complete mechanism of DMS oxidation is still under investigation although considerable effort has been made during the last decades to clarify its oxidation scheme via field measurements, kinetic experiments and model results (e.g. Berresheim et al., 1995 and references therein). From numerous field and laboratory experiments it is clear that the atmospheric distribution of DMS oxidation products critically depends on various factors like meteorology (temperature, humidity) and the oxidation power of the atmosphere (e.g. NO_x, HO_x).

Concurrent field data both on the oxidants involved in the DMS oxidation and including at least the major oxidation products of DMS in gaseous and particulate phase are clearly needed to validate the laboratory results. Such data is particularly difficult to obtain since the levels of free radicals and DMS oxidation products are often very low (in sub-pmol·mol^{-1} level) and state-of-the-art analytical techniques are needed. To our knowledge only few studies have fulfilled all the above requirements. Especially in the marine boundary layer only one study has been performed so far (the SCATE experiment in coastal Antarctica; Berresheim and Eisele, 1998). The MINOS campaign conducted in Crete in July-August 2001, provides the first comprehensive dataset to understand the factors controlling the distribution and the fate of DMS oxidation products under moderately polluted marine conditions. Intensive measurements of gaseous DMS, DMSO, SO_2, H_2SO_4 and MSA and particulate sulfate (SO_4^{2-}) and methanesulfonate (MS^-) have been performed during the campaign in conjunction with parallel measurements of OH (Berresheim et al., 2003) and NO_3 radicals (Vrekoussis et al., 2003).

2 Experimental

2.1 Location

All measurements reported here have been performed at Finokalia (35°24' N, 25°60' E), a remote location on the northern coast of Crete, Greece during the MINOS campaign. The station is located at the top of an elevation (150 m) facing the sea within the sector of 270° to 90°. No significant anthropogenic sources exist at a distance shorter than 20 km within the above-mentioned sector (Mihalopoulos et al., 1997; Kouvarakis et al., 2000). Details on the meteorological conditions encountered during the campaign and air masses origin are presented in companion papers (Salisbury et al., 2003; Berresheim et al., 2003). We note only the quasi absence of clouds during the campaign with the exception of 1st of August when the sky was partial cloudy.

2.2 Sampling and analysis

DMS: Air sampling and subsequent analysis of DMS were carried out as described in previous studies (Nguyen et al., 1990; Sciare et al., 2001; Kouvarakis and Mihalopoulos, 2002). Briefly, air was collected into 6-liter stainless steel canisters and pressurized up to 5 bars using a pump. As described by Kouvarakis and Mihalopoulos (2002), the use of such canisters minimizes possible DMS destruction by oxidants. Immediately after sampling DMS analysis was isothermally performed (95°C) by gas chromatography (HP 5890A GC equipped with a Chromosil 310 column and a flame photometric detector; FPD). The detection limit was typically 0.2 ng (DMS), which under our sampling conditions corresponds to 2 pmol·mol^{-1} of DMS. The accuracy of the analysis was 10%. In total, 490 atmospheric DMS measurements were performed during the campaign (28 July to 22 August 2001) with an average sampling interval of 1 h.

DMSO and SO_2: The technique used for the sampling of atmospheric DMSO and SO_2 is based on the nebulization/reflux principle (Cofer et al., 1985; Sciare and Mihalopoulos, 2000). Two samplers were operated in parallel during the campaign (one for DMSO and the other for SO_2 and other water soluble gases) with an average flow rate of 16 L min^{-1}. After sampling the aliquots for DMSO and SO_2 analysis were kept refrigerated at 4°C in the dark and analyzed within a month period. Conservation tests have shown that both DMSO and SO_2 were stable under these storage conditions.

DMSO was reduced by sodium borohydride ($NaBH_4$) to DMS and subsequently analysed by gas chromatography. Recovery and reproducibility of the analysis were of the order of 90% and better than 10%, respectively. Details on the DMSO collection and analysis have been described by Sciare and Mihalopoulos (2000). In total, 201 atmospheric DMSO measurements have been performed during the campaign with an average sampling step of 3 hrs. The detection limit for a mean sampling volume of 3 m^3 was found to be 0.03 pmol·mol^{-1} and the precision was estimated to be 20%.

Gaseous SO_2 (n=226) was trapped in aqueous droplets produced inside the sample mist chamber and analyzed as sulfate by Ion Chromatography (see details below). The Cofer mist technique has been successfully compared in the field against the alternative Na_2CO_3-impregnated filter technique (Sciare et al., 2003). The detection limit for a mean sampling volume of 3 m^3 was found to be 20 pmol·mol^{-1} and the precision was estimated to be 15%.

Aerosol (sulfate and methanesulfonate) sampling and analysis: Bulk aerosol samples were collected on 0.5 μm PTFE filters. In total, 226 aerosol samples were collected during the sampling period in parallel with DMSO and SO_2 and subsequently analysed for the main anions and cations by Ion Chromatography.

For the analysis of anions (chloride: Cl^-; bromide: Br^-; nitrate: NO_3^-; sulfate: SO_4^{2-}; oxalate: $C_2O_4^{2-}$) a

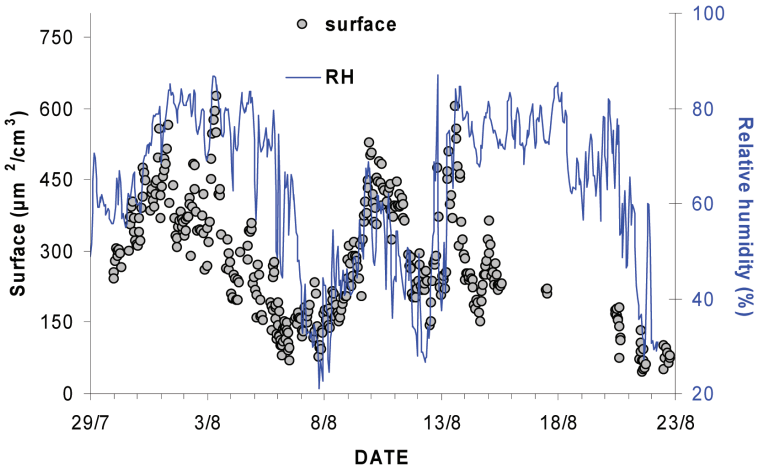

Fig. 1. Variation of aerosol surface data and RH during the campaign.

Dionex AS4A-SC column with ASRS-I suppressor in autosuppression mode of operation was used and isocratic elution at 2.0 ml min^{-1} of Na_2CO_3/$NaHCO_3$ eluent. MS^- was analyzed using a Dionex DX-500 Ion Chromatography with an AS11 analytical column and NaOH (0.1–3.5 mM) as eluent in gradient mode. For the cations (sodium: Na^+; ammonium: NH_4^+; potassium: K^+; magnesium: Mg^{2+} and calcium: Ca^{2+}) a CS12-SC column was used with CSRS-I suppressor. Separation was achieved under isocratic conditions with 20 mM MSA eluent and flow rate of 1.0 ml min^{-1}: The reproducibility of the measurements was better than 2% and the detection limit ranged from around 5 ppb for the main anions and cations to below 0.2 ppb for MS^-. More details on the analytical technique can be found in Baboukas et al. (2000) and Kouvarakis and Mihalopoulos (2002).

Atmospheric H_2SO_4, MSA, and OH measurements: H_2SO_4, MSA, and OH concentrations in the gas phase were measured by Deutscher Wetterdienst (DWD) using chemical ionization mass spectrometry (CIMS) based on methods previously developed by Eisele and coworkers (Tanner et al., 1997; Eisele and Tanner, 1993). The DWD CIMS system has been described in detail by Berresheim et al. (2000, 2002). At 5 min signal integration, the overall 2-sigma precisions and detection limits for both the H_2SO_4 and MSA measurements were 21% and 3×10^4 molecules cm^{-3}. For OH radicals the corresponding values were 22% and 2.4×10^5 molecules cm^{-3} (Berresheim et al., 2003). Ancillary measurements: An optical particle counter (PCS 2010, Palas Inc.) was used to measure the size distribution of particles between 270 nm and 9 μm. The size distribution below 200 nm was measured with a TSI 3040 diffusion battery. The optical particle counter was installed in a measurement container which was deployed approximately 30 m below the meteorological station. The total aerosol surface density was calculated from the data of both instruments. The gap between 200 and 270 nm could be closed by comparison with an aircraft based PCASP instrument (Schneider et al., 2003), which yielded a contribution of about 30% to the total surface from particles between 200 and 270 nm. The particles with diameters smaller than 200 nm contributed about 60% to the total aerosol surface density.

Meteorological parameters like temperature, humidity, wind speed and direction were measured continuously by an automated meteorological station.

Table 1. Observed mean mixing ratios of DMS and the other gaseous and particulate S species during the campaign (SD=standard deviation, n=number of data). The mean values for MSA and H_2SO_4 have been calculated using hourly data

Compound	n	Mean	SD	Median
(pmol·mol^{-1})				
DMS	490	21.7	21.5	15.2
DMSO	201	1.66	1.74	1.20
MS^-	224	11.5	5.8	10.4
SO_2	220	1030	600	910
nss$-SO_4^{2-}$	220	1440	760	1300
(molecules cm^{-3})				
MSAg	3453	5.2×10^5	2.5×10^6	1.1×10^5
H_2SO_4g	674	1.0×10^7	1.7×10^7	1.8×10^6

3 Results and Discussion

Table 1 summarizes the range, mean and standard deviation of concentrations of DMS and the other gaseous and particulate S species during the campaign. In the present section the variations and the factors controlling the levels of the gaseous and particulate S species measured during the campaign will be discussed. Preceding the presentation of the levels of the S compounds and the discussion of the factors controlling their

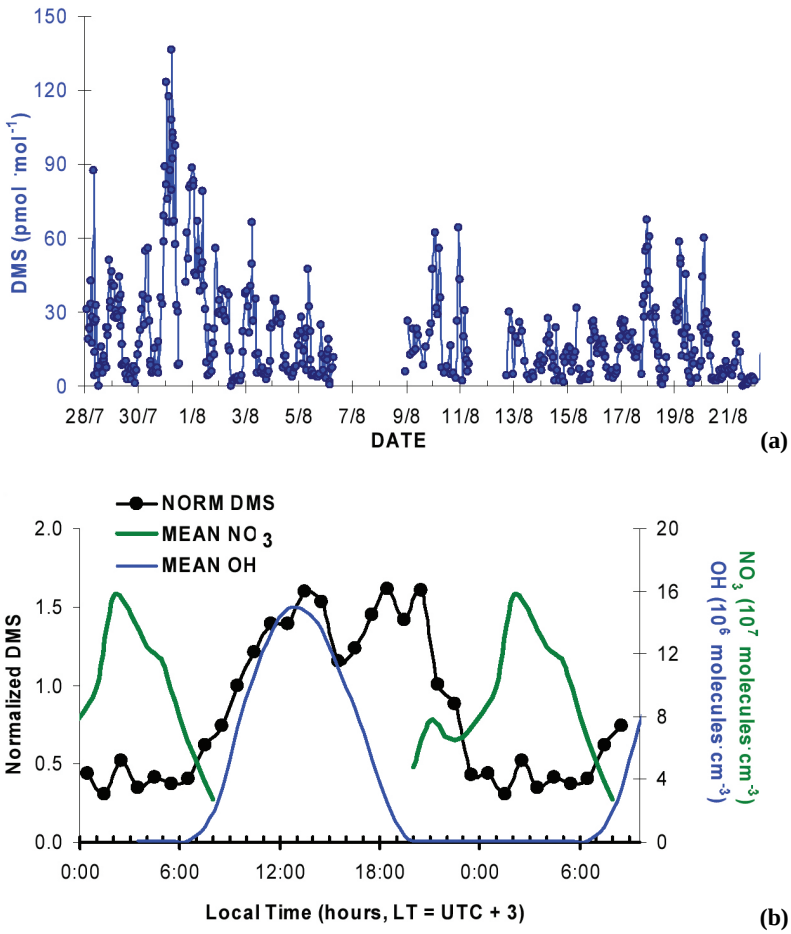

Fig. 2. Atmospheric concentrations of **(a)** DMS during the campaign and **(b)** diurnal variation of normalized DMS (Normalized DMS=$DMS_{(time\,y,day\,x)}$/mean $DMS_{(day\,x)}$) in relation to the mean OH and NO_3 levels measured during the campaign. Time refers to local time (LT=UT+3)

levels, a summary of the aerosol surface and RH data is given in Fig. 1. Aerosol surface concentrations during the MINOS campaign ranged from 49 to 605 $\mu m^2\,cm^{-3}$ (272±120) and followed the variation of RH quite well (r^2=0.21, $p<0.001$).

3.1 DMS

DMS levels (Fig. 2a) range from values near the detection limit (2.0 pmol·mol^{-1}) to 136 pmol·mol^{-1} (on 31 July) with a mean value of 21.7 pmol·mol^{-1}. They are in good agreement with the data reported for a coastal site in Israel measured during August 1995 (22 pmol·mol^{-1}; Ganor et al., 2000). The DMS values observed during MINOS fall in the lower range of summer values reported for the Finokalia area for the period 1997–1999 (57.5±33.8 pmol·mol^{-1}; Kouvarakis and Mihalopoulos, 2002). The relatively low DMS values observed during MINOS can be attributed to two possible factors: 1) Lower wind speed during the campaign (7.1 m s^{-1}) than the summertime mean value derived from long term measurements (9 m s^{-1}; Kouvarakis et al., 2000). Such difference in wind speed can affect the transfer (or piston) velocity (K_w) of DMS and consequently DMS sea-to-air fluxes by almost 50% (Liss and Merlivat, 1986). 2) Land breeze occurrence during nighttime (see also Berresheim et al, 2003, this issue). This rather unusual situation for the area, resulted in a dilution of oceanic DMS flux during night by a factor of 30% compared to the daytime flux (see the companion modelling paper by Kanakidou et al., manuscript in preparation). Figure 2b depicts the diurnal variation of normalized DMS (Normalized DMS=$DMS_{(time\,y,day\,x)}$/mean $DMS_{(day\,x)}$) during the experiment.

Measurements of OH and NO_3 radicals, averaged during the entire campaign, are also reported in Fig. 2b. DMS presents a well-defined diurnal variation with high values during daytime and a clear minimum (in most cases near to the detection limit) during night with a day to night amplitude of 5.4. Based on the observed levels of OH and NO_3 radicals, the nighttime DMS minimum is attributed to dilution due to

land-to-sea breeze occurrence and the reaction of DMS with NO_3 radicals. The minimum observed around 15:00 (LT), can be explained by/attributed to the DMS+OH reaction. Finally the increase observed around 18:00 can be explained by the absence of significant levels of both radicals. A detailed modeling study of the diurnal variation of DMS is presented in Kanakidou et al. (manuscript in preparation). The DMS diurnal variation is in agreement with previous observations during summertime at Finokalia (Kouvarakis and Mihalopoulos, 2002) and observations in the marine atmosphere under continental influence (e.g. Andreae et al., 1985). However, it is different from those observed in the remote marine atmosphere (Andreae et al., 1985; Ayers et al., 1995; Sciare et al., 2001) where a clear DMS minimum is observed around noon time.

3.2 DMSO

DMSO is one of the major intermediate DMS oxidation products. In the gas phase it is produced through the OH-addition channel, with a yield of approximately 30% of the total reaction at 298 K and as high as 70% at lower temperatures (Hynes and Wine, 1996; Arsene et al., 1999). In the case of XO-initiated atmospheric DMS oxidation (X = Br, Cl), DMSO yield is expected to be around 100% (Barnes et al., 1991; Ingham et al., 1999). DMSO is also produced via multiphase reactions of DMS with O_3 (Lee and Zhu, 1994). To our knowledge the present gaseous DMSO measurements are the first to have been conducted in the Mediterranean area. DMSO ranges from below the detection limit (0.03 pmol·mol^{-1}) to 10.1 pmol·mol^{-1} with a mean value of 1.7 pmol·mol^{-1} (Fig. 3a). DMSO appeared to generally follow the DMS variability, especially during the first week of the campaign (28 July–3 August). However, using the full data sets no statistically significant correlation was found between these two compounds (r^2=0.1), reflecting different lifetime and removal pathways.

The lowest DMSO levels (close to the detection limit) were observed between 6–12 August as well as during the end of the campaign. The values at the end of the campaign are ascribed to very low wind speed and consequently decreased atmospheric DMS levels (Fig. 2a). On the other hand, low DMSO values between 10–11 August despite the relative high DMS levels can be attributed to high RH and aerosol surface during that period (Fig. 1). The above results indicate a significant role of heterogeneous reactions in controlling the DMSO levels, in agreement with the conclusions by Davis et al. (1998, 1999) and Sciare et al. (2001).

The observed DMSO levels fall between those reported by Putaud et al. (1999) for a Northern Atlantic coastal area (0.45 pmol·mol^{-1}) and by Sciare et al. (2000) during summertime at Amsterdam Island, situated in a similar latitude in the southern hemisphere. The DMSO values during MINOS are also lower than the values reported by Nowak et al. (2001; 4.5–11.5 pmol·mol^{-1}) over the tropical Pacific where the authors reported the possibility of an additional source of DMSO in the tropics.

The values reported for coastal Antarctica by Berresheim et al. (1998; 0.2–25 pmol·mol^{-1}) and Legrand et al. (2001; 0.4–57 pmol·mol^{-1}) are significantly higher than our data. A possible explanation is the lower temperature during these experiments that favour DMSO production through the addition channel of the OH-initiated DMS reaction. Furthermore, Davis et al. (1998) showed that the elevated DMSO levels reported by Berresheim et al. (1998), were short-term spikes due to downward mixing of air from the middle troposphere enriched in DMSO compared to the ground. Therefore, these results suggested a change in DMS chemistry and relative product yields with altitude. We have indication that similar situation occurred during MINOS. As in Berresheim et al. (1998), DMSO spikes seem to be associated most of the time with low dew point (Fig. 3b). However, clearly additional field BL and FT measurements of DMSO are needed in the Mediterranean area.

Figure 3c shows the diurnal variation of normalized DMSO during the MINOS experiment. Normalized DMSO shows a minimum around 14:00 (L.T) as well as a maximum around 19:00 (L.T). The DMSO minimum occurs almost simultaneously with the OH maximum and therefore, may be attributable to the relatively fast DMSO reaction with OH radicals (about 15 times faster than DMS+OH; Urbanski et al., 1998). The DMSO maximum at 19:00 coincides with the DMS maximum that may be explained by the very low levels of OH radicals. Furthermore, for the same time period a decrease in dew point is observed which could indicate intrusion of free tropospheric air, an event that also favors high DMSO levels. However, the existing data set does not allow a clear distinction between these two factors.

During nighttime, when DMS reacts with NO_3 radicals, DMSO levels are also decreasing and a decay rate of $3.5\times10^{-5}\,s^{-1}$ was estimated from the data presented in Fig. 3c. Based on the rate constant of DMSO with NO_3 radicals of $(1.7\pm0.3)\times10^{-13}\,cm^3\,molecule^{-1}\,s^{-1}$ and $(5\pm3.8)\times10^{-13}\,cm^3\,molecule^{-1}\,s^{-1}$ (about 2–6 times slower than DMS+NO_3; Barnes et al., 1989; Falbe-Hansen et al., 2000) and the mean NO_3 levels of 10^8 molecule cm^{-3} observed during the campaign (Vrekoussis et al., 2003), a decay rate of $(1.7–5)\times10^{-5}\,s^{-1}$ is estimated. Thus without neglecting possible dilution of DMSO by the land-breeze situation, the DMSO decrease during night could be mainly explained by reaction with NO_3 radicals.

3.3 MS^-

MS^- in the aerosol phase (Fig. 4a) ranges between 1.1 and 37.5 pmol·mol^{-1} (mean value of 11.5 pmol·mol^{-1}) within the range of summertime values reported for the area (Kouvarakis and Mihalopoulos, 2002; Kubilay et al., 2002). During the campaign the mean DMS/MS^- ratio was 2.5, close to the value of 3 reported by Mihalopoulos et al. (1992) for

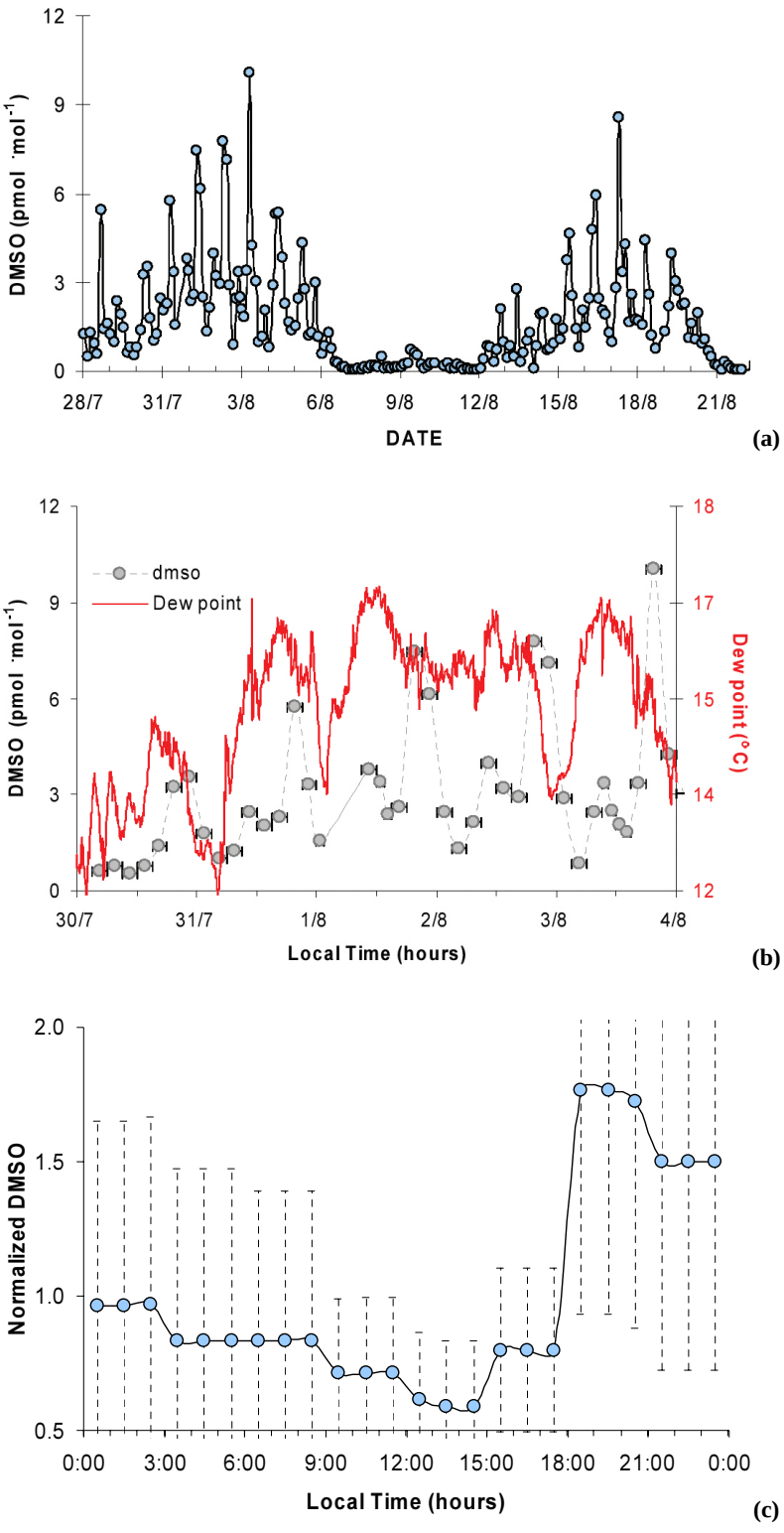

Fig. 3. Atmospheric concentrations of **(a)** DMSO, **(b)** variation of DMSO and dew point and **(c)** diurnal variation of DMSO during the campaign.

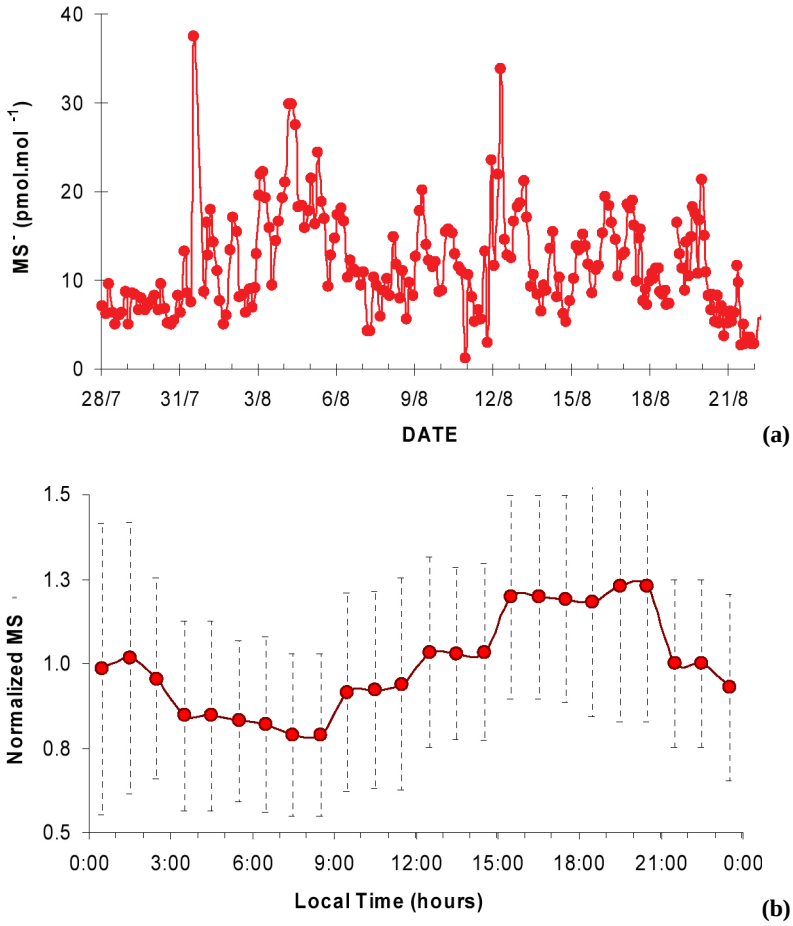

Fig. 4. Atmospheric concentrations of **(a)** MS^- and **(b)** diurnal profile of MS^- during the campaign.

summertime measurements at a coastal area in French Brittany under continental influence. These values are, however, significantly lower than those reported for the remote marine atmosphere by Ayers et al. (1995; ratio approximately 8 at Cape Grim) and by Wylie and de Mora (1996; ratio ranging from 8–13 in New Zealand). According to laboratory experiments by Yin et al. (1990) and Patroescu et al. (1999), the low DMS/MS^- ratios observed during the campaign could indicate that the oxidation of DMS to MSA under the high NO_x regime during MINOS was more efficient compared to the remote marine conditions.

Of particular interest is the relation of MS^- with DMSO during MINOS. Davis et al. (1998, 1999) proposed that heterogeneous reactions of DMSO can account for a significant part of the observed MS^-. During the periods with high RH associated in most times with high aerosol surface (in particular, 5 August and 14–18 August) DMSO significantly correlated with MS^- (DMSO/MS^- slope=0.3, r^2=0.5, $p<0.05$). On the other hand, during periods with low RH and low aerosol surface (7 and 22 August) the slope becomes smaller by a factor of 8–20, probably indicating much lower production of MS^- via multiphase reactions.

Figure 4b presents the diurnal variation of normalized MS^- during the campaign. Similar to DMS and DMSO, MS^- presents a clear maximum in the evening. From the data presented in Fig. 4b a decay rate of $7.9\times10^{-6}\,s^{-1}$ was estimated for MS^- during nighttime. This value is very close to the value of $6.6\times10^{-6}\,s^{-1}$ estimated by Huebert et al. (1996), in the equatorial Pacific marine boundary layer during the same period of the year, under clean conditions. These authors attributed the nighttime decrease of MS^- by dry deposition and entrainment velocity from the free troposphere. The good agreement between the two data sets obtained under clean and semi-polluted marine conditions, indicates that despite the important reaction of DMS with NO_3 during nighttime, MS^- is not expected to be an important oxidation product of the NO_3 initiated oxidation of DMS in the atmosphere and that this reaction is expected to lead mainly to SO_2 via the initial production of CH_3SCH_2 which is identical to the first intermediate sulfur species produced in the DMS+OH abstraction channel (Kanakidou et al., manuscript in preparation).

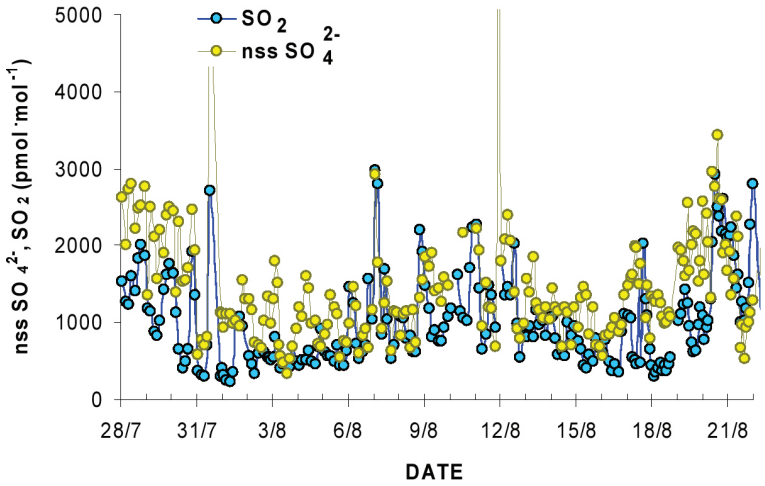

Fig. 5. Variability of SO_2 and nss-SO_4^{2-} during the campaign.

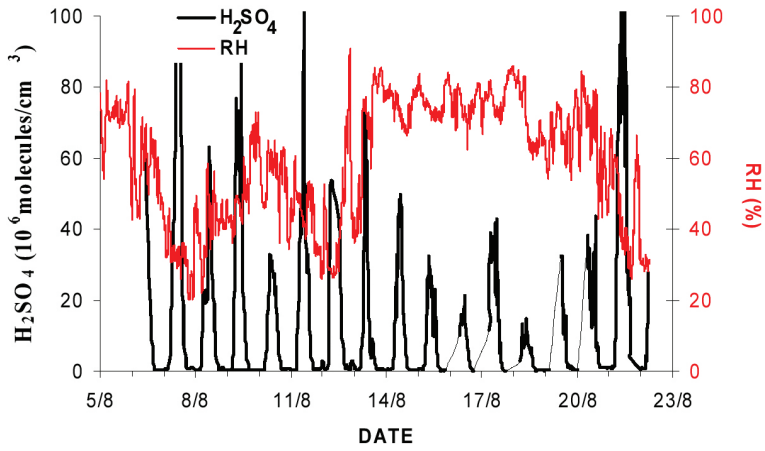

Fig. 6. Variation of atmospheric molecular concentrations of H_2SO_4 and RH during the campaign.

3.4 SO_2 and nss-SO_4^{2-}

SO_2 ranged between 220 and 2970 pmol·mol^{-1} (mean of 1030 pmol·mol^{-1}; Fig. 5). Nss-SO_4^{2-} (calculated using Na^+ as a seasalt tracer) ranged between 330 and 7100 pmol·mol^{-1} (mean of 1440 pmol·mol^{-1}) (Fig. 5). SO_2 and nss-SO_4^{2-} levels are in agreement with the summertime values reported for Crete (Kouvarakis and Mihalopoulos, 2002; Sciare et al., 2003). Both SO_2 and nss-SO_4^{2-} significantly correlated (r^2=0.79 and 0.81, respectively) with total SO_x (sum of nss-SO_4+SO_2) with slopes of 0.48 and 0.52, respectively, indicating aged air masses (Luria et al., 1996; Sciare et al., 2003). Based on a sourcereceptor model, Sciare et al. (2003) concluded that the majority of SO_2 and nss-SO_4^{2-} in the area originates from long-range transport from anthropogenic sources in central and E. Europe.

3.5 H_2SO_4

In the following paragraph the contribution of H_2SO_4 to the levels of particulate nss-SO_4^{2-} is discussed. Figure 6 shows the H_2SO_4 measurements for the whole campaign. H_2SO_4 ranged between the detection limit of $<1\times10^5$ and 9.0×10^7 molecules cm^{-3} (mean value=1.0×10^7 molecules cm^{-3}). These values are by almost a factor of 10 higher than those reported for remote marine environments in earlier studies (1.6×10^6 molecules cm^{-3}: Jefferson et al., 1998; 1.5×10^6 molecules cm^{-3}: Berresheim et al., 2002).

Meteorological factors such as relative humidity and dew point, and physicochemical parameters like aerosol surface have been reported as important factors controlling the variation of gaseous H_2SO_4 levels. Although the highest H_2SO_4 levels were observed during periods with low RH and dew

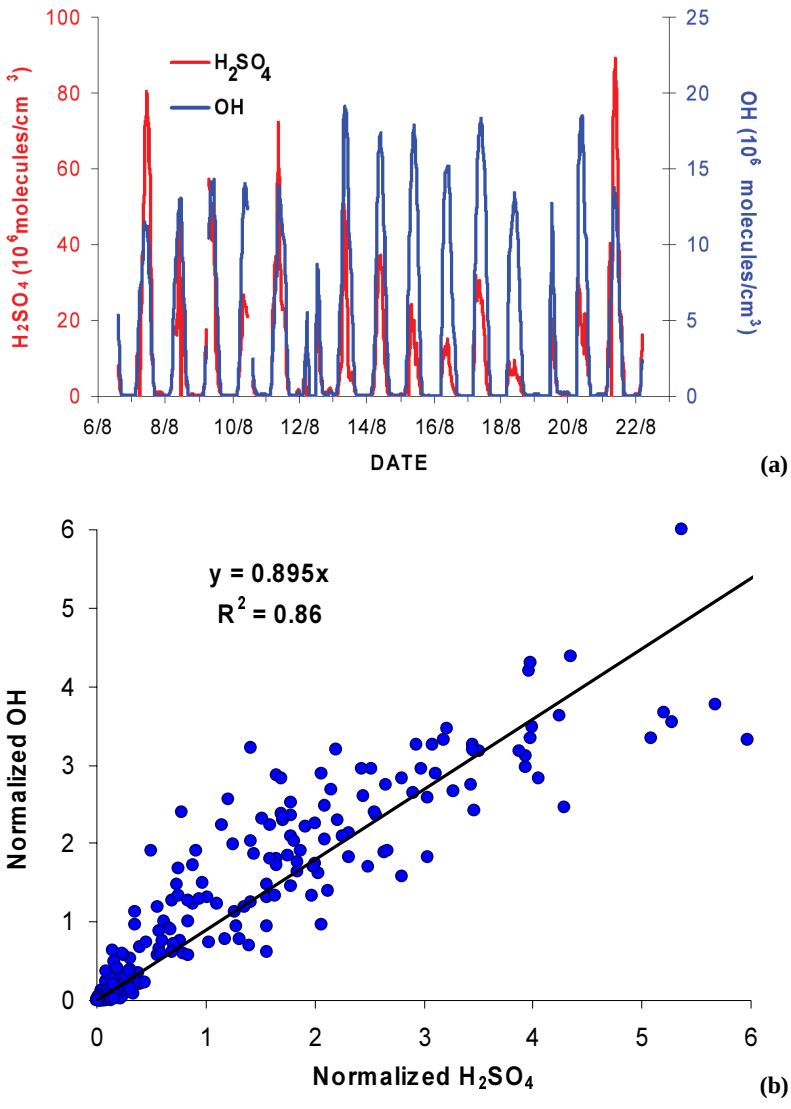

Fig. 7. Variation of **(a)** hourly means atmospheric molecular concentrations of H_2SO_4 and OH radicals and **(b)** regression between normalized H_2SO_4 and OH levels during the campaign.

point (Fig. 6) no statistically significant correlation was observed between the H_2SO_4 data and these parameters when the data is analysed for the entire measurement period.

Figure 7a presents the hourly mean variation of H_2SO_4 during the campaign in conjunction with corresponding hourly mean OH values (averaged from the 5 min time-integrated data presented by Berresheim et al., 2003). The diurnal profile of H_2SO_4 follows quite well that of OH radicals indicating a clear photochemical origin for H_2SO_4, i.e. via the reaction of gaseous SO_2 with OH radicals. The same tendency is observed when the comparison between H_2SO_4 and OH is performed on a daily-average basis. This is clearly seen in Fig. 7b demonstrating an important correlation between the normalized hourly average values of H_2SO_4 and OH for the whole campaign (slope=0.895, r^2=0.86).

To better understand the fate of H_2SO_4, we can assume that H_2SO_4 was in steady state conditions with its precursor SO_2. Under this assumption the reaction of SO_2+OH radicals is considered as the only source of H_2SO_4 and condensation (of H_2SO_4) onto the existing aerosol surface as the only sink:

$$k[SO_2][OH]=k_{cs}[H_2SO_4] \quad (1)$$

where $k=8.5\times10^{-13}$ cm^3 molecule^{-1} s^{-1} at 298 K (DeMore et al., 1997) and k_{cs} the pseudo-first order rate constant for the H_2SO_4 condensational sink that can be calculated from the Fuchs-Sutugin equation including the sticking coefficient and the aerosol surface concentration (Fuchs and Sutugin, 1970). The OH and H_2SO_4 concentrations were averaged over 2–3 h corresponding to the time-averaged measurements of SO_2. Thus, by plotting the H_2SO_4 production rate by SO_2 reaction with OH radicals (k [SO_2] [OH]) as a function of

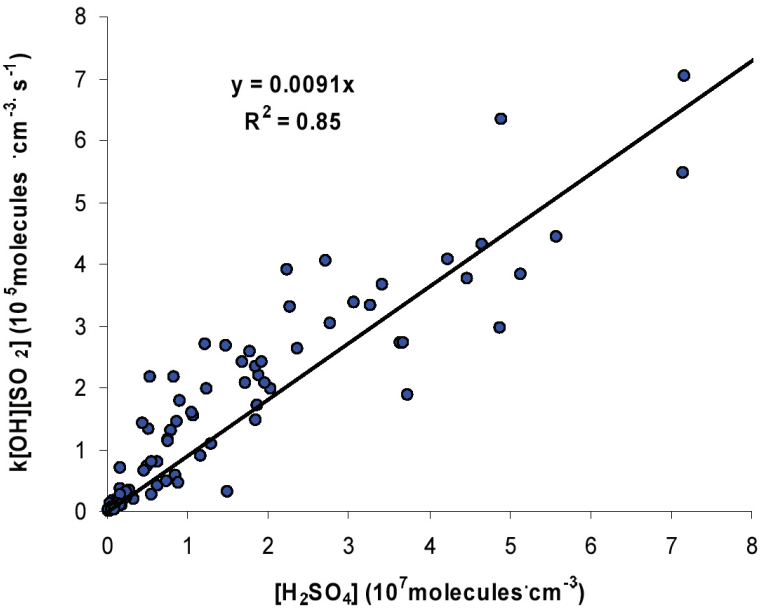

Fig. 8. Regression between k [OH] [SO_2] and [H_2SO_4] during the campaign. The slope corresponds to k_{cs} of H_2SO_4. OH and H_2SO_4 values averaged corresponding to SO_2 sampling integration times (typically 2–3 h).

Table 2. Daily variation of k_{cs} (s^{-1}) and the corresponding r^2, obtained by plotting k [SO_2] [OH] as a function of [H_2SO_4]

Date	k_{cs} (s^{-1})	r^2
7/8/2001	0.0089	0.94
8/8/2001	0.0084	0.94
9/8/2001	0.0086	0.95
10/8/2001	0.0133	0.98
11/8/2001	0.0122	0.98
12/8/2001	0.0071	0.70
13/8/2001	0.0087	0.86
14/8/2001	0.0114	0.98
15/8/2001	0.0148	0.94
16/8/2001	0.0167	0.74
17/8/2001	0.0129	0.71
18/8/2001	0.0149	0.98
19/8/2001	0.0144	1.00
20/8/2001	0.0134	0.84
21/8/2001	0.0072	0.98

[H_2SO_4], a significant correlation was obtained on a daily basis (Table 2 and Fig. 8) and when all the 2–3 hourly data are plotted together (Fig. 8). Regression slopes correspond to kcs ranging between 0.0071 s^{-1} and 0.0167 s^{-1} (mean value of 0.011 s^{-1}) during the experiment. It is interesting to note that the daily averaged kcs values follows remarkably the variation of RH (r^2=0.6, data not shown), indicating shorter lifetime under increased RH conditions, probably due to hygroscopic aerosol growth and thus increased aerosol surface (Fig. 1). The very short lifetime for H_2SO_4 (1–2.5 min) is in agreement with the assumption of steady state condition between production and loss to particles and consistent with the good correlation between OH and H_2SO_4 shown above.

Based on the kcs values derived from the regression analysis, the surface aerosol data, and the use of the Fuchs-Sutugin equation, the mass accommodation coefficient "α" can be calculated for H_2SO_4 at a mean temperature during MINOS of 300 K. The calculated mean value of 0.52±0.28 (geometric mean of 0.49) is in excellent agreement with the value of 0.5 derived by Jefferson et al. (1997) at 298 K on the basis of laboratory experiments. The use of an accomodation coefficient of 0.5 is supported by field studies of H_2SO_4 production and loss (Eisele and Tanner, 1993) and model results of aerosols nucleation and growth (Kerminen and Wexler, 1995).

The above data set can be used to investigate whether the observed H_2SO_4 levels can account for the observed levels of particulate nss-SO_4^{2-}. From mass transfer considerations the aerosol sulphate concentration derived from condensation of gas phase H_2SO_4 on to the aerosol surface can be estimated by the following equation:

$$\text{nss}-SO_4^{2-}=k_{cs}[H_2SO_4]\Delta t \quad (2)$$

By using mean values of $1.2\times10^{-2}\,s^{-1}$ and 10^7 molecules cm^{-3} for k_{cs} and H_2SO_4 respectively, the time needed to produce measured nss-SO_4^{2-} levels of 1.4 nmol.mol^{-1} is 3.5 days, which falls within the range of estimated residence times of submicronic aerosols in the marine boundary layer. Thus condensational uptake of H_2SO_4 could dominate nss-SO_4^{2-} formation in the Mediterranean area at some distance from major industrial sources

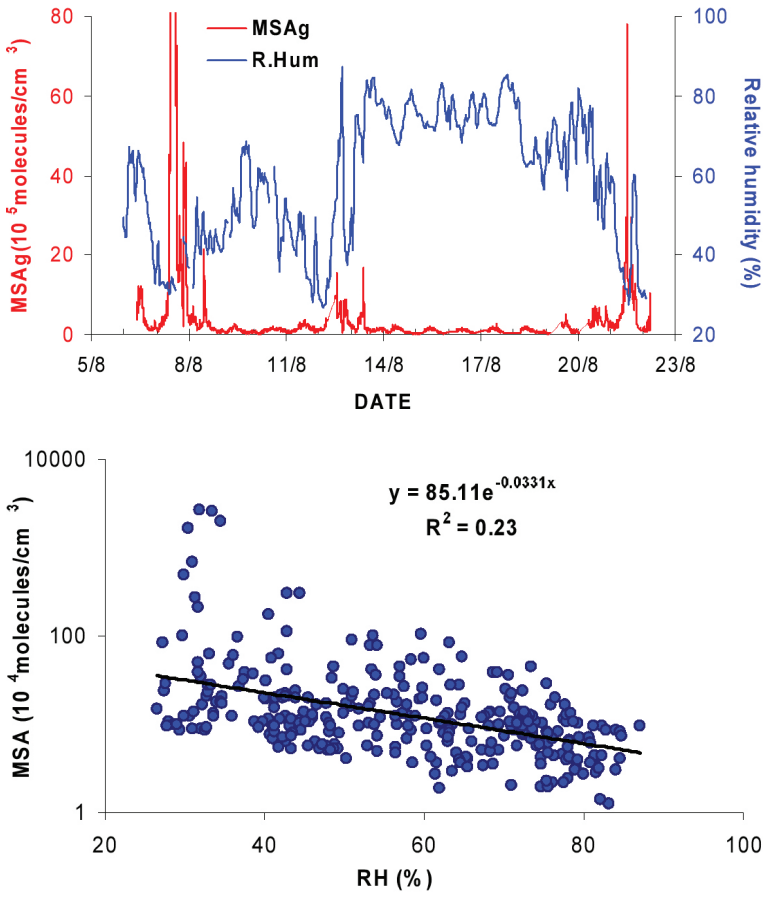

Fig. 9. Variation of **(a)** atmospheric molecular concentrations of MSAg and RH and **(b)** regression between MSAg and RH during the campaign.

Table 3. MSA yield (Y) derived during the experiment by using the equation k [DMS] [OH] [Y]=k_{cs} [MSA] and the corresponding r^2

Date	MSA yield* per mil (‰)	r^2	MSA yield** per mil (‰)	r^2
10/8/2001	1.2	0.98	–	–
11/8/2001	1.5	0.67	–	–
14/8/2001	3.4	0.88	4.2	0.56
15/8/2001	2.8	0.93	3.0	0.75
16/8/2001	–	–	1.0	0.65
17/8/2001	–	–	1.3	0.78
18/8/2001	–	–	1.0	0.88

*=k_{cs} derived using the Fuchs-Sutugin equation (Fuchs and Sutugin, 1970) **=k_{cs} derived from H_2SO_4 data but corrected using an accommodation coefficient a" for MSA equal to 0.075 as derived from the equation of De Bruyn et al. (1994).

with direct nss-SO_4^{2-} emissions. Box-model simulations (Kanakidou et al., manuscript in preparation) indicate that homogeneous reactions can account for about two thirds of the observed nss-SO_4^{2-} levels in the area, with heterogeneous reaction of SO_2 on aerosol surface accounting for the rest. Note however that this situation is specific to the region during summer-time due to the absence of clouds.

3.6 MSA(g)

Figure 9 shows the results of the gas phase MSA measurements for the whole campaign. MSA ranged from below the detection limit (3×10^4 molecules cm^{-3}) to 3.7×10^7 molecules cm^{-3}. The average value recorded during the campaign (5.2×10^5 molecules cm^{-3}) is

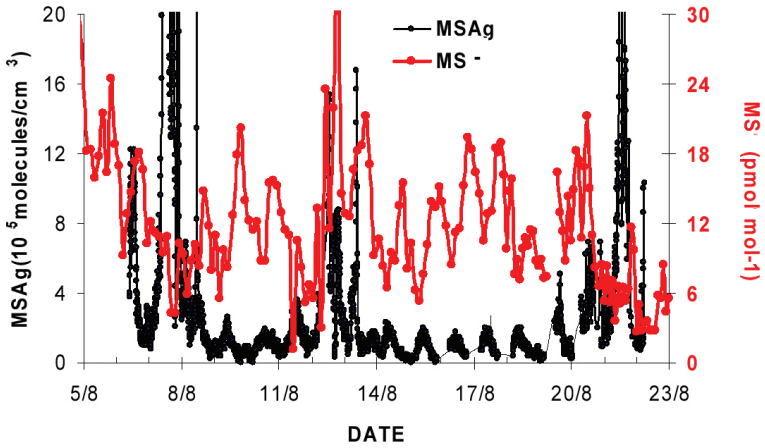

Fig. 10. Variation of atmospheric molecular concentrations of MSAg and aerosol MS^- during the campaign.

slightly lower than those recorded during previous studies (9.5×10^5 molecules cm^{-3}: Jefferson et al., 1998; 1.2×10^6 molecules cm^{-3}: Berresheim et al., 2002). However these studies have been conducted under significantly lower temperatures, and laboratory experiments (Arsene et al., 1999) have shown that temperature significantly influences the MSA yields from DMS oxidation.

MSA appears to be more sensitive than H_2SO_4 to variations of RH, dew point and aerosol surface, probably due to its higher vapor pressure (Kreidenweis and Seinfeld, 1988, Berresheim et al., 2002). This is seen in Figs. 9a and b, presenting the variation of MSA and RH during the experiment (Fig. 9a) as well as the regression between these two parameters (Fig. 9b; r^2=0.23, $p<0.001$). Similar relations with RH have been observed when MSA is also plotted against aerosol surface. MSA clearly decreases with increasing RH, dew point and aerosol surface. Especially all the significant elevations in MSA concentration occur under low RH (smaller than 30%), in agreement with the observations over the equatorial Pacific by Mauldin et al. (1999) or on the western coast of Ireland at Mace Head (Berresheim et al., 2002). When compared to H_2SO_4 the different behavior can be explained by the higher vapor pressure of MSA (Kreidenweis and Seinfeld, 1988). The larger volatility of MSA compared to H_2SO_4 can be also seen in Fig. 10, presenting the variation of MSA and MS^- during the experiment. An inverse correlation can be seen between gaseous and particulate MSA levels almost during the whole campaign. Especially at low RH the partitioning is such that the MSA resides mainly in the gas phase. Conversely, at high RH values most of the MS^- exists in particulate form.

Figures 11a and b present two well distinguished cases of diurnal variation of MSA during the campaign: In the first, MSA (Fig. 11a) follows closely the OH variation, as in the case of H_2SO_4. In the second, in addition to the MSA peak occurring in conjunction with OH, several secondary peaks also can be distinguished (Fig. 11b). A closer look of these cases demonstrates that the first type of profile with a single maximum occurred mainly under periods with high RH (14–18 August). The covariation of MSA with OH radicals indicates a clear photochemical origin of MSA, i.e., possibly via the reaction of gaseous DMS, its precursor, with OH radicals. It indicates also that MSA is not a product of the nighttime reaction of DMS with NO_3 radicals, in agreement with the conclusions presented in Sect. 3.3. The MSA peaks observed in the second type of profile coincide with periods of low RH or dew point and are in very good agreement with the general tendency of MSA relative to RH observed during the entire campaign.

The above data set and the simultaneous measurements of DMS, gaseous MSA and particulate MS^- offer a unique opportunity to estimate the yield (Y) of MSA from the OH initiated oxidation of DMS. Laboratory experiments showed that this yield is highly dependant on NO_x and can range from less than a few percents (Patroescu et al., 1999) to values as high as 50% (Hatakeyama et al.,1982). For this purpose, we assume that MSA is in steady state conditions with its precursor DMS. Under this assumption the reaction of DMS+OH radicals is considered as the only source of MSA, while condensation onto the existing aerosol surface is the only sink for MSA:

$$k[DMS][OH][Y]=k_{cs}[MSA] \quad (3)$$

where k_{OH+DMS} is taken from Hynes et al. (1986) and kcs is the pseudo-first order rate constant for the MSA condensational sink and is derived using two different approaches. The first using the Fuchs-Sutugin equation (Fuchs and Sutugin, 1970) and the second using the k_{cs} derived from H_2SO_4 but corrected using an accommodation coefficient "α" for MSA equal to 0.075 (De Bruyn et al., 1994) at T=300 K. The second approach is used since aerosol surface data are not available for the whole period. Table 3 presents the MSA yield as well as the correlation coefficient estimated by

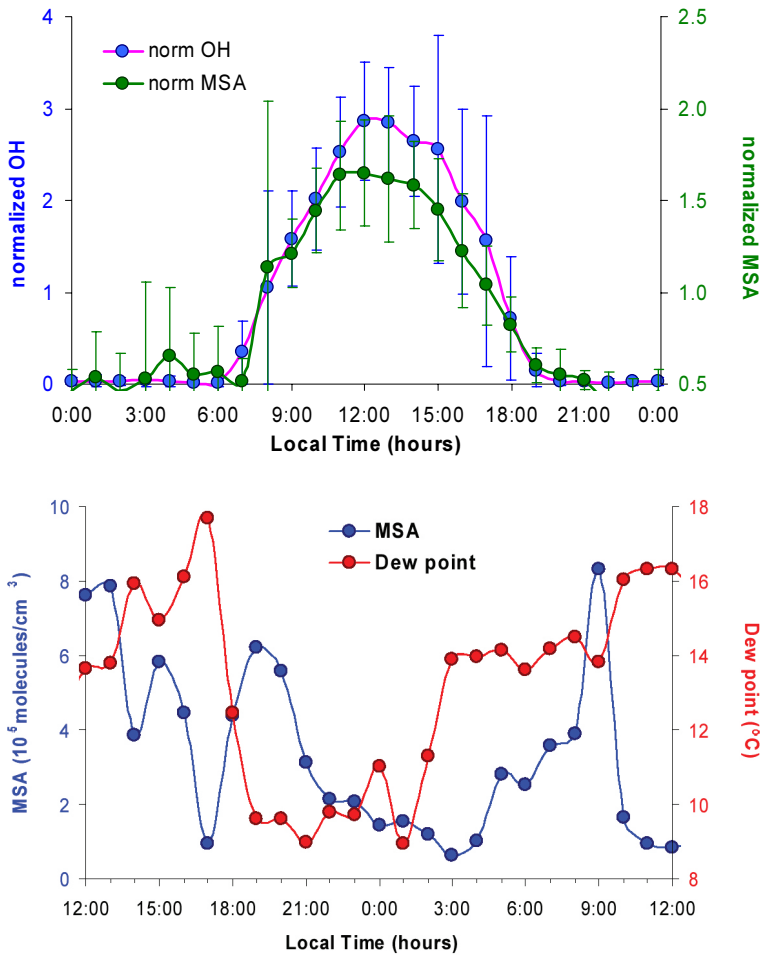

Fig. 11. Normalized profiles of MSA during two cases: **(a)** during periods with elevated (>40%) RH and **(b)** during a day with low RH (12–13 August). The variation of dew point is also plotted during the second case.

applying the above Eq. (3) for each day. A quite good agreement was found in the MSA yield using these approaches.

Statistically significant correlations were observed only during days with no important changes in RH, dew point and air masses origin, since these parameters can influence the MSA levels.

By considering only the periods with statistically significant correlation coefficients (r^2), the calculated MSA yield is very low and ranges between (10^{-3} to 4.2×10^{-3}, mean of 1.9×10^{-3}), but it is in good agreement with the value of 5×10^{-3} deduced by Davis et al. (1999) for the equatorial Pacific boundary layer.

To investigate whether the observed MSA levels can account for the observed particulate levels MS^-, as in the case of H_2SO_4, by assuming mass transfer considerations, the MS^- levels from condensation of MSA on to the aerosol surface can be derived by the following equation:

$$MS^- = k_{cs}[MSA]\Delta t \quad (4)$$

Given the important variability of MSA as a function of RH we have considered two different cases: 1) the days with particularly low RH (7 and 21–22 August) and 2) the rest of the campaign.

For the first case, the time needed to reproduce the measured MS^- levels range from 16 h on 7 August to 2.6 days for the period 21–22 August. These lifetimes fall within the range of residence time of submicronic aerosols in the marine boundary layer, indicating that the measured MS^- levels are in equilibrium with the corresponding gas phase MSA.

For the remaining part of the campaign, the time needed to reproduce the measured MS^- levels has been calculated to be around 19 days, period significantly higher than the residence time of submicronic aerosols. Thus gaseous to particle conversion of MSA can not account for the observed levels of MS^-. The above results are in agreement with those of Davis et al. (1998, 1999) who proposed that multiphase reactions involving DMSO can account for the observed MS^- levels. Bardouki et al. (2002) conducted aqueous phase experiments

and reported that DMSO reacts very fast with OH radicals leading to MS^- formation with almost 100% yield.

By using the DMSO concentrations observed during the campaign the production rate of MS^- via multiphase reactions can account for the measured MS^- levels. Details on these calculations can be found in Kanakidou et al. (manuscript in preparation).

4 Conclusions-Implications for the S cycle in the area

During the MINOS campaign conducted in July–August 2001 in a coastal location in Crete, measurements of DMS and its major gas and aerosol phase oxidation products as well as the dominant radical species (OH and NO_3) involved in DMS oxidation have been performed. The analysis of this extensive data set shows that:

1. DMS variability during the MINOS campaign in the coastal area of Crete was mainly governed by NO_3 radicals and dilution due to land-to-sea breeze occurrence. DMS presents a clear night-time minimum and thus DMS diurnal variation is opposite to that observed in the remote marine atmosphere of the southern hemisphere.

2. A very short lifetime has been calculated for H_2SO_4 (1–2 min) with respect to condensation on the existing aerosol surface (k_{cs}). Based on the calculated k_{cs}, the surface aerosol data and the use of the Fuchs-Sutugin equation, the average mass accommodation coefficient "α" for H_2SO_4 at 300 K was calculated to be 0.52 ± 0.28. Finally, condensational uptake of H_2SO_4 is proved to play a very important role in the nss-SO_4^{2-} budget over the study area.

3. MSA seems to be very sensitive to variations of RH, dew point and aerosol surface compared to H_2SO_4. This can be explained by the higher vapor pressure of MSA compared to H_2SO_4. From the simultaneous measurements of DMS, gaseous MSA and OH radicals, a MSA yield from the OH-initiated oxidation of DMS can be calculated, which ranges between 10^{-3} and 4.2×10^{-3} (mean of 1.9×10^{-3}) representing the lowest values reported so far in the literature. These values should be considered as upper limits since laboratory work clearly shows enhancement of MSA yield in the presence of NO_x as in our case.

4. With the exception of the periods with low RH, gas to particle conversion of MSA can not account for the observed levels of MS^-. This result is consistent with the observations by Davis et al. (1998, 1999) based on two campaigns performed in Antarctica and equatorial Pacific. Given the completely different conditions under which these three experiments have been performed, the above results indicate that under humidity conditions typical of the marine boundary layer (RH higher than 40%), multiphase reactions involving DMSO could account for the observed MS^- levels. These reactions should be taken into account in any attempt to model the biogenic sulfur cycle.

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