Mullite interaction with bismuth oxide from minerals and sol-gel processes

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Abstract

The high temperature treatment of kaolinite-muscovite alternate layers doped by bismuth oxide was studied by TGA and DTA, X ray diffraction, and electron microscopy (SEM). Thermal analyses shows that the two main transformation stages are the dehydroxylation of phyllosilicates and the structural reorganization of the whole assembly. During dehydroxylation, a progressive decrease of the structural order of kaolinite and muscovite occurs. It is more significant at temperature above 1000°C and mullite, glass and Al-rich oxides are formed. Mullite exhibits an accentuated acicular morphology along 3 preferential orientations in relation to the remaining structure in the basal (001) _musc planes of the former muscovite. With addition of bismuth oxide, SEM observations point to the strong accentuation of mullite growth along the c axis, at the very low temperature of 1050°C. This behavior was studied with various mullite-bismuth oxide compounds at temperatures ranging from 1000°C to 1400°C. Sol-gel processing were performed with TEOS (C₂H₅O)₄Si, aluminum nitrate Al(NO₃)₃.9H₂O and bismuth nitrate Bi(NO₃)₃. From Rietveld refinements of X ray diffraction patterns it was found that bismuth doesn’t form a specific phase with mullite during the mullite nucleation and growth. At low Bi₂O₃ content, mullite coexists with amorphous silico-aluminate phases, even at high temperature. When the Bi₂O₃ content increases above 12mol%, mullite is no longer crystallized. Above 750°C, aluminum bismuth oxide coexists with silicon bismuth oxide. The liquidus temperature is below 1100°C and decreases with Bi₂O₃ additions. Above the liquidus, only corundum and a liquid coexist. A tentative mullite-Bi₂O₃ binary phase diagram is then proposed. These results evidence the Bi₂O₃ role on the early and large mullite growth in kaolinite-muscovite materials. Particularly, a very low quantity of a low viscous phase favor the local mobility of diffusing species and promotes the nucleation and growth of very large mullite crystals. In this way, highly textured and dense ceramics are obtained at low sintering temperature. Thin substrates (<500µm) present specific mechanical behaviors with the deviation of cracks along mullite layers.

Keywords: Muscovite, Kaolinite, Mullite, Bi₂O₃.

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Introduction

Mullite is an important high-technology ceramics for high temperature applications [1]. The addition of cations in the reaction mixture of mullite allows the improvement of the physical properties of sintered materials. Ca²⁺ or Ba²⁺ can added [2], but the addition of Bi³⁺ appears to be a promising way for mullite formation in compositions. It was also observed the effects of some oxides (SnO₂, Sb₂O₃) in the phase formation and the microstructural evolution of mullite ceramics [3]. In general, Bi₂O₃ doping have a much better effects on densification than SnO₂ and Sb₂O₃, whereas the latter is an effective additive to decrease the temperature of mullite phase formation in diphasic gels. The addition of other oxides which form a solid solution with mullite such as Fe₂O₃ or TiO₂ induces the increase of both the unit cell volume and of the crystallite size, when they are added in large proportion (up to 10mol%). In the case of Fe₂O₃, it leads to the decrease of the crystallization temperature (~200°C). Simultaneously the crystal morphology is also changed from parallelelepiedic to acicular [4].

In mullite ceramics, some additives were considered to reduce the viscosity of the SiO₂-rich liquid phase, favoring a higher mobility of the diffusing species. The dissolution of Al₂O₃ species into the SiO₂-rich liquid phase controls the rate of the mullite formation [5]. Also, the aluminum compounds might act as nuclei for crystallization of mullite crystal and the mullite formation temperature is significantly decreased. The aim of the study is to determine the phase formation against temperature, from mullite-bismuth oxide compounds, which are synthesized using diphasic gels. The role of existing liquid phases is observed in the case of mullite crystallization in muscovite-kaolinite compounds.

Experimental

Mullite powders are usually prepared by methods such as co-precipitation using salts [2], sol-gel [1,6] and spray pyrolysis [7-9]. In this study, mullite compositions were synthesized with sol-gels using TEOS (C₂H₅O)₄Si, aluminum nitrate Al(NO₃)₃.9H₂O and bismuth nitrate Bi(NO₃)₃. Aluminum nitrate and bismuth nitrate were dissolved into aqueous solution of nitrate acid. TEOS was diluted in ethanol. Solutions were subsequently mixed at 60°C and become gel after
4 days. Gels were dried at 110°C for 24 hour. In table I the stoechiometry of used compounds are presented and reported in the ternary diagram SiO₂-Al₂O₃-Bi₂O₃ in figure 1 [10-12]. All compositions are along and beside the pseudo-binary mullite-bismuth oxide.

<table>
<thead>
<tr>
<th>ref</th>
<th>Stoechiometry</th>
<th>Bi₂O₃ wt% - mol%</th>
</tr>
</thead>
<tbody>
<tr>
<td>1</td>
<td>3Al₂O₃ . 2SiO₂</td>
<td>0 – 0</td>
</tr>
<tr>
<td>2</td>
<td>3Al₂O₃ . 2(Si₀.₈₄.Bi₀.₁₆) O₂</td>
<td>15- 19</td>
</tr>
<tr>
<td>3</td>
<td>3(Al₁.₈.Bi₀.₂)O₃ . 2SiO₂</td>
<td>26 - 33</td>
</tr>
<tr>
<td>4</td>
<td>3Al₂O₃ . 2(Si₀.₇.Bi₀.₃) O₂</td>
<td>26 - 42</td>
</tr>
<tr>
<td>5</td>
<td>3Al₂O₃ . 2SiO₂ + 5 mol% Bi₂O₃</td>
<td>5.2 – 5</td>
</tr>
<tr>
<td>6</td>
<td>3Al₂O₃ . 2SiO₂ + 20 mol% Bi₂O₃</td>
<td>18 – 20</td>
</tr>
<tr>
<td>7</td>
<td>3Al₂O₃ . 2SiO₂ + 40 mol% Bi₂O₃</td>
<td>30 - 40</td>
</tr>
</tbody>
</table>

Tab. I : Compositions of mullite compounds

X-ray diffractions were performed with a D5000 (Bruker) and pattern refinements were obtained by the Rietveld method, using the FullProf software. The thermal behavior of gels after calcinations at different temperatures up to 1100°C were studied by DTA (Setsys, SETARAM) under air atmosphere at a constant heating rate of 10°C min⁻¹. Curve backgrounds were corrected by preliminary experiments with α alumina.

Results and discussion

DTA study

Differential thermal analysis (DTA) of mullite gel exhibits typical endothermic peaks at 550-650°C, due to the loss of hydroxyls and an exothermic peak at ~980°C related to the crystallization of Al-Si spinel. A second weak exothermic peak at ~1250°C corresponds to a secondary mullitization process and cristobalite crystallization [13,14].

DTA of samples containing bismuth (4, 5, 6 and 7 in table I), were performed after heating at 500°C to preliminary remove all volatile species. The DTA curve of sample 4 in figure 2 presents four exothermic peaks at 530, 892, 958 and 1044°C.

With sample 5, the DTA curve in figure 2 indicates four exothermic peaks at 530, 905, 971°C and one weak peak between 1000 and 1100°C. DTA curve of sample 4 differs to that of sample 5 in the intensity of the two main exothermic peaks. With sample 5 the crystallization process is slightly more accentuated and occurs at a higher temperature.

The sample 6 in figure 2 exhibits three DTA exothermic peaks at 526, 885 and 1210°C and one small endothermic peak at 730°C.

In the case of sample containing 7, the DTA curve (figure 2) shows two endothermic peaks at 725°C and 875°C and two exothermic peaks at 815°C and 1293°C. In this sample with the highest bismuth content, some peaks should be related to the transition of monoclinic α-phase→δ (fcc-face central cubic) of bismuth oxide, which occurs at 705-1013°C. The wide variation of the transition temperature is related to the purity of samples. The melting point of Bi₂O₃ is at 840°C [15].

Fig. 2: DTA curve of samples 4, 5, 6 and 7.
bismuth oxide. In sample 7 (40 mol.% bismuth), we observe three peaks, which are originated from transitions of bismuth oxide (735°C, 815°C and 875°C). The high temperature peaks at 1210°C for mullite with 20 mol.% Bi and at 1193°C in mullite with 40 mol.% Bi are due to the structural transformation of Al-Si spinel to new phases. In general, the curve slope at high temperature increases with the bismuth content.

X ray experiments

In a first approach, dried gels of samples 1, 2 and 3 were heated at 1350°C for 3 hours before XRD measurements. The Rietveld method was used to determine the mullite structural characteristics with bismuth additions in initial compositions. In general, the orthorhombic mullite structure and the stoichiometry 3Al2O3 2SiO2 are always obtained. When the bismuth content increases, but a significant background variation is observed from an increasing content of liquid phase.

For a better understanding of transformations with temperature, X-ray diffractometry was used with heated samples 4, 5, 6 and 7, at typical temperatures near the onset and end of peaks in DTA curves of figure 2. All thermal cycles with a 30 min holding time were followed by a quenching stage to retain a steady state of phases. Temperature values for samples are reported in table II.

<table>
<thead>
<tr>
<th>Sample</th>
<th>Temperature</th>
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<tbody>
<tr>
<td>4</td>
<td>1000°C, 1100°C</td>
</tr>
<tr>
<td>5</td>
<td>1000°C, 1100°C</td>
</tr>
<tr>
<td>6</td>
<td>1000°C, 1100°C, 1250°C</td>
</tr>
<tr>
<td>7</td>
<td>770°C, 845°C, 930°C, 1200°C</td>
</tr>
</tbody>
</table>

Tab. II : Quenching temperatures of samples 4, 5, 6, 7

Patterns of sample 5, containing 5wt% Bi2O3, heated at 1000°C and 1100°C are in figure 3. It is seen
that only orthorhombic mullite with the stoechiometry 3Al₂O₃ 2SiO₂ is identified.

Figure 4 shows X-ray patterns of sample 4 containing 26wt% of Bi₂O₃ (3Al₂O₃2(Si₀.₇Bi₀.₃)O₂) heated at 1000°C and 1100°C. At 1100°C, peaks from mullite, aluminum bismuth oxide and aluminum bismuth oxide are clearly recognized. Identified phases are orthorhombic mullite Al₂(Al₂.₈Si₁.₂)O₉.₅₄ (pattern 01-084-1205), cubic bismuth silicon oxide Bi₁₂Si₀.₇₅O₂₀ (pattern 01-084-0090) and aluminium bismuth oxide Al₄Bi₂O₉ (pattern 00-025-1048). At 1100°C, peaks from sample 4 indicate that aluminum bismuth oxide vanishes and that the relative content of bismuth silicon oxide increases in comparison to that of mullite.

For both patterns the large variation of the background indicate the presence of an amorphous phase or a liquid. Correspondingly, the smooth endothermic variation in the DTA curve of figure 2 above 958°C should be related to the liquid formation.

In figure 5, sample 6 with 18wt.% bismuth heating at 1000°C has peaks from bismuth silicon oxide and aluminum silicon oxide. But a new phase is evidenced with a major peak at 10.44°2θ. This new phase also occurs in sample 4 at 1000°C and in all samples containing more than 5wt%, below 1100°C. The stoechiometry of this phase is supposed to be similar to a Aluminium Silicon Bismuth oxide, but further work are necessary to precise the structural characteristics. At the 1000°C temperature, the background variation indicates the occurrence of a liquid, in accordance to the DTA curve, which shows a significant endothermic variation at 885°C.

At 1100°C and 1250°C, sample 6 show the progressive formation of sharp peaks from a single phase, very similar to corundum (pattern 00-010-0173). Structural characteristics of this phase were identified after heating at 1200°C for 14 hours holding time and sampling a single crystal. The obtained structural parameters are the following: system hexagonal (rhomboedric, R-3C or R3C); a = 4.735 Å; b = 4.735 Å; c = 12.941 Å; α = 90°; β = 90°; γ = 120°. These data are close to that of corundum (α-Al₂O₃).

For a higher Bi₂O₃ content (figure 6), sample 7 with 30wt% Bi₂O₃ heated at 770°C has peaks from bismuth silicon oxide, aluminum silicon oxide and bismuth silicate (sillenite Bi₁₂SiO₂₀, pattern 00-037-085). After heating at 845°C and 930°C, peaks from sillenite vanished and peaks from bismuth silicon oxide and aluminum bismuth oxide are present. At 1200°C, peaks from the same phase already observed in sample 6, are very similar to corundum peaks.

The major peak at 10.44°2θ is also observed at 770°C, 845°C and 930°C, as in samples 6. It must be precisely identified in a future work. From these experiments, we note the occurrence of a liquid at temperatures above 930°C that corresponds to the strong endothermic variation in the DTA curve of figure 2.

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Three-phases diagram

A literature review reveals the lack of information about the SiO₂ – Al₂O₃ – Bi₂O₃ ternary system. Only two-phase diagrams are reported: Al₂O₃ – SiO₂ [11], Al₂O₃-Bi₂O₃ [10], Bi₂O₃ – SiO₂ [12]. The Al₂O₃ and SiO₂ rich regions of two binary diagrams are in figure 7 a-b respectively. We observe in the Al₂O₃ – SiO₂ system the presence of the Al₂Bi₂O₉ phase which was observed in compositions 6 and 7 containing 20-40mol% Bi₂O₃. Beside, the Alumina rich region indicates a first liquid at 1070°C. This temperature is
very similar to the temperature range for liquid phases detection in compositions 4, 5, 6 and 7 in figures 4 and 5. The Bi$_2$O$_3$ – SiO$_2$ system is different as it contains the Bi$_4$Si$_3$O$_{12}$ phase, which is different from the Bi$_{12}$Si$_{0.87}$O$_{20}$ detected in samples 6 and 7 at 1000°C and 1100°C. But the first liquid appears at the 1020°C temperature, close to the liquid temperature of our samples.

A tentative pseudo-binary diagram mullite-Bi$_2$O$_3$ is presented in figure 8. The liquid limit is slightly below 1000°C, for a large Bi$_2$O$_3$ content. Above approximately 1100°C, corundum coexists with the liquid and this behavior is similarly observed in the Al$_2$O$_3$-Bi$_2$O$_3$ system above 1070°C. For compositions outside the mullite - Bi$_2$O$_3$ axis (samples 2, 3 and 4) the crystallized phases are different (composition plotted in gray shade in figure 7), but the liquid temperature is not different.

Mullite formation in phyllosilicate mixes

The effect of Bi$_2$O$_3$ addition on phase formation and microstructural evolution of mullite ceramics from kaolinite-muscovite minerals was investigated. In this material, Bi$_2$O$_3$ was added as a thin coating onto muscovite flakes, before the muscovite-kaolinite mixing, using a bismuth nitrate aqueous solution. The Bi$_2$O$_3$ to mullite ratio in sintered material is close to 5mol%. But, it can be supposed that a local high concentration of bismuth oxide exist at crystallite interfaces. In these restricted zones, local interactions could be described by the behavior of compounds containing high concentration of Bi$_2$O$_3$ in the pseudo binary diagram of figure 8.

With muscovite-kaolinite compound, SEM observations point to the acceleration of mullitization with the presence of Bi$_2$O$_3$. In figure 9, the surface of muscovite sintered at 1250°C for 1 hour holding time shows large anisotropic mullite grains, which average length attains 30µm.

In this material, whereas the probable occurrence of crystallized compounds of Bi$_2$O$_3$ with alumina and silica, no reactions are found by x-ray in sintered materials. Also, Bi$_2$O$_3$ was not detected even for samples with high Bi$_2$O$_3$ addition. The most interesting point is the significant mullite crystallization and growth at low temperature. Simultaneously, Bi$_2$O$_3$ highly favors the sintering densification and this effect is attributed to the specific role of a local low-viscosity liquid phase in the systems.

Conclusion

Mullite doped by bismuth was synthesized by a sol-gel process. Effect of Bi$_2$O$_3$ on phase formation and microstructural evolution of mullite ceramics were investigated by different methods. They all point to the important role of Bi$_2$O$_3$ on mullite formation. At weak bismuth content, mullite is crystallized in all mixes below 1000°C and above. When bismuth content exceed about 10mol%, mullite is no longer crystallized and both bismuth silicon oxide, aluminum bismuth oxide, sillenite and corundum are present in variable contents. Mullite is detected in sample with high bismuth content at temperature exceeding 1200°C and for compositions beside the mullite Bi$_2$O$_3$ axis. For all studied compositions, a liquid phase appears at temperature in the 950-1000°C temperature range. In silico-aluminate compositions, the occurrence of a liquid when Bi$_2$O$_3$ is added highly favors the mullite growth at temperature below 1200°C. It is favored by local concentrations at interfaces of a transient liquid phase, which enhance the mobility of species.

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References